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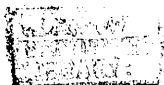
**X-RAY DIFFRACTION ANALYSIS OF ORGANOMETALLIC
COMPOUNDS WITH CATALYTIC PROPERTIES
VOLUME I**

**Submitted to the University of Glasgow in partial
fulfilment of the requirements for the degree of
Doctor of Philosophy in the Faculty of Science**

by

Graeme Douglas

**Chemistry Department
September 1990
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PREFACE

The study, by use of X-ray diffraction analysis, of the structures of organometallic complexes with catalytic properties has proved to be of great importance in understanding the chemistry of catalytic processes. It is this technique that has been vital in determining the structures of the complexes reported in this thesis.

The thesis is divided into three sections. Section 1 discusses the structural features of new dppm-stabilised triplatinum complexes derived from the parent cluster $[\text{Pt}_3(\text{CO})(\text{dppm})_3]^{2+}$ (dppm = $\text{Ph}_2\text{PCH}_2\text{PPh}_2$). The structural work was carried out to determine how closely Pt_3 clusters could model the behaviour of similar Pt_3 species which occur on platinum surfaces. The new complexes are derived by chemical modification of the $\text{M}_3(\text{dppm})_3$ nucleus which is unusually stable to fragmentation into species containing one or two metal atoms. The structure analyses reported here have been of critical importance in the development of this new field of cluster chemistry.

Brief reviews on binuclear dppm complexes and on platinum cluster chemistry have been carried out and earlier work on $\text{M}_3(\text{dppm})_3$ species is also reported. A brief analysis on the conformations of the dppm ligand has also been carried out.

The work reported in Section 2 relates to the addition of an alkyne to the complex $[\text{CpM}(\text{SR})(\text{hfb})_2]$, where $\text{M} = \text{Mo}$ or W , $\text{Cp} = \eta^5\text{-C}_5\text{H}_5$, $\text{R} = \text{an alkyl or aryl group}$ and $\text{hfb} = \text{F}_3\text{CC}=\text{CCF}_3$. The products of this reaction can adopt at least twelve different structural forms. The form isolated

depends on several factors including the natures of the thiolato substituent R, of the metal atom and of the incoming ligand. This work is directed to enhancing our understanding of transition metal catalysed oligomerisation and polymerisation reactions of alkynes.

The structures of other related products are also reported including those of the thermolysis reaction of a $[\text{CpM}(\text{SR})(\text{hfb})_2]$ complex and those of the addition of ligands other than alkynes to the starting complex.

Reported in Section 3 are three unrelated structure analyses. The sulphimide complex of palladium is of interest because of its similarity to *cis*- $[\text{PtCl}_2(\text{NH}_3)_2]$, a chemotherapeutic but toxic compound.

$[\text{Ru}(\text{SC}(\text{NH}_2)_2)_6][\text{CF}_3\text{SO}_3]_2$ is the first characterised thiourea complex of ruthenium(II). Its structure was determined because the mode of attachment of thiourea to the ruthenium was unclear.

A pilot study of electron deformation density analysis is also reported. Two data sets were collected for $\text{C}_4\text{Cl}_4\text{O}_2\text{S}$, one at 295K the other at 150K using profile analysis of the Bragg reflections. Although the data did not lead to a satisfactory deformation density the analysis gave a clearer picture of the experimental and computational requirements for such work.

A brief summary of the experimental techniques and equations used is given in Appendix 1.

Appendix 2, bound in a separate volume, contains the calculated and observed structure factors of all the non-published structures reported in this thesis.

**SECTION I: STRUCTURAL STUDIES OF PLATINUM CLUSTERS
CONTAINING DIPHOSPHINOMETHANE LIGANDS**

CHAPTER 1: INTRODUCTION

Although phosphine ligands have been used to stabilise complexes of transition metals for many years it was only in the 1970s that the unusual properties of diphosphinomethanes, $R_2PCH_2PR_2$, began to be systematically exploited. During the last fifteen years $R_2PCH_2PR_2$ -stabilised complexes of many transition metals have been synthesized and several hundred structural studies, almost exclusively on binuclear species, have been reported.

A new departure in this field was the synthesis in 1983 of the trinuclear dicationic cluster $[Pd_3(CO)-(dppm)_3]^{2+}$, where dppm is $Ph_2PCH_2PPh_2$, the most commonly-used diphosphinomethane ligand, by Professor Puddephatt's group at the University of Western Ontario.¹ This synthesis and that of the platinum analogue has allowed Puddephatt's group to develop a new chemistry of tripalladium and triplatinum species. The nine structure analyses described below have been an integral part of this development. The strategic objectives of the structural work have been (i) characterisation of key new derivatives, (ii) provision of the experimental basis for the assessment of structure and bonding and (iii) clarification of the parallels between the behaviour of M_3 species ($M = Pd, Pt$) and those of the corresponding metal surfaces.

These topics are discussed and the new structures, together with the experimental details of their structure determinations, are described in Chapters 5-8. Before this, brief reviews, with the emphasis on structural features, are first given of binuclear complexes containing dppm and

related ligands (Chapter 2), of platinum cluster chemistry (Chapter 3) and of the earlier work on $M_3(dppm)_3$, $M = Pd$ and Pt , species (Chapter 4).

A number of papers on the new structures reported here have already been published.²⁻⁶

CHAPTER 2: DINUCLEAR BIS(DIPHENYLPHOSPHINO)METHANE COMPLEXES OF PALLADIUM AND PLATINUM

Tertiary phosphine ligands have been widely used to stabilise the lower oxidation states of transition metals.⁷ These ligands can be easily synthesized with a variety of substituents on phosphorus and are thus capable of displaying a wide range of electronic and steric properties.⁸ With diphosphines, $R_2P(CH_2)_nPR_2$, the length of the carbon chain connecting the phosphorus atoms is a further factor in determining ligand behaviour.

As noted above a substantial structural literature is now available on diphosphinomethane complexes, especially for dppm ($R = Ph$) but increasingly for other members of the family, notably for dmpm ($R = Me$), and four modes of ligation are known: (a) incorporation in a four-membered chelate ring, 2(i), as in $trans-[RhHCl(dppm)_2]^+$,⁹ (b) bridging a metal-metal bond, 2(ii), as in $[Pt_2Cl_2(\mu-dppm)_2]$,¹⁰ (c) bridging two metal atoms which are not directly bonded to one another, 2(iii), as in $[Rh(CO)_4-(CN)_2(\mu-dppm)_2]$ ¹¹ and (d) coordination through only one phosphorus atom, 2(iv), as in $trans-[Pd(Bu^tNC)_2(dppm)_2]^{2+}$ ¹² (see Figure 2(a)). Similar ligation modes are of course available to diphosphines with two or more methylene groups between the phosphorus atoms but such ligands tend to prefer the chelate arrangement 2(i) to the bridging structures 2(ii) and 2(iii) for steric reasons whereas the reverse is true of diphosphinomethanes.

In consequence, dppm and its relatives show a strong tendency to stabilise dinuclear complexes, as may be seen

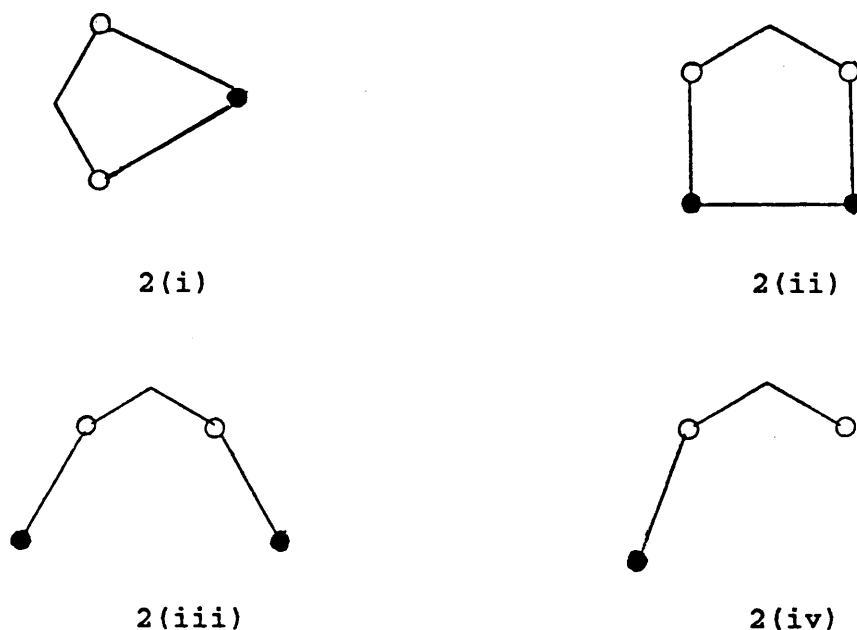


Figure 2(a). Modes of bonding of diphosphenomethane to a metal atom. M = ● P = O .

from a recent review by Puddephatt.¹³ Such complexes are particularly diverse for platinum and palladium and so far nine distinct structural arrangements of the nucleus M_2L_n , M = Pt or Pd, L = dppm or dmpm and n = 2 or 3 have been described (see 2(v)-2(xiii) in Figure 2(b)).

In 2(v) two metals with square-planar coordination are directly bonded to one another and are also linked by a pair of bridging diphosphine ligands, an arrangement typified by $[Pt_2Cl_2(dppm)_2]$.¹⁰ The cation $[Pt_2Me_2(\mu-H)(\mu-dppm)_2]^+$,¹⁴ type 2(vi), formally arises from protonation and consequent weakening of the Pt-Pt bond of 2(v). In 2(vii) $\mu-H$ is replaced by a larger bridging group such as CH_2^{2+} , μ -methylene, and the M-M distance is now much longer and clearly non-bonding, e.g. in $[Pt_2I_2(CH_2)(dppm)_2]$ ¹⁵ the

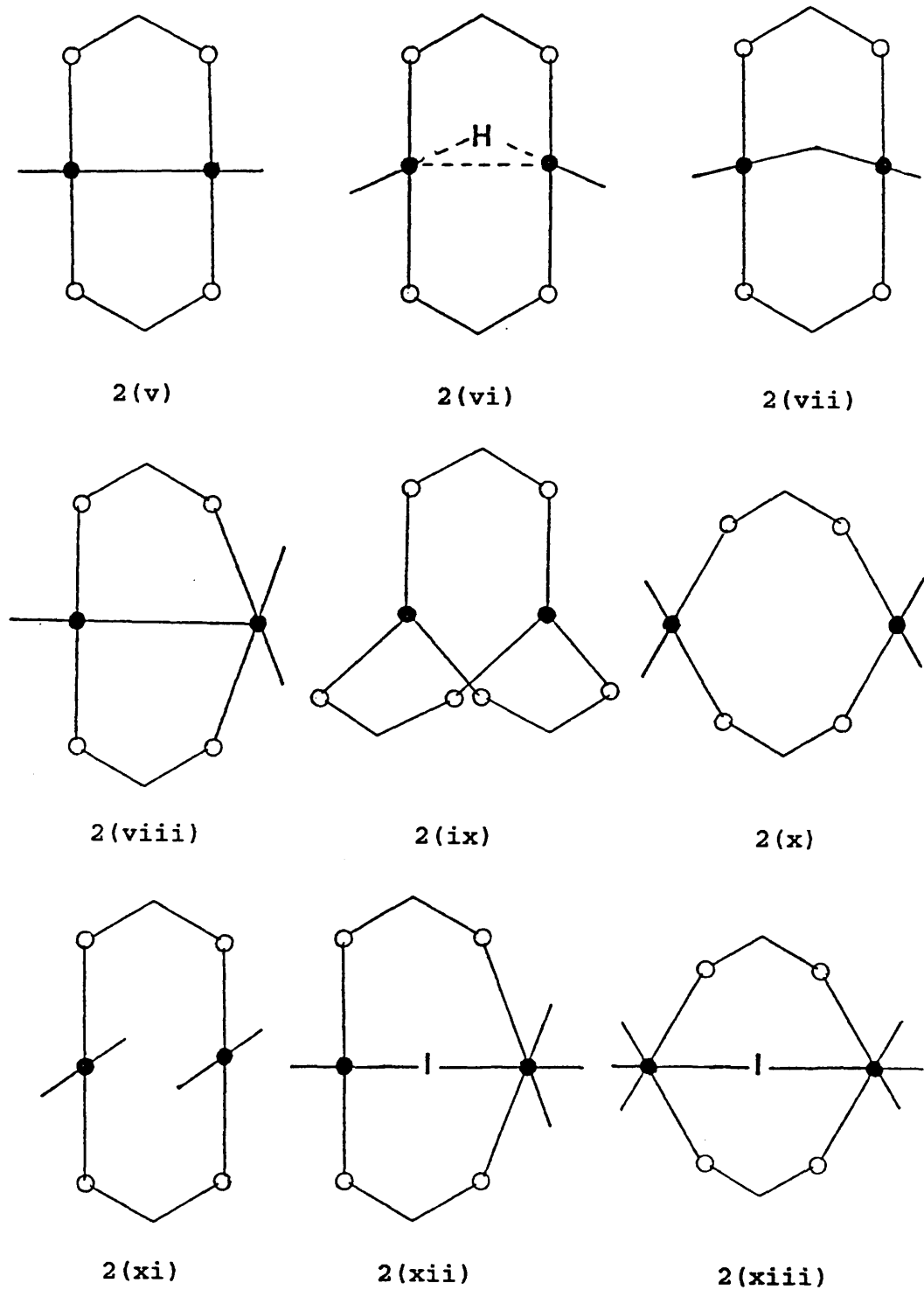


Figure 2(b). Structural arrangements of the nucleus M_2L_n , $M = \text{Pt or Pd}$, $L = \text{dppm or dmpm}$ and $n = 2 \text{ or } 3$.

Pt...Pt distance is 3.35 Å, compared with 2.77 Å and 2.65 Å for species of types 2(vi) and 2(v) respectively. Typical Pt-Pt bonding distances are in the range 2.6-2.8 Å.¹⁶ Both terminal ligands lie on the opposite side of the Pt₂P₄ plane from the bridging group in 2(vi) and 2(vii) which have therefore come to be known as A-frame complexes.

Bridging dppm ligands can also stabilise donor-acceptor metal-metal bonds: in [Pt₂Me₃(μ-dppm)₂]⁺¹⁷ which has skeleton 2(viii) the electron-rich PtMe₂ centre donates an electron-pair to the PtMe centre. 2(ix) is the only type of binuclear complex with three μ-dppm ligands and is exemplified by [Pd₂(μ-dppm)₃].¹⁸ Such compounds have a rich chemistry with addition possible at one or both metal centres.¹³

Binuclear complexes, [X₂M(μ-dppm)₂MX₂], in which a pair of μ-dppm ligands is the link between the metal atoms can have *cis* 2(x) or *trans* 2(xi) coordination at the metal atoms: [Pt₂Me₄(μ-dppm)₂] has the *cis* stereochemistry¹⁹ whereas in [Pt₂(C≡CR)₄(μ-dppm)₂]²⁰ there is a *trans* coordination of the metal atoms.

Reaction of *cis*-[Pt₂(dmpm)₂Me₄]²¹ with diiodine forms the mixed-valence Pt(II)-Pt(IV) complex [Pt₂I(dmpm)₂Me₄]⁺ of type 2(xii)²² while reaction with methyl iodide gives [Pt₂IME₆(dmpm)₂]⁺,²³ a Pt(IV)-Pt(IV) complex of type 2(xiii). Steric factors preclude similar reactions with the analogous dppm complexes.

Much of the interest in the M₂(dppm)₂ system arises from its stability both with and without direct M-M bonding since reactions involving M-M bond formation and scission

can occur without fragmentation into mononuclear fragments. Poilblanc ²⁴ has emphasized that such reactions are relevant to homogeneous catalysis by bimetallic systems and may also serve as simple models for catalysis at metal surfaces. The ability of the system to accommodate bridging groups is also significant. Indeed $[\text{Pd}_2\text{Cl}_2(\mu\text{-dppm})_2]$ catalyses the cyclotrimerisation of alkynes ²⁵ while $[\text{Pt}_2\text{H}_2(\mu\text{-H})(\mu\text{-dppm})_2]^+$ is a short-lived catalyst for the reduction of CO to CH_4 .¹³

CHAPTER 3: THE STRUCTURES OF PLATINUM CLUSTER COMPLEXES

STABILISED BY $\text{Ph}_2\text{PCH}_2\text{PPh}_2$

3.1 Introduction

As we have seen, although dinuclear transition metal complexes stabilised by dppm have been extensively studied, very little was known about corresponding trinuclear $\text{M}_3(\text{dppm})_3$ species prior to 1983.¹ Since then the chemistry and structural characterisation of such complexes, particularly with $\text{M} = \text{Pd}$ or Pt , has become an active area of research. Partly, this has been prompted by the consideration that the behaviour of these trinuclear complexes offers a model for that of metal surfaces. Additionally, there has been a continuing theoretical debate about the nature of the bonding in palladium and platinum cluster compounds. The structural results described below provide insights into both of these topics.

3.2 Some Comments on the Chemistry of Metal Surfaces

The chemisorption of small molecules, such as alkenes, alkynes and H_2S , on metal surfaces has been widely investigated because of its relevance to the catalytic properties of such surfaces. Techniques used to characterize the surface species include temperature-programmed desorption (TDS), low energy electron diffraction (LEED), and Auger electron spectroscopy (AES).²⁶⁻²⁸

The results for Pt(111) surfaces are particularly relevant to the structural studies discussed below. Many species are thought to bind in a μ_3 -fashion to the Pt_3 triangular units present on such surfaces. Acetylene, for example, is adsorbed to give $\text{Pt}_3(\mu_3-\eta^2\text{-HCCH})$ species at low temperatures. On heating $\text{Pt}_3(\mu_3-\eta^2\text{-C=CH}_2)$ units are found and these appear to react with surface-bound hydrogen so that $\text{Pt}_3(\mu_3\text{-CCH}_3)$ units are ultimately observed.²⁶ Likewise, hydrogen sulphide is adsorbed on the Pt(111) surface to form a layer of sulphur with emission of hydrogen gas.^{27,28} LEED analysis suggests that the surface layer consists of $\text{Pt}_3(\mu_3\text{-S})$ species.²⁸ The involvement of Pt_3 triangular units in the attachment of small molecules on the Pt(111) surface makes the study of triangulo- Pt_3 clusters of particular interest. An ability to mimic reactions happening on these surfaces could lead to a better understanding of the catalytic processes which occur.

3.3 Current Views on Bonding in Platinum Clusters

For most transition metal cluster complexes it has long been known that the geometry of the cluster is related to the number of skeletal electron pairs required to fill all the bonding metal-metal molecular orbitals.³⁰ This has become known as the polyhedral skeletal electron pair theory.³¹ This theory successfully explained the geometry of many metal clusters, particularly those containing conical $\text{M}(\text{CO})_3$ and $\text{M}(\eta^5\text{-C}_5\text{H}_5)$ fragments, because of their isolobal relationship³² to main group fragments such as BH and

CH₂.³³ In its initial form the theory was not usually applicable to platinum and palladium clusters, and this has triggered a theoretical discussion which still continues.

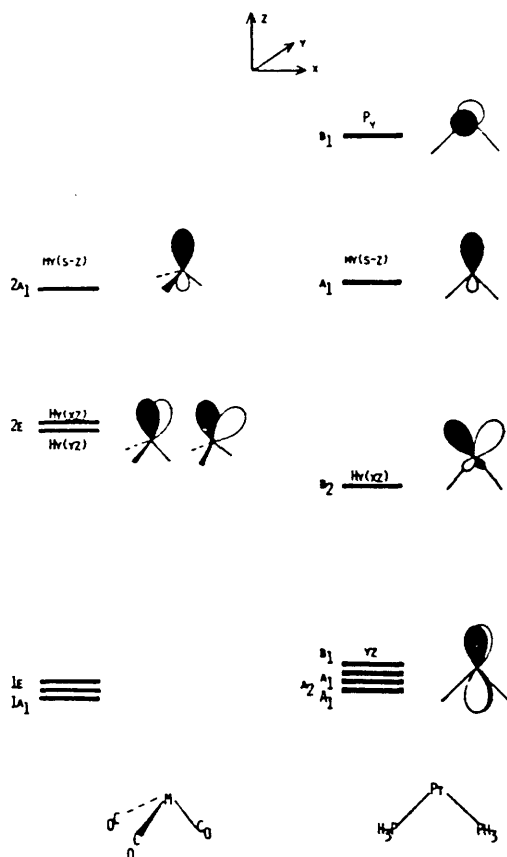


Figure 3(a). A comparison of the frontier molecular orbitals of Pt(PH₃)₂ and M(CO)₃.³⁴

In 1982 Evans and Mingos³⁴ extended the polyhedral skeletal electron pair theory to cover the non-conical ML₂ fragments which are found in many platinum and palladium clusters. They compared the frontier molecular orbitals of M(CO)₃ and Pt(PH₃)₂ fragments (see Figure 3(a)),³⁴ and concluded that the Hy(s-z), Hy(xz) and p_y orbitals of the Pt(PH₃)₂ fragment had the potential to form radial and tangential molecular orbitals equivalent to those formed by

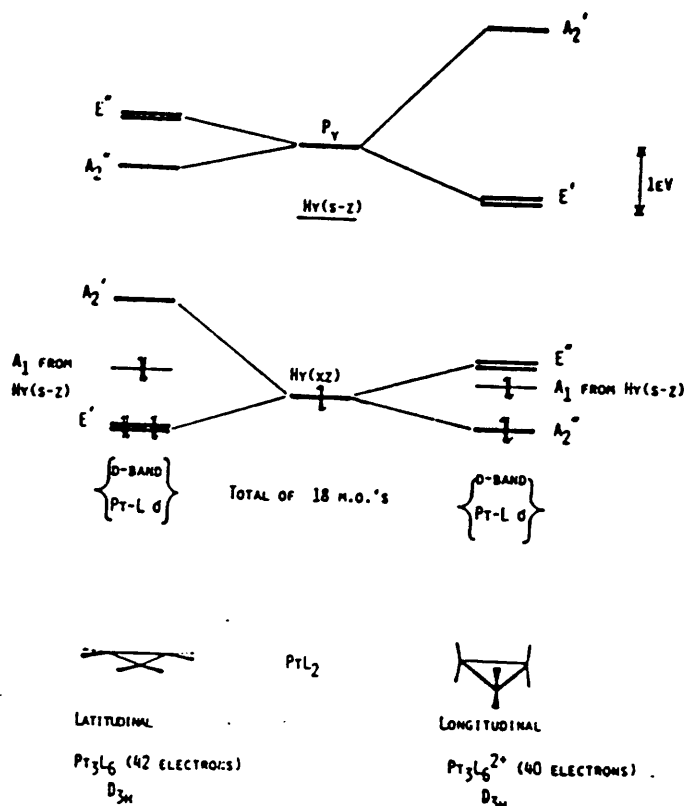


Figure 3(b). A comparison of the bonding molecular orbital diagrams for latitudinal and longitudinal geometries of $P_3(PH_3)_6$.³⁴

$M(CO)_3$ fragments in a similar polyhedral cluster. However, those bonding molecular orbitals derived from p_y orbitals were not sufficiently stable for electron occupation. This meant the bonding would be dominated by the molecular orbitals formed from the $Hy(s-z)$ and $Hy(xz)$ orbitals which depend heavily on the conformation of the PtL_2 fragments relative to the principle axis of the polyhedron. Figure 3(b)³⁴ shows that for a triangular $Pt_3(PH_3)_6$ species the latitudinal geometry needs 42 valence electrons and the

longitudinal geometry 40 valence electrons to fill the bonding molecular orbitals. Evans and Mingos showed that the more stable geometry is that with more filled bonding molecular orbitals derived from $H_y(xz)$, hence $Pt_3(PH_3)_6$ is expected to adopt a latitudinal geometry.

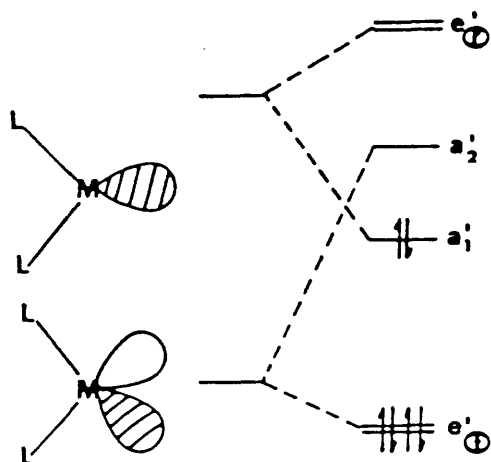


Figure 3(c). The molecular orbital diagram for a ML_2 fragment.³⁵

Mealli³⁵ used qualitative perturbation theory arguments to study the M_3L_6 species by comparing it to the isolobally equivalent cyclopropane.³² Each metal atom of the Pt_3 triangle contributes one σ and one π orbital, to produce two tangential and one radial bonding, and one tangential and two radial antibonding molecular orbital combinations. The resulting molecular orbital diagram (see Figure 3(c))³⁵ is similar to that found by Evans and Mingos.³⁴ The cluster is stable when the a_1' orbital is of lower energy than a_2' . For the L_6Pt_3 model with all ligands terminal (see Figure 3(d))³⁵ this ordering does not occur, unless there are additional ligands present to stabilize the cluster. This is because the σ orbitals from each L_2Pt

fragment are destabilized and hence the radial bonding orbital, a_1' , is higher in energy than the tangential antibonding combination, a_2' , of the π orbitals.

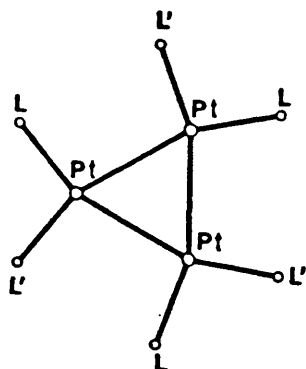


Figure 3(d). Latitudinal Pt_3L_6 with all ligands terminal.

When the $[Pt_3L_6]$ cluster is μ_3 -capped by CO^{2+} Mealli showed that the bonding molecular orbitals are stabilized by added sp-orbital character. This reduces direct metal-metal overlap slightly, but there is a much larger gain in stability due to metal-carbon bonding. A similar effect is noted due to the bridging hydrogen ligands in $[Pt_3(CO)_6(\mu-H)_3]$.³⁵

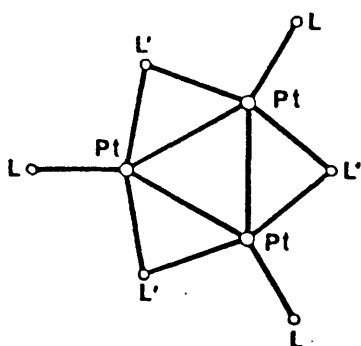


Figure 3(e). Latitudinal Pt_3L_6 with three bridging ligands.

Mealli also showed that an alternative geometry for the 42-electron Pt_3L_6 cluster, with three of the six ligands

acting as bridges between metals (see Figure 3(e))³⁵ is stable, without requiring the presence of additional ligands. This geometry has been observed experimentally in $[\text{Pt}_3(\mu\text{-CO})_3(\text{CO})_3]_n^{2-}$ for example.³⁶

3.4 Structural Types of M_3 Clusters ($\text{M} = \text{Pd}, \text{Pt}$)

In 1985 Mingos and Wardle published a comprehensive review of platinum cluster chemistry.²⁹ This, together with a survey using the Cambridge Structural Database to identify structurally characterised triangulo- M_3 ($\text{M} = \text{Pd}, \text{Pt}$) clusters reveals that the great majority of tripalladium and triplatinum complexes are 42-electron, planar species of the $[\text{M}_3\text{L}_3(\mu\text{-L}')_3]$ bridged type discussed by Mealli. In general, either the terminal L ligands or the bridging L' ligands, or in some cases both, are fairly bulky. Pt-Pt distances lie in the range 2.618-2.816 Å.^{37,38} The longer bonds tend to be associated with bulkier $\mu\text{-P}$ and $\mu\text{-S}$ bridging ligands, e.g. as in $[\text{Pt}_3(\text{PPh}_3)_3(\mu\text{-SO}_2)(\mu\text{-Ph})(\mu\text{-PPh}_2)]$.³⁸

Sometimes an additional ligand has added to the basic $[\text{M}_3\text{L}_3(\mu\text{-L}')_3]$ skeleton to give a 44-electron cluster. Examples include AuPR_3^+ , $\text{R}=\text{C}_6\text{H}_{11}$,³⁹ and Tl^+ ,⁴⁰ acting as μ_3 -capping groups to the M_3 triangle. Alternatively, an extra terminal ligand, e.g. $\text{P}(\text{C}_6\text{H}_{11})_3$ in $[\text{Pt}_3(\text{P}(\text{C}_6\text{H}_{11})_3)_4(\mu\text{-CO})_3]$,⁴¹ is found attached to a single metal atom. A few examples are also known of 46-electron species with two additional ligands, e.g. the bis-capped $[\text{Pt}_3(\text{Au}(\text{P}(\text{p}\text{-C}_6\text{H}_4\text{F})_3))(\mu\text{-Cl})(\mu\text{-SO}_2)_2(\text{P}(\text{C}_6\text{H}_{11})_3)_3]^+ 42$ or $[\text{Pt}_3(\text{Au}(\text{PPh}_3)-(\text{PPh}_3)_4(\mu\text{-CO})_3]^+ 43$ which contains additional μ_3 -capping

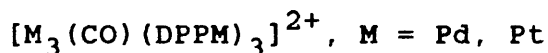
and terminal ligands. The addition of extra ligands to the basic 42-electron $M_3L_3(\mu-L')_3$ species results in electrons occupying non-bonding and anti-bonding orbitals. This generally is associated with a lengthening of Pt-Pt bonds e.g. compare $[Pt_3(P(C_6H_{11})_3(\mu-CO)_3)]^{44}$ (42e) with $[Pt_3(\mu_3-Au(P(C_6H_{11})_3))(P(C_6H_{11})_3)_3(\mu-CO)_3]^{39}$ (44e) and $[Pt_3-(P(C_6H_{11})_3)_4(\mu-CO)_3]^{41}$ (44e) where average Pt-Pt bond lengths are 2.654, 2.696 and 2.708 Å respectively.

In some cases complete Pt-Pt bond rupture has occurred, e.g. in the two isomers of the 44-electron species $[Pt_3(PPh_3)_2(Ph)(\mu-PPh_2)_3]^{45}$

Other structures not falling into the above categories include the 42-electron $M_3L_6(\mu-L')_3$ species $[Pt_3H_3(PBu_3^t)_3(\mu-H)_3]^{46}$ and the 44-electron $[Pt_3(cod)_3(\mu_3-SnCl_3)_2]^{47}$ (cod = cyclo-octadiene).

Attempts to develop the chemistry of clusters based on the $M_3L_3(\mu-L')_3$ skeleton have been greatly hampered by their tendency to fragment into mononuclear species. The $M_3(dppm)_3$ species discussed in subsequent chapters do not have this disadvantage.

CHAPTER 4: SALTS OF THE 42-ELECTRON DICATIONIC CLUSTER



The exploration of this new area of platinum metal cluster chemistry began in 1983 with the synthesis from palladium acetate, the diphosphine ligand and carbon monoxide of $[Pd_3(CO)(dppm)_3][CF_3CO_2]_2 \cdot 3Me_2CO$ (1) and its subsequent X-ray characterisation.¹ Since 1983 X-ray analyses have been completed on $[M_3(CO)(dppm)_3]^{2+}$, $M = Pd$, chloride (2) and iodide (3) salts^{48,49} and $M = Pt$, PF_6^- (4) and thiocyanate (5) salts.^{50,51}

All these complexes contain an equilateral triangle of metal atoms with each M-M bond bridged by a dppm ligand so that the M_3P_6 core is roughly planar, corresponding to the latitudinal geometry for such species, as defined by Mingos.³⁴ The single carbonyl ligand triply bridges one face of the M_3 triangle.

Mingos pointed out that for a 42-electron M_3P_6 cluster the expected geometry is latitudinal, since such a skeleton permits filling of all the bonding cluster orbitals. This view seems to imply that the dicationic cluster is chemically saturated. However, salts of the dication behave as 1:1 rather than 1:2 electrolytes and softer anions displace harder ones, with the anion affinity order $SCN^- \approx I^- > Br^- > Cl^- > CF_3CO_2^- > PF_6^-$. Furthermore, X-ray studies of (1)-(5) reveal that the anion is attached to the cluster in every case, except that of the PF_6^- salt (4), suggesting an electron count of 44 rather than 42.

Pertinent structural details in (1)-(5) are summarised in Table 4.1. In (4), where the PF_6^- anions do not interact

with the M_3 triangle, the Pt-CO distances are experimentally equal [2.08(1)-2.10(1) Å]. In (5) the SCN^- anion lies above the face of the M_3 triangle opposite to CO and is attached weakly to one platinum atom only [Pt-S 2.656(4) Å]. The Pt-CO bond lengths [2.04(1), 2.17(1) and 2.18(1) Å] indicate that the triply bridging CO is more strongly bound to the SCN^- -substituted metal atom. In contrast, in the Pd/I $^-$ complex (3), although the Pd-I distances show some variation [2.951(1)-3.083(1) Å], the iodo ligand is triply-bridging, as is the carbonyl ligand [Pd-C 2.14(1)-2.19(1) Å]. The Pd/Cl $^-$ complex (2) appears to have a similar bis- μ_3 -Cl/CO structure but Cl/CO disorder obscures the details. In (1) both oxygen atoms of a $CF_3CO_2^-$ anion are involved in weak Pd-O interactions [Pd-O 2.77(2)-3.24(2) Å] and the Pd-CO distances are irregular [2.08(1)-2.18(1) Å].

Individual M-M bond lengths (Table 4.1) do not appear to reflect the differences in anion coordination: for example, contrast the regularity of the Pt-Pt distances in the unsymmetrical SCN^- complex (5) with their variation in the PF_6^- complex (4). The only trend apparent in the mean M-M distances can be ascribed to the difference in the Pauling single bond metallic radii for Pd and Pt, respectively 1.283 and 1.295 Å.⁵² Likewise, the slightly greater mean M-P distances in the palladium complexes is expected and can be ascribed, according to Ibers,⁵³ to the greater hard acid character of palladium relative to platinum.

These results leave unanswered several important questions. Is the interaction of the anion with the

dication mainly ionic, i.e. a kind of ion pairing, or is there a significant covalent component? If the latter, how are the extra electrons accommodated without occupation of anti-bonding molecular orbitals and disruption of the cluster skeleton? How significant is the difference between the μ_3 -iodo and mono-hapto SCN^- interactions (and its effect on the μ -CO bonding)? These problems were the starting point for structural studies described in the following sections of this thesis.

Of chemical interest were the questions: how well do reactions involving the Pt_3 clusters mimic those on platinum surfaces and how are the differences, if any, explained?

Table 4.1. Comparison of geometries of $[M_3X(CO)(dppm)_3]^{n+}$ structures with different anions X^- . Distances are in Å, angles in degrees.

M	Pt ²⁺	Pt ²⁺	Pd ²⁺	Pd ²⁺	Pd ²⁺
X	-	SCN ⁻	CF ₃ CO ₂ ⁻	Cl ⁻	I ⁻
n	2	1	1	1	1
	42	44	44	44	44
Pt(1)-Pt(2)	2.638(1)	2.620(1)	2.576(1)	2.604(1)	2.599(1)
Pt(1)-Pt(3)	2.650(1)	2.625(1)	2.607(2)	2.583(1)	2.599(1)
Pt(2)-Pt(3)	2.613(1)	2.623(1)	2.610(2)	2.586(1)	2.591(1)
Pt(1)-P(1)	2.284(2)	2.313(4)	2.304(4)	2.300(2)	2.320(2)
Pt(1)-P(6)	2.304(2)	2.306(4)	2.303(3)	2.321(2)	2.316(2)
Pt(2)-P(2)	2.277(2)	2.294(4)	2.296(3)	2.314(2)	2.317(2)
Pt(2)-P(3)	2.271(2)	2.271(4)	2.324(3)	2.309(2)	2.309(2)
Pt(3)-P(4)	2.294(2)	2.280(4)	2.336(3)	2.303(2)	2.301(2)
Pt(3)-P(5)	2.262(2)	2.280(4)	2.340(3)	2.310(2)	2.325(2)
mean Pt-Pt	2.634(19)	2.623(3)	2.598(19)	2.591(11)	2.596(5)
mean Pt-P	2.282(15)	2.291(16)	2.317(19)	2.310(8)	2.315(9)
Pt(1)-C	2.095(9)	2.042(13)	2.175(11)	2.14 2.14	2.137(8)
Pt(2)-C	2.089(8)	2.165(13)	2.141(12)	2.25 2.21	2.192(8)
Pt(3)-C	2.080(9)	2.175(13)	2.080(10)	2.17 2.14	2.160(8)
Pt(1)-C-O	131.8(6)	143.8(11)	130.3(9)	-	143.4(6)
Pt(2)-C-O	133.1(7)	129.8(10)	133.2(10)	-	131.9(6)
Pt(3)-C-O	134.9(7)	128.9(10)	141.6(12)	-	132.3(6)
C=O	1.154(9)	1.165(14)	1.105(14)	-	1.114(9)
Pt(1)-I	-	2.656(4)	- 2.83(2)	2.74 2.94	2.951(1)
Pt(2)-I	-	3.380(4)	2.77(2) -	3.17 2.76	3.083(1)
Pt(3)-I	-	3.243(4)	2.92(2) 3.24(2)	2.79 2.83	3.031(1)
CH ₂ to CO ^a	uud	ddu	ddu	(ddu)	ddu
mean P-CH ₂	1.840(11)	1.842(18)	1.833(12)	1.839(4)	1.836(8)
mean P-C(Ph)	1.806(8)	1.820(6)	1.821(13)	1.829(9)	1.819(6)
MMP	91.0(1)- 97.3(1)	93.6(1)- 95.8(1)	92.9(1)- 97.7(1)	93.7- 95.8	93.5(1)- 96.3(1)
MPC	107.6(3)-111.6(3)	106.4(5)-110.9(5)	107.9(4)-111.6(4)	108.2-110.8	107.7(3)-110.1(3)
PCP	109.0(4)-112.8(4)	111.0(7)-112.0(7)	108.7(6)-113.2(6)	109.5-111.9	110.1(4)-111.7(4)
MPCP		52.2(6) -41.9(5)	-45.2(5) 49.7(5)	-48.2 45.5	-46.9(3) 47.1(3)
		-44.7(6) 46.2(6)	42.9(5) -47.9(5)	-52.1 42.1	50.3(4) -48.7(4)
		39.8(6) -51.8(6)	41.8(5) -48.4(5)	45.7 -48.7	-39.3(3) 49.5(4)
mean Pt-C	2.088(7)	2.127(74)	2.132(48)	2.18	2.163(28)

^aCl and CO are disordered.

^bSee Chapter 8.

CHAPTER 5: $[\text{Pt}_3(\mu_3\text{-CO})(\mu\text{-DPPM})_3\text{L}]^{2+}$ COMPLEXES

The 44-electron cations $[\text{Pt}_3(\mu_3\text{-CO})(\mu\text{-dppm})_3\text{L}]^{2+}$, L = phosphine or phosphite, have recently been synthesised by Professor Puddephatt by ligand addition to the parent cation $[\text{Pt}_3(\mu_3\text{-CO})(\mu\text{-dppm})_3]^{2+}$, (1).⁵⁰ This reaction can be regarded as a model for non-dissociative chemisorption of such species as PH_3 and PF_3 on platinum surfaces.⁵⁴

When L is a less sterically demanding ligand, e.g. phosphite or PMe_2Ph , the new complexes are thermally stable, whereas with more bulky phosphines, such as PMePh_2 or PPh_3 , dissociation occurs in solution at room temperature. Detailed N.M.R. studies suggest that at low temperatures the additional L ligand is terminally bonded but, on heating, a novel fluxional process, involving rapid migration of L around the Pt_3 triangle, gives a time-averaged $\mu_3\text{-PR}_3$ spectrum.

Accordingly, the structure of a typical member of the group of new 44-electron complexes, namely that with $\text{L} = \text{P(OPh)}_3$, (2), has been determined in order to establish the nature of the attachment of the P(OPh)_3 ligand to the cluster and how this affects the cluster bonding (see Tables 5.1-5.3). Also described briefly here is the structure of a related species with $\text{L} = \text{CNC}_6\text{H}_{11}$, (3), although in this case severe disorder obscures important structural details (see Tables 5.1, 5.4 and 5.5). The structures of these 44-electron dicationic species afford an interesting comparison with that of the anionic adducts $[\text{M}(\mu_3\text{-CO})(\text{X})(\mu\text{-dppm})_3]^+$, (4), ($\text{X} = \text{CF}_3\text{CO}_2^-$, Cl^- , SCN^- , I^- , and $\text{M} = \text{Pd}$ or Pt), where the weak cluster-anion bonding

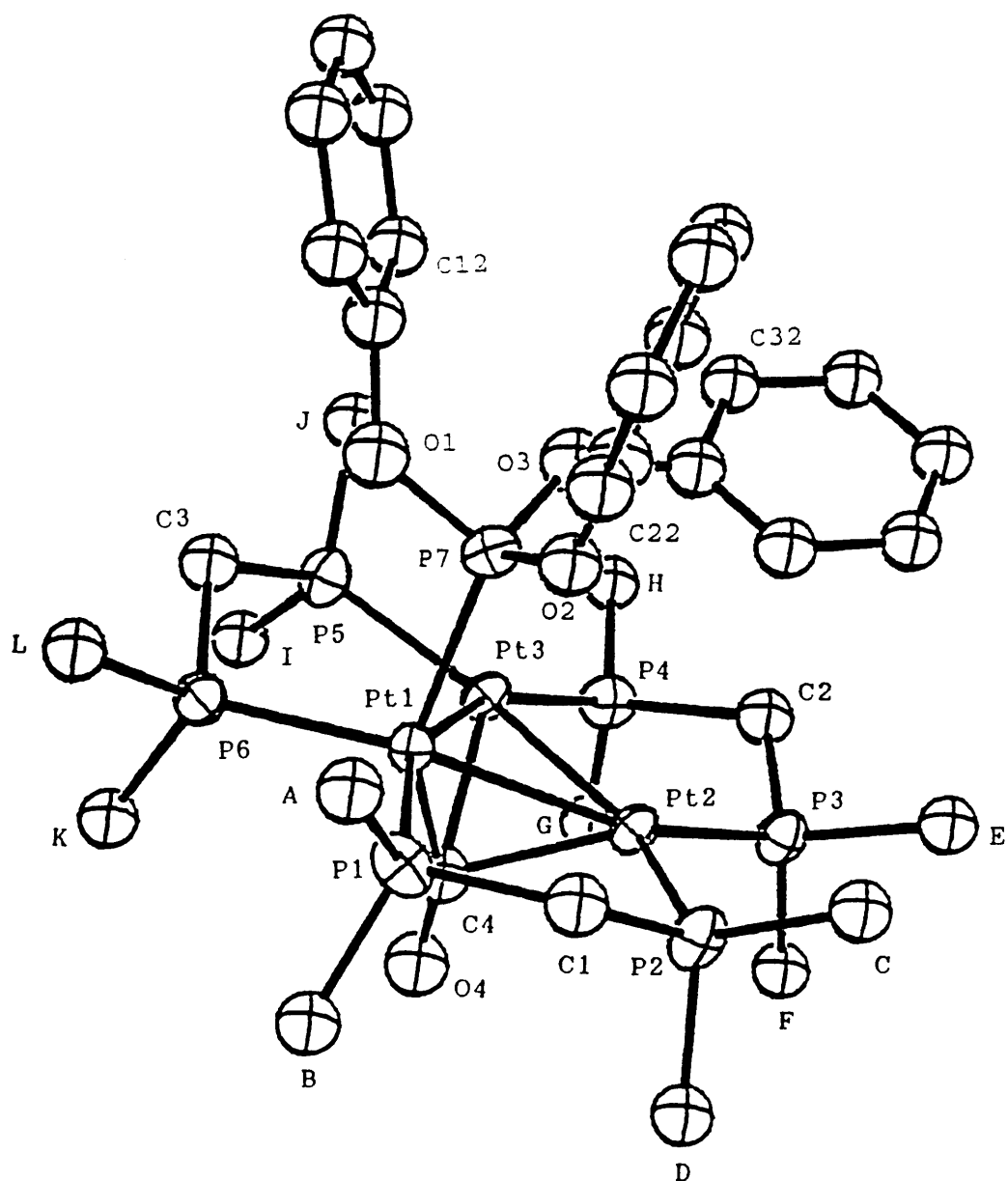
appears to be partially ionic (see Table 4.1).^{1,48,49,51}

The structure of (2) is shown in Figure 5(a). X-ray diffraction analysis of a crystal of the PF_6^- salt of (2), showed it to consist of two PF_6^- anions for every one cation (2). Two regions of diffuse electron density were also present in the asymmetric unit. These were tentatively ascribed to disordered solvent molecules (see Experimental). All non-bonded contacts between the various species present in the crystal were consistent with normal van der Waals radii.⁵²

The incoming P(OPh)_3 ligand is attached only to Pt(1), with the Pt(1)-P(7) bond vector at an angle of 85.4° to the Pt_3 plane. The Pt(1)-P(7) bond length of $2.293(8) \text{ \AA}$ is in the range considered typical for platinum-phosphite complexes.^{55,56} The phosphite conformation appears to be dictated by the need to reduce steric congestion with the $\text{Pt}_3(\mu\text{-dppm})_3$ nucleus. Thus, two of the phosphite phenyl groups point away from the cluster centre [Pt(1)-P(7)-C-O torsion angles respectively $-163.2(23)$ and $166.1(24)^\circ$ for C(11) and C(21)]. The third ring lies across the face of the Pt_3 triangle [Pt(1)-P(7)-O(3)-C(31) = $-76.2(21)^\circ$] and the shortest interligand contact involving this ring is the C(35)...C(E2) distance of 3.28 \AA . The conformation of the $\text{Pt}_3(\mu\text{-dppm})_3$ nucleus also shows evidence of steric congestion. The envelope-shaped $\text{Pt}_2\text{P}_2\text{C}$ rings all have their methylene flaps lying on the same side of the Pt_3 triangle as the P(OPh)_3 ligand (see Figure 5(a)). This arrangement is the more favourable one for attachment of a bulky terminal ligand normal to the Pt_3 face (see Chapter 8).

The mean Pt-Pt bond length in (2) is comparable to that

Figure 5(a). A view of the $[\text{Pt}_3(\mu_3\text{-CO})(\text{P}(\text{OPh})_3)(\text{dppm})_3]^{2+}$ cation. For clarity only the ipso carbons of phenyl rings A-L and no hydrogen atoms are displayed. Carbon atoms are shown as spheres of arbitrary size. Probability ellipsoids(50%) are shown for all other atoms.

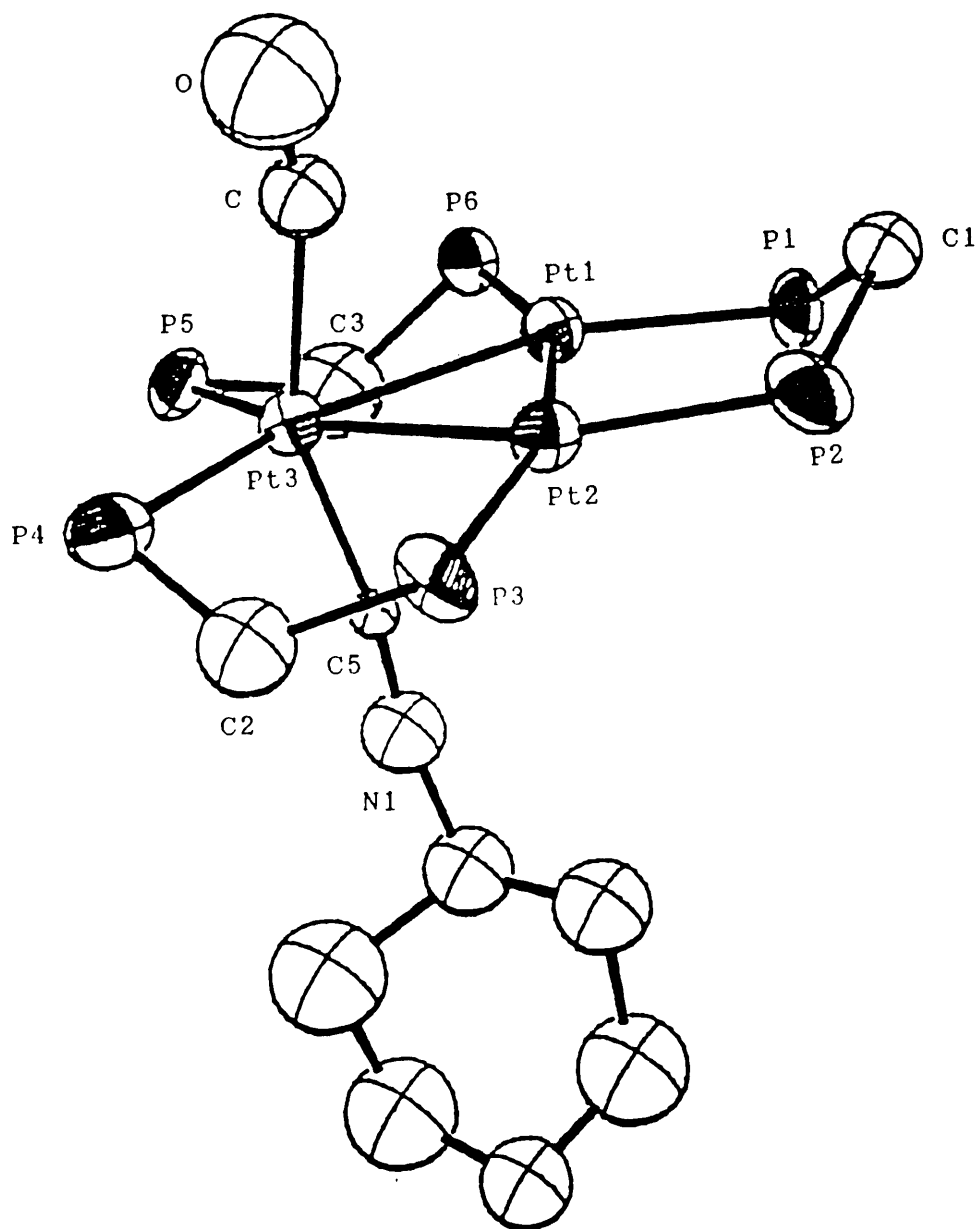


in (1) and in the 44-electron anion adducts of (1) and its palladium analogue (see Tables 4.1 and 5.1). The individual Pt-Pt distances show slight irregularities [2.626(2)-2.656(2) Å] but these do not seem to be attributable to the addition of an extra ligand to Pt(1). However, the greater coordination number of Pt(1) may explain the slight lengthening of the Pt(1)-P(dppm) distances [2.341(8) and 2.335(8) Å] compared with those involving Pt(2) and Pt(3) of 2.288(8)-2.323(8) Å.

The most dramatic effect of the addition of the P(OPh)₃ ligand to the Pt₃(μ₃-CO)(μ-dppm)₃ nucleus is shown by the carbonyl bonding. In (1) the CO carbon atom is equidistant from the three metal atoms (see Table 4.1), whereas in (2) the Pt(1)-CO bond length of 1.93(3) Å is 0.23(4) and 0.36(4) Å shorter than the corresponding distances involving Pt(2) and Pt(3) (see Table 5.1). Likewise, the Pt(1)-C-O angle of 155(2)° is some 30° closer to linearity than the corresponding Pt(2) and Pt(3) angles. A similar, but less marked, asymmetry in the μ₃-CO-Pt₃ bonding is found in the thiocyanate adduct (see Table 4.1). The bonding implications of this result are considered below.

The finer details of the structure of the cluster [Pt₃(CO)(CNC₆H₁₁)(μ-dppm)₃]²⁺ (3) are obscured by problems with disorder of occluded solvent and one PF₆⁻ anion but also by CO/CNC₆H₁₁ disorder. The model used in the final refinement included the Pt₃dppm₃(CNC₆H₁₁)²⁺ fragment of the cation as well as one PF₆⁻ anion and the phosphorus atom of the other PF₆⁻ anion. In the final difference synthesis the two highest peaks had positions consistent with those of a terminal carbonyl group attached to the same platinum atom

Figure 5(b). A view of the proposed structure of complex (3). For clarity all phenyl rings have been removed.



as the isonitrile substituent (see Figure 5(b)). However, these positions are also consistent with those of the CN part of another isonitrile group and there were also peaks, though much less pronounced, consistent with five of the six carbon atoms of an attached cyclohexane ring. This is clearly due to CO/CNC₆H₁₁ disorder with the model shown in Figure 5(b) likely to be the major isomer. However, no model could be found which refined satisfactorily.

Although the mean Pt-Pt distance is comparable with those for related 44-electron triangulo-Pt₃(dppm)₃ systems (Tables 4.1 and 5.1) the bond between the two metal atoms not attached to the C-donor ligands [2.584(4) Å] is 0.04-0.06 Å shorter than the remaining two. There appears to be less steric congestion in (3) than in (2): the conformation of the Pt₃(μ-dppm)₃ is such that only two dppm methylene carbon atoms lie on the nitrile side of the Pt₃ plane.

A recent paper by Evans provides an elegant rationalisation of these results in terms of extended Huckel (EHMO) theory.⁵⁷ This predicts that the addition of a second terminal CO to a Pt₃(μ₃-CO)(μ-dppm)₃²⁺ cluster will be stabilised by the π-acidic properties of the μ-CO ligand (Figure 5(c)). The Mulliken population analysis predicts little change in Pt-Pt bond lengths but a strengthening of the (μ₃-CO)-Pt bonding to the metal atom bearing the terminal carbonyl at the expense of the other (μ₃-CO)-Pt bonds. From Tables 4.1 and 5.1 it is apparent that the parent 42-electron cation and the adducts with L = I⁻, CF₃CO₂⁻, SCN⁻, P(OPh)₃ and CNC₆H₁₁ show a gradual transition from symmetrical μ₃-CO towards a terminal CO with little change in Pt-Pt bond lengths.

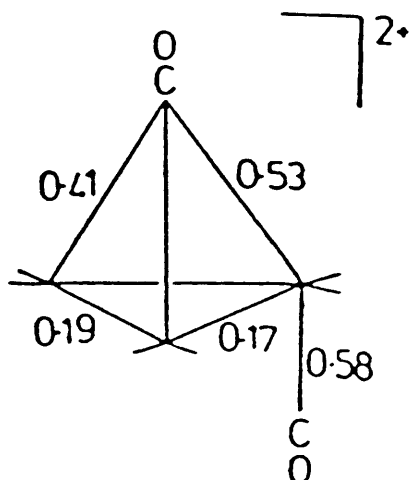


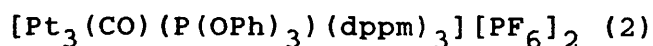
Figure 5(c). Mulliken population analysis of addition of terminal CO to the $\text{Pt}_3(\mu_3\text{-CO})(\text{dppm})_3^{2+}$ cluster.⁵⁷

Evans also compared the stability of the cluster with CO added terminally with that of a $\text{Pt}_3(\mu_3\text{-CO})_2(\mu\text{-dppm})_3^{2+}$ 44-electron cluster. The greater stability due to extra Pt-C bonding in the latter is more than offset by weaker metal-metal bonding but the difference in stability is not great. Infra-red and N.M.R. studies confirm that there is a rapid equilibrium between the two types of CO bonding in solution.⁵⁸ Although $\text{Pt}_3(\mu_3\text{-X})_2(\mu\text{-dppm})_3$ 44-electron clusters with μ_3 -bridging ligands containing first or second row donor atoms have not been obtained so far the structure of $[\text{Pt}_3(\mu_3\text{-CO})(\mu_3\text{-SnF}_3)(\mu\text{-dppm})_3]^+$, for example, shows that such systems can be stabilised with bulkier bridging donor atoms (see Chapter 7).

Experimental

Broadly similar methods were used for all the analyses described in this Section. Full details of a typical

analysis are given below for complex (2) and details pertinent to each of the remaining analyses are given in their respective chapters. Table 5.6 summarises crystal data for complexes (2) and (3).

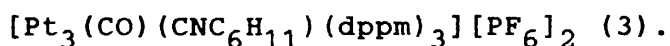


Data Collection. All measurements were made using Mo-K α radiation on an Enraf Nonius CAD4F diffractometer equipped with a graphite monochromator. The crystal was a red plate, size 0.36x0.24x0.16 mm, mounted in a capillary. Cell dimensions were determined by a least-squares refinement of the setting angles of 23 automatically centred reflections with $8 < \theta < 12^\circ$ (see Table 5.6). The intensities of 12262 reflections with $2 < \theta < 21^\circ$ were measured by continuous $\omega/2\theta$ scans of 0.80° in ω , increased by 25% at each end to allow for background. The intensity standards, measured every two hours, showed no significant change. After correction for Lorentz, polarisation and absorption effects (empirical transmission factors on F 0.81-1.17),⁵⁹ and merging of equivalent reflections, 9840 unique reflections were obtained; of these 3967 had $I \geq 3\sigma(I)$ and were subsequently used.

Structure Analysis. The heavy atom positions were found from a Patterson synthesis and all other non-hydrogen atoms from subsequent difference syntheses. Two carbon atoms and one oxygen atom were introduced into the calculations to model two regions of diffuse electron density ascribed to disordered solvent. These atoms were included in the final refinement but are not allowed for in Table 5.6 because it

was not possible to establish the identity of the disordered solvent. Anisotropic displacement parameters were used for Pt and P atoms and isotropic displacement parameters were used for all the remaining atoms. A fixed contribution for the scattering of the hydrogen atoms was added to the structure factors assuming a C-H distance of 0.96 \AA and a fixed isotropic $U(H) = 0.05 \text{ \AA}^2$. The phenyl rings were idealised, with C-C distances 1.380 \AA and C-C-C angles 120° , and refined as groups. Similarly, P-F bond distances were fixed at 1.53 \AA with F-P-F angles either 90° or 180° and each PF_6^- anion was refined as a group. Final parameters were obtained from full-matrix least-squares minimisation of $\sum w(|F_o| - |F_c|)^2$ with $w^{-1} = \sigma^2 + 0.00023F^2$ (where σ is derived from counting statistics), by adjustment of 351 parameters which converged with $R = 0.0475$ and $R_w = 0.0545$.

All calculations were performed using the GX package,⁶⁰ on a GOULD-SEL 3227 computer. Neutral atom scattering factors and complex anomalous dispersion corrections were taken from ref. 61.



Data Collection. Crystal mounted on a glass fibre.

Structure Analysis. The model used in the final refinement was $\text{C}_{82}\text{F}_6\text{NP}_8\text{Pt}_3$. Anisotropic displacement parameters were used for all Pt and P atoms with the exception of P(8). It and all other atoms used isotropic displacement parameters. Group refinement was used for all phenyl rings and for the P(7)F_6^- anion with bond distances and angles idealised as

above.

Table 5.1. Comparison of geometries of $[M_3(X)(CO)(dppm)_3]^{n+}$ structures (distances in Å, angles in °) where X may be attached to one, two, or three M atoms.

H	Pt	Pt *	Pt *	Pt *	Pt	Pt	Pt
I	P(OPh) ₃	ClC ₆ H ₄	CO/C1/SnF ₃ ⁻	dopa	S ₂ CNMe ₂	(SnF ₃ ⁻) ₂	SnMe ₂ (PO ₂ F ₂) ⁻
n	2	2	1/0	2	1	0	1
	44	44	44	46	46	44	42
Pt(1)-Pt(2)	2.656(2)	2.584(4)	2.634(3)	2.628(1)	2.612(1)	2.639(1)	2.609(1)
Pt(1)-Pt(3)	2.626(2)	2.626(3)	-	2.648(1)	2.622(1)	2.609(2)	2.615(1)
Pt(2)-Pt(3)	2.636(2)	2.646(3)	-	2.620(1)	2.605(1)	2.622(1)	2.635(1)
Pt(1)-P(1)	2.341(8)	2.205(13)	2.268(15)	2.289(3)	2.298(3)	2.301(5)	2.272(4)
Pt(1)-P(6)	2.335(8)	2.251(15)	2.315(15)	2.342(3)	2.356(3)	2.303(5)	2.297(4)
Pt(2)-P(2)	2.323(8)	2.227(16)	-	2.246(3)	2.246(3)	2.304(6)	2.284(4)
Pt(2)-P(3)	2.290(8)	2.373(14)	-	2.256(3)	2.244(3)	2.313(5)	2.264(4)
Pt(3)-P(4)	2.288(8)	2.460(15)	-	2.291(3)	2.309(3)	2.278(5)	2.288(4)
Pt(3)-P(5)	2.300(8)	2.406(15)	-	2.319(3)	2.357(3)	2.308(5)	2.284(4)
mean Pt-Pt	2.639(15)	2.619(35)	2.634(3)	2.632(14)	2.613(9)	2.623(15)	2.620(14)
mean Pt-P	2.313(23)	2.32(14)	2.292(24)	2.291(31)	2.300(50)	2.301(12)	2.282(12)
Pt(1)-C	1.93(3)	-	2.142	2.049(8)	2.011(11)	-	-
Pt(2)-C	2.16(3)	-	-	-	2.671(11)	-	-
Pt(3)-C	2.27(3)	-	-	2.061(9)	2.023(11)	-	-
Pt(1)-C-O	154.6(21)	-	134.8	140.3(7)	137.1(8)	-	-
Pt(2)-C-O	121.1(19)	-	-	-	-	-	-
Pt(3)-C-O	119.9(20)	-	-	138.6(7)	140.9(8)	-	-
C=O	1.21(4)	-	1.129	1.194(10)	1.217(14)	-	-
Pt(1)-I	2.293(8)	1.88(5)	2.773(2)	2.355(3)	2.480(4)	2.746(2)	2.808(2)
Pt(2)-I	-	-	-	-	3.58 3.54	2.801(2)	2.699(2)
Pt(3)-I	-	-	-	2.372(3)	2.503(3)	2.783(2)	2.830(2)
CH ₂ to CO	ddd	ddu	ddd	ddd	udd	uud	uuu (to Sn)
mean P-CH ₂	1.832(20)	1.84(16)	1.844(57)	1.830(4)	1.840(3)	1.836(30)	1.848(16)
mean P-C(Ph)	1.832(5)	1.82(6)	1.819(57)	1.805(5)	1.832(3)	1.832(24)	1.826(6)
MP	93.2(2)- 98.2(2)	93.0(4)- 97.7(4)	95.8(4)	91.2(1)- 97.8(1)	89.3(1)- 98.0(1)	93.3(2)- 97.6(2)	89.5(1)- 97.6(1)
MP	106.1(9)-112.2(9)	101.6(18)-110.8(16)	106.4(17)-106.5(15)	107.9(5)-111.7(3)	105.7(4)-110.5(4)	107.6(7)-110.9(6)	102.4(8)-112.7(5)
PCP	111.6(13)-121.5(14)	106.9(27)-124.8(34)	114.2(25)	97.4(10)-109.2(9)	109.1(6)-114.7(6)	112.9(10)-113.4(10)	107.5(8)-110.5(7)
MP	-34.7(11) 28.7(12)	51.9(19) -43.8(20)	49.5(22) -49.4(21)	-	34.4(5) -52.0(5)	39.6(8) -44.8(8)	-40.2(6) 48.1(6)
	-49.2(11) 50.6(11)	-38.0(30) 41.9(36)	-	-	-37.6(4) 57.2(5)	46.6(8) -48.2(9)	-34.7(6) 52.3(7)
	-28.0(11) 46.2(12)	-51.1(24) 57.7(25)	-	-	-47.1(5) 50.8(5)	-44.5(8) 46.0(8)	-39.1(6) 52.0(7)
mean Pt-C	2.12(17)	-	2.142	2.055(9)	2.017(11)	-	-

*Disordered - see page 25

*Disordered - see page 74

Table 5.2. Fractional atomic coordinates and displacement parameters (\AA^2) of the crystal structure of compound (2).

	X/A	Y/B	Z/C	u ^a
PT(1)	-.24987(7)	.12158(6)	.28551(4)	.027
PT(2)	-.25422(7)	.13492(6)	.18332(4)	.029
PT(3)	-.16941(7)	.04082(6)	.24780(4)	.029
P(1)	-.3194(4)	.2190(4)	.2892(3)	.036
P(2)	-.3386(4)	.2251(4)	.1622(3)	.037
P(3)	-.2249(4)	.1021(4)	.1091(3)	.035
P(4)	-.1223(4)	-.0011(4)	.1848(3)	.037
P(5)	-.1260(4)	-.0255(4)	.3253(3)	.037
P(6)	-.1820(4)	.0953(4)	.3764(3)	.034
P(7)	-.3424(4)	.0409(4)	.2643(3)	.036
P(F1)	.2041(7)	.3890(7)	.5347(5)	.174
P(F2)	-.7078(7)	.1426(8)	.3778(8)	.324
O(1)	-.3440(10)	.0058(10)	.3175(7)	.046(6)
O(2)	-.4200(10)	.0755(9)	.2366(7)	.039(5)
O(3)	-.3423(10)	-.0278(9)	.2312(7)	.047(5)
O(4)	-.1176(11)	.1904(10)	.2663(7)	.049(6)
F(11)	.248(2)	.325(2)	.561(2)	.34(3)
F(12)	.161(2)	.453(2)	.508(2)	.31(2)
F(13)	.268(3)	.434(3)	.570(1)	.38(3)
F(14)	.140(3)	.344(3)	.500(1)	.44(4)
F(15)	.169(2)	.392(1)	.579(1)	.36(3)
F(16)	.239(2)	.386(1)	.490(1)	.26(2)
F(21)	-.689(3)	.103(1)	.334(2)	.54(6)
F(22)	-.727(3)	.182(1)	.422(2)	.79(9)
F(23)	-.694(2)	.211(1)	.354(2)	.43(3)
F(24)	-.721(2)	.075(1)	.402(2)	.27(2)
F(25)	-.626(2)	.140(4)	.415(2)	.36(3)
F(26)	-.789(2)	.145(4)	.341(2)	.39(3)
C(1)	-.3746(13)	.2393(12)	.2194(9)	.025(7)
C(2)	-.1950(14)	.0157(13)	.1159(10)	.031(7)
C(3)	-.1607(15)	.0038(14)	.3786(10)	.039(8)
C(4)	-.1715(16)	.1558(14)	.2621(10)	.038(8)
C(11)	-.3770(16)	-.0553(12)	.3255(8)	.039(8)
C(12)	-.3498(13)	-.1185(16)	.3177(12)	.061(10)
C(13)	-.3789(21)	-.1778(12)	.3311(15)	.076(11)
C(14)	-.4353(16)	-.1739(12)	.3523(8)	.113(15)
C(15)	-.4625(13)	-.1107(16)	.3601(12)	.152(19)
C(16)	-.4334(21)	-.0514(12)	.3467(15)	.093(13)
C(21)	-.4933(14)	.0486(22)	.2264(10)	.052(9)
C(22)	-.5404(24)	.0941(15)	.2384(10)	.107(14)
C(23)	-.6118(21)	.0743(16)	.2322(16)	.112(15)
C(24)	-.6360(14)	.0091(22)	.2138(10)	.131(17)
C(25)	-.5889(24)	-.0364(15)	.2018(10)	.104(14)
C(26)	-.5175(21)	-.0166(16)	.2080(16)	.060(10)
C(31)	-.3567(16)	-.0389(15)	.1769(7)	.030(7)
C(32)	-.3453(9)	-.1045(15)	.1613(11)	.053(9)
C(33)	-.3658(18)	-.1207(7)	.1067(13)	.076(11)
C(34)	-.3977(16)	-.0714(15)	.0677(7)	.074(11)
C(35)	-.4091(9)	-.0058(15)	.0833(11)	.061(10)
C(36)	-.3886(18)	.0104(7)	.1379(13)	.056(10)
C(A1)	-.3922(11)	.2186(11)	.3175(10)	.031(7)
C(A2)	-.3944(12)	.1659(13)	.3522(13)	.041(8)
C(A3)	-.4510(9)	.1631(9)	.3733(7)	.056(9)
C(A4)	-.5052(11)	.2131(11)	.3597(10)	.061(10)
C(A5)	-.5030(12)	.2658(13)	.3249(13)	.051(9)
C(A6)	-.4464(9)	.2685(9)	.3038(7)	.056(9)
C(B1)	-.2659(16)	.2955(13)	.3197(7)	.030(7)
C(B2)	-.2747(8)	.3256(10)	.3649(8)	.041(8)
C(B3)	-.2328(17)	.3822(9)	.3887(10)	.048(8)
C(B4)	-.1820(16)	.4087(13)	.3674(7)	.056(9)
C(B5)	-.1732(8)	.3786(10)	.3223(8)	.070(10)
C(B6)	-.2151(17)	.3219(9)	.2984(10)	.047(9)
C(C1)	-.4271(15)	.2249(20)	.1043(10)	.046(8)
C(C2)	-.4443(14)	.2740(14)	.0637(5)	.067(10)
C(C3)	-.5134(10)	.2743(13)	.0234(10)	.067(10)
C(C4)	-.5653(15)	.2255(20)	.0237(10)	.067(10)
C(C5)	-.5481(14)	.1764(14)	.0643(5)	.078(11)
C(C6)	-.4790(10)	.1761(13)	.1046(10)	.056(9)
C(D1)	-.2941(18)	.3070(14)	.1591(7)	.041(8)
C(D2)	-.2243(20)	.3066(12)	.1554(14)	.047(9)
C(D3)	-.1865(11)	.3675(18)	.1580(11)	.078(11)
C(D4)	-.2185(18)	.4289(14)	.1644(7)	.077(12)
C(D5)	-.2883(20)	.4293(12)	.1682(14)	.108(15)
C(D6)	-.3261(11)	.3683(18)	.1655(11)	.077(11)

Table 5.2(cont).

	X/A	Y/B	Z/C	U
C(E1)	-.2987(13)	.1017(17)	.0411(8)	.031(7)
C(E2)	-.3148(12)	.0454(12)	.0069(12)	.042(8)
C(E3)	-.3703(18)	.0492(11)	-.0435(10)	.063(10)
C(E4)	-.4097(13)	.1094(17)	-.0597(8)	.050(9)
C(E5)	-.3936(12)	.1657(12)	-.0255(12)	.058(10)
C(E6)	-.3381(18)	.1619(11)	.0249(10)	.050(9)
C(F1)	-.1509(11)	.1515(13)	.0957(8)	.041(8)
C(F2)	-.1438(20)	.1465(17)	.0451(6)	.065(10)
C(F3)	-.0873(17)	.1809(11)	.0348(9)	.068(10)
C(F4)	-.0379(11)	.2203(13)	.0751(8)	.081(12)
C(F5)	-.0451(20)	.2252(17)	.1257(6)	.075(11)
C(F6)	-.1016(17)	.1908(11)	.1360(9)	.065(10)
C(G1)	-.0384(11)	.0404(15)	.1821(10)	.038(7)
C(G2)	-.0169(20)	.0250(17)	.1382(6)	.050(9)
C(G3)	.0507(17)	.0474(8)	.1374(11)	.073(10)
C(G4)	.0969(11)	.0853(15)	.1805(10)	.087(12)
C(G5)	.0754(20)	.1007(17)	.2243(6)	.095(13)
C(G6)	.0077(17)	.0783(8)	.2251(11)	.069(10)
C(H1)	-.0981(16)	-.0912(10)	.1843(14)	.033(7)
C(H2)	-.1398(12)	-.1381(19)	.1465(11)	.060(9)
C(H3)	-.1177(19)	-.2059(16)	.1495(7)	.101(14)
C(H4)	-.0538(16)	-.2268(10)	.1903(14)	.055(9)
C(H5)	-.0121(12)	-.1798(19)	.2281(11)	.066(10)
C(H6)	-.0342(19)	-.1121(16)	.2251(7)	.056(9)
C(I1)	-.0251(14)	-.0209(20)	.3585(9)	.038(8)
C(I2)	.0154(15)	-.0722(11)	.3923(7)	.063(10)
C(I3)	.0895(11)	-.0621(16)	.4219(12)	.061(10)
C(I4)	.1231(14)	-.0008(20)	.4178(9)	.093(13)
C(I5)	.0827(15)	.0505(11)	.3840(7)	.073(11)
C(I6)	.0086(11)	.0404(16)	.3544(12)	.053(9)
C(J1)	-.1525(17)	-.1158(11)	.3209(11)	.028(7)
C(J2)	-.1478(20)	-.1526(16)	.3672(9)	.090(13)
C(J3)	-.1701(10)	-.2204(13)	.3628(8)	.093(13)
C(J4)	-.1971(17)	-.2514(11)	.3122(11)	.095(13)
C(J5)	-.2017(20)	-.2146(16)	.2660(9)	.084(12)
C(J6)	-.1794(10)	-.1468(13)	.2704(8)	.065(10)
C(K1)	-.0943(16)	.1410(18)	.4084(6)	.049(8)
C(K2)	-.0857(11)	.2064(19)	.3911(11)	.047(8)
C(K3)	-.0217(14)	.2428(8)	.4173(10)	.073(11)
C(K4)	.0337(16)	.2139(18)	.4607(6)	.084(12)
C(K5)	.0251(11)	.1484(19)	.4780(11)	.066(10)
C(K6)	-.0389(14)	.1120(8)	.4518(10)	.074(11)
C(L1)	-.2229(14)	.1061(7)	.4304(10)	.037(8)
C(L2)	-.2603(18)	.0533(10)	.4447(11)	.045(8)
C(L3)	-.2926(11)	.0638(10)	.4837(6)	.055(9)
C(L4)	-.2874(14)	.1272(7)	.5084(10)	.072(10)
C(L5)	-.2499(18)	.1801(10)	.4941(11)	.087(12)
C(L6)	-.2177(11)	.1696(10)	.4551(6)	.063(10)
O(S)	.2324(20)	.1313(20)	.5260(14)	.203(16)
C(S1)	-.518(4)	-.037(4)	-.493(4)	.24(4)
C(S2)	-.477(4)	.001(5)	-.458(3)	.23(3)

Table 5.2(cont).

	X/A	Y/B	Z/C	U
H(1A)	-.39163	.28714	.21768	.050
H(1B)	-.42214	.21197	.21015	.050
H(2A)	-.17596	.00452	.08481	.050
H(2B)	-.23844	-.01572	.10745	.050
H(3A)	-.12431	-.00686	.41443	.050
H(3B)	-.20540	-.02223	.37611	.050
H(12)	-.31054	-.12076	.30308	.050
H(13)	-.36059	-.22193	.32423	.050
H(14)	-.45720	-.21629	.35915	.050
H(15)	-.50371	-.10951	.37298	.050
H(16)	-.45371	-.00834	.35184	.050
H(22)	-.52265	.13708	.25447	.050
H(23)	-.64294	.10174	.24441	.050
H(24)	-.68383	-.00789	.21044	.050
H(25)	-.60446	-.08215	.18647	.050
H(26)	-.48419	-.04681	.19651	.050
H(32)	-.32173	-.13943	.18800	.050
H(33)	-.35903	-.16766	.09536	.050
H(34)	-.41363	-.08418	.02919	.050
H(35)	-.43085	.02755	.05563	.050
H(36)	-.39360	.05578	.14830	.050
H(A2)	-.35708	.13215	.36190	.050
H(A3)	-.45366	.12797	.39695	.050
H(A4)	-.54562	.21274	.37295	.050
H(A5)	-.54100	.30166	.31398	.050
H(A6)	-.44442	.30587	.27896	.050
H(B2)	-.30930	.30610	.38061	.050
H(B3)	-.23854	.40221	.42122	.050
H(B4)	-.15325	.44804	.38483	.050
H(B5)	-.13871	.39782	.30781	.050
H(B6)	-.20946	.30169	.26720	.050
H(C2)	-.40906	.30620	.06247	.050
H(C3)	-.52633	.30769	-.00565	.050
H(C4)	-.61512	.22585	-.00500	.050
H(C5)	-.58666	.14249	.06370	.050
H(C6)	-.46939	.14100	.13182	.050
H(D2)	-.20113	.26379	.15207	.050
H(D3)	-.13776	.36744	.15615	.050
H(D4)	-.19286	.47129	.16638	.050
H(D5)	-.31136	.47147	.17250	.050
H(D6)	-.37472	.36783	.16842	.050
H(E2)	-.28575	.00384	.01845	.050
H(E3)	-.37913	.00994	-.06734	.050
H(E4)	-.44654	.11167	-.09522	.050
H(E5)	-.42054	.20725	-.03732	.050
H(E6)	-.32720	.20112	.04846	.050
H(F2)	-.17641	.11860	.01802	.050
H(F3)	-.08141	.17769	.00010	.050
H(F4)	-.00220	.24470	.06824	.050
H(F5)	-.00919	.25261	.15436	.050
H(F6)	-.10420	.19354	.17225	.050
H(G2)	-.04746	.00045	.10852	.050
H(G3)	.06865	.03692	.10897	.050
H(G4)	.14701	.09899	.18331	.050
H(G5)	.10927	.12458	.25717	.050
H(G6)	-.00685	.08811	.25671	.050
H(H2)	-.18317	-.12444	.11784	.050
H(H3)	-.14612	-.23958	.12365	.050
H(H4)	-.03850	-.27502	.19347	.050
H(H5)	.03209	-.19533	.25753	.050
H(H6)	-.00496	-.08023	.25176	.050
H(I2)	-.00919	-.11421	.39535	.050
H(I3)	.11630	-.09735	.44607	.050
H(I4)	.17408	.00606	.43858	.050
H(I5)	.10639	.09258	.38035	.050
H(I6)	-.01907	.07575	.32959	.050
H(J2)	-.12949	-.13342	.40162	.050
H(J3)	-.16665	-.24860	.39356	.050
H(J4)	-.21293	-.30016	.30744	.050
H(J5)	-.22209	-.23652	.22939	.050
H(J6)	-.18501	-.12133	.23746	.050
H(K2)	-.12425	.22705	.36187	.050
H(K3)	-.01667	.28989	.40688	.050

Table 5.2(cont).

	X/A	Y/B	Z/C	U
H(K4)	.07772	.24105	.48040	.050
H(K5)	.06448	.12938	.50898	.050
H(K6)	-.04310	.06654	.46396	.050
H(L2)	-.26179	.00897	.42882	.050
H(L3)	-.31409	.02617	.49640	.050
H(L4)	-.30578	.13388	.53794	.050
H(L5)	-.24510	.22436	.51190	.050
H(L6)	-.19279	.20715	.44433	.050

^aThroughout this thesis U values quoted with standard deviations are refined isotropic displacement parameters and define an isotropic temperature factor of the form $\exp(-8\pi^2 U \sin^2 \theta / \lambda^2)$. U valued without e.s.d.s are equivalent isotropic displacement parameters defined as:-

$$U = 1/3 \sum_{i=1}^3 \sum_{j=1}^3 U_{ij} a_i^* a_j^* (a_i \cdot a_j).$$

For H-atoms only isotropic U were employed.

Anisotropic displacement parameters (\AA^2)

	U11	U22	U33	U12	U13	U23
PT(1)	.0260(8)	.0272(7)	.0238(7)	.0009(7)	.0077(6)	.0002(6)
PT(2)	.0282(8)	.0262(7)	.0278(7)	.0050(7)	.0102(6)	.0045(6)
PT(3)	.0268(7)	.0274(7)	.0294(7)	.0056(7)	.0102(6)	.0048(6)
P(1)	.021(5)	.039(5)	.042(5)	.002(4)	.010(4)	-.009(4)
P(2)	.034(5)	.035(5)	.038(5)	.018(5)	.010(4)	.008(4)
P(3)	.035(5)	.032(5)	.032(4)	.010(4)	.012(4)	.005(4)
P(4)	.028(5)	.033(5)	.046(5)	.008(4)	.012(4)	.004(4)
P(5)	.036(5)	.041(5)	.032(4)	.008(5)	.018(4)	.000(4)
P(6)	.032(5)	.036(5)	.031(4)	-.005(4)	.015(4)	-.003(4)
P(7)	.037(5)	.030(5)	.036(5)	-.001(5)	.014(4)	.004(4)
P(F1)	.251(21)	.122(13)	.133(12)	-.018(16)	.091(14)	.005(11)
P(F2)	.100(14)	.164(20)	.587(45)	-.032(15)	-.072(21)	.151(28)

Throughout this thesis the form of the anisotropic displacement factor was:-

$$\exp[-2\pi^2 \sum_{i=1}^3 \sum_{j=1}^3 U_{ij} h_i h_j a_i^* a_j^*].$$

Table 5.3. Selected bond distances (Å) and angles (°) of complex (2).

PT(1) - PT(2)	2.656(2)	PT(1) - PT(3)	2.626(2)
PT(1) - P(1)	2.341(8)	PT(1) - P(6)	2.335(8)
PT(1) - P(7)	2.293(8)	PT(1) - C(4)	1.93(3)
PT(2) - PT(3)	2.636(2)	PT(2) - P(2)	2.323(8)
PT(2) - P(3)	2.290(8)	PT(2) - C(4)	2.16(3)
PT(3) - P(4)	2.288(8)	PT(3) - P(5)	2.300(8)
PT(3) - C(4)	2.27(3)	P(1) - C(1)	1.81(3)
P(1) - C(A1)	1.80(3)	P(1) - C(B1)	1.83(3)
P(2) - C(1)	1.88(3)	P(2) - C(C1)	1.84(3)
P(2) - C(D1)	1.83(3)	P(3) - C(2)	1.76(3)
P(3) - C(E1)	1.85(3)	P(3) - C(F1)	1.85(3)
P(4) - C(2)	1.89(3)	P(4) - C(G1)	1.83(3)
P(4) - C(H1)	1.81(3)	P(5) - C(3)	1.83(3)
P(5) - C(I1)	1.84(3)	P(5) - C(J1)	1.82(3)
P(6) - C(3)	1.82(3)	P(6) - C(K1)	1.83(4)
P(6) - C(L1)	1.85(3)	P(7) - O(1)	1.56(2)
P(7) - O(2)	1.57(2)	P(7) - O(3)	1.59(2)
O(1) - C(11)	1.40(4)	O(2) - C(21)	1.44(4)
O(3) - C(31)	1.37(3)	O(4) - C(4)	1.21(4)

PT(2) - PT(1) - PT(3)	59.9(1)	PT(2) - PT(1) - P(1)	98.2(2)
PT(2) - PT(1) - P(6)	148.8(2)	PT(2) - PT(1) - P(7)	94.2(2)
PT(2) - PT(1) - C(4)	53.4(8)	PT(3) - PT(1) - P(1)	156.9(2)
PT(3) - PT(1) - P(6)	93.2(2)	PT(3) - PT(1) - P(7)	90.5(2)
PT(3) - PT(1) - C(4)	57.4(8)	P(1) - PT(1) - P(6)	105.1(3)
P(1) - PT(1) - P(7)	99.1(3)	P(1) - PT(1) - C(4)	104.7(9)
P(6) - PT(1) - P(7)	102.1(3)	P(6) - PT(1) - C(4)	100.1(8)
P(7) - PT(1) - C(4)	141.8(8)	PT(1) - PT(2) - PT(3)	59.5(1)
PT(1) - PT(2) - P(2)	95.1(2)	PT(1) - PT(2) - P(3)	153.3(2)
PT(1) - PT(2) - C(4)	45.9(8)	PT(3) - PT(2) - P(2)	154.6(2)
PT(3) - PT(2) - P(3)	93.9(2)	PT(3) - PT(2) - C(4)	55.5(8)
P(2) - PT(2) - P(3)	111.5(3)	P(2) - PT(2) - C(4)	107.9(8)
P(3) - PT(2) - C(4)	122.9(8)	PT(1) - PT(3) - PT(2)	60.6(1)
PT(1) - PT(3) - P(4)	156.3(2)	PT(1) - PT(3) - P(5)	95.0(2)
PT(1) - PT(3) - C(4)	45.7(8)	PT(2) - PT(3) - P(4)	95.7(2)
PT(2) - PT(3) - P(5)	155.2(2)	PT(2) - PT(3) - C(4)	51.6(7)
P(4) - PT(3) - P(5)	108.7(3)	P(4) - PT(3) - C(4)	120.3(7)
P(5) - PT(3) - C(4)	115.5(7)	PT(1) - P(1) - C(1)	106.2(9)
PT(1) - P(1) - C(A1)	123.5(8)	PT(1) - P(1) - C(B1)	115.7(10)
C(1) - P(1) - C(A1)	98.4(12)	C(1) - P(1) - C(B1)	107.9(11)
C(A1) - P(1) - C(B1)	103.1(12)	PT(2) - P(2) - C(1)	110.0(8)
PT(2) - P(2) - C(B1)	124.3(12)	PT(2) - P(2) - C(D1)	111.7(11)
C(1) - P(2) - C(C1)	99.5(12)	C(1) - P(2) - C(D1)	102.4(12)
C(C1) - P(2) - C(D1)	106.2(14)	PT(2) - P(3) - C(2)	110.4(9)
PT(2) - P(3) - C(E1)	118.8(9)	PT(2) - P(3) - C(F1)	116.6(8)
C(2) - P(3) - C(E1)	101.9(14)	C(2) - P(3) - C(F1)	105.6(12)
C(E1) - P(3) - C(F1)	101.8(11)	PT(3) - P(4) - C(2)	106.1(9)
PT(3) - P(4) - C(G1)	116.4(10)	PT(3) - P(4) - C(H1)	121.1(11)
C(2) - P(4) - C(G1)	105.0(12)	C(2) - P(4) - C(H1)	105.6(14)
C(G1) - P(4) - C(H1)	101.2(14)	PT(3) - P(5) - C(3)	112.2(9)
PT(3) - P(5) - C(I1)	113.0(10)	PT(3) - P(5) - C(J1)	118.6(10)
C(3) - P(5) - C(I1)	103.2(12)	C(3) - P(5) - C(J1)	100.4(13)
C(I1) - P(5) - C(J1)	107.7(16)	PT(1) - P(6) - C(3)	106.6(9)
PT(1) - P(6) - C(K1)	117.4(8)	PT(1) - P(6) - C(L1)	121.0(9)
C(3) - P(6) - C(K1)	107.3(15)	C(3) - P(6) - C(L1)	103.7(11)
C(K1) - P(6) - C(L1)	99.4(11)	PT(1) - P(7) - O(1)	109.8(8)
PT(1) - P(7) - O(2)	110.7(8)	PT(1) - P(7) - O(3)	124.0(8)
O(1) - P(7) - O(2)	106.9(11)	O(1) - P(7) - O(3)	96.9(11)
O(2) - P(7) - O(3)	106.8(10)	P(7) - O(1) - C(11)	131.1(15)
P(7) - O(2) - C(21)	130.1(20)	P(7) - O(3) - C(31)	131.3(18)
P(1) - C(1) - P(2)	121.5(14)	P(3) - C(2) - P(4)	111.6(13)
P(5) - C(3) - P(6)	114.6(15)	PT(1) - C(4) - PT(2)	80.7(11)
PT(1) - C(4) - PT(3)	76.8(10)	PT(1) - C(4) - O(4)	154.6(21)
PT(2) - C(4) - PT(3)	72.9(8)	PT(2) - C(4) - O(4)	121.1(19)
PT(3) - C(4) - O(4)	119.9(20)		

Table 5.4. Fractional atomic coordinates and displacement parameters (\AA^2) of the crystal structure of complex (3).

	X/A	Y/B	Z/C	U
PT(1)	0.37496(16)	0.22406(14)	0.29094(9)	0.036
PT(2)	0.23808(17)	0.30919(15)	0.22589(10)	0.038
PT(3)	0.26417(17)	0.12769(14)	0.27562(10)	0.038
F(1)	0.4150(12)	0.3554(10)	0.2898(6)	0.050
F(2)	0.2758(12)	0.4455(11)	0.2027(7)	0.064
F(3)	0.0950(10)	0.3341(10)	0.1628(7)	0.052
F(4)	0.1340(11)	0.1057(11)	0.2222(7)	0.059
F(5)	0.0375(11)	-0.0098(9)	0.3417(6)	0.046
F(6)	0.4735(11)	0.1012(10)	0.3468(6)	0.049
F(7)	0.1620(12)	0.2204(10)	0.5349(6)	0.091
F(8)	0.781(3)	0.200(3)	0.102(2)	0.25(2)
F(1)	0.094(2)	0.330(2)	0.517(2)	0.14(2)
F(2)	0.243(3)	0.255(3)	0.547(1)	0.14(2)
F(3)	0.215(4)	0.208(3)	0.479(1)	0.21(2)
F(4)	0.109(4)	0.233(3)	0.591(1)	0.18(2)
F(5)	0.230(2)	0.111(2)	0.553(2)	0.17(2)
F(6)	0.081(3)	0.186(3)	0.523(1)	0.22(3)
N(1)	0.105(3)	0.255(3)	0.355(2)	0.06(1)
C(1)	0.410(4)	0.421(3)	0.223(2)	0.05(2)
C(2)	0.050(4)	0.253(4)	0.208(2)	0.07(2)
C(3)	0.389(5)	0.038(4)	0.386(3)	0.08(2)
C(5)	0.170(3)	0.213(3)	0.324(2)	0.02(1)
C(A1)	0.545(3)	0.336(4)	0.294(3)	0.03(1)
C(A2)	0.623(5)	0.281(5)	0.259(1)	0.05(1)
C(A3)	0.724(5)	0.263(2)	0.262(2)	0.10(2)
C(A4)	0.748(3)	0.300(4)	0.301(3)	0.13(3)
C(A5)	0.671(5)	0.354(5)	0.327(1)	0.14(3)
C(A6)	0.569(5)	0.372(2)	0.333(2)	0.06(2)
C(B1)	0.341(6)	0.440(3)	0.336(3)	0.04(1)
C(B2)	0.286(4)	0.404(2)	0.378(2)	0.05(2)
C(B3)	0.220(3)	0.468(2)	0.413(2)	0.07(2)
C(B4)	0.209(6)	0.567(3)	0.408(3)	0.15(3)
C(B5)	0.264(4)	0.603(2)	0.366(2)	0.14(3)
C(B6)	0.330(3)	0.540(2)	0.330(2)	0.10(2)
C(C1)	0.179(5)	0.562(3)	0.231(3)	0.04(1)
C(C2)	0.186(4)	0.654(2)	0.221(2)	0.07(2)
C(C3)	0.120(4)	0.733(2)	0.248(2)	0.10(2)
C(C4)	0.045(5)	0.719(3)	0.286(3)	0.12(3)
C(C5)	0.037(4)	0.627(2)	0.297(2)	0.09(2)
C(C6)	0.104(4)	0.548(2)	0.269(2)	0.06(2)
C(D1)	0.290(6)	0.478(6)	0.131(2)	0.07(2)
C(D2)	0.217(6)	0.562(3)	0.108(4)	0.15(3)
C(D3)	0.232(4)	0.593(5)	0.055(4)	0.13(3)
C(D4)	0.323(6)	0.539(5)	0.026(2)	0.22(5)
C(D5)	0.396(5)	0.454(3)	0.049(4)	0.14(3)
C(D6)	0.380(4)	0.424(5)	0.102(4)	0.10(2)
C(E1)	0.113(3)	0.331(4)	0.113(2)	0.07(2)
C(E2)	0.212(5)	0.314(6)	0.089(4)	0.08(2)
C(E3)	0.234(5)	0.311(3)	0.036(4)	0.20(5)
C(E4)	0.158(3)	0.324(4)	0.006(2)	0.19(4)
C(E5)	0.059(5)	0.341(6)	0.030(4)	0.16(4)
C(E6)	0.037(5)	0.344(3)	0.083(4)	0.12(3)
C(F1)	-0.015(4)	0.454(2)	0.191(3)	0.04(1)
C(F2)	-0.035(5)	0.535(2)	0.152(2)	0.06(2)
C(F3)	-0.117(3)	0.625(4)	0.161(2)	0.09(2)
C(F4)	-0.179(4)	0.635(2)	0.208(3)	0.05(2)
C(F5)	-0.150(5)	0.554(3)	0.247(2)	0.12(3)
C(F6)	-0.079(3)	0.464(4)	0.238(2)	0.11(3)
C(G1)	0.046(3)	0.060(2)	0.270(2)	0.05(1)
C(G2)	0.000(4)	0.056(4)	0.324(2)	0.08(2)
C(G3)	-0.043(5)	0.022(4)	0.352(1)	0.09(2)
C(G4)	-0.101(3)	-0.007(2)	0.325(2)	0.07(2)
C(G5)	-0.034(4)	-0.003(4)	0.272(2)	0.07(2)
C(G6)	-0.011(5)	0.001(4)	0.244(1)	0.12(3)
C(H1)	0.193(5)	0.033(4)	0.179(3)	0.07(2)
C(H2)	0.242(5)	-0.067(5)	0.193(2)	0.11(2)
C(H3)	0.287(3)	-0.135(3)	0.156(3)	0.06(2)
C(H4)	0.280(5)	-0.099(4)	0.103(3)	0.10(2)

Table 5.4(cont).

	X/A	-Y/B	Z/C	U
C(H5)	0.233(5)	0.003(5)	0.088(2)	0.14(3)
C(H6)	0.188(3)	0.069(3)	0.126(3)	0.11(3)
C(I1)	0.425(4)	-0.133(2)	0.325(3)	0.03(1)
C(I2)	0.438(3)	-0.157(5)	0.274(3)	0.08(2)
C(I3)	0.504(5)	-0.252(5)	0.262(1)	0.10(2)
C(I4)	0.556(4)	-0.324(2)	0.300(3)	0.09(2)
C(I5)	0.543(3)	-0.301(5)	0.351(3)	0.10(2)
C(I6)	0.478(5)	-0.205(5)	0.363(1)	0.09(2)
C(J1)	0.249(3)	-0.046(2)	0.387(2)	0.06(2)
C(J2)	0.207(5)	-0.007(4)	0.435(1)	0.08(2)
C(J3)	0.137(5)	-0.041(4)	0.467(2)	0.06(2)
C(J4)	0.110(3)	-0.112(2)	0.453(2)	0.12(3)
C(J5)	0.152(5)	-0.150(4)	0.406(1)	0.10(2)
C(J6)	0.222(5)	-0.117(4)	0.373(2)	0.10(2)
C(K1)	0.590(4)	0.003(2)	0.313(2)	0.04(1)
C(K2)	0.645(4)	-0.088(3)	0.340(1)	0.05(1)
C(K3)	0.729(3)	-0.159(3)	0.315(2)	0.06(2)
C(K4)	0.759(4)	-0.137(2)	0.262(2)	0.04(1)
C(K5)	0.704(4)	-0.046(3)	0.235(1)	0.10(2)
C(K6)	0.520(3)	0.025(3)	0.261(2)	0.07(2)
C(L1)	0.527(5)	0.129(4)	0.399(2)	0.06(2)
C(L2)	0.629(5)	0.115(2)	0.394(2)	0.04(1)
C(L3)	0.667(3)	0.138(4)	0.433(3)	0.09(2)
C(L4)	0.602(5)	0.175(4)	0.477(2)	0.09(2)
C(L5)	0.500(5)	0.189(2)	0.482(2)	0.10(2)
C(L6)	0.462(3)	0.166(4)	0.445(3)	0.06(2)
C(N1)	0.031(5)	0.021(4)	0.397(3)	0.07(2)
C(N2)	-0.060(5)	0.283(5)	0.402(3)	0.13(3)
C(N3)	-0.140(5)	0.357(5)	0.444(3)	0.09(2)
C(N4)	-0.195(4)	0.457(4)	0.428(2)	0.06(2)
C(N5)	-0.092(7)	0.480(5)	0.424(4)	0.14(3)
C(N6)	0.006(5)	0.422(4)	0.377(3)	0.08(2)

Anisotropic displacement parameters (\AA^2).

	U11	U22	U33	U12	U13	U23
PT(1)	0.036(2)	0.034(1)	0.032(2)	-0.016(1)	-0.005(1)	-0.002(1)
PT(2)	0.035(1)	0.035(1)	0.035(2)	-0.011(1)	-0.006(1)	-0.002(1)
PT(3)	0.033(1)	0.033(1)	0.040(2)	-0.013(1)	-0.001(1)	-0.006(1)
P(1)	0.06(1)	0.04(1)	0.05(1)	-0.03(1)	0.00(1)	-0.01(1)
P(2)	0.06(1)	0.07(1)	0.05(1)	-0.04(1)	0.02(1)	-0.02(1)
P(3)	0.02(1)	0.06(1)	0.08(1)	-0.03(1)	-0.03(1)	0.00(1)
P(4)	0.04(1)	0.06(1)	0.07(1)	-0.01(1)	0.00(1)	-0.02(1)
P(5)	0.05(1)	0.04(1)	0.04(1)	-0.02(1)	0.00(1)	-0.02(1)
P(6)	0.04(1)	0.05(1)	0.05(1)	-0.02(1)	0.00(1)	-0.02(1)
P(7)	0.08(1)	0.11(1)	0.06(1)	-0.02(1)	0.00(1)	-0.02(1)

Table 5.5. Selected bond distances (Å) and angles (°) of complex (3).

PT(1) - PT(2)	2.584(4)
PT(1) - P(1)	2.205(13)
PT(2) - PT(3)	2.646(3)
PT(2) - P(3)	2.373(14)
PT(3) - P(5)	2.406(15)
P(1) - C(1)	1.80(6)
P(1) - C(B1)	1.79(6)
P(2) - C(C1)	1.88(6)
P(3) - C(2)	1.56(6)
P(3) - C(F1)	1.84(5)
P(4) - C(G1)	1.77(5)
P(5) - C(3)	1.84(6)
P(5) - C(J1)	1.76(5)
P(6) - C(K1)	1.89(5)
N(1) - C(5)	1.18(7)
C(N1) - C(N2)	1.59(10)
C(N2) - C(N3)	1.63(11)
C(N4) - C(N5)	1.47(11)

PT(1) - PT(3)	2.626(3)
PT(1) - P(6)	2.251(15)
PT(2) - P(2)	2.227(16)
PT(3) - P(4)	2.460(15)
PT(3) - C(5)	1.88(5)
P(1) - C(A1)	1.79(5)
P(2) - C(1)	1.95(6)
P(2) - C(D1)	1.83(5)
P(3) - C(E1)	1.80(5)
P(4) - C(2)	2.00(6)
P(4) - C(H1)	1.92(7)
P(5) - C(I1)	1.81(5)
P(6) - C(3)	1.88(7)
P(6) - C(L1)	1.84(5)
N(1) - C(N1)	1.53(8)
C(N1) - C(N6)	1.38(9)
C(N3) - C(N4)	1.35(9)

PT(2) - PT(1) - PT(3)	61.0(1)
PT(2) - PT(1) - P(6)	157.5(4)
PT(3) - PT(1) - P(6)	96.7(4)
PT(1) - PT(2) - PT(3)	60.3(1)
PT(1) - PT(2) - P(3)	157.5(4)
PT(3) - PT(2) - P(3)	97.7(4)
PT(1) - PT(3) - PT(2)	58.7(1)
PT(1) - PT(3) - P(5)	93.0(4)
PT(2) - PT(3) - P(4)	94.3(4)
PT(2) - PT(3) - C(5)	77.8(13)
P(4) - PT(3) - C(5)	96.3(13)
PT(1) - P(1) - C(1)	108.8(17)
PT(1) - P(1) - C(B1)	115.6(22)
C(1) - P(1) - C(B1)	111.4(25)
PT(2) - P(2) - C(1)	110.8(16)
PT(2) - P(2) - C(D1)	112.2(25)
C(1) - P(2) - C(D1)	103.9(30)
PT(2) - P(3) - C(2)	109.2(23)
PT(2) - P(3) - C(F1)	114.9(20)
C(2) - P(3) - C(F1)	103.1(28)
PT(3) - P(4) - C(2)	101.6(18)
PT(3) - P(4) - C(H1)	111.9(24)
C(2) - P(4) - C(H1)	112.7(28)
PT(3) - P(5) - C(3)	107.2(19)
PT(3) - P(5) - C(J1)	115.7(14)
C(3) - P(5) - C(J1)	98.5(26)
PT(1) - P(6) - C(3)	106.2(20)
PT(1) - P(6) - C(L1)	122.1(19)
C(3) - P(6) - C(L1)	101.3(27)
C(3) - N(1) - C(N1)	170.1(47)
P(3) - C(2) - P(4)	124.8(34)
PT(3) - C(5) - N(1)	169.9(37)
P(1) - C(A1) - C(A6)	121.5(47)
P(1) - C(B1) - C(B6)	123.2(47)
P(2) - C(C1) - C(C5)	115.4(29)
P(2) - C(D1) - C(D6)	119.8(59)
P(3) - C(E1) - C(E5)	125.1(68)
P(3) - C(F1) - C(F5)	118.4(41)
P(4) - C(G1) - C(G6)	118.2(39)
P(4) - C(H1) - C(H6)	124.8(42)
P(5) - C(I1) - C(I6)	119.6(49)
P(5) - C(J1) - C(J6)	115.0(37)
P(6) - C(K1) - C(K6)	118.8(29)
P(6) - C(L1) - C(L5)	119.0(44)
N(1) - C(N1) - C(N6)	108.6(52)
C(N1) - C(N2) - C(N3)	97.5(52)
C(N3) - C(N4) - C(N5)	97.2(54)

PT(2) - PT(1) - P(1)	96.2(4)
PT(3) - PT(1) - P(1)	156.4(4)
P(1) - PT(1) - P(6)	106.2(6)
PT(1) - PT(2) - P(2)	94.4(5)
PT(3) - PT(2) - P(2)	154.6(5)
P(2) - PT(2) - P(3)	107.7(6)
PT(1) - PT(3) - P(4)	153.0(4)
PT(1) - PT(3) - C(5)	80.5(13)
PT(2) - PT(3) - P(5)	151.4(4)
P(4) - PT(3) - P(5)	114.0(5)
P(5) - PT(3) - C(5)	94.4(14)
PT(1) - P(1) - C(A1)	119.5(20)
C(1) - P(1) - C(A1)	96.6(28)
C(A1) - P(1) - C(B1)	103.4(32)
PT(2) - P(2) - C(C1)	114.2(21)
C(1) - P(2) - C(C1)	107.8(27)
C(C1) - P(2) - C(D1)	107.3(32)
PT(2) - P(3) - C(E1)	119.1(33)
C(2) - P(3) - C(E1)	106.4(36)
C(E1) - P(3) - C(F1)	102.8(33)
PT(3) - P(4) - C(G1)	119.3(18)
C(2) - P(4) - C(G1)	106.1(21)
C(G1) - P(4) - C(H1)	105.3(26)
PT(3) - P(5) - C(I1)	121.7(24)
C(3) - P(5) - C(I1)	112.5(28)
C(I1) - P(5) - C(J1)	98.8(23)
PT(1) - P(6) - C(K1)	113.8(19)
C(3) - P(6) - C(K1)	109.1(23)
C(K1) - P(6) - C(L1)	103.3(26)
P(1) - C(1) - P(2)	106.9(27)
P(5) - C(3) - P(6)	109.1(34)
P(1) - C(A1) - C(B2)	118.4(46)
P(1) - C(B1) - C(B2)	115.5(32)
P(2) - C(C1) - C(C2)	123.2(47)
P(2) - C(D1) - C(D2)	119.3(57)
P(3) - C(E1) - C(E2)	114.8(65)
P(3) - C(F1) - C(F2)	121.6(48)
P(4) - C(G1) - C(G2)	121.7(33)
P(4) - C(H1) - C(H2)	115.2(47)
P(5) - C(I1) - C(I2)	120.4(44)
P(5) - C(J1) - C(J2)	124.9(33)
P(6) - C(K1) - C(K2)	121.2(42)
P(6) - C(L1) - C(L2)	121.0(36)
N(1) - C(N1) - C(N2)	98.3(47)
C(N2) - C(N1) - C(N6)	114.9(54)
C(N2) - C(N3) - C(N4)	119.7(61)

Table 5.6. Crystallographic details of the structure analyses of compounds (2) and (3).

Compound	[Pt ₃ (CO)(P(OPh) ₃)(dppm) ₃][PF ₆]	[Pt ₃ (CO)(CNC ₆ H ₁₁)(dppm) ₃][PF ₆] ₂
Formula	C ₉₄ H ₈₁ F ₁₂ O ₄ P ₉ Pt ₃	C ₈₃ H ₇₇ F ₁₂ NOP ₈ Pt ₃ *
Formula Wt.	2366.7	2165.6
Crystal habit	red plate	red plate
Crystal size, mm	0.36x0.24x0.16	0.64x0.52x0.44
Crystal system	monoclinic	triclinic
Space group	P 2 ₁ /c	P $\bar{1}$
a, Å	19.239(6)	14.300(11)
b, Å	19.461(10)	14.614(7)
c, Å	26.137(10)	26.088(7)
α , °	—	79.79(4)
β , °	110.28(3)	79.99(5)
γ , °	—	65.91(5)
Obtained from	23 refln, 8< θ <12°	25 refln, 10< θ <15°
V, Å ³	9179(7)	4867(5)
Z	4	2
F(000)	4616	2104
d calc, gcm ⁻³	1.713	1.478
T, K	293	297
μ (Mo-K α), cm ⁻¹	48.4	45.4
Absorption factors on F ²	0.81-1.17	0.73-1.45
Scan width, °	0.80	0.80
Max count time, s	120	95
Total refln measured	12262	12472
Unique refln	9840	11885
R _{INT}	0.050	0.030
Miller indices	h	0-15
measured,	k	-15-15
	l	-27-27
2 θ range, °	4-42	4-44
Unique refln \geq 3 σ (I)	3967	5803
No. of parameters	351	292
R	0.0475	0.1017
R _w	0.0545	0.1530
$ \Delta\rho _{\max}$, eÅ ⁻³	1.31	3.54
Δ/σ_{\max}	0.099	0.058

* Actual model used in final refinement was C₈₂F₆NP₈Pt₃

CHAPTER 6: 46-ELECTRON DERIVATIVES OF $[\text{Pt}_3(\text{CO})(\text{DPPM})_3]^{2+}$

The addition reactions of small organosulphur molecules and of alkynes to $[\text{Pt}_3(\text{CO})(\text{dppm})_3]^{2+}$ (1),⁵⁰ are particularly important as models for reactions on Pt(111) surfaces because of the participation of these small molecules in catalytic processes. The strong affinity of sulphur for some transition metals can be responsible for the poisoning of heterogeneous catalysts, while alkynes participate in many catalytic reactions. Professor Puddephatt's group has synthesised a number of 46-electron cluster products involving addition of sulphur or an alkyne to the parent 42-electron cluster. Their characterisation by X-ray analysis described below and elsewhere,^{51,62} has established that there are two structurally distinct types of 46-electron cluster, dependant upon the nature of the incoming ligand L.

When L is a bidentate, 4-electron donor ligand, such as $\text{Me}_2\text{NCS}_2^-$ (Tables 6.1-6.3) or a diphosphine,⁶³ it bridges one of the Pt-Pt bonds longitudinally. The same bond is also bridged by a doubly, rather than triply, bridging carbonyl. Apart from rupture of one Pt-CO bond the $\text{Pt}_3(\text{CO})\text{P}_6$ cluster unit remains intact.

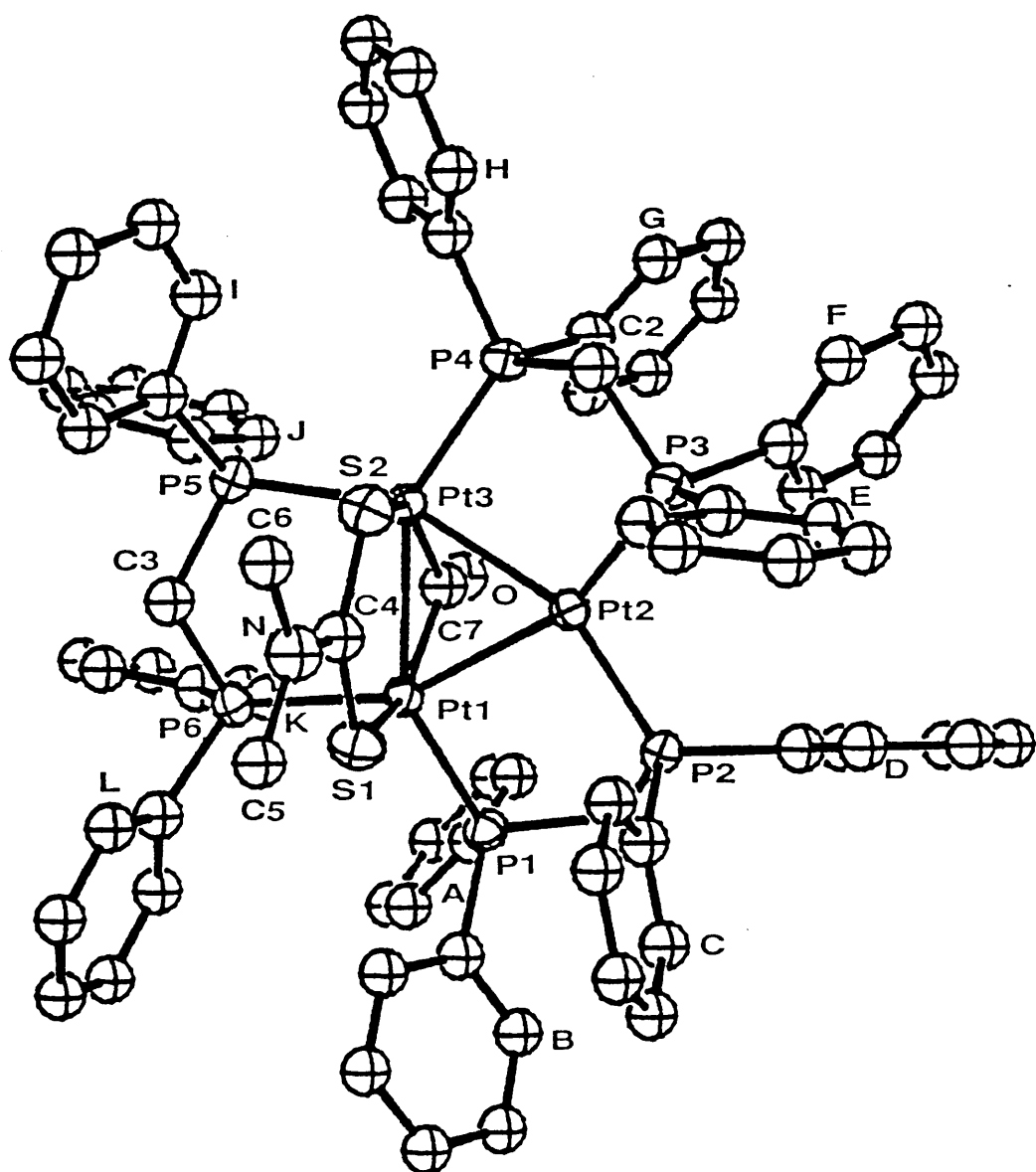
When $\text{L} = \text{HCCH}$ (Tables 6.1, 6.4 and 6.5) or an organosulphur ligand,^{51,62} the capping CO ligand is displaced by L which triply bridges the M_3 triangle. In this case the cluster geometry is more severely disrupted: two M-M bonds are ruptured and only the presence of the bridging dppm ligands stops disintegration of the cluster into mononuclear fragments.

Attachment of the dithiocarbamate ligand to the parent cluster (1) ⁵⁰ produced the 46-electron species $[\text{Pt}_3(\text{CO})(\text{dppm})_3(\text{S}_2\text{CNMe}_2)]^+$ (2), (see Figure 6(a)). X-ray analysis of a crystal of the PF_6^- salt establishes that the unit cell contains equal numbers of the cation (2) and the anion $[\text{PF}_6]^-$ (see Experimental). Interionic distances are consistent with normal van der Waals radii.⁵²

The incoming dithiocarbamate ligand is attached through Pt(1)-S(1) and Pt(3)-S(2) bonds of length 2.480(4) and 2.503(3) Å respectively, markedly shorter than the Pt-SCN ionic contact in Chapter 4 but still long for a Pt-S covalent bond. The lengths of such bonds are generally in the range 2.31-2.41 Å.⁵¹ The C_2NCS_2 skeleton is approximately planar [C(5)NC(4)S(2) and C(6)NC(4)S(1) torsion angles are 176.3(15) and -176.1(14)° respectively], and the angle between the C(5)C(6)NC(4)S(1)S(2)Pt(1)Pt(3) and Pt(1)Pt(2)Pt(3) planes is 93.2°.

The average Pt-Pt bond length in (2), of 2.613(9) Å is comparable with those in (1) and in the 44-electron species discussed in previous chapters (see Tables 4.1 and 5.1), despite the presence of four more electrons than in the parent cation (1). The major effect of the extra electrons donated by the bidentate $\text{S}_2\text{CNMe}_2^-$ ligand is on the Pt-CO bonding. The bonds from carbonyl to Pt(1) and Pt(3), which are attached to the dithiocarbamate, have shortened significantly to 2.011(11) and 2.023(11) Å, compared with the corresponding mean Pt-C distance of 2.088(9) Å in (1), while the CO ligand is no longer attached to Pt(2). The carbonyl is now directly above the Pt(1)-Pt(3) edge and normal to the Pt_3 face (the angle between the Pt_3 and

Figure 6(a). The structure of cation (2). Hydrogen atoms are omitted and carbon atoms are presented as spheres of arbitrary size. 50% probability ellipsoids are used for all other atoms.



Pt(1)Pt(3)C(7)O planes being 90.1°).

The conformation of the $\text{Pt}_3(\text{dppm})_3$ nucleus does not appear to be sterically constrained by the incoming ligand $\text{L} = \text{S}_2\text{CNMe}_2^-$ as it is by $\text{L} = \text{P}(\text{OPh})_3$ (see Chapter 5): the Pt_3 face to which the dithiocarbamate is attached has four equatorial and two axial phenyl rings whereas there are six equatorial phenyl groups on the Pt_3 face carrying the phosphite (see Chapter 8). There appear to be three types of Pt-P bonds. The mean Pt(2)-P bond length is $2.245(3) \text{ \AA}$, that for the Pt(1)-P and Pt(3)-P bonds *cis* to Pt(2) is longer at $2.304(3) \text{ \AA}$, and the mean length of the Pt(1)-P and Pt(3)-P bonds which are *trans* to Pt(2) is longer still at $2.356(3) \text{ \AA}$.

A structure similar to (2) is $[\text{Pt}_3(\text{CO})(\text{dmpm})_4]^{2+}$ (3) ($\text{dmpm} = \text{bis}(\text{dimethylphosphino})\text{methane}$ or $\text{Me}_2\text{PCH}_2\text{PMe}_2$),⁶³ also a 46-electron cluster. In this case the Pt_3 face with the extra dmpm ligand attached has six equatorial methyl groups (see Chapter 8). In (3) the CO ligand is μ_2 -bonded, as in (2), and the Pt(1)Pt(3)C(5)O and Pt(1)Pt(3)P(7)P(8) planes are approximately normal to the Pt_3 plane (102° and 91° respectively). The average Pt-Pt bond length of $2.632(14) \text{ \AA}$ is comparable with the values for (1) and (2), and the latitudinal Pt-P bonds fall into the same three categories described for (2), average distances being 2.25, 2.29 and 2.33 \AA respectively. The longest Pt-P bonds in (3) involve the longitudinal dmpm ligand, mean length 2.36 \AA , which is consistent with the long Pt-S bonds found in (2).

These differences in the latitudinal Pt-P bonds are probably due to steric factors. The four-coordinate platinum atoms give the shortest bonds, as expected;

however, the reason for the difference in length of the two sets of Pt-P bonds involving the more highly substituted metal atoms is not so obvious. It is possible that the pseudo-axial Me/Ph substituents on P(5) and P(6) are subject to steric repulsion by the CO ligand. This would also explain why the effect is greater in (2) than (3). However, the different *trans*-influence experienced by the two sets of Pt-P bonds may also be significant.

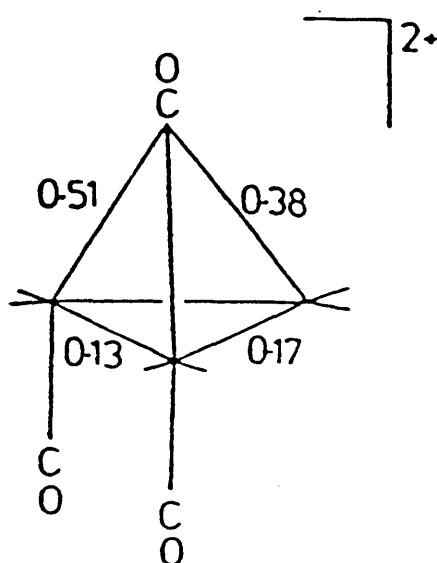


Figure 6(b). Mulliken population analysis of addition of two terminal CO ligands to the $\text{Pt}_3(\mu_3\text{-CO})(\text{dppm})_3^{2+}$ cation.

Structures (2) and (3) can be regarded as the products of the addition of two terminal 2-electron donor ligands to (1). Evans has performed EHMO calculations on such a 46-electron species formed by adding two terminal CO ligands to the parent cluster.⁵⁷ He predicts that population of the $\mu_3\text{-CO } \pi^*$ orbitals stabilises the attachment of incoming ligands to the M_3 triangle, resulting in the slipping of the $\mu_3\text{-CO}$ towards the pair of more highly substituted

platinum atoms (see Figure 6(b)). He also predicts only slight weakening of Pt-Pt bonds. A further calculation on $[\text{Pt}_3(\text{dppm})_3(\mu_2\text{-CO})(\text{CO})_2]^{2+}$, where $\mu_2\text{-CO}$ is directly above the Pt-Pt edge, suggests that this model is 0.2 eV more stable than the $\mu_3\text{-CO}$ isomer. This accords with the structures of (2) and (3).

Reaction of the monosulphur species H_2S and NCS^- with (1) gives products which are structurally different from (2) and (3). Thiocyanate reacts with $[\text{Pd}_3(\mu_3\text{-CO})(\text{dppm})_3]^{2+}$, (4)¹ first by attachment of the sulphur to the vacant face of the Pd_3 triangle, followed by cleavage of the S-C bond with transfer of CN to a Pd atom, loss of the CO and rupture of two Pd-Pd bonds.⁵¹ When $\text{M} = \text{Pt}$, the second step is very slow and the intermediate SCN^- complex can be isolated (see Chapter 4). For $\text{M} = \text{Pd}$ however, the second step is immediate and only $[\text{Pd}_3(\text{S})(\text{CN})(\text{dppm})_3]^+$, (5) can be isolated.⁵¹

Reaction of H_2S with (1) gives a similar reaction,⁶² the displacement of CO by sulphur, the rupture of two Pt-Pt bonds and attachment of hydrogen to the isolated platinum atom to give $[\text{Pt}_3(\text{S})(\text{H})(\text{dppm})_3]^+$, (6). Similar reactions with molecules of type H_2E , $\text{E} = \text{Se}, \text{S}$, with (1) and (4), and HSR with (1), have also been observed using N.M.R.⁶²

All the above structures are of the form $[\text{M}_3(\mu_3\text{-S})(\text{X})-(\text{dppm})_3]^+$ (see Figure 6(c)). The crystal structures of (5) and (6) show clearly there is only one M-M bond [length 2.579(1) and 2.597(4) Å respectively]. The other M-M distances [3.508(1) and 3.508(1) Å, and 3.574(2) and 3.678(7) Å respectively] are clearly non-bonding (Table 6.1). The side of the Pt_3 triangle with sulphur attached

has six equatorial dppm phenyl rings (see Chapter 8) in both (5) and (6) and ligand X is almost trans to the sulphur atom: $176.0(4)^\circ$ for S-Pd-CN in (5) and $179(2)^\circ$ for S-Pt-H in (6). The metal-sulphur bonds in (5) are equal at $2.297(3) \text{ \AA}$ whereas in (6) the Pt-S(trans to H) bond length of $2.375(6) \text{ \AA}$ is markedly longer than the others [$2.294(5)$ and $2.311(3) \text{ \AA}$] due to the high *trans* influence of hydrides.^{51,62}

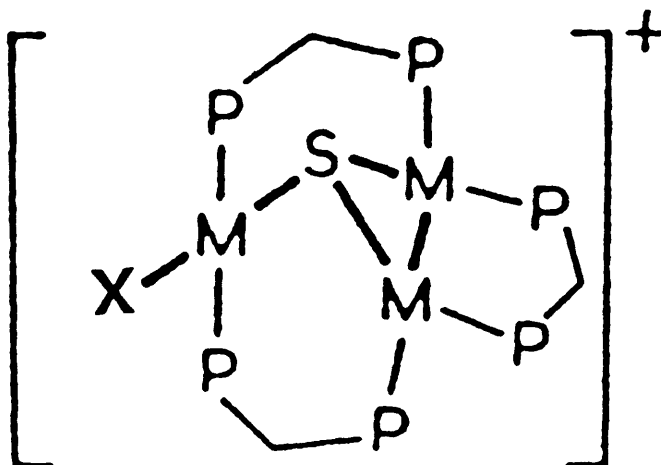


Figure 6(c). The structure of $[M_3(\mu_3-S)(X)(dppm)_3]^+$ cations.

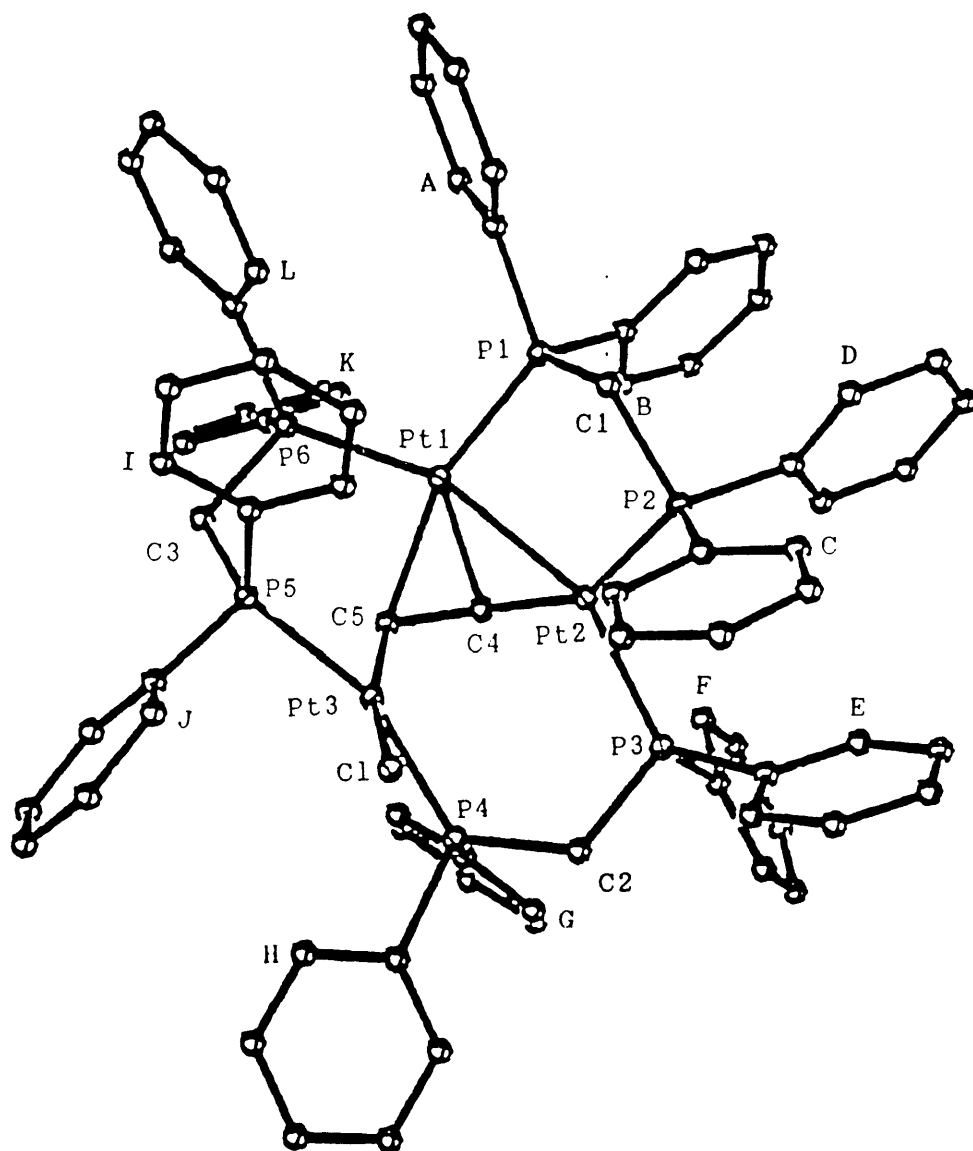
There are interesting similarities between the reaction of (1) with organosulphur compounds and the chemisorption of such compounds on Pt(111) surfaces. H_2S also displaces CO at Pt(111) surfaces with loss of hydrogen to give $Pt_3(\mu_3-S)$ units in which the Pt-S bond length of $2.28(3) \text{ \AA}$ found by LEED is comparable to the corresponding distances in μ_3 -metal clusters.²⁸ However, addition of H_2S to (1) involves rupture of M-M bonds whereas there is no evidence that formation of $Pt_3(\mu_3-S)$ units on Pt(111) surfaces leads to weakening of the surface Pt-Pt bonds.^{27,28}

The structures discussed here also provide confirmation of the important role of the capping CO ligand in stabilising these clusters. In 46-electron species which retain the CO ligand the M-M bonds retain their integrity whereas loss of the CO ligand and its empty π^* orbitals in the μ_3 -S complexes involves rupture of two M-M bonds since anti-bonding cluster orbitals are now populated.

Addition of acetylene to (1) gives ultimately the cationic cluster $[\text{Pt}_3\text{Cl}(\text{HC}\equiv\text{CH})(\text{dppm})_3]^+$, (7) whose structure is shown in Figure 6(d). X-ray analysis of a crystal of the PF_6^- salt of (8) reveals that there are two distinct PF_6^- sites to one cation in each asymmetric unit. Both anion sites straddle crystallographic diad axes and one is disordered (see Experimental), giving overall a 1:1 ratio of cations to anions. Interionic contact distances are not significantly shorter than the sum of the appropriate van der Waals radii (Tables 6.1, 6.4 and 6.5).

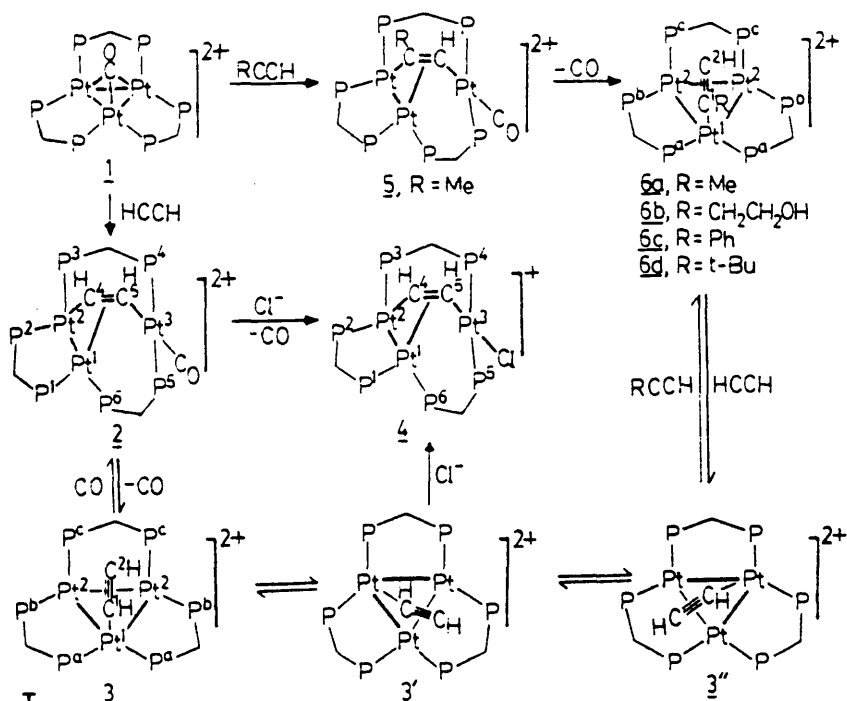
The Pt_3 triangle contains only one metal-metal bond, between Pt(1) and Pt(2) of length $2.631(2) \text{ \AA}$. The other Pt...Pt distances [$3.232(2)$ and $3.277(2) \text{ \AA}$] are clearly non-bonding. The $\text{HC}\equiv\text{CH}$ ligand lies over one face of the Pt_3 triangle. Each acetylenic carbon atom is σ -bonded to a platinum atom [Pt(2)-C(4) $2.05(3) \text{ \AA}$ and Pt(3)-C(5) $1.99(3) \text{ \AA}$] and there is π -bonding between the acetylene and Pt(1) [Pt-C(4) $2.17(3) \text{ \AA}$ and Pt-C(5) $2.21(3) \text{ \AA}$]. The chloro ligand is attached to the isolated Pt(3) atom and is *trans* to the acetylene [Pt(3)-Cl $2.408(9) \text{ \AA}$]. The conformation of dppm ligands suggests there is little steric congestion within the cluster, the acetylene side of the Pt_3 triangle having two equatorial and four axial dppm

Figure 6(d). A view of the cation $[\text{Pt}_3\text{Cl}(\text{HC}=\text{CH})(\text{dppm})_3]^+$, (7). Hydrogen atoms are omitted for clarity and all other atoms are displayed as spheres of arbitrary size.



phenyl rings (see Chapter 8). Ring J is disordered (see Experimental), adopting two orientations [torsion angles Pt(3)-P(5)-C(J1)-C(J2) and Pt(3)-P(5)-C(SJ1)-C(SJ2) are 86(4) and 54(4)° respectively]. The Pt-P distances in (7) vary significantly from 2.241(9) to 2.333(8) Å but reasons for this are not obvious.

Scheme I shows the process by which acetylene dislodges the CO ligand in (1) to a terminal position. Subsequent replacement by chloride gives (7).³ Reaction of acetylene with (1) and also with $[\text{Pt}_3(\text{dppm})_3\text{H}]^+$, (8),⁶⁴ again suggests analogies with surface chemistry. The $\text{Pt}_3(\text{HC}\equiv\text{CH})$ unit in (7) is an example of a $\mu_3-\eta^2-||$ bonding arrangement which is also favoured on Pt(111) surfaces. The alternative $\mu_3-\eta^2-|$ arrangement is also found in clusters but not on surfaces.^{3,26} N.M.R. studies of (8) have shown formation of the $\mu_3-\eta^2-\text{C}\equiv\text{CH}_2$ complex,⁶⁴ also found on Pt(111) surfaces after addition of acetylene.²⁶



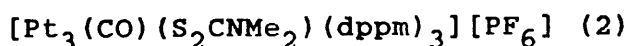
Scheme I. 3

The order of affinity of these small molecules for Pt_3 triangular units in clusters is the same as for $\text{Pt}(111)$ surfaces: i.e. $\text{H}_2\text{S} > \text{HC}\equiv\text{CH} > \text{CO}$. Acetylene can displace CO, and H_2S can displace both CO and $\text{HC}\equiv\text{CH}$, while the reverse reactions do not occur.^{3,64} It is possible, however, for one alkyne to replace another, as shown by the displacement of acetylene by propyne.

The breaking of Pt-Pt bonds in (7) can be ascribed to the population of antibonding orbitals, as with (5) and (6). In contrast, with $\text{Pt}_3(\mu_3\text{-S})$ units, however, it has been suggested that when acetylene binds to a $\text{Pt}(111)$ surface there is weakening of the Pt-Pt bonds,⁶⁵ but because metal atoms on a surface are held much tighter by lower layers and because they have a greater degree of coordinative unsaturation the effect is less marked.

Experimental

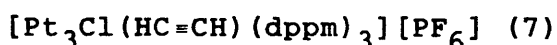
Because broadly similar methods were used for all analyses, full details are only given for the one analysis (see Experimental in Chapter 5). Any differences in procedure from that analysis are given below. Table 6.6 is a summary of the crystal data for the newly determined structures in this chapter.



Data Collection. The crystal was mounted on a glass fibre and coated in araldite.

Structure Analysis. The phenyl hydrogen atoms contributed

to the group refinement of the C_6H_5 rings. The methyl and methylenic hydrogen atoms were constrained to ride on the parent carbon atoms with $C-H = 0.96 \text{ \AA}$. Hydrogen atom displacement parameters, $U(H)$, were set at 1.2 times those of their parent carbon atoms and were not refined. Anisotropic displacement parameters were used for Pt, S, P and F atoms and isotropic parameters for all other atoms.



Data Collection. The crystal was mounted on a glass fibre and coated in araldite.

Structure Analysis. Ring J (see Figure 6(d)) proved to be disordered and both orientations were refined as rigid groups (each with an occupancy of 0.5 and an isotropic group displacement parameter). All hydrogen atoms were placed in calculated positions, except those of the alkyne and of disordered ring J. Phenyl hydrogen atoms were included in group refinements of their respective rings, while the methylenic hydrogen atoms were constrained to ride on their parent carbon atoms. Pt, Cl and P atoms had anisotropic displacement parameters, with isotropic parameters for all others. $U(H)$ values were set at and constrained to 1.2 times that of the parent carbon atoms. The two PF_6^- sites straddle two-fold axes but are crystallographically distinct. One phosphorus atom is on the axis (with occupancy 0.5) and as a consequence has only three independent fluorine atoms [F(71), F(72) and F(73)] associated with it. The other phosphorus site (occupancy 0.5) is disordered and is positioned just off the axis with

five associated fluorine sites: F(81), F(82), F(83), F(84), F(85), with occupancies of 1.0, 0.5, 0.5, 0.5 and 0.5 respectively.

Table 6.1. A comparison of selected bond distances (Å) of $[\text{Pt}_3(\mu_3\text{-Y})(\text{X})(\text{dppm})_3]^{n+}$ structures.

M	Pt ⁶²	Pt	a _{Pd} ⁵¹
X	H	Cl	CN
Y	S	HCCH	S
Pt-Pt	2.597(4)	2.631(2)	2.579(1)
Pt...Pt	3.574(2)	3.232(2)	3.508(1)
Pt...Pt	3.678(7)	3.277(2)	-
Pt ^{iso} -P	2.290(4)	2.33(1)	2.320(2)
	2.289(5)	2.25(1)	-
Pt-P(<i>trans</i> to Pt)	2.258(3)	2.24(1)	2.297(2)
	2.272(6)	2.28(1)	-
Pt-P(<i>cis</i> to Pt)	2.234(3)	2.25(1)	2.270(2)
	2.244(3)	2.28(1)	-
Pt ^{iso} -X	2.15(8)	2.408(9)	2.059(12)
Pt ^{iso} -S	2.375(6)		2.297(3)
Pt-S	2.294(5)		2.297(3)
	2.311(3)		-
Pt ^{iso} -C _σ		1.99(3)	
Pt-C _σ		2.05(3)	
Pt-C _π		2.17(3)	
Pt-C _π		2.21(3)	

^aTwo Pt atoms are crystallographically equivalent. Only one distance is listed for equivalent bonds.

Table 6.2. Fractional atomic coordinates and displacement parameters (\AA^2) of the crystal structure of compound (2).

	X/A	Y/B	Z/C	U
PT(1)	.30328(2)	-.23601(3)	-.12626(1)	.028
PT(2)	.24238(2)	-.12841(3)	-.19044(1)	.029
PT(3)	.17035(2)	-.21360(3)	-.12533(1)	.027
S(1)	.32803(13)	-.08617(23)	-.07958(10)	.043
S(2)	.17351(13)	-.05513(22)	-.08085(10)	.040
P(1)	.39886(13)	-.26247(22)	-.17095(10)	.035
P(2)	.33917(13)	-.09260(22)	-.22866(9)	.032
P(3)	.15871(13)	-.03796(21)	-.22327(10)	.034
P(4)	.07265(12)	-.18084(21)	-.16901(9)	.031
P(5)	.15349(13)	-.28968(21)	-.05147(9)	.033
P(6)	.30893(13)	-.32700(22)	-.05547(9)	.033
P(7)	.72740(14)	.40578(27)	.36422(12)	.081
F(1)	.8054(2)	.4080(7)	.3614(3)	.131
F(2)	.6494(2)	.4036(7)	.3671(3)	.202
F(3)	.7307(6)	.4664(7)	.4097(2)	.166
F(4)	.7240(6)	.3452(7)	.3187(2)	.162
F(5)	.7314(2)	.3116(5)	.3935(3)	.117
F(6)	.7234(2)	.4999(5)	.3350(3)	.154
O	.2245(3)	-.3876(6)	-.1822(3)	.045(2)
N	.2627(4)	.0438(7)	-.0312(3)	.049(3)
C(1)	.3886(5)	-.2068(8)	-.2298(4)	.043(3)
C(2)	.0797(5)	-.0522(8)	-.1893(4)	.037(3)
C(3)	.2362(5)	-.2876(8)	-.0200(4)	.041(3)
C(4)	.2550(5)	-.0262(8)	-.0625(4)	.038(3)
C(5)	.3295(6)	.0714(10)	-.0106(4)	.065(4)
C(6)	.2043(6)	.0955(11)	-.0118(5)	.071(4)
C(7)	.2291(5)	-.3121(8)	-.1596(4)	.041(3)
C(A1)	.4145(6)	-.3908(4)	-.1856(4)	.040(3)
C(A2)	.4638(6)	-.4423(10)	-.1607(2)	.055(3)
C(A3)	.4699(3)	-.5421(10)	-.1669(4)	.068(4)
C(A4)	.4267(5)	-.5904(4)	-.1981(4)	.065(4)
C(A5)	.3775(5)	-.5388(10)	-.2230(2)	.072(4)
C(A6)	.3714(3)	-.4390(10)	-.2167(5)	.060(4)
C(B1)	.4821(5)	-.2182(11)	-.1496(3)	.046(3)
C(B2)	.5381(7)	-.2215(11)	-.1787(3)	.067(4)
C(B3)	.5998(5)	-.1839(4)	-.1634(4)	.078(4)
C(B4)	.6056(5)	-.1429(10)	-.1188(3)	.084(5)
C(B5)	.5496(7)	-.1396(8)	-.0897(3)	.095(5)
C(B6)	.4878(4)	-.1772(6)	-.1051(4)	.060(3)
C(C1)	.3930(6)	-.0004(8)	-.2014(2)	.036(3)
C(C2)	.4623(7)	.0054(5)	-.2095(4)	.051(3)
C(C3)	.5005(3)	.0790(7)	-.1888(4)	.059(3)
C(C4)	.4694(5)	.1470(7)	-.1600(2)	.055(3)
C(C5)	.4001(6)	.1413(4)	-.1519(4)	.068(4)
C(C6)	.3620(3)	.0676(8)	-.1726(3)	.049(3)
C(D1)	.3371(5)	-.0551(11)	-.2908(3)	.041(3)
C(D2)	.3321(7)	.0436(10)	-.3007(4)	.065(4)
C(D3)	.3310(5)	.0754(6)	-.3471(5)	.112(6)
C(D4)	.3348(4)	.0085(9)	-.3836(3)	.098(5)
C(D5)	.3398(7)	-.0902(8)	-.3737(4)	.087(5)
C(D6)	.3409(5)	-.1219(7)	-.3273(5)	.058(3)
C(E1)	.1743(6)	.0950(6)	-.2205(4)	.036(3)
C(E2)	.1822(7)	.1521(9)	-.2604(3)	.065(4)
C(E3)	.1969(4)	.2505(7)	-.2559(3)	.079(4)
C(E4)	.2035(5)	.2918(5)	-.2115(3)	.079(4)
C(E5)	.1956(6)	.2347(8)	-.1716(2)	.060(4)
C(E6)	.1809(3)	.1363(7)	-.1761(3)	.046(3)
C(F1)	.1303(6)	-.0600(10)	-.2843(2)	.038(3)
C(F2)	.0745(7)	-.0117(5)	-.3033(4)	.050(3)
C(F3)	.0505(3)	-.0363(8)	-.3479(4)	.062(4)
C(F4)	.0823(5)	-.1091(8)	-.3734(2)	.063(4)
C(F5)	.1381(5)	-.1574(4)	-.3543(4)	.062(4)
C(F6)	.1621(3)	-.1328(10)	-.3097(4)	.043(3)
C(G1)	.0540(7)	-.2467(9)	-.2238(2)	.035(3)
C(G2)	.0014(6)	-.2141(4)	-.2526(4)	.043(3)
C(G3)	-.0130(3)	-.2616(9)	-.2947(4)	.054(3)
C(G4)	.0252(6)	-.3417(8)	-.3080(2)	.060(3)
C(G5)	.0778(5)	-.3743(6)	-.2791(4)	.062(4)
C(G6)	.0922(4)	-.3268(10)	-.2371(4)	.052(3)
C(H1)	-.0129(4)	-.1871(7)	-.1424(4)	.040(3)
C(H2)	-.0511(5)	-.1057(6)	-.1310(2)	.056(3)
C(H3)	-.1163(5)	-.1168(5)	-.1137(4)	.077(4)

Table 6.2(cont).

	X/A	Y/B	Z/C	U
C(H4)	-.1432(3)	-.2092(6)	-.1078(3)	.069(4)
C(H5)	-.1051(5)	-.2906(5)	-.1191(3)	.059(4)
C(H6)	-.0399(6)	-.2796(6)	-.1364(5)	.049(3)
C(I1)	.0969(6)	-.2270(4)	-.0102(5)	.037(3)
C(I2)	.0301(7)	-.2146(9)	-.0251(3)	.054(3)
C(I3)	-.0145(4)	-.1606(10)	.0018(4)	.068(4)
C(I4)	.0076(5)	-.1189(4)	.0438(4)	.061(4)
C(I5)	.0744(6)	-.1313(9)	.0588(2)	.071(4)
C(I6)	.1190(3)	-.1854(10)	.0318(4)	.056(3)
C(J1)	.1258(6)	-.4172(6)	-.0493(4)	.043(3)
C(J2)	.1240(3)	-.4728(7)	-.0900(3)	.051(3)
C(J3)	.1045(6)	-.5698(9)	-.0881(3)	.071(4)
C(J4)	.0869(5)	-.6112(5)	-.0454(4)	.085(5)
C(J5)	.0887(4)	-.5556(8)	-.0046(2)	.102(5)
C(J6)	.1082(7)	-.4586(9)	-.0065(4)	.075(4)
C(K1)	.3024(6)	-.4614(5)	-.0569(4)	.040(3)
C(K2)	.3123(3)	-.5090(7)	-.0994(3)	.050(3)
C(K3)	.3091(6)	-.6096(7)	-.1014(3)	.058(3)
C(K4)	.2961(5)	-.6631(4)	-.0610(4)	.078(4)
C(K5)	.2862(4)	-.6155(7)	-.0186(2)	.081(4)
C(K6)	.2893(7)	-.5147(8)	-.0165(4)	.066(4)
C(L1)	.3833(5)	-.3097(12)	-.0159(4)	.039(3)
C(L2)	.3861(4)	-.2347(8)	.0167(5)	.050(3)
C(L3)	.4452(7)	-.2188(7)	.0427(4)	.070(4)
C(L4)	.5014(4)	-.2780(10)	.0360(3)	.076(4)
C(L5)	.4985(5)	-.3531(6)	.0034(5)	.075(4)
C(L6)	.4394(7)	-.3689(9)	-.0226(3)	.060(3)
H(11)	.36552	-.25264	-.25014	.051
H(12)	.43301	-.19341	-.24234	.051
H(21)	.04093	-.03583	-.20885	.045
H(22)	.08076	-.00931	-.16233	.045
H(31)	.23322	-.32991	.00699	.049
H(32)	.24472	-.22193	-.00940	.049
H(51)	.33508	.12282	.01227	.071
H(52)	.35700	.08909	-.03719	.071
H(53)	.34787	.01328	.00370	.071
H(61)	.20995	.14690	.01109	.079
H(62)	.17516	.04746	.00235	.079
H(63)	.18108	.12396	-.03856	.079
H(A2)	.49375	-.40876	-.13906	.071
H(A3)	.50409	-.57797	-.14962	.084
H(A4)	.43100	-.65981	-.20243	.077
H(A5)	.34750	-.57240	-.24462	.082
H(A6)	.33716	-.40318	-.23404	.068
H(B2)	.53406	-.24997	-.20974	.082
H(B3)	.63882	-.18624	-.18366	.096
H(B4)	.64860	-.11676	-.10818	.101
H(B5)	.55360	-.11112	-.05870	.104
H(B6)	.44885	-.17488	-.08477	.073
H(C2)	.48386	-.04191	-.22957	.059
H(C3)	.54862	.08301	-.19445	.070
H(C4)	.49594	.19825	-.14562	.064
H(C5)	.37853	.18852	-.13187	.077
H(C6)	.31378	.06357	-.16697	.056
H(D2)	.32942	.09016	-.27535	.077
H(D3)	.32748	.14405	-.35401	.122
H(D4)	.33401	.03062	-.41587	.121
H(D5)	.34246	-.13670	-.39908	.097
H(D6)	.34436	-.19061	-.32046	.071
H(E2)	.17752	.12338	-.29130	.077
H(E3)	.20240	.29019	-.28361	.093
H(E4)	.21371	.36027	-.20835	.089
H(E5)	.20021	.26350	-.14073	.072

Table 6.2(cont).

	X/A	Y/B	Z/C	U
H(E6)	.17533	.09666	-.14840	.053
H(F2)	.05245	.03892	-.28560	.057
H(F3)	.01174	-.00276	-.36115	.077
H(F4)	.06558	-.12626	-.40435	.072
H(F5)	.16016	-.20804	-.37200	.075
H(F6)	.20088	-.16634	-.29646	.054
H(G2)	-.02523	-.15835	-.24341	.050
H(G3)	-.04967	-.23885	-.31474	.067
H(G4)	.01514	-.37471	-.33722	.071
H(G5)	.10437	-.43003	-.28836	.072
H(G6)	.12885	-.34951	-.21706	.061
H(H2)	-.03234	-.04142	-.13519	.069
H(H3)	-.14285	-.06015	-.10584	.088
H(H4)	-.18857	-.21690	-.09575	.081
H(H5)	-.12383	-.35494	-.11498	.069
H(H6)	-.01332	-.33623	-.14435	.060
H(I2)	.01471	-.24355	-.05433	.063
H(I3)	-.06096	-.15194	-.00856	.077
H(I4)	-.02340	-.08131	.06256	.074
H(I5)	.08979	-.10235	.08796	.087
H(I6)	.16546	-.19398	.04219	.067
H(J2)	.13620	-.44397	-.11976	.061
H(J3)	.10322	-.60844	-.11652	.087
H(J4)	.07331	-.67868	-.04408	.100
H(J5)	.07644	-.58446	.02511	.119
H(J6)	.10945	-.41996	.02184	.089
H(K2)	.32134	-.47190	-.12748	.061
H(K3)	.31603	-.64287	-.13098	.070
H(K4)	.29392	-.73320	-.06245	.088
H(K5)	.27717	-.65258	.00956	.093
H(K6)	.28247	-.48161	.01308	.078
H(L2)	.34705	-.19348	.02137	.061
H(L3)	.44722	-.16660	.06544	.079
H(L4)	.54247	-.26698	.05413	.089
H(L5)	.53757	-.39424	-.00126	.096
H(L6)	.43742	-.42113	-.04533	.075

Anisotropic displacement parameters (\AA^2).

	U11	U22	U33	U12	U13	U23
PT(1)	.0270(2)	.0284(3)	.0274(2)	.0031(2)	.0000(2)	.0012(2)
PT(2)	.0281(2)	.0279(3)	.0295(2)	.0017(2)	0.0000(2)	.0037(2)
PT(3)	.0275(2)	.0274(3)	.0270(2)	.0016(2)	.0008(2)	-.0007(2)
S(1)	.040(2)	.040(2)	.048(2)	0.000(1)	-.004(1)	-.007(2)
S(2)	.043(2)	.036(2)	.043(2)	.004(1)	0.000(1)	-.009(1)
P(1)	.033(1)	.035(2)	.037(2)	.003(1)	.005(1)	.002(1)
P(2)	.033(1)	.033(2)	.029(2)	-.001(1)	.003(1)	.003(1)
P(3)	.036(1)	.026(2)	.039(2)	.001(1)	-.002(1)	.002(1)
P(4)	.030(1)	.028(2)	.037(2)	.002(1)	-.001(1)	-.002(1)
P(5)	.037(2)	.032(2)	.030(2)	-.002(1)	.003(1)	.000(1)
P(6)	.037(1)	.032(2)	.031(2)	.001(1)	-.003(1)	.004(1)
P(7)	.067(3)	.077(3)	.100(3)	.014(2)	-.017(2)	-.006(3)
F(1)	.070(6)	.165(11)	.158(9)	.011(7)	.009(6)	-.007(9)
F(2)	.052(5)	.227(16)	.326(17)	.008(8)	-.016(8)	.089(14)
F(3)	.166(10)	.201(14)	.133(9)	.050(10)	.002(8)	-.085(10)
F(4)	.271(15)	.122(10)	.093(7)	-.038(10)	-.045(9)	-.016(7)
F(5)	.125(8)	.094(8)	.131(8)	.003(7)	-.001(6)	.009(7)
F(6)	.208(12)	.100(9)	.154(10)	.042(9)	-.014(9)	.038(8)

Table 6.3. Selected bond distances (Å) and angles (°) of complex (2).

PT(1) - PT(2)	2.612(1)	PT(1) - PT(3)	2.622(1)
PT(1) - S(1)	2.480(4)	PT(1) - P(1)	2.298(3)
PT(1) - P(6)	2.356(3)	PT(1) - C(7)	2.011(11)
PT(2) - PT(3)	2.605(1)	PT(2) - P(2)	2.246(3)
PT(2) - P(3)	2.244(3)	PT(3) - S(2)	2.503(3)
PT(3) - P(4)	2.309(3)	PT(3) - P(5)	2.357(3)
PT(3) - C(7)	2.023(11)	S(1) - C(4)	1.722(11)
S(2) - C(4)	1.719(10)	P(1) - C(1)	1.836(12)
P(1) - C(A1)	1.828(8)	P(1) - C(B1)	1.834(12)
P(2) - C(1)	1.837(12)	P(2) - C(C1)	1.809(12)
P(2) - C(D1)	1.829(10)	P(3) - C(2)	1.841(10)
P(3) - C(E1)	1.844(9)	P(3) - C(F1)	1.829(7)
P(4) - C(2)	1.854(12)	P(4) - C(G1)	1.824(9)
P(4) - C(H1)	1.846(8)	P(5) - C(3)	1.838(11)
P(5) - C(I1)	1.831(13)	P(5) - C(J1)	1.826(9)
P(6) - C(3)	1.831(11)	P(6) - C(K1)	1.841(7)
P(6) - C(L1)	1.840(11)	O - C(7)	1.217(14)
N - C(4)	1.310(15)	N - C(5)	1.473(16)
N - C(6)	1.456(16)		
PT(2) - PT(1) - PT(3)	59.7(1)	PT(2) - PT(1) - S(1)	89.4(1)
PT(2) - PT(1) - P(1)	94.3(1)	PT(2) - PT(1) - P(6)	154.5(1)
PT(2) - PT(1) - C(7)	69.2(4)	PT(3) - PT(1) - S(1)	95.0(1)
PT(3) - PT(1) - P(1)	147.0(1)	PT(3) - PT(1) - P(6)	95.2(1)
PT(3) - PT(1) - C(7)	49.7(4)	S(1) - PT(1) - P(1)	105.4(1)
S(1) - PT(1) - P(6)	88.7(1)	S(1) - PT(1) - C(7)	144.2(4)
P(1) - PT(1) - P(6)	110.6(1)	P(1) - PT(1) - C(7)	104.5(4)
P(6) - PT(1) - C(7)	98.7(4)	PT(1) - PT(2) - PT(3)	60.3(1)
PT(1) - PT(2) - P(2)	94.3(1)	PT(1) - PT(2) - P(3)	157.3(1)
PT(3) - PT(2) - P(2)	154.6(1)	PT(3) - PT(2) - P(3)	98.0(1)
P(2) - PT(2) - P(3)	107.3(1)	PT(1) - PT(3) - PT(2)	60.0(1)
PT(1) - PT(3) - S(2)	95.0(1)	PT(1) - PT(3) - P(4)	146.5(1)
PT(1) - PT(3) - P(5)	96.1(1)	PT(1) - PT(3) - C(7)	49.3(3)
PT(2) - PT(3) - S(2)	87.6(1)	PT(2) - PT(3) - P(4)	89.3(1)
PT(2) - PT(3) - P(5)	154.7(1)	PT(2) - PT(3) - C(7)	69.2(4)
S(2) - PT(3) - P(4)	96.7(1)	S(2) - PT(3) - P(5)	86.6(1)
S(2) - PT(3) - C(7)	143.5(4)	P(4) - PT(3) - P(5)	115.8(1)
P(4) - PT(3) - C(7)	110.2(3)	P(5) - PT(3) - C(7)	102.4(4)
PT(1) - S(1) - C(4)	112.6(4)	PT(3) - S(2) - C(4)	111.6(4)
PT(1) - P(1) - C(1)	110.5(4)	PT(1) - P(1) - C(A1)	114.5(5)
PT(1) - P(1) - C(B1)	119.6(4)	C(1) - P(1) - C(A1)	102.1(6)
C(1) - P(1) - C(B1)	104.4(5)	C(A1) - P(1) - C(B1)	103.9(7)
PT(2) - P(2) - C(1)	105.7(4)	PT(2) - P(2) - C(C1)	116.0(4)
PT(2) - P(2) - C(D1)	120.8(4)	C(1) - P(2) - C(C1)	107.1(6)
C(1) - P(2) - C(D1)	103.2(6)	C(C1) - P(2) - C(D1)	102.7(5)
PT(2) - P(3) - C(2)	109.9(4)	PT(2) - P(3) - C(E1)	113.9(4)
PT(2) - P(3) - C(F1)	120.8(5)	C(2) - P(3) - C(E1)	102.8(6)
C(2) - P(3) - C(F1)	103.0(5)	C(E1) - P(3) - C(F1)	104.5(6)
PT(3) - P(4) - C(2)	106.6(4)	PT(3) - P(4) - C(G1)	121.0(5)
PT(3) - P(4) - C(H1)	121.6(4)	C(2) - P(4) - C(G1)	102.7(5)
C(2) - P(4) - C(H1)	103.9(5)	C(G1) - P(4) - C(H1)	98.6(6)
PT(3) - P(5) - C(3)	106.8(4)	PT(3) - P(5) - C(I1)	116.7(4)
PT(3) - P(5) - C(J1)	119.7(4)	C(3) - P(5) - C(I1)	102.8(6)
C(3) - P(5) - C(J1)	105.0(6)	C(I1) - P(5) - C(J1)	104.0(5)
PT(1) - P(6) - C(3)	106.2(4)	PT(1) - P(6) - C(K1)	120.3(5)
PT(1) - P(6) - C(L1)	118.5(5)	C(3) - P(6) - C(K1)	104.5(6)
C(3) - P(6) - C(L1)	104.2(5)	C(K1) - P(6) - C(L1)	101.3(7)
C(4) - N - C(5)	123.3(10)	C(4) - N - C(6)	121.5(10)
C(5) - N - C(6)	115.1(10)	P(1) - C(1) - P(2)	112.9(6)
P(3) - C(2) - P(4)	109.1(6)	P(5) - C(3) - P(6)	114.7(6)
S(1) - C(4) - S(2)	125.3(7)	S(1) - C(4) - N	116.5(8)
S(2) - C(4) - N	118.1(8)	PT(1) - C(7) - PT(3)	81.1(5)
PT(1) - C(7) - O	137.1(8)	PT(3) - C(7) - O	140.9(8)

Table 6.4. Fractional atomic coordinates and displacement parameters (\AA^2) of the crystal structure of compound (7).

	X/A	Y/B	Z/C	U
PT (1)	.23817 (6)	.34653 (4)	-.19976 (5)	.051
PT (2)	.30558 (6)	.26932 (5)	-.20803 (4)	.050
PT (3)	.36899 (6)	.36581 (5)	-.15848 (5)	.057
CL (1)	.4536 (4)	.3741 (3)	-.2129 (4)	.092
P (1)	.1974 (4)	.3204 (3)	-.2741 (3)	.069
P (2)	.2994 (4)	.2552 (4)	-.2949 (3)	.077
P (3)	.3793 (4)	.2199 (3)	-.1864 (3)	.050
P (4)	.4182 (4)	.3045 (3)	-.1134 (3)	.048
P (5)	.3365 (4)	.4406 (4)	-.1792 (4)	.092
P (6)	.2046 (4)	.4215 (3)	-.1753 (3)	.061
P (7)	.7500	.0364 (17)	.0000	.298
P (8)	.2754 (11)	.0532 (7)	-.4851 (8)	.083
F (71)	.8135 (24)	.0353 (17)	.9790 (18)	.32 (3)
F (72)	.7604 (24)	.0801 (20)	1.0437 (19)	.37 (3)
F (73)	.7633 (30)	-.0078 (23)	1.0354 (22)	.44 (4)
F (81)	.2193 (15)	.0524 (11)	-.4609 (13)	.201 (13)
F (82)	.2718 (16)	-.0036 (13)	-.4840 (12)	.089 (12)
F (83)	.3343 (24)	.0445 (17)	-.4897 (17)	.135 (17)
F (84)	.2500	.1080 (22)	-.5000	.274 (24)
F (85)	.2774 (21)	.0478 (16)	-.4229 (17)	.135 (17)
C (1)	.2502 (14)	.3024 (10)	-.3219 (10)	.062 (9)
C (2)	.4399 (11)	.2537 (9)	-.1572 (9)	.042 (8)
C (3)	.2659 (15)	.4574 (12)	-.1498 (12)	.086 (11)
C (4)	.2760 (12)	.3002 (10)	-.1402 (10)	.053 (9)
C (5)	.2920 (12)	.3483 (11)	-.1286 (10)	.056 (9)
C (A1)	.1464 (9)	.3582 (15)	-.3099 (18)	.100 (13)
C (A2)	.0947 (14)	.3679 (12)	-.2846 (9)	.075 (10)
C (A3)	.0531 (14)	.3966 (7)	-.3082 (14)	.105 (13)
C (A4)	.0631 (9)	.4156 (12)	-.3571 (15)	.144 (17)
C (A5)	.1147 (14)	.4059 (9)	-.3824 (7)	.134 (16)
C (A6)	.1564 (15)	.3772 (10)	-.3588 (17)	.136 (16)
C (B1)	.1522 (12)	.2679 (8)	-.2616 (8)	.059 (9)
C (B2)	.1448 (16)	.2488 (13)	-.2124 (7)	.076 (10)
C (B3)	.1121 (11)	.2071 (11)	-.2054 (8)	.084 (11)
C (B4)	.0869 (10)	.1844 (7)	-.2478 (8)	.091 (11)
C (B5)	.0944 (15)	.2035 (12)	-.2970 (6)	.099 (12)
C (B6)	.1270 (9)	.2452 (10)	-.3039 (8)	.086 (11)
C (C1)	.3625 (12)	.2618 (10)	-.3356 (13)	.099 (12)
C (C2)	.3789 (18)	.2300 (14)	-.3746 (8)	.144 (16)
C (C3)	.4286 (22)	.2390 (11)	-.4032 (14)	.170 (20)
C (C4)	.4619 (11)	.2797 (10)	-.3928 (11)	.150 (17)
C (C5)	.4456 (19)	.3115 (13)	-.3538 (10)	.157 (18)
C (C6)	.3959 (23)	.3026 (10)	-.3252 (15)	.092 (12)
C (D1)	.2656 (17)	.2008 (12)	-.3115 (11)	.093 (12)
C (D2)	.2515 (7)	.1916 (10)	-.3628 (11)	.126 (15)
C (D3)	.2244 (15)	.1484 (15)	-.3761 (8)	.148 (17)
C (D4)	.2115 (15)	.1144 (11)	-.3382 (10)	.105 (13)
C (D5)	.2256 (9)	.1237 (11)	-.2869 (10)	.128 (15)
C (D6)	.2526 (18)	.1669 (16)	-.2736 (7)	.102 (13)
C (E1)	.4197 (17)	.1863 (8)	-.2362 (12)	.055 (9)
C (E2)	.3976 (8)	.1423 (7)	-.2533 (9)	.103 (12)
C (E3)	.4232 (14)	.1185 (9)	-.2947 (8)	.136 (15)
C (E4)	.4710 (14)	.1387 (7)	-.3191 (10)	.095 (12)
C (E5)	.4931 (8)	.1827 (8)	-.3020 (7)	.069 (10)
C (E6)	.4675 (17)	.2064 (10)	-.2606 (10)	.069 (10)
C (F1)	.3607 (7)	.1706 (7)	-.1423 (7)	.046 (8)
C (F2)	.3045 (10)	.1690 (11)	-.1235 (14)	.085 (11)
C (F3)	.2890 (10)	.1342 (10)	-.0872 (11)	.108 (13)
C (F4)	.3298 (7)	.1011 (7)	-.0696 (7)	.085 (11)
C (F5)	.3860 (10)	.1028 (11)	-.0884 (13)	.102 (13)
C (F6)	.4015 (9)	.1375 (9)	-.1248 (11)	.079 (10)
C (G1)	.3873 (10)	.2775 (8)	-.0547 (5)	.047 (8)
C (G2)	.4029 (14)	.2317 (6)	-.0372 (12)	.070 (10)
C (G3)	.3788 (10)	.2134 (8)	.0079 (12)	.100 (12)
C (G4)	.3393 (9)	.2410 (7)	.0357 (5)	.093 (12)
C (G5)	.3237 (13)	.2868 (6)	.0182 (12)	.068 (10)
C (G6)	.3478 (9)	.3051 (9)	-.0269 (12)	.058 (9)
C (H1)	.4877 (13)	.3247 (15)	-.0860 (7)	.063 (9)
C (H2)	.4971 (9)	.3745 (14)	-.0868 (13)	.079 (11)
C (H3)	.5481 (14)	.3934 (6)	-.0670 (11)	.134 (16)

Table 6.4(cont).

	X/A	Y/B	Z/C	U
C(H4)	.5897(11)	.3627(13)	-.0464(7)	.091(11)
C(H5)	.5803(11)	.3130(12)	-.0456(12)	.085(11)
C(H6)	.5292(16)	.2940(8)	-.0653(10)	.083(11)
C(I1)	.3305(12)	.4607(18)	-.2448(10)	.084(11)
C(I2)	.3232(19)	.5091(15)	-.2580(13)	.111(13)
C(I3)	.3169(14)	.5223(9)	-.3095(16)	.124(15)
C(I4)	.3179(11)	.4870(15)	-.3478(9)	.166(19)
C(I5)	.3252(18)	.4386(12)	-.3346(14)	.134(16)
C(I6)	.3315(12)	.4255(12)	-.2831(17)	.115(14)
C(J1)	.3920(24)	.4802(22)	-.1597(14)	.059(8)
C(J2)	.4373(30)	.4888(21)	-.1937(13)	.059
C(J3)	.4860(19)	.5137(10)	-.1770(17)	.059
C(J4)	.4895(24)	.5300(22)	-.1263(14)	.059
C(J5)	.4442(30)	.5213(21)	-.0923(13)	.059
C(J6)	.3955(19)	.4965(10)	-.1090(17)	.059
C(SJ1)	.3845(31)	.4922(23)	-.1307(10)	.068(8)
C(SJ2)	.4441(34)	.4872(10)	-.1344(23)	.068
C(SJ3)	.4802(16)	.5185(22)	-.1073(24)	.068
C(SJ4)	.4567(31)	.5547(23)	-.0766(10)	.068
C(SJ5)	.3971(34)	.5597(10)	-.0729(23)	.068
C(SJ6)	.3610(16)	.5285(22)	-.1000(24)	.068
C(K1)	.1562(11)	.4150(7)	-.1188(9)	.046(8)
C(K2)	.1212(17)	.3742(10)	-.1163(10)	.073(10)
C(K3)	.0843(12)	.3677(9)	-.0745(7)	.091(11)
C(K4)	.0824(10)	.4020(6)	-.0352(8)	.081(11)
C(K5)	.1174(15)	.4428(10)	-.0378(9)	.091(12)
C(K6)	.1543(11)	.4493(9)	-.0796(6)	.096(12)
C(L1)	.1669(8)	.4658(15)	-.2180(18)	.063(10)
C(L2)	.1857(16)	.4693(10)	-.2689(20)	.078(11)
C(L3)	.1599(16)	.5023(10)	-.3023(11)	.123(15)
C(L4)	.1153(8)	.5318(13)	-.2848(15)	.163(19)
C(L5)	.0965(16)	.5283(8)	-.2340(17)	.191(22)
C(L6)	.1223(16)	.4953(13)	-.2006(8)	.152(17)
H(11)	.27245	.33038	-.33244	.125
H(12)	.23031	.28940	-.35165	.125
H(21)	.46220	.26739	-.18519	.159
H(22)	.46385	.23129	-.13813	.159
H(31)	.25861	.49134	-.15630	.169
H(32)	.26819	.45218	-.11292	.169
H(A2)	.08776	.35470	-.25053	.090
H(A3)	.01712	.40336	-.29056	.126
H(A4)	.03407	.43554	-.37352	.173
H(A5)	.12169	.41909	-.41642	.160
H(A6)	.19234	.37041	-.37636	.163
H(B2)	.16235	.26457	-.18294	.091
H(B3)	.10690	.19379	-.17116	.101
H(B4)	.06421	.15536	-.24294	.110
H(B5)	.07686	.18768	-.32643	.119
H(B6)	.13225	.25847	-.33819	.103
H(C2)	.35570	.20171	-.38179	.173
H(C3)	.43994	.21682	-.43026	.204
H(C4)	.49650	.28585	-.41268	.179
H(C5)	.46878	.33983	-.34657	.188
H(C6)	.38449	.32469	-.29810	.110
H(D2)	.26045	.21518	-.38917	.152
H(D3)	.21462	.14194	-.41177	.178
H(D4)	.19268	.08438	-.34746	.126
H(D5)	.21654	.10006	-.26055	.153
H(D6)	.26239	.17330	-.23795	.122
H(E2)	.36434	.12822	-.23635	.124
H(E3)	.40780	.08789	-.30659	.164
H(E4)	.48880	.12219	-.34788	.114
H(E5)	.52636	.19670	-.31895	.083
H(E6)	.48295	.23703	-.24873	.082
H(F2)	.27610	.19197	-.13574	.102
H(F3)	.24992	.13310	-.07410	.129
H(F4)	.31908	.07701	-.04428	.102
H(F5)	.41436	.07978	-.07613	.123

Table 6.4(cont).

	X/A	Y/B	Z/C	U
H(F6)	.44055	.13860	-.13784	.095
H(G2)	.43039	.21259	-.05653	.084
H(G3)	.38957	.18156	.02003	.120
H(G4)	.32255	.22827	.06713	.112
H(G5)	.29621	.30594	.03754	.082
H(G6)	.33699	.33695	-.03905	.070
H(H2)	.46818	.39584	-.10115	.095
H(H3)	.55470	.42801	-.06758	.161
H(H4)	.62521	.37587	-.03263	.109
H(H5)	.60918	.29157	-.03122	.102
H(H6)	.52268	.25941	-.06478	.100
H(I2)	.32245	.53364	-.23133	.133
H(I3)	.31180	.55596	-.31872	.149
H(I4)	.31348	.49617	-.38368	.199
H(I5)	.32592	.41408	-.36130	.161
H(I6)	.33661	.39178	-.27393	.139
H(K2)	.12244	.35032	-.14362	.087
H(K3)	.05992	.33927	-.07278	.109
H(K4)	.05677	.39743	-.00614	.097
H(K5)	.11614	.46670	-.01050	.109
H(K6)	.17872	.47770	-.08136	.115
H(L2)	.21672	.44884	-.28102	.093
H(L3)	.17299	.50473	-.33766	.148
H(L4)	.09739	.55471	-.30805	.196
H(L5)	.06548	.54877	-.22183	.229
H(L6)	.10922	.49290	-.16517	.183

Anisotropic displacement parameters (\AA^2).

	U11	U22	U33	U12	U13	U23
PT(1)	.0520(9)	.0522(8)	.0502(7)	-.0061(8)	-.0117(7)	.0001(7)
PT(2)	.0585(9)	.0504(8)	.0403(7)	-.0043(8)	-.0081(7)	-.0083(7)
PT(3)	.0472(8)	.0499(8)	.0748(9)	.0000(9)	.0034(8)	.0077(8)
CL(1)	.060(6)	.094(7)	.123(8)	.017(6)	.030(6)	.040(6)
P(1)	.089(8)	.075(7)	.044(5)	.005(7)	-.025(5)	.002(5)
P(2)	.081(7)	.098(8)	.053(5)	.004(7)	-.007(6)	-.014(6)
P(3)	.056(6)	.042(5)	.052(5)	-.002(5)	.000(4)	-.010(4)
P(4)	.044(6)	.063(6)	.036(5)	-.003(5)	-.004(4)	-.014(4)
P(5)	.048(7)	.060(7)	.167(10)	-.012(6)	-.004(7)	.028(7)
P(6)	.035(6)	.057(6)	.091(6)	-.007(5)	-.010(5)	.006(5)
P(7)	.26(6)	.22(4)	.41(6)	.00	.06(5)	.00
P(8)	.10(2)	.07(1)	.07(2)	.04(1)	.01(1)	.02(1)

Table 6.5. Selected bond distances (Å) and angles (°) of complex (7).

PT(1) - PT(2)	2.631(2)	PT(1) - P(1)	2.245(9)
PT(1) - P(6)	2.283(9)	PT(1) - C(4)	2.17(3)
PT(1) - C(5)	2.21(3)	PT(2) - P(2)	2.271(9)
PT(2) - P(3)	2.241(9)	PT(2) - C(4)	2.05(3)
PT(3) - CL(1)	2.408(9)	PT(3) - P(4)	2.333(8)
PT(3) - P(5)	2.246(10)	PT(3) - C(5)	1.99(3)
P(1) - C(1)	1.80(4)	P(1) - C(A1)	1.81(4)
P(1) - C(B1)	1.80(3)	P(2) - C(1)	1.85(4)
P(2) - C(C1)	1.80(4)	P(2) - C(D1)	1.73(4)
P(3) - C(2)	1.83(3)	P(3) - C(E1)	1.83(4)
P(3) - C(F1)	1.81(3)	P(4) - C(2)	1.86(3)
P(4) - C(G1)	1.82(2)	P(4) - C(H1)	1.83(4)
P(5) - C(3)	1.85(4)	P(5) - C(I1)	1.78(4)
P(5) - C(J1)	1.75(6)	P(6) - C(3)	1.84(4)
P(6) - C(K1)	1.84(3)	P(6) - C(L1)	1.85(5)
P(7) - F(71)	1.56(15)	P(7) - F(72)	1.66(12)
P(7) - F(73)	1.55(12)	P(8) - F(81)	1.43(5)
P(8) - F(81')	1.39(4)	P(8) - F(82)	1.56(5)
P(8) - F(83)	1.38(7)	P(8) - F(84)	1.66(6)
P(8) - F(85)	1.60(5)	C(4) - C(5)	1.40(4)

PT(2) - PT(1) - P(1)	85.4(3)	PT(2) - PT(1) - P(6)	160.9(3)
PT(2) - PT(1) - C(4)	49.5(8)	PT(2) - PT(1) - C(5)	75.8(8)
P(1) - PT(1) - P(6)	112.3(4)	P(1) - PT(1) - C(4)	125.5(8)
P(1) - PT(1) - C(5)	161.0(8)	P(6) - PT(1) - C(4)	117.9(8)
P(6) - PT(1) - C(5)	86.7(8)	C(4) - PT(1) - C(5)	37.3(11)
PT(1) - PT(2) - P(2)	100.3(3)	PT(1) - PT(2) - P(3)	155.8(2)
PT(1) - PT(2) - C(4)	53.5(8)	P(2) - PT(2) - P(3)	100.9(4)
P(2) - PT(2) - C(4)	152.1(9)	P(3) - PT(2) - C(4)	106.8(8)
CL(1) - PT(3) - P(4)	88.0(3)	CL(1) - PT(3) - P(5)	92.5(4)
CL(1) - PT(3) - C(5)	165.2(9)	P(4) - PT(3) - P(5)	159.4(4)
P(4) - PT(3) - C(5)	93.8(9)	P(5) - PT(3) - C(5)	90.9(9)
PT(1) - P(1) - C(1)	112.8(11)	PT(1) - P(1) - C(A1)	121.2(15)
PT(1) - P(1) - C(B1)	110.0(8)	C(1) - P(1) - C(A1)	104.3(17)
C(1) - P(1) - C(B1)	107.0(13)	C(A1) - P(1) - C(B1)	100.0(14)
PT(2) - P(2) - C(1)	106.7(10)	PT(2) - P(2) - C(C1)	120.2(12)
PT(2) - P(2) - C(D1)	114.6(11)	C(1) - P(2) - C(C1)	101.8(14)
C(1) - P(2) - C(D1)	103.6(16)	C(C1) - P(2) - C(D1)	107.8(16)
PT(2) - P(3) - C(2)	111.7(9)	PT(2) - P(3) - C(E1)	120.9(11)
PT(2) - P(3) - C(F1)	115.3(7)	C(2) - P(3) - C(E1)	98.9(15)
C(2) - P(3) - C(F1)	107.5(11)	C(E1) - P(3) - C(F1)	100.5(11)
PT(3) - P(4) - C(2)	111.6(9)	PT(3) - P(4) - C(G1)	120.9(9)
PT(3) - P(4) - C(H1)	113.4(13)	C(2) - P(4) - C(G1)	107.5(11)
C(2) - P(4) - C(H1)	103.1(14)	C(G1) - P(4) - C(H1)	98.2(11)
PT(3) - P(5) - C(3)	115.0(11)	PT(3) - P(5) - C(I1)	122.2(16)
PT(3) - P(5) - C(J1)	104.8(20)	C(3) - P(5) - C(I1)	104.0(15)
C(3) - P(5) - C(J1)	111.7(21)	C(I1) - P(5) - C(J1)	97.8(19)
PT(1) - P(6) - C(3)	108.7(12)	PT(1) - P(6) - C(K1)	109.5(7)
PT(1) - P(6) - C(L1)	125.8(15)	C(3) - P(6) - C(K1)	103.6(13)
C(3) - P(6) - C(L1)	102.7(15)	C(K1) - P(6) - C(L1)	104.3(14)
P(1) - C(1) - P(2)	110.4(15)	P(3) - C(2) - P(4)	115.1(15)
P(5) - C(3) - P(6)	113.2(17)	PT(1) - C(4) - PT(2)	77.0(9)
PT(1) - C(4) - C(5)	72.8(16)	PT(2) - C(4) - C(5)	118.8(20)
PT(1) - C(5) - PT(3)	100.7(11)	PT(1) - C(5) - C(4)	69.9(16)
PT(3) - C(5) - C(4)	112.3(20)		

Table 6.6 Crystallographic and experimental details of the structure analyses of compounds (2) and (7).

Compound	[Pt ₃ (CO)(S ₂ CN(CH ₃) ₂)(dppm) ₃][PF ₆]	[Pt ₃ (HC≡CH)(Cl)(dppm) ₃][PF ₆]
Formula	C ₇₉ H ₇₂ F ₆ N ₆ O ₇ Pt ₃ S ₂	C ₇₇ H ₆₈ ClF ₆ P ₇ Pt ₃
Formula Wt.	2031.7	1944.9
Crystal habit	purple plate	orange plate
Crystal size, mm	0.84x0.72x0.72	0.48x0.28x0.16
Crystal system	monoclinic	orthorhombic
Space group	P 2 ₁ /c	Pnab (No. 60)
a, Å	19.583(3)	22.987(4)
b, Å	13.661(3)	27.408(9)
c, Å	28.245(5)	25.701(11)
β, °	90.63(2)	—
Obtained from	23 refln, 12<θ<14°	23 refln, 10<θ<14°
V, Å ³	7556(3)	16192(9)
Z	4	8
F(000)	3944	7520
d calc, gcm ⁻³	1.786	1.596
T, K	295	295
μ(Mo-Kα), cm ⁻¹	58.6	54.5
Absorption factors on F ²	0.87-1.20	0.73-1.42
Scan width, °	0.80	(0.72+0.35tanθ)
Max count time, s	90	90
Total refln measured	12547	12136
Unique refln	9218	11189
R _{INT}	0.039	0.19
Miller indicies	h	0-25
measured,	k	-14- 3
	l	-29-29
2θ range, °	4-44	4-46
Unique refln ≥ 3σ(I)	6699	3925
No. of parameters	328	301
R	0.033	0.059
R _w	0.044	0.069
Δρ _{max} , eÅ ⁻³	0.98	1.05
Δ/σ _{max}	0.023	0.203

CHAPTER 7: HETEROMETALLIC SPECIES OF HIGHER NUCLEARITY

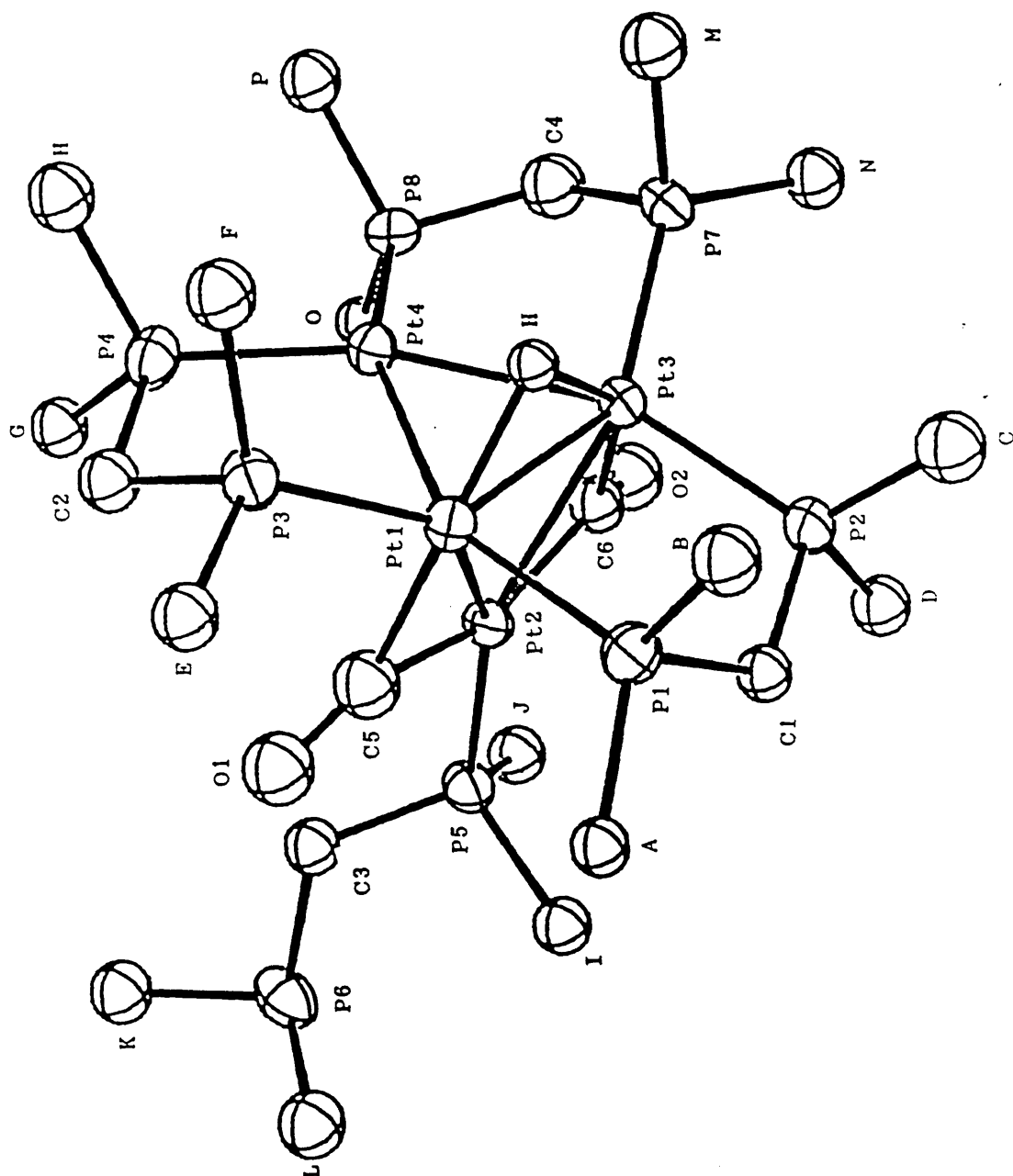
7.1 Introduction

The observation that the properties of both heterogeneous and homogeneous platinum catalysts can be usefully modified by the addition of other metals, notably gold and tin, underlies current interest in heterometallic platinum cluster chemistry.²⁹ In an attempt to develop this chemistry Professor Puddephatt has synthesised new types of clusters derived from the parent cation $[\text{Pt}_3(\text{CO})(\text{dppm})_3]^{2+}$, (1),⁵⁰ by addition of extra metal atoms. The characterisation of these new species by X-ray analysis is described here. The complexes fall into three categories: a) $[\text{Pt}_4\text{H}(\text{CO})_2(\mu\text{-dppm})_3(\text{dppm})]^+$, (2),⁴ a 58-electron tetra-platinum cluster, b) $[\text{Pt}_3(\mu_3\text{-S})(\text{AuPPh}_3)(\mu_3\text{-AgCl})(\mu\text{-dppm})_3]^+$ a pentanuclear species with gold and silver heteroatoms,⁵ and c) several compounds containing $\mu_3\text{-SnX}_3$ units attached to the basic Pt_3 triangle.⁶ Apart from the strategic objective of establishing how additional metal atoms can be attached to the parent cation (1), the X-ray studies were necessary in order to understand the reactions involved in the formation of the new complexes. As will become apparent these reactions are often of a novel and unpredictable nature.

7.2 $[\text{Pt}_4(\mu\text{-H})(\mu\text{-CO})_2(\mu\text{-dppm})_3(\text{dppm})][\text{PF}_6]$ (2)

Crystals of the PF_6^- salt of (2) were obtained by reaction of $[\text{Pt}(\text{CF}_3\text{CO}_2)_2(\text{dppm})]$ with CO followed by anion

Figure 7(a). The structure of complex (2). For clarity only the ipso carbons of the phenyl rings are shown and all hydrogen atoms except the μ -H atom are omitted. Probability ellipsoids (50%) are displayed for all atoms shown.



exchange. X-ray analysis establishes that these crystals contain the monocationic tetranuclear cluster (2) (see Figure 7(a)) and the anion PF_6^- , in equal numbers. Interionic distances are consistent with normal van der Waals radii.⁵² The structure of the cation (2) (see Tables 7.1 and 7.4) is based on the $\text{Pt}_3(\mu\text{-dppm})_3$ unit also found in (1) but with the $\mu_3\text{-CO}$ unit replaced by H^- and $\text{Pt}(\text{CO})_2(\text{dppm})$ units which lie above and below the Pt_3 face, doubly bridging the $\text{Pt}(1)\text{-Pt}(3)$ bond.

The Pt_4 framework can be described as a distorted tetrahedron or "butterfly" structure, similar to other 58-electron Pt_4 clusters,^{29,66-68} and is almost isomorphous with the neutral cluster $[\text{Pt}_4(\text{CO})_2(\text{Ph}_2\text{PCH}_2\text{P}(\text{O})\text{Ph}_2)(\mu\text{-dppm})_3]$. The hinge torsion angle $\text{Pt}(2)\text{-Pt}(1)\text{-Pt}(3)\text{-Pt}(4)$ is $83.7(1)^\circ$ which is larger than the 70.5° of a regular tetrahedron but less than the 89.4 and 98.6° found in the different crystalline forms of the butterfly cluster $[\text{Pt}_4(\text{CO})_5(\text{PMe}_2\text{Ph})_4]$.⁶⁶ The $\text{Pt}(1)\text{Pt}(3)\text{Pt}(4)$ face is edge-bridged by three $\mu\text{-dppm}$ ligands with six equatorial phenyl rings on the $\text{Pt}(\text{CO})_2(\text{dppm})$ unit side of the Pt_3 triangle, thereby relieving steric congestion.

The $\text{Pt}(2)\text{-Pt}(1)$ and $\text{Pt}(2)\text{-Pt}(3)$ edges are bridged by $\mu\text{-CO}$ ligands displaced slightly [$0.09(2)$ and $0.20(2)$ Å respectively], out of the $\text{Pt}(1)\text{Pt}(2)\text{Pt}(3)$ plane towards $\text{Pt}(4)$. The $\text{Pt}(2)\text{-CO}$ bonds are significantly shorter [$1.97(2)$ and $1.98(2)$ Å] than those from the carbonyl to $\text{Pt}(1)$ and $\text{Pt}(3)$ [$2.09(2)$ and $2.15(2)$ Å]. This asymmetry could be of relevance in the fluxional behaviour of (2) in solution. The Pt-H distances of $1.69(12)$ and $2.02(13)$ Å, are based on a least-squares refinement of the hydrogen

atom position, initially obtained from a difference synthesis and are clearly of limited accuracy. The Pt-P distances lie in the range 2.250(5)-2.266(5) Å, except for those of Pt(1)-P(1) and Pt(3)-P(2) whose bond lengths are 2.360(4) and 2.329(5) Å respectively. The steric congestion brought about by the additional longitudinal bridging of the Pt(1)-Pt(3) bond may, at least partially, explain this.

Five of the six Pt-Pt vectors [2.613(1)-2.750(1) Å] are bonding, with the sixth Pt(2)...Pt(4) distance of 3.082(1) Å clearly non-bonding. The longest Pt-Pt bonds are those bridged by the carbonyls, followed by the hinge bond Pt(1)-Pt(3), while the shortest are those bonded to the Pt(4) atom, which has a lower coordination number than the other metal atoms (see Table 7.4). The differences in Pt-Pt bond lengths in (2) follow the trends discussed in Chapter 3: Pt-Pt bonds are lengthened by μ_2 -bridging atoms, in this case carbon and hydrogen, and the bulkier the bridging atom the greater the lengthening.

Evans and Mingos³⁴ predicted that a stable Pt₄ tetrahedral cluster would have 56 bonding electrons with latitudinal ligands. On addition of further electrons, population of antibonding molecular orbitals would occur with a consequent disruption of the Pt₄ framework by lengthening or even breaking of a Pt-Pt bond. This clearly has occurred, creating the "butterfly" structure seen in here in cluster (2) and in other 58-electron complexes.^{29,66-68}

7.3 $[\text{Pt}_3(\mu_3\text{-S})(\text{AuPPh}_3)(\mu_3\text{-AgCl})(\mu\text{-dppm})_3][\text{PF}_6] \quad (3)$

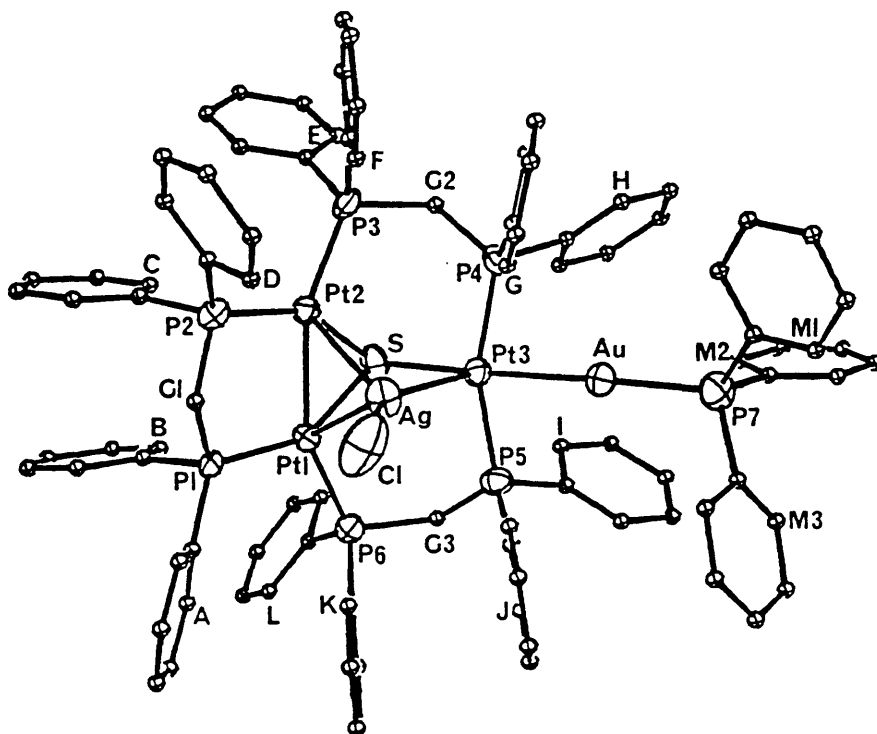
Cluster $[\text{Pt}_3(\mu_3\text{-S})(\text{AuPPh}_3)(\mu_3\text{-AgCl})(\mu\text{-dppm})_3]^+ \quad (3)$, was formed when attempting to replace the hydride ligand in $[\text{Pt}_3\text{H}(\mu_3\text{-S})(\mu\text{-dppm})_3]^+ \quad (4)$,⁶² by AuPPh_3^+ , derived from $[\text{AuCl}(\text{PPh}_3)]$ and using AgNO_3 to remove the chloride. X-ray analysis of a crystal of the PF_6^- salt of (3) revealed the presence of four ion pairs in a centrosymmetric triclinic unit cell. There are therefore two crystallographically independent cations, A and B (see Figure 7(b)), as well as two independent anions. The analysis is based on the Type II reduced triclinic cell, and the Niggli matrix elements $[A = 311.1, B = 347.9, C = 802.5, D = -44.5, E = -77.2 \text{ and } F = -1.9 \text{ \AA}^2]$ indicate that no transformation to higher symmetry is possible (see Experimental).⁶⁹ Interionic distances are consistent with van der Waals radii.⁵²

The cations A and B are chemically identical and their minor structural differences (see Tables 7.3 and 7.4) can be attributed to crystal packing. Each cation has a Pt_3 triangular face incorporating only one Pt-Pt bond. Each Pt-Pt edge is bridged latitudinally by a dppm ligand and opposite faces of the Pt_3 triangle are μ_3 -bridged by AgCl and S groups. In addition, the isolated Pt(3) atom has attached a terminal AuPPh_3 ligand. The most obvious differences in the cations occur in the distances and angles involving the Ag and Cl atoms (see Table 7.4), the orientations of the PPh_3 phenyl rings and the orientations of the dppm phenyl rings, in particular ring L (see Figure 7(b)).

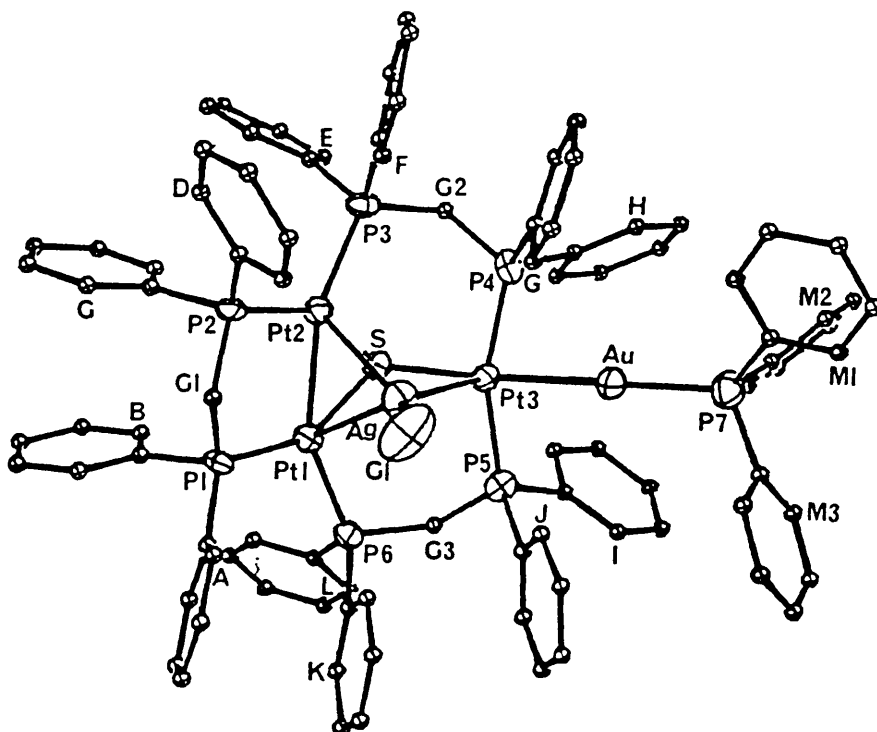
The $\text{Pt}_3(\mu_3\text{-S})\text{P}_6$ skeleton in (3) shows only minor

Figure 7(b). Views of cations A and B of complex (3). Hydrogen atoms are omitted and the carbon atoms are represented by spheres of arbitrary size. Probability ellipsoids (50%) represent all other atoms shown.

A



B



changes from (4), with the average Pt(1)-Pt(2) bond length only slightly longer than the equivalent bond in (4) [2.617(2) Å compared to 2.597(4) Å], as are the non-bonded Pt(1)...Pt(3) and Pt(2)...Pt(3) distances of 3.631(2)-3.726(2) Å compared with 3.574(2) and 3.678(7) Å in (4). These increases are probably due to the extra steric demand of the added μ_3 -AgCl ligand. Clusters $[\text{Pd}_3(\mu_3\text{-S})(\text{CN})(\text{dppm})_3]^+$, (5),⁵¹ and $[\text{Pt}_3\text{Cl}(\text{HC}\equiv\text{CH})(\text{dppm})_3]^+$, (6), which are structurally related to (4), also have shorter non-bonded M-M distances than in (3) (see Table 6.1), and seem to confirm this view.

The sulphur atom in (3) is 0.12 Å further from Pt(3) than it is from Pt(1) and Pt(2) [mean distances 2.389(6) and 2.272(5) Å respectively]. This may indicate a strong *trans*-influence exerted by the AuPPh_3 ligand on the Pt-S bond, similar to the effect of the hydride ligand in (4). In (5), however, the low *trans*-influence of CN^- leads to all Pd-S distances being equal.

Complex (3) affords the first example of an AgCl ligand triply bridging three metal atoms. The AgCl ligand is slightly closer to Pt(3) [Pt-Ag distances 2.780(4) and 2.768(4) Å for cations A and B respectively], than it is to Pt(1) and Pt(2) [Pt-Ag distances 2.808(4)-2.861(4) Å]. In B the silver atom is equidistant from Pt(1) and Pt(2) but in A the distances differ by 0.05 Å. These distances are similar to values found in complexes where a silver atom bridges two or more platinum atoms [2.74-3.06 Å].^{70,71}

The Ag-Cl distances of 2.39(1) and 2.41(1) Å are comparable with the shortest Ag-Cl distances in $[\text{Pt}_2\text{Ag}_2\text{Cl}_4(\text{C}_6\text{F}_5)_4]^{2-}$, where the chlorine atom bridges the

Ag-Pt bonds.⁷¹ The corresponding Pt-Ag-Cl angles in A and B with variations of up to 9° (see Table 7.4) suggest marked differences in the environments of the AgCl units in the different cations.

The conformations of the dppm ligands create a hydrophobic cavity in which the AgCl sits. In (4)⁶² and (5)⁵¹ the sulphur atom side of the Pt₃ triangle has six equatorial phenyl rings but in (3) it has only four equatorial and two axial rings (see Chapter 8), with the C(1) methylene being on the AgCl ligand side. The C(1) methylene group along with phenyl rings A, D, F, G, J, and K surround the AgCl unit, the shortest Ag...H and Cl...H distances being 2.79 and 2.77 Å respectively for cation A and 2.87 and 2.69 Å for cation B.

The bonding of AuPPh₃ to the cluster core is through a particularly short Pt-Au single bond [mean length 2.576(2) Å]. In general, gold tends to form Au-Au bonds or AuM₂ and AuM₃ clusters,^{72,73} and any single bonds that are formed are usually ligand bridged.⁷⁴ This is clearly not the case in cluster (3), since the linearity of the Pt-Au-P unit [mean angle 174°] rules out any close contacts between the gold atom and surrounding phenyl rings [shortest Au-H(phenyl) distance 2.92 Å]. The Pt-Au bond distances in triangular^{72,75} and linear⁷⁴ PtAu clusters range from 2.60-2.67 Å with 2.600(3) Å in PtCl(AuPPh₃)₂(PEt₃)₂]⁺⁷² the shortest distance previously reported. While in butterfly⁷⁶ and tetrahedral⁷⁷ PtAu clusters the Pt-Au bonds are all longer than 2.71 Å.

The synthesis of complexes containing Pt-Au or Pt-Ag bonds has been growing rapidly in recent years.⁷²⁻⁷⁸ This

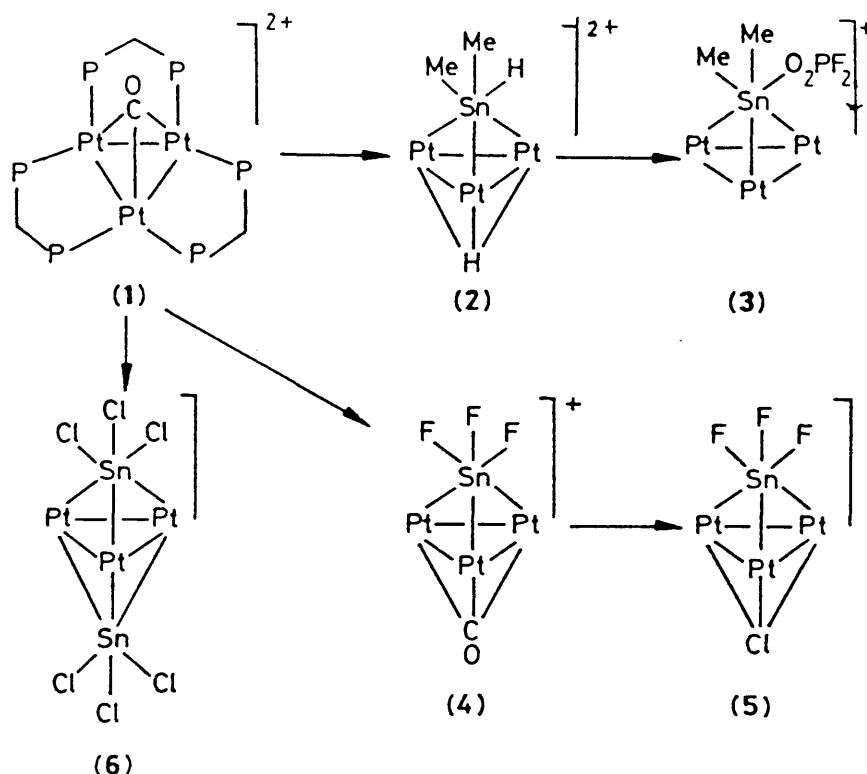
initial interest was because these complexes may serve as models for heterobimetallic catalysts where, in the case of gold added to platinum catalysts, changes in selectivity and catalytic activity were observed.⁷⁹ In addition to this primary aim, many interesting and often unexpected products have been formed in these reactions, e.g. complex (3) which displays some particularly interesting structural features.

Cluster (3) is the first PtAgAu cluster to be characterised, also the $\text{Pt}_3(\mu_3\text{-AgCl})$ unit with only one Pt-Pt bond has not been observed before. The mode of attachment of the AgCl is also of chemical interest in building of higher nuclearity clusters. If (3) is formed from the addition of electron-poor AgCl to electron-rich $[\text{Pt}_3(\mu_3\text{-S})(\text{AuPPh}_3)(\text{dppm})_3]^+$ it is clear that the AgCl prefers to add via the electron rich platinum atoms rather than to the sulphur which has an exposed donor electron pair.⁸⁰

7.4 Higher Nuclearity Clusters Containing Tin(II)

The use of Pt-Sn complexes and Pt-Sn/support combinations in homogeneous and heterogeneous catalysis respectively, has met with some success.^{81,82} On platinum surface catalysts tin(II), as well as tin(0), species are known to exist.⁸³ Therefore, it was hoped to react tin(II) species with the cluster cation (1) in an attempt to model the coordination which occurs on such surfaces. Prior to this work the only Pt_3Sn complexes characterised by X-ray diffraction analysis were $[\text{Pt}_3(\mu_3\text{-SnCl}_3)_2(\text{cod})_3]$ (cod =

cyclo-octa-1,5-diene)⁴⁷ and $[\text{Pt}_3(\text{CO})_3(\mu_2\text{-Sn}(\text{NR})_2)_3]$ ($\text{R} = \text{SiMe}_3$).⁸⁴ The first reactions carried out were between (1) and organotin compounds. Details are shown in Scheme II.⁶

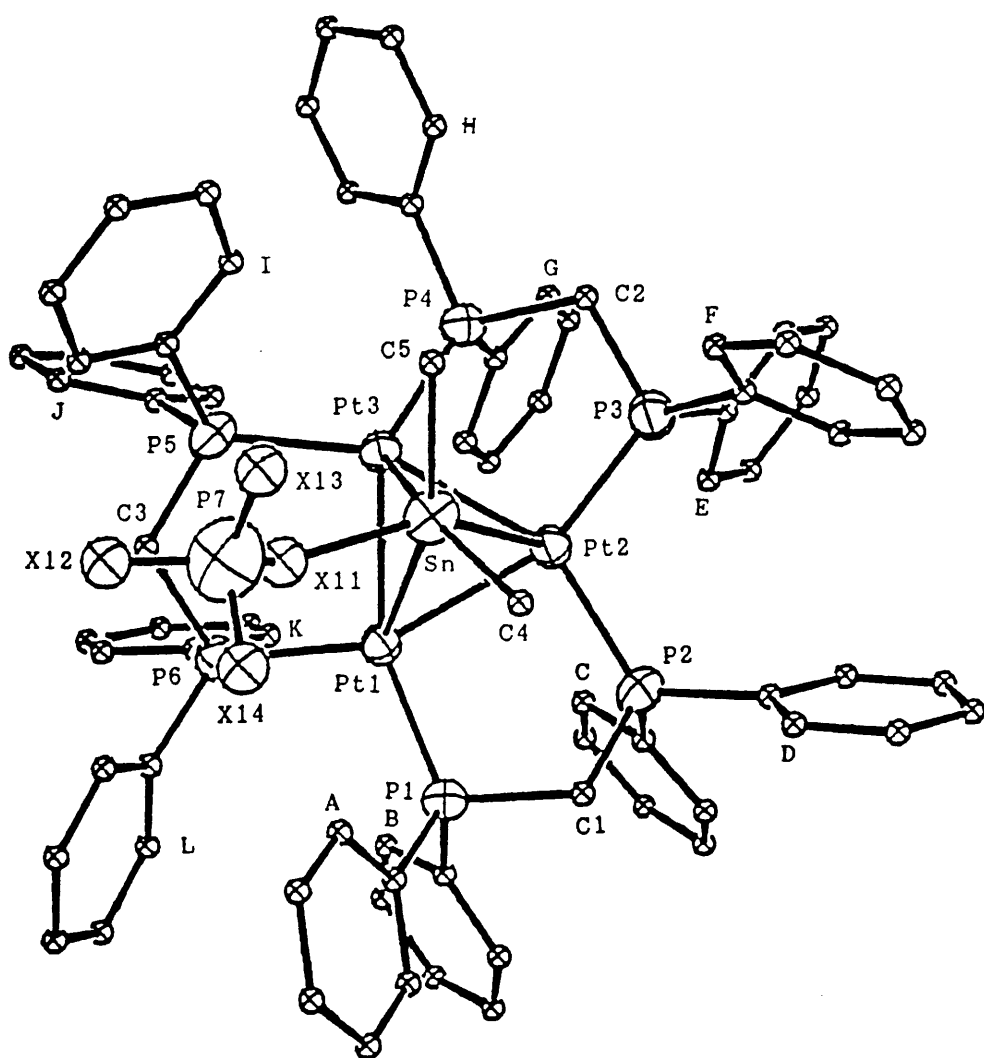


Scheme II.

Reaction of Me_2SnH_2 with the PF_6^- salt of the parent cation (1) initially formed $[\text{Pt}_3(\mu_3\text{-SnMe}_2\text{H})(\mu_3\text{-H})(\text{dppm})_3]^{2+}$ which was characterised spectroscopically. X-ray analysis indicates that on slow crystallisation there was partial reaction with a PF_6^- group to give PO_2F_2^- and that this attached itself to the tin atom, with rupture of the Sn-H and Pt-H bonds, to give crystals consisting of well separated cations $[\text{Pt}_3(\mu_3\text{-SnMe}_2(\text{PO}_2\text{F}_2))(\mu\text{-dppm})_3]^+$, (7), (see Figure 7(c)) and PF_6^- anions in equal quantities (see Tables 7.5 and 7.6).

In the Pt_3 triangle of cation (7) all three Pt-Pt

Figure 7(c). Structure of cation (7). Hydrogen atoms are omitted and carbon atoms represented by spheres of arbitrary size. Otherwise 50% probability ellipsoids are displayed. The O and F atoms of the PO_2F_2^- are represented by X(11)-X(14) as they could not be distinguished from one another in the structure analysis.



vectors are bonding and the mean Pt-Pt bond length of 2.620(14) Å is similar to that in the parent cation (see Table 4.1). The ability of $\mu_3\text{-SnMe}_2(\text{PO}_2\text{F}_2)$ to replace CO in stabilising the 42-electron cluster (7) is particularly significant. The Sn-Pt bonds vary in length from 2.702(2) to 2.766(2) Å, slightly less than the 2.80(1) Å found in $[\text{Pt}_3(\mu_3\text{-SnCl}_3)_2(\text{cod})_3]$.⁴⁷ The roughly octahedral coordination around the tin atom is completed by two methyl groups [Pt-C, 2.08(3) and 2.17(3) Å] and the PO_2F_2^- unit [Sn-X(11), 2.12(2) Å, X = O or F]. The presence of the PO_2F_2^- unit was finally confirmed by N.M.R. studies since X-ray analysis was unable to distinguish between the oxygen and the fluorine atoms attached to the phosphorus atom. P-X distances range from 1.48(3) to 1.53(2) Å, the longest being that involving the atom connecting the ligand to the tin atom.

Possibly because the $\mu_3\text{-SnMe}_2(\text{PO}_2\text{F}_2)$ unit is much bulkier than the carbonyl, the dppm phenyl rings on its side of the Pt_3 triangle adopt a six equatorial mode in order to relieve steric congestion. In (1) there are four equatorial and two axial phenyl rings on the carbonyl side. However the Pt-P distances in (7) are very similar to those in (1) [mean distances 2.282(12) Å in (7) and 2.282(15) Å in (1)].

Reaction of (1) with Ph_3SnH also gave an unexpected product.⁶ X-ray analysis carried out in this department by Dr.M.C. Jennings revealed a disordered structure of unusual type, in which cleavage of all three Sn-C(phenyl) bonds had occurred and were replaced by much more electronegative F^- ions. The cubic unit cell contains six cations $[\text{Pt}_3(\mu_3-$

$\text{SnF}_3)(\mu_3\text{-CO})(\mu\text{-dppm})_3]^+$, (8), six PF_6^- anions and two neutral clusters $[\text{Pt}_3(\mu_3\text{-SnF}_3)(\mu_3\text{-Cl})(\mu\text{-dppm})_3]$, (9). Clusters (8) and (9) occupy crystallographically equivalent sites and differ only in the μ_3 -capping ligand, CO or Cl^- . Cluster (8) is shown in Figure 7(d).

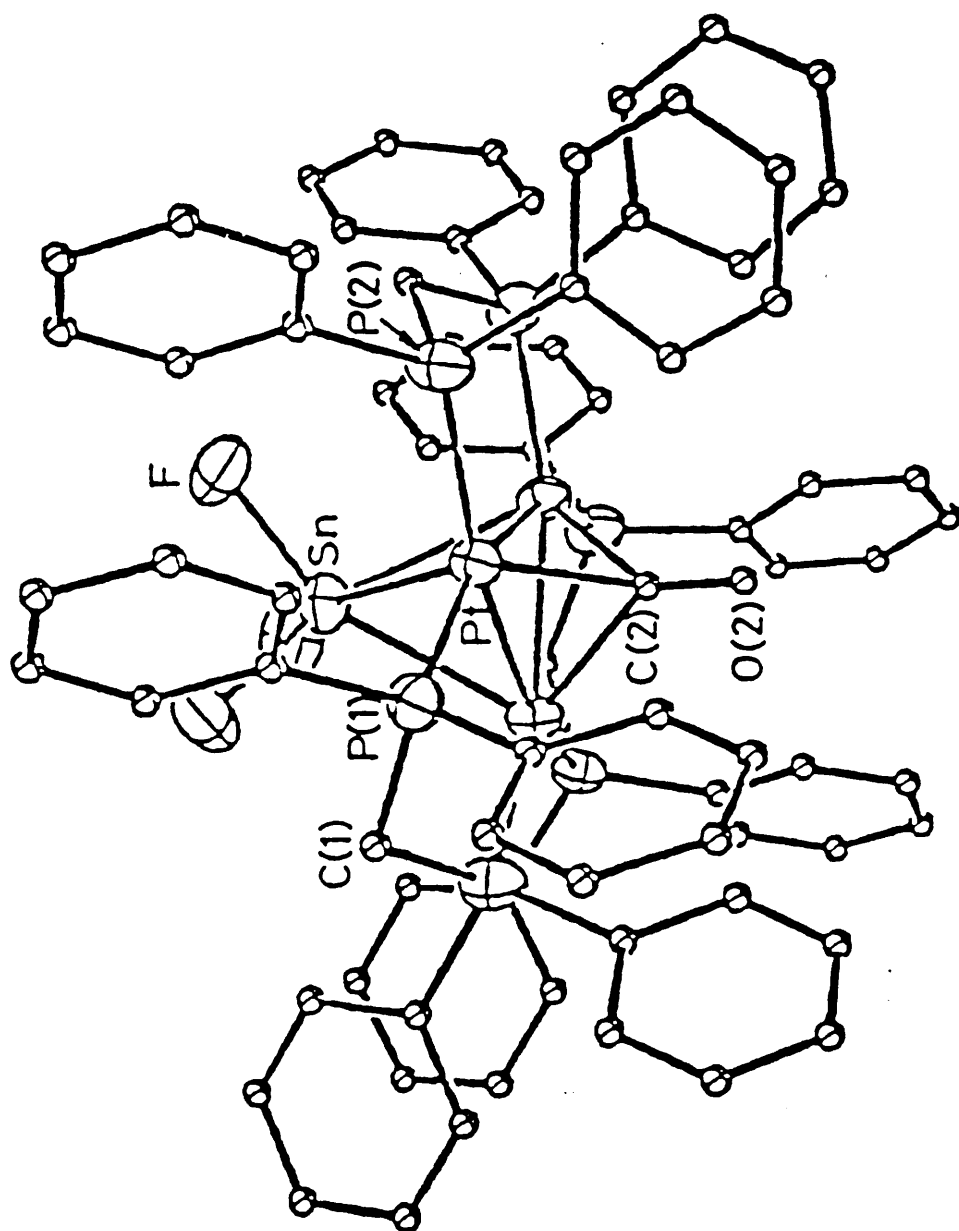
An attempt to synthesise cluster (8) by reaction of (1) with SnF_3^- provided crystals which on X-ray analysis gave almost identical unit cell dimensions to the previous crystal and a subsequent structural analysis showed them to be so (see Tables 7.7 and 7.8).

While the Pt-Pt bond distances in (8) and (9) are comparable with those in (7) [mean distance 2.634(4) and 2.620(14) Å respectively], the Pt-Sn distances do show a significant lengthening [distances 2.773(3) Å in (8) and (9) and 2.702(2)-2.766(2) Å in (7)], which may be ascribed to the increased number of bonding electrons in (8) and (9), i.e. 44 instead of 42.

The Pt-P bond distances in (8) and (9) also compare favourably with those of (7) [mean distances, 2.292(24) and 2.282(12) Å]. The much greater steric bulk of the $\mu_3\text{-SnF}_3^-$ ligand compared to the $\mu_3\text{-CO}$ and $\mu_3\text{-Cl}^-$ ligands is clearly indicated by the dppm ligands having six equatorial phenyl rings on its side of the Pt_3 triangle.

Reactions of (1) with excess SnCl_3^- or SnF_3^- gave 44-electron clusters of the type $[\text{Pt}_3(\mu_3\text{-SnX}_3)_2(\mu\text{-dppm})_3]$. In view of the disorder shown in the attempts to obtain crystals of (8), an X-ray analysis of $[\text{Pt}_3(\mu_3\text{-SnF}_3)_2(\mu\text{-dppm})_3]$, (10), recrystallised from dichloromethane, was carried out in order to study, more clearly, the $\mu_3\text{-SnF}_3$ bonding to the Pt_3 triangle. The crystal structure

Figure 7(d). A view of cation (8). Hydrogen atoms are omitted and the carbon atoms and oxygen atom are presented as arbitrary spheres. All other atoms are represented by 50% probability ellipsoids.



contains three molecules of dichloromethane for every molecule of cluster (10). Two of these solvent molecules are heavily disordered (see Experimental). Intermolecular distances are consistent with van der Waals radii.⁵² Cluster (10) is a neutral 44-electron cluster based on a Pt_3 triangle with SnF_3^- units μ_3 -bonded to each face and dppm ligands latitudinally bridge each of the Pt-Pt edges (see Figure 7(e) and Tables 7.9 and 7.10).

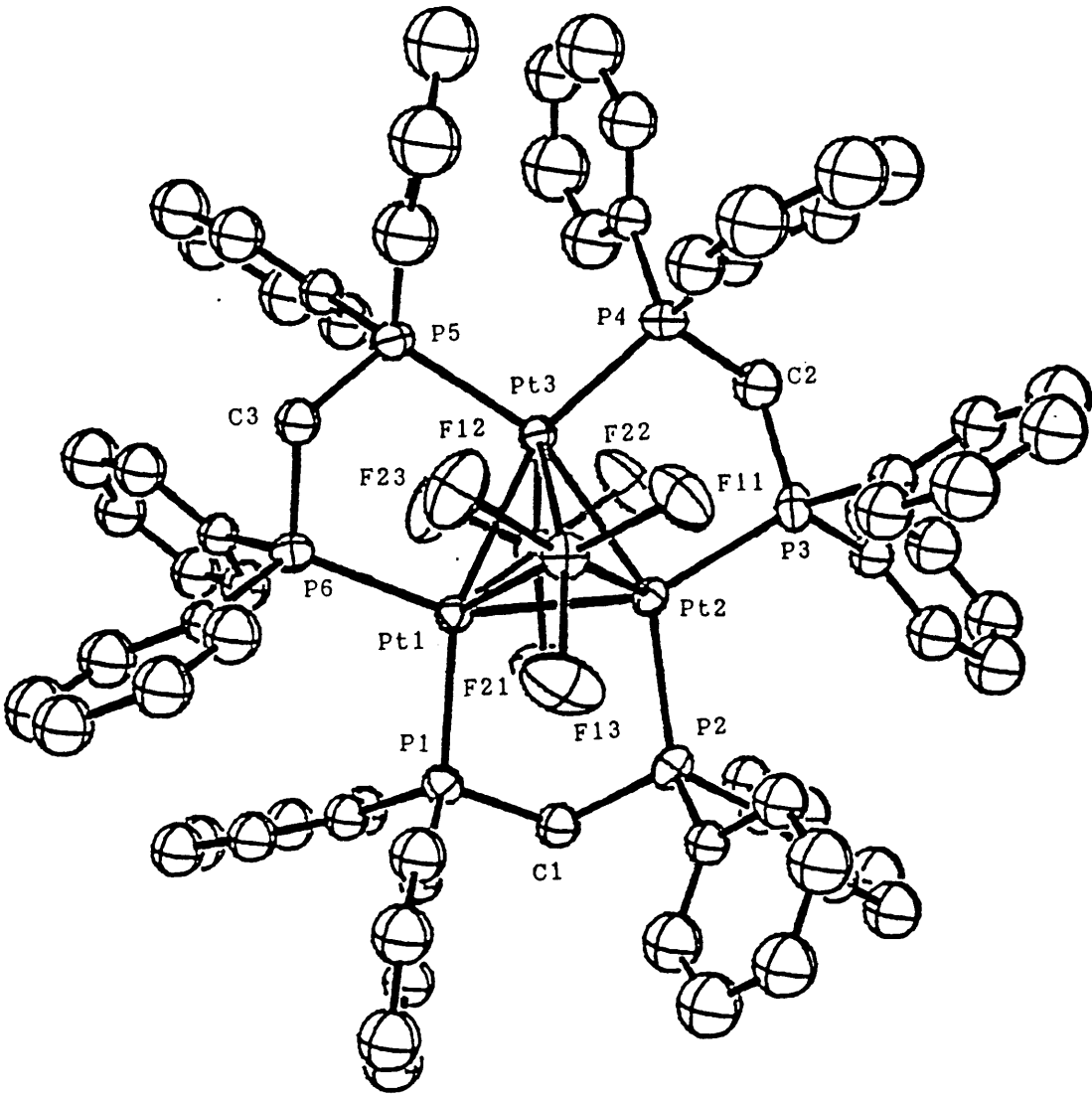
The Pt-Pt bond distances in (10) [2.609(2)-2.639(1) Å with mean distance 2.623(15) Å] are comparable to those found in (1) (see Table 4.1), despite the addition of an extra 2-electron donor ligand. However, comparison of Pt-Sn distances in (7), (8)/(9) and (10) shows that an additional two electrons in the cluster bonding does result in an increase in Pt-Sn bond lengths [mean distances 2.736(32), 2.773(3) and 2.778(65) Å respectively].

Three Sn-F bonds complete the roughly octahedral coordination around each tin atom. Distances range from 1.901(11) to 1.995(12) Å. When viewed along the normal to the Pt_3 plane, it is clear that the two sets of fluorine substituents virtually eclipse each other.

The dppm ligands have conformations such that one face has four equatorial and two axial phenyl rings surrounding it while the other has two equatorial and four axial rings, as is expected where two identical μ_3 -units are bonded to the Pt_3 triangle. The Pt-P distances range from 2.278(5) to 2.313(5) Å and are therefore comparable with other similar compounds (see Table 5.1).

It appears that the $\text{Pt}_3(\mu_3\text{-SnX}_3)$ unit is the most favourable form of attachment for tin(II) complexes to Pt_3

Figure 7(e). The structure of cluster (10). Hydrogen atoms are omitted for clarity. 50% probability ellipsoids are displayed for all atoms shown.

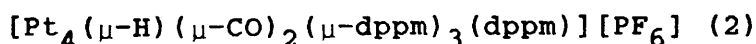


triangles, particularly if X is an electronegative substituent such as F⁻. As noted previously it appears that the SnX₃⁻ unit can stabilise the Pt₃(dppm)₃²⁺ framework in a similar fashion to the carbonyl ligand in (1). On addition of a further ligand to the metal cluster, the SnX₃⁻ unit can accomodate additional electrons in empty π* orbitals, hence the slight lengthening of the Pt-Sn bonds. The more electronegative X is, the less there is need to occupy antibonding orbitals, since X can pull electron density from the cluster bonding.

These model compounds suggest Pt₃(μ₃-SnX₃) units may occur on Pt-Sn/support catalyst surfaces where, for example, X could be the oxygen of the tin(II) aluminate support.⁸³

7.5 Experimental

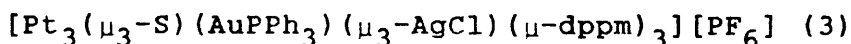
Because fairly similar methods were used in all of the analyses in this section, full details of only one analysis has been given (see the Experimental in Chapter 5). Details of the analyses of structures in this chapter which differ from the model analysis are given below. Summaries of the crystal data of each is given in Table 7.11.



Data Collection. The crystal was mounted on a glass fibre.

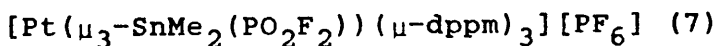
Structure analysis. The heavy atom positions were found using direct methods.⁸⁵ Anisotropic displacement parameters were used for Pt, P and F atoms and isotropic parameters

for all other atoms. The hydridic hydrogen atom was detected in a difference synthesis and its position and isotropic displacement parameters were refined. All the other hydrogen atoms were positioned geometrically (C-H = 0.96 Å).



Data Collection. The crystal was mounted on a glass fibre. In a Type II reduced triclinic cell there are two crystallographically independent ion pairs but the Niggli matrix elements [$A = 311.1$, $B = 347.9$, $C = 802.5$, $D = -44.5$, $E = -77.2$ and $F = -1.9 \text{ Å}^2$] show that no transformation to higher symmetry is possible.⁶⁹ Collection of the $36\text{-}44^\circ$ 2θ shell was discontinued because the reflections were too weak.

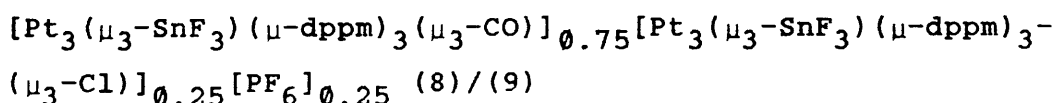
Structure Analysis. Heavy atom positions were determined using direct methods.⁸⁵ All the hydrogen atoms were positioned geometrically (C-H = 0.96 Å) and given a common isotropic displacement parameter $U(\text{H}) = 0.10 \text{ Å}^2$. Anisotropic displacement parameters were used for Au, Ag, Pt, Cl, S and P atoms and isotropic parameters for the F and C atoms. The structure was refined using full-matrix least-squares, adjusting the parameters of cation A and the two PF_6^- anions and then those of cation B in alternate calculations (see 7.6 Program Modifications).



Data Collection. The crystal was mounted on a glass fibre.

A shell of $4-16^\circ$ in 2θ with $h+k = 2n+1$ was collected to check that those reflections were systematically absent. Otherwise, measurements were made only for reflections with $h+k = 2n$.

Structure Analysis. All the hydrogen positions were calculated ($C-H = 0.96 \text{ \AA}$) and given $U(H) = 1.2U(C_{iso}) \text{ \AA}$. There were two distinct PF_6^- sites, each with the phosphorus atom on a two-fold axes and associated with only three independent fluorine atoms. The oxygen and fluorine atoms of the $PO_2F_2^-$ ligand could not be distinguished by X-ray analysis and each was assigned a fluorine scattering factor. The absolute structure was established by refining η , the coefficient of the imaginary component of the anomalous scattering factor.⁸⁶ For the coordinates in Table 7.5 the refined η value was $1.05(3)$. The platinum, tin and phosphorus atoms were given anisotropic displacement parameters, isotropic displacement parameters being used for the remaining atoms.



Data Collection. The crystal was mounted on a glass fibre. An initial shell of data, 2θ range $4-16^\circ$, was collected in order to determine the space group. At higher angle only unique reflections were collected. A Gaussian absorption correction was carried out.⁸⁷

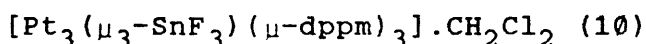
Structure Analysis. The heavy atom positions were found using direct methods.⁸⁵ Anisotropic thermal parameters were used for the Pt, Sn, P(1), P(2) and F(1) atoms, and

isotropic parameters were used for all other atoms.

The cation is positioned around a three fold axis with the tin atom and the carbonyl/ Cl^- both on the axes. The positions of the carbonyl group and the chlorine atom were fixed with the chlorine and oxygen atoms sitting on the same position.

All hydrogen atoms were positioned geometrically, with the phenyl hydrogen positions refined as part of the rigid phenyl groups, while the methylene hydrogens were set to ride on their attached carbon atoms. All hydrogen atoms were given an isotropic displacement parameter $U(\text{H}) = 0.10 \text{ \AA}^2$ which was not refined. The PF_6^- phosphorus atom lay at the intersection of three two fold-axes and each of the three fluorine sites was on a two-fold axis.

For the coordinates given in Table 7.7 the refined η value was 1.09(6).



Data Collection. The crystal was mounted on a glass fibre.

Structure Analysis. A fixed contribution was used for the scattering of geometrically positioned hydrogen atoms, calculated using $\text{C-H} = 0.96 \text{ \AA}$ and $U(\text{H}) = 0.075 \text{ \AA}^2$. Anisotropic displacement parameters were used for all non-hydrogen atoms except C atoms which were assigned isotropic parameters.

There were three molecules of solvent detected in difference syntheses. One molecule was clearly identified as CH_2Cl_2 and refined as such; it is included in Table 7.11. The other two molecules, however, were heavily

disordered. Distances and angles between electron density peaks suggested they were also CH_2Cl_2 molecules but the disorder could not be modelled satisfactorily. Five atomic sites for each molecule were used in the refinement. They were assigned chlorine scattering factors, and given a fixed isotropic displacement parameter of 0.16 \AA^2 and only their positions and occupancies refined [occupancies were $0.51(2)$, $0.54(2)$, $0.38(2)$, $0.36(2)$ and $0.28(2)$ for one molecule and $0.51(2)$, $0.49(2)$, $0.24(2)$, $0.33(2)$ and $0.29(2)$ for the other]. These molecules were not accounted for in Table 7.11.

7.6 Program Modifications

The size of the structures discussed above led to some problems with the GX program package. In particular $[\text{Pt}_3(\text{AuPPh}_3)(\text{AgCl})(\text{S})(\text{dppm})_3][\text{PF}_6]$ with 390 atoms in the asymmetric unit caused difficulty since GX has a limit of 200 atoms per asymmetric unit. Accordingly the following new features were added to the RBLs refinement program.

- (1) Provision for 400 rather than 350 refined parameters.
- (2) Output of a file of fixed structure factor contributions.
- (3) Addition of fixed contributions calculated in (2) to structure factors.
- (4) Addition of fixed contributions and creation of an augmented set of fixed contributions.

Thus for the Pt_3AuAg structure hydrogen contributions were first calculated, then contributions for molecule A and the two PF_6^- anions were added using (4) and finally the

augmented contributions were used in the refinement of molecule B. Although this facility was added to overcome a specific difficulty the ability to use fixed contributions has proved generally useful, especially in dealing with disordered structures where trial full matrix calculations on disordered models can be quickly carried out with the ordered portion of the structure included as a fixed contribution.

Table 7.1. Fractional atomic coordinates and displacement parameters (\AA^2) of the crystal structure of complex (2).

	X/A	Y/B	Z/C	U
PT(1)	.30279(3)	.09876(2)	.33122(2)	.027
PT(2)	.24673(3)	.13707(2)	.40280(2)	.028
PT(3)	.17853(3)	.08408(2)	.34727(2)	.028
PT(4)	.21659(3)	.15875(2)	.30775(2)	.028
P(1)	.3342(2)	.0271(1)	.3605(1)	.035
P(2)	.1926(2)	.0172(1)	.3879(2)	.036
P(3)	.3692(2)	.1278(1)	.2807(1)	.031
P(4)	.2839(2)	.2106(1)	.2785(1)	.036
P(5)	.2608(2)	.1687(1)	.4683(1)	.033
P(6)	.3493(2)	.2396(2)	.5184(2)	.047
P(7)	.0831(2)	.0882(1)	.3150(1)	.035
P(8)	.1194(2)	.1910(1)	.3067(1)	.033
P(9)	.1916(3)	.4053(2)	.4777(3)	.112
F(1)	.1951(10)	.3530(3)	.4838(8)	.266
F(2)	.1457(10)	.3987(10)	.4401(5)	.275
F(3)	.2376(10)	.4119(10)	.5153(5)	.179
F(4)	.1882(10)	.4577(3)	.4716(8)	.252
F(5)	.1358(8)	.4075(3)	.5090(7)	.353
F(6)	.2475(8)	.4031(3)	.4464(7)	.300
O(1)	.3833(5)	.1625(4)	.3870(4)	.052(3)
O(2)	.1046(6)	.1354(4)	.4162(4)	.053(3)
C(1)	.2775(7)	.0105(4)	.3994(5)	.030(4)
C(2)	.3670(7)	.1908(5)	.2836(5)	.036(4)
C(3)	.3141(7)	.2182(5)	.4676(5)	.037(4)
C(4)	.0569(8)	.1486(5)	.3173(5)	.045(5)
C(5)	.3334(8)	.1432(5)	.3800(6)	.046(5)
C(6)	.1543(8)	.1258(5)	.4020(5)	.036(4)
C(A1)	.4080(5)	.0236(7)	.3912(5)	.030(4)
C(A2)	.4511(4)	.0590(5)	.3864(5)	.045(5)
C(A3)	.5070(6)	.0584(4)	.4093(3)	.054(5)
C(A4)	.5199(5)	.0224(7)	.4371(5)	.053(5)
C(A5)	.4768(4)	-.0130(5)	.4419(5)	.057(5)
C(A6)	.4209(6)	-.0124(4)	.4190(3)	.059(5)
C(B1)	.3352(10)	-.0225(5)	.3236(5)	.040(4)
C(B2)	.3413(6)	-.0141(4)	.2800(6)	.054(5)
C(B3)	.3453(7)	-.0506(6)	.2515(3)	.072(6)
C(B4)	.3431(10)	-.0954(5)	.2666(5)	.078(6)
C(B5)	.3370(6)	-.1037(4)	.3102(6)	.082(7)
C(B6)	.3330(7)	-.0673(6)	.3387(3)	.058(5)
C(C1)	.1745(11)	-.0394(5)	.3656(6)	.047(5)
C(C2)	.1702(10)	-.0771(7)	.3926(4)	.078(6)
C(C3)	.1647(4)	-.1209(5)	.3756(6)	.108(8)
C(C4)	.1635(11)	-.1271(5)	.3315(6)	.118(9)
C(C5)	.1678(10)	-.0894(7)	.3044(4)	.094(7)
C(C6)	.1733(4)	-.0456(5)	.3215(6)	.065(6)
C(D1)	.1548(9)	.0176(4)	.4407(5)	.038(4)
C(D2)	.1876(5)	.0109(5)	.4786(7)	.065(6)
C(D3)	.1568(8)	.0148(7)	.5177(5)	.078(6)
C(D4)	.0931(9)	.0254(4)	.5188(5)	.081(7)
C(D5)	.0602(5)	.0320(5)	.4808(7)	.076(6)
C(D6)	.0911(8)	.0282(7)	.4418(5)	.067(6)
C(E1)	.4550(4)	.1147(6)	.2843(6)	.036(4)
C(E2)	.4766(8)	.0738(6)	.2670(2)	.046(5)
C(E3)	.5378(9)	.0594(3)	.2749(6)	.066(6)
C(E4)	.5774(4)	.0859(6)	.3001(6)	.061(5)
C(E5)	.5558(8)	.1268(6)	.3175(2)	.053(5)
C(E6)	.4946(9)	.1413(3)	.3096(6)	.044(4)
C(F1)	.3544(8)	.1174(6)	.2228(3)	.040(4)
C(F2)	.3920(7)	.1400(5)	.1932(3)	.049(5)
C(F3)	.3811(4)	.1344(3)	.1496(3)	.066(6)
C(F4)	.3326(8)	.1062(6)	.1357(3)	.064(6)
C(F5)	.2950(7)	.0836(5)	.1654(3)	.061(5)
C(F6)	.3059(4)	.0892(3)	.2089(3)	.038(4)
C(G1)	.2870(10)	.2672(4)	.3047(6)	.037(4)
C(G2)	.3175(5)	.2729(6)	.3437(7)	.056(5)
C(G3)	.3116(8)	.3139(7)	.3660(4)	.065(6)
C(G4)	.2751(10)	.3492(4)	.3494(6)	.084(7)
C(G5)	.2445(5)	.3435(6)	.3104(7)	.064(5)
C(G6)	.2505(8)	.3025(7)	.2881(4)	.062(5)
C(H1)	.2767(8)	.2245(6)	.2214(4)	.045(4)
C(H2)	.3188(9)	.2541(7)	.2015(4)	.064(5)
C(H3)	.3130(5)	.2634(3)	.1581(3)	.081(7)
C(H4)	.2652(8)	.2431(6)	.1344(4)	.087(7)
C(H5)	.2232(9)	.2134(7)	.1543(4)	.090(7)

Table 7.1(cont).

	X/A	Y/B	Z/C	U
C(H6)	.2289(5)	.2042(3)	.1977(3)	.054(5)
C(I1)	.2971(3)	.1281(6)	.5052(6)	.036(4)
C(I2)	.2907(8)	.1330(3)	.5493(7)	.042(4)
C(I3)	.3198(8)	.1020(5)	.5767(3)	.071(6)
C(I4)	.3553(3)	.0662(6)	.5600(6)	.083(7)
C(I5)	.3617(8)	.0613(3)	.5160(7)	.067(6)
C(I6)	.3326(8)	.0923(5)	.4886(3)	.046(5)
C(J1)	.1902(7)	.1875(4)	.4971(6)	.042(4)
C(J2)	.1480(10)	.1535(3)	.5089(7)	.053(5)
C(J3)	.0933(6)	.1649(4)	.5308(3)	.077(6)
C(J4)	.0808(7)	.2105(4)	.5407(6)	.083(7)
C(J5)	.1231(10)	.2445(3)	.5289(7)	.071(6)
C(J6)	.1778(6)	.2330(4)	.5070(3)	.061(5)
C(K1)	.3888(6)	.2907(3)	.4964(4)	.048(5)
C(K2)	.4435(10)	.2882(4)	.4722(8)	.069(6)
C(K3)	.4698(7)	.3278(4)	.4553(6)	.088(7)
C(K4)	.4415(6)	.3700(3)	.4627(4)	.084(7)
C(K5)	.3868(10)	.3726(4)	.4870(8)	.091(7)
C(K6)	.3605(7)	.3329(4)	.5038(6)	.068(6)
C(L1)	.4166(5)	.2009(6)	.5282(8)	.051(5)
C(L2)	.4428(9)	.1722(7)	.4975(4)	.052(5)
C(L3)	.4947(11)	.1454(4)	.5077(7)	.064(6)
C(L4)	.5204(5)	.1473(6)	.5485(8)	.081(7)
C(L5)	.4942(9)	.1760(7)	.5792(4)	.103(8)
C(L6)	.4423(11)	.2028(4)	.5691(7)	.083(7)
C(M1)	.0764(11)	.0743(7)	.2585(4)	.043(4)
C(M2)	.0173(9)	.0750(4)	.2395(7)	.055(5)
C(M3)	.0111(7)	.0648(7)	.1962(7)	.079(7)
C(M4)	.0638(11)	.0540(7)	.1720(4)	.108(9)
C(M5)	.1229(9)	.0533(4)	.1911(7)	.14(1)
C(M6)	.1291(7)	.0634(7)	.2343(7)	.085(7)
C(N1)	.0139(4)	.0569(4)	.3366(4)	.041(4)
C(N2)	.0136(8)	.0096(4)	.3315(7)	.062(5)
C(N3)	-.0368(9)	-.0161(3)	.3466(6)	.075(6)
C(N4)	-.0869(4)	.0056(4)	.3668(4)	.072(6)
C(N5)	-.0865(8)	.0529(4)	.3719(7)	.069(6)
C(N6)	-.0361(9)	.0786(3)	.3568(6)	.055(5)
C(O1)	.1056(9)	.2323(4)	.3500(5)	.031(4)
C(O2)	.1555(6)	.2495(4)	.3738(6)	.048(5)
C(O3)	.1443(7)	.2809(6)	.4065(3)	.056(5)
C(O4)	.0833(9)	.2950(4)	.4153(5)	.058(5)
C(O5)	.0333(6)	.2777(4)	.3915(6)	.054(5)
C(O6)	.0445(7)	.2464(6)	.3589(3)	.047(5)
C(P1)	.0880(5)	.2209(7)	.2590(6)	.042(4)
C(P2)	.0853(10)	.2685(7)	.2579(4)	.056(5)
C(P3)	.0599(9)	.2906(3)	.2225(7)	.065(6)
C(P4)	.0370(5)	.2651(7)	.1883(6)	.099(8)
C(P5)	.0397(10)	.2175(7)	.1894(4)	.117(9)
C(P6)	.0652(9)	.1954(3)	.2248(7)	.079(7)
H(1)	.245(6)	.076(4)	.301(4)	.02(4)
H(11)	.28623	.02799	.42492	.050
H(12)	.28450	-.02168	.40524	.050
H(21)	.39208	.20367	.26071	.050
H(22)	.38420	.20084	.31074	.050
H(31)	.29111	.24354	.45516	.050
H(32)	.34875	.21036	.44886	.050
H(41)	.04000	.15434	.34558	.050
H(42)	.02383	.15281	.29644	.050
H(A2)	.44211	.08399	.36704	.050
H(A3)	.53698	.08306	.40596	.050
H(A4)	.55882	.02198	.45308	.050
H(A5)	.48571	-.03811	.46124	.050
H(A6)	.39085	-.03714	.42233	.050
H(B2)	.34268	.01700	.26947	.050
H(B3)	.34937	-.04485	.22116	.050
H(B4)	.34588	-.12077	.24679	.050
H(B5)	.33549	-.13489	.32071	.050
H(B6)	.32877	-.07305	.36903	.050
H(C2)	.17106	-.07278	.42333	.050
H(C3)	.16164	-.14712	.39448	.050
H(C4)	.15960	-.15757	.31970	.050
H(C5)	.16698	-.09371	.27378	.050

Table 7.1(cont).

	X/A	Y/B	Z/C	U
H(C6)	.17638	-.01939	.30265	.050
H(D2)	.23183	.00346	.47789	.050
H(D3)	.17948	.01004	.54410	.050
H(D4)	.07151	.02802	.54592	.050
H(D5)	.01588	.03950	.48155	.050
H(D6)	.06826	.03295	.41539	.050
H(E2)	.44899	.05534	.24942	.050
H(E3)	.55275	.03089	.26281	.050
H(E4)	.61990	.07589	.30565	.050
H(E5)	.58331	.14532	.33506	.050
H(E6)	.47961	.16981	.32162	.050
H(F2)	.42574	.15958	.20286	.050
H(F3)	.40722	.15005	.12901	.050
H(F4)	.32488	.10222	.10545	.050
H(F5)	.26111	.06393	.15571	.050
H(F6)	.27961	.07345	.22950	.050
H(G2)	.34279	.24841	.35525	.050
H(G3)	.33275	.31790	.39311	.050
H(G4)	.27093	.37776	.36484	.050
H(G5)	.21919	.36808	.29878	.050
H(G6)	.22926	.29859	.26095	.050
H(H2)	.35208	.26841	.21795	.050
H(H3)	.34235	.28412	.14429	.050
H(H4)	.26128	.24961	.10421	.050
H(H5)	.19004	.19928	.13784	.050
H(H6)	.19975	.18358	.21151	.050
H(I2)	.26595	.15784	.56082	.050
H(I3)	.31515	.10526	.60730	.050
H(I4)	.37532	.04453	.57911	.050
H(I5)	.38640	.03638	.50445	.050
H(I6)	.33714	.08890	.45798	.050
H(J2)	.15665	.12177	.50198	.050
H(J3)	.06391	.14125	.53903	.050
H(J4)	.04279	.21846	.55592	.050
H(J5)	.11444	.27620	.53583	.050
H(J6)	.20719	.25673	.49881	.050
H(K2)	.46327	.25882	.46716	.050
H(K3)	.50789	.32612	.43848	.050
H(K4)	.45986	.39767	.45100	.050
H(K5)	.36712	.40194	.49207	.050
H(K6)	.32247	.33467	.52072	.050
H(L2)	.42494	.17097	.46908	.050
H(L3)	.51303	.12559	.48624	.050
H(L4)	.55657	.12869	.55554	.050
H(L5)	.51208	.17728	.60763	.050
H(L6)	.42397	.22271	.59046	.050
H(M2)	-.01934	.08261	.25623	.050
H(M3)	-.02995	.06534	.18290	.050
H(M4)	.05950	.04686	.14191	.050
H(M5)	.15962	.04569	.17421	.050
H(M6)	.17022	.06288	.24754	.050
H(N2)	.04854	-.00550	.31749	.050
H(N3)	-.03690	-.04907	.34306	.050
H(N4)	-.12181	-.01234	.37731	.050
H(N5)	-.12130	.06795	.38596	.050
H(N6)	-.03585	.11147	.36035	.050
H(O2)	.19795	.23967	.36772	.050
H(O3)	.17909	.29287	.42300	.050
H(O4)	.07558	.31685	.43799	.050
H(O5)	-.00906	.28745	.39762	.050
H(O6)	.00984	.23445	.34233	.050
H(P2)	.10125	.28621	.28174	.050
H(P3)	.05810	.32378	.22176	.050
H(P4)	.01928	.28061	.16372	.050
H(P5)	.02364	.19991	.16566	.050
H(P6)	.06690	.16237	.22560	.050

Table 7.1(cont).

Anisotropic displacement parameters (\AA^2).

	U11	U22	U33	U12	U13	U23
PT(1)	.0270(3)	.0260(3)	.0283(4)	-.0001(3)	.0010(3)	.0015(3)
PT(2)	.0258(4)	.0323(3)	.0248(4)	-.0021(3)	-.0013(3)	-.0012(3)
PT(3)	.0269(4)	.0274(3)	.0304(4)	-.0036(3)	-.0033(3)	.0015(3)
PT(4)	.0308(4)	.0247(3)	.0287(4)	-.0004(3)	-.0019(3)	.0017(3)
P(1)	.035(3)	.029(2)	.041(3)	.004(2)	-.005(2)	.002(2)
P(2)	.031(3)	.033(2)	.045(3)	-.004(2)	-.009(2)	.008(2)
P(3)	.033(3)	.032(2)	.029(3)	-.004(2)	.003(2)	-.003(2)
P(4)	.046(3)	.028(2)	.033(3)	-.002(2)	.003(2)	.002(2)
P(5)	.034(3)	.039(3)	.027(3)	-.007(2)	.001(2)	-.002(2)
P(6)	.043(3)	.058(3)	.040(3)	-.011(2)	.001(3)	-.013(2)
P(7)	.033(2)	.033(2)	.038(3)	-.007(2)	-.008(2)	-.002(2)
P(8)	.037(3)	.037(2)	.026(3)	-.002(2)	-.005(2)	.000(2)
P(9)	.065(4)	.077(5)	.194(9)	.000(4)	-.015(6)	-.059(5)
F(1)	.35(3)	.09(1)	.35(4)	-.02(2)	-.19(3)	-.06(2)
F(2)	.20(2)	.16(2)	.47(4)	.05(1)	-.22(3)	-.10(2)
F(3)	.21(2)	.13(1)	.20(2)	-.01(1)	-.12(2)	-.05(1)
F(4)	.21(2)	.10(1)	.44(4)	.01(1)	-.18(3)	-.05(2)
F(5)	.17(2)	.53(5)	.36(5)	-.02(3)	.16(3)	-.09(4)
F(6)	.16(2)	.56(5)	.18(3)	.10(3)	.03(2)	-.02(3)

Table 7.2. Selected bond distances (\AA) and angles ($^\circ$) of complex (2).

PT(1) - PT(2)	2.750(1)	PT(1) - PT(3)	2.705(1)
PT(1) - PT(4)	2.620(1)	PT(1) - P(1)	2.360(4)
PT(1) - P(3)	2.266(5)	PT(1) - C(5)	2.089(17)
PT(1) - H(1)	1.69(12)	PT(2) - PT(3)	2.720(1)
PT(2) - P(5)	2.250(5)	PT(2) - C(5)	1.971(18)
PT(2) - C(6)	1.980(17)	PT(3) - PT(4)	2.613(1)
PT(3) - P(2)	2.329(5)	PT(3) - P(7)	2.254(5)
PT(3) - C(6)	2.146(16)	PT(3) - H(1)	2.02(13)
PT(4) - P(4)	2.257(5)	PT(4) - P(8)	2.256(5)
P(1) - C(1)	1.766(16)	F(1) - C(A1)	1.830(14)
P(1) - C(B1)	1.836(16)	P(2) - C(1)	1.839(16)
P(2) - C(C1)	1.821(17)	P(2) - C(D1)	1.822(17)
P(3) - C(2)	1.826(15)	P(3) - C(E1)	1.855(10)
P(3) - C(F1)	1.847(11)	P(4) - C(2)	1.853(16)
P(4) - C(G1)	1.832(14)	P(4) - C(H1)	1.826(14)
P(5) - C(3)	1.825(15)	P(5) - C(I1)	1.811(17)
P(5) - C(J1)	1.819(18)	P(6) - C(3)	1.848(17)
P(6) - C(K1)	1.828(12)	P(6) - C(L1)	1.835(16)
P(7) - C(4)	1.837(16)	P(7) - C(M1)	1.803(13)
P(7) - C(N1)	1.845(11)	P(8) - C(4)	1.834(17)
P(8) - C(O1)	1.824(14)	P(8) - C(P1)	1.837(20)
O(1) - C(5)	1.212(21)	O(2) - C(6)	1.170(20)

Table 7.2(cont).

PT(2) - PT(1) - PT(3)	59.8(1)	PT(2) - PT(1) - PT(4)	70.0(1)
PT(2) - PT(1) - P(1)	99.5(2)	PT(2) - PT(1) - P(3)	132.5(2)
PT(2) - PT(1) - C(5)	45.6(5)	PT(2) - PT(1) - H(1)	107.3(41)
PT(3) - PT(1) - PT(4)	58.7(1)	PT(3) - PT(1) - P(1)	93.6(2)
PT(3) - PT(1) - P(3)	141.8(2)	PT(3) - PT(1) - C(5)	105.3(5)
PT(3) - PT(1) - H(1)	48.3(42)	PT(4) - PT(1) - P(1)	152.3(2)
PT(4) - PT(1) - P(3)	89.5(2)	PT(4) - PT(1) - C(5)	90.4(5)
PT(4) - PT(1) - H(1)	66.3(38)	P(1) - PT(1) - P(3)	114.8(2)
P(1) - PT(1) - C(5)	100.2(5)	P(1) - PT(1) - H(1)	94.3(38)
P(3) - PT(1) - C(5)	94.6(5)	P(3) - PT(1) - H(1)	102.0(42)
C(5) - PT(1) - H(1)	150.9(40)	PT(1) - PT(2) - PT(3)	59.3(1)
PT(1) - PT(2) - P(5)	146.9(2)	PT(1) - PT(2) - C(5)	49.2(5)
PT(1) - PT(2) - C(6)	110.3(5)	PT(3) - PT(2) - P(5)	150.7(2)
PT(3) - PT(2) - C(5)	108.4(5)	PT(3) - PT(2) - C(6)	51.4(5)
P(5) - PT(2) - C(5)	99.5(6)	P(5) - PT(2) - C(6)	102.0(5)
C(5) - PT(2) - C(6)	157.8(7)	PT(1) - PT(3) - PT(2)	60.9(1)
PT(1) - PT(3) - PT(4)	59.0(1)	PT(1) - PT(3) - P(2)	96.1(2)
PT(1) - PT(3) - P(7)	141.0(2)	PT(1) - PT(3) - C(6)	106.8(5)
PT(1) - PT(3) - H(1)	38.6(34)	PT(2) - PT(3) - PT(4)	70.6(1)
PT(2) - PT(3) - P(2)	93.4(2)	PT(2) - PT(3) - P(7)	136.5(2)
PT(2) - PT(3) - C(6)	46.2(5)	PT(2) - PT(3) - H(1)	98.9(34)
PT(4) - PT(3) - P(2)	154.5(2)	PT(4) - PT(3) - P(7)	91.3(2)
PT(4) - PT(3) - C(6)	88.8(4)	PT(4) - PT(3) - H(1)	63.2(31)
P(2) - PT(3) - P(7)	113.4(2)	P(2) - PT(3) - C(6)	94.1(5)
P(2) - PT(3) - H(1)	101.6(32)	P(7) - PT(3) - C(6)	96.2(5)
P(7) - PT(3) - H(1)	107.8(35)	C(6) - PT(3) - H(1)	142.8(33)
PT(1) - PT(4) - PT(3)	62.3(1)	PT(1) - PT(4) - P(4)	96.6(2)
PT(1) - PT(4) - P(8)	155.7(2)	PT(3) - PT(4) - P(4)	158.8(2)
PT(3) - PT(4) - P(8)	94.0(2)	P(4) - PT(4) - P(8)	107.0(2)
PT(1) - P(1) - C(1)	108.1(5)	PT(1) - P(1) - C(A1)	119.2(7)
PT(1) - P(1) - C(B1)	116.8(6)	C(1) - P(1) - C(A1)	101.9(7)
C(1) - P(1) - C(B1)	102.8(8)	C(A1) - P(1) - C(B1)	105.8(9)
PT(3) - P(2) - C(1)	108.5(5)	PT(3) - P(2) - C(C1)	121.1(7)
PT(3) - P(2) - C(D1)	115.2(5)	C(1) - P(2) - C(C1)	100.5(9)
C(1) - P(2) - C(D1)	104.7(8)	C(C1) - P(2) - C(D1)	104.8(8)

Table 7.3. Fractional atomic positions and displacement parameters (\AA^2) of the crystal structure of compound (3).

Cation A				
	X/A	Y/B	Z/C	U
AU	.61167(9)	.13628(8)	.27931(7)	.046
PT(1)	.94537(8)	.18177(7)	.26225(6)	.033
PT(2)	.87511(8)	.18892(7)	.17450(6)	.033
PT(3)	.73610(8)	.18777(7)	.25710(6)	.034
AG	.82940(16)	.07792(14)	.22862(13)	.051
CL	.8534(7)	-.0494(5)	.2267(5)	.100
S	.8438(5)	.2528(4)	.2398(4)	.041
P(1)	1.0445(5)	.1160(5)	.2433(4)	.036
P(2)	.9576(5)	.1166(5)	.1418(4)	.040
P(3)	.7799(5)	.2242(5)	.1175(4)	.039
P(4)	.6570(5)	.2013(5)	.1868(4)	.040
P(5)	.7757(5)	.1789(5)	.3374(4)	.039
P(6)	.9569(5)	.2004(5)	.3420(4)	.040
P(7)	.5019(6)	.0998(5)	.3048(4)	.051
C(1)	1.0134(19)	.0638(18)	.1859(15)	.051(11)
C(2)	.6947(18)	.2571(16)	.1447(14)	.040(10)
C(3)	.8651(18)	.2280(16)	.3626(14)	.039(10)
C(A1)	1.0965(12)	.0499(23)	.2822(17)	.055(12)
C(A2)	1.1391(25)	.0776(13)	.3249(16)	.061(12)
C(A3)	1.1818(25)	.0324(26)	.3543(8)	.082(15)
C(A4)	1.1818(12)	-.0405(23)	.3411(17)	.062(12)
C(A5)	1.1393(25)	-.0682(13)	.2984(16)	.082(14)
C(A6)	1.0966(25)	-.0230(26)	.2689(8)	.098(16)
C(B1)	1.1234(16)	.1773(17)	.2364(8)	.033(9)
C(B2)	1.1188(18)	.2499(15)	.2497(17)	.043(10)
C(B3)	1.1803(24)	.2953(11)	.2483(15)	.067(13)
C(B4)	1.2464(16)	.2682(17)	.2337(8)	.071(13)
C(B5)	1.2509(18)	.1956(15)	.2205(17)	.095(16)
C(B6)	1.1894(24)	.1502(11)	.2218(15)	.067(13)
C(C1)	1.0319(14)	.1663(16)	.1221(8)	.034(9)
C(C2)	1.0307(20)	.2402(14)	.1302(18)	.058(12)
C(C3)	1.0896(24)	.2818(11)	.1194(16)	.093(16)
C(C4)	1.1498(14)	.2495(16)	.1006(8)	.087(15)
C(C5)	1.1511(20)	.1755(14)	.0925(18)	.083(15)
C(C6)	1.0921(24)	.1339(11)	.1032(16)	.080(14)
C(D1)	.9220(24)	.0508(20)	.0934(10)	.039(10)
C(D2)	.8801(18)	-.0065(17)	.1044(8)	.050(11)
C(D3)	.8491(14)	-.0580(11)	.0687(9)	.077(14)
C(D4)	.8598(24)	-.0522(20)	.0220(10)	.078(14)
C(D5)	.9017(18)	.0052(17)	.0109(8)	.075(14)
C(D6)	.9327(14)	.0567(11)	.0466(9)	.046(11)
C(E1)	.8063(13)	.3034(10)	.0891(10)	.046(11)
C(E2)	.7693(17)	.3683(13)	.0912(15)	.054(11)
C(E3)	.7939(21)	.4251(14)	.0688(18)	.064(12)
C(E4)	.8555(13)	.4170(10)	.0443(10)	.082(15)
C(E5)	.8925(17)	.3521(13)	.0422(15)	.074(14)
C(E6)	.8680(21)	.2953(14)	.0646(18)	.066(13)
C(F1)	.7364(27)	.1662(27)	.0649(12)	.066(13)
C(F2)	.7318(15)	.0927(27)	.0672(15)	.065(13)
C(F3)	.7000(22)	.0488(12)	.0273(19)	.088(15)
C(F4)	.6729(27)	.0784(27)	-.0149(12)	.104(17)
C(F5)	.6775(15)	.1519(27)	-.0173(15)	.140(22)
C(F6)	.7092(22)	.1958(12)	.0226(19)	.090(15)
C(G1)	.6160(25)	.1206(18)	.1500(11)	.036(16)
C(G2)	.6376(17)	.0530(24)	.1628(12)	.033(9)
C(G3)	.6102(17)	-.0073(15)	.1334(16)	.058(12)
C(G4)	.5612(25)	-.0001(18)	.0912(11)	.078(14)
C(G5)	.5396(17)	.0675(24)	.0784(12)	.089(15)
C(G6)	.5670(17)	.1278(15)	.1078(16)	.069(13)
C(H1)	.5791(22)	.2606(14)	.2002(18)	.035(10)
C(H2)	.5020(27)	.2473(19)	.1843(11)	.070(13)
C(H3)	.4485(16)	.2956(24)	.1976(13)	.072(13)
C(H4)	.4722(22)	.3572(14)	.2267(18)	.093(16)
C(H5)	.5493(27)	.3705(19)	.2426(11)	.071(13)
C(H6)	.6028(16)	.3222(24)	.2293(13)	.059(12)
C(I1)	.7134(21)	.2343(12)	.3712(17)	.050(11)
C(I2)	.7027(12)	.3042(9)	.3594(8)	.046(11)
C(I3)	.6511(18)	.3471(13)	.3802(15)	.074(14)
C(I4)	.6103(21)	.3199(12)	.4127(17)	.074(14)
C(I5)	.6210(12)	.2500(9)	.4245(8)	.068(13)
C(I6)	.6725(18)	.2072(13)	.4037(15)	.060(12)
C(J1)	.7856(14)	.0951(14)	.3637(15)	.036(10)

Table 7.3(cont).

	X/A	Y/B	Z/C	U
C(J2)	.8115(23)	.0916(17)	.4120(13)	.047(11)
C(J3)	.8211(18)	.0256(22)	.4306(10)	.073(13)
C(J4)	.8048(14)	-.0369(14)	.4009(15)	.077(14)
C(J5)	.7789(23)	-.0333(17)	.3526(13)	.062(12)
C(J6)	.7693(18)	.0326(22)	.3341(10)	.059(12)
C(K1)	.9859(11)	.1254(16)	.3795(14)	.029(9)
C(K2)	.9745(19)	.0558(21)	.3588(8)	.036(10)
C(K3)	.9942(22)	-.0014(13)	.3858(13)	.056(12)
C(K4)	1.0254(11)	.0110(16)	.4335(14)	.051(11)
C(K5)	1.0368(19)	.0806(21)	.4543(8)	.083(15)
C(K6)	1.0171(22)	.1378(13)	.4273(13)	.052(11)
C(L1)	1.0227(24)	.2781(19)	.3610(9)	.040(10)
C(L2)	1.0948(18)	.2736(17)	.3874(17)	.058(12)
C(L3)	1.1399(17)	.3350(26)	.4005(15)	.082(14)
C(L4)	1.1128(24)	.4007(19)	.3872(9)	.079(14)
C(L5)	1.0407(18)	.4051(17)	.3607(17)	.092(16)
C(L6)	.9956(17)	.3438(26)	.3476(15)	.074(14)
C(M11)	.4376(25)	.0435(23)	.2605(9)	.046(11)
C(M12)	.3857(21)	-.0042(19)	.2732(8)	.070(13)
C(M13)	.3351(13)	-.0430(12)	.2381(8)	.074(14)
C(M14)	.3363(25)	-.0340(23)	.1904(9)	.086(15)
C(M15)	.3882(21)	.0137(19)	.1777(8)	.088(15)
C(M16)	.4388(13)	.0525(12)	.2128(8)	.077(14)
C(M21)	.4436(25)	.1746(14)	.3241(20)	.061(12)
C(M22)	.4673(12)	.2452(18)	.3241(14)	.066(13)
C(M23)	.4210(23)	.3004(14)	.3374(13)	.066(13)
C(M24)	.3511(25)	.2848(14)	.3507(20)	.077(14)
C(M25)	.3274(12)	.2142(18)	.3507(14)	.108(18)
C(M26)	.3737(23)	.1590(14)	.3374(13)	.093(16)
C(M31)	.5265(12)	.0510(21)	.3582(14)	.047(11)
C(M32)	.5098(19)	.0756(14)	.4027(19)	.063(12)
C(M33)	.5346(23)	.0383(16)	.4421(12)	.071(13)
C(M34)	.5761(12)	-.0236(21)	.4370(14)	.093(16)
C(M35)	.5929(19)	-.0481(14)	.3925(19)	.075(14)
C(M36)	.5680(23)	-.0109(16)	.3531(12)	.053(11)
H(11)	.98237	.02366	.19121	.100
H(12)	1.05781	.04617	.17328	.100
H(21)	.65380	.26269	.11876	.100
H(22)	.70855	.30325	.16143	.100
H(31)	.87215	.22493	.39665	.100
H(32)	.85752	.27737	.35610	.100
H(A2)	1.13911	.12830	.33420	.100
H(A3)	1.21135	.05159	.38409	.100
H(A4)	1.21143	-.07201	.36157	.100
H(A5)	1.13931	-.11890	.28911	.100
H(A6)	1.06707	-.04219	.23920	.100
H(B2)	1.07279	.26869	.25975	.100
H(B3)	1.17706	.34582	.25746	.100
H(B4)	1.28912	.29987	.23281	.100
H(B5)	1.29690	.17682	.21035	.100
H(B6)	1.19263	.09972	.21258	.100
H(C2)	.98886	.26269	.14336	.100
H(C3)	1.08875	.33327	.12509	.100
H(C4)	1.19079	.27842	.09309	.100
H(C5)	1.19286	.15304	.07928	.100
H(C6)	1.09299	.08246	.09762	.100
H(D2)	.87252	-.01052	.13695	.100
H(D3)	.81994	-.09785	.07642	.100
H(D4)	.83828	-.08799	-.00286	.100
H(D5)	.90929	.00919	-.02159	.100
H(D6)	.96186	.09654	.03895	.100
H(E2)	.72637	.37383	.10824	.100
H(E3)	.76797	.47016	.07030	.100
H(E4)	.87246	.45651	.02870	.100
H(E5)	.93541	.34655	.02511	.100
H(E6)	.89382	.25023	.06306	.100
H(F2)	.75058	.07211	.09660	.100
H(F3)	.69666	-.00226	.02896	.100
H(F4)	.65074	.04794	-.04272	.100
H(F5)	.65877	.17250	-.04674	.100
H(F6)	.71259	.24688	.02094	.100
H(G2)	.67156	.04797	.19224	.100
H(G3)	.62524	-.05429	.14233	.100
H(G4)	.54228	-.04202	.07072	.100
H(G5)	.50570	.07254	.04896	.100
H(G6)	.55205	.17482	.09887	.100

Table 7.3(cont).

	X/A	Y/B	Z/C	U
H(H2)	.48552	.20446	.16397	.100
H(H3)	.39483	.28630	.18652	.100
H(H4)	.43496	.39076	.23595	.100
H(H5)	.56577	.41332	.26290	.100
H(H6)	.65640	.33145	.24040	.100
H(I2)	.73096	.32310	.33665	.100
H(I3)	.64361	.39568	.37195	.100
H(I4)	.57450	.34969	.42718	.100
H(I5)	.59270	.23115	.44717	.100
H(I6)	.68005	.15857	.41189	.100
H(J2)	.82288	.13507	.43268	.100
H(J3)	.83909	.02317	.46418	.100
H(J4)	.81135	-.08279	.41384	.100
H(J5)	.76750	-.07683	.33201	.100
H(J6)	.75132	.03507	.30050	.100
H(K2)	.95282	.04724	.32554	.100
H(K3)	.98634	-.04977	.37132	.100
H(K4)	1.03924	-.02875	.45234	.100
H(K5)	1.05855	.08929	.48759	.100
H(K6)	1.02501	.18628	.44183	.100
H(L2)	1.11365	.22789	.39672	.100
H(L3)	1.19011	.33108	.41884	.100
H(L4)	1.14411	.44338	.39626	.100
H(L5)	1.02183	.45085	.35138	.100
H(L6)	.94540	.34688	.32923	.100
H(M12)	.38485	-.01052	.30639	.100
H(M13)	.29896	-.07617	.24695	.100
H(M14)	.30094	-.06089	.16599	.100
H(M15)	.38893	.02001	.14450	.100
H(M16)	.47489	.08564	.20393	.100
H(M22)	.51586	.25602	.31488	.100
H(M23)	.43743	.34950	.33747	.100
H(M24)	.31886	.32314	.36004	.100
H(M25)	.27874	.20333	.35998	.100
H(M26)	.35715	.10986	.33732	.100
H(M32)	.48097	.11866	.40624	.100
H(M33)	.52307	.05544	.47306	.100
H(M34)	.59340	-.04959	.46440	.100
H(M35)	.62172	-.09127	.38896	.100
H(M36)	.57965	-.02803	.32217	.100

Cation B

	X/A	Y/B	Z/C	U
AU	.47691 (9)	.37094 (8)	.78804 (7)	.042
PT (1)	.80396 (8)	.30868 (7)	.75340 (6)	.034
PT (2)	.71765 (8)	.28921 (7)	.66925 (6)	.033
PT (3)	.59501 (8)	.30886 (7)	.76117 (6)	.032
AG	.68732 (16)	.41019 (14)	.73040 (13)	.050
CL	.6944 (8)	.5373 (5)	.7260 (5)	.093
S	.6962 (5)	.2375 (4)	.7360 (4)	.037
P (1)	.8971 (5)	.3734 (5)	.7298 (4)	.037
P (2)	.7958 (5)	.3542 (5)	.6322 (4)	.035
P (3)	.6131 (5)	.2436 (4)	.6164 (4)	.034
P (4)	.5070 (5)	.2763 (4)	.6953 (4)	.036
P (5)	.6446 (5)	.3340 (5)	.8391 (4)	.038
P (6)	.8248 (5)	.3085 (5)	.8356 (4)	.039
P (7)	.3707 (6)	.4150 (5)	.8152 (4)	.051
C (1)	.8593 (19)	.4129 (17)	.6756 (15)	.047 (11)
C (2)	.5363 (17)	.2155 (15)	.6482 (13)	.029 (9)
C (3)	.7362 (17)	.2898 (15)	.8593 (13)	.028 (9)
C (A1)	.9487 (24)	.4493 (21)	.7671 (8)	.045 (11)
C (A2)	.9387 (15)	.5206 (24)	.7585 (13)	.056 (12)
C (A3)	.9825 (18)	.5736 (12)	.7875 (15)	.067 (13)
C (A4)	1.0362 (24)	.5552 (21)	.8251 (8)	.068 (13)
C (A5)	1.0463 (15)	.4839 (24)	.8337 (13)	.065 (13)
C (A6)	1.0025 (18)	.4309 (12)	.8047 (15)	.051 (11)
C (B1)	.9739 (23)	.3207 (18)	.7134 (21)	.047 (11)
C (B2)	.9707 (13)	.2467 (18)	.7131 (11)	.064 (13)
C (B3)	1.0300 (25)	.2061 (12)	.6993 (17)	.092 (16)
C (B4)	1.0925 (23)	.2395 (18)	.6860 (21)	.105 (17)
C (B5)	1.0956 (13)	.3135 (18)	.6863 (11)	.117 (19)
C (B6)	1.0363 (25)	.3541 (12)	.7001 (17)	.097 (16)
C (C1)	.8623 (27)	.2996 (14)	.6041 (20)	.050 (11)
C (C2)	.9169 (26)	.3285 (13)	.5808 (17)	.087 (15)

Table 7.3(cont).

	X/A	Y/B	Z/C	U
C(C3)	.9641(12)	.2837(15)	.5576(11)	.120(19)
C(C4)	.9568(27)	.2100(14)	.5579(20)	.138(22)
C(C5)	.9022(26)	.1811(13)	.5812(17)	.121(20)
C(C6)	.8550(12)	.2259(15)	.6043(11)	.065(13)
C(D1)	.7520(10)	.4148(20)	.5879(16)	.041(10)
C(D2)	.7524(22)	.3990(10)	.5395(18)	.059(12)
C(D3)	.7228(22)	.4471(19)	.5074(10)	.075(14)
C(D4)	.6927(10)	.5110(20)	.5236(16)	.080(14)
C(D5)	.6923(22)	.5268(10)	.5720(18)	.073(13)
C(D6)	.7219(22)	.4787(19)	.6042(10)	.073(14)
C(E1)	.6393(23)	.1617(10)	.5819(16)	.040(10)
C(E2)	.6263(15)	.0934(10)	.5946(10)	.057(12)
C(E3)	.6558(16)	.0346(10)	.5712(12)	.054(11)
C(E4)	.6982(23)	.0441(10)	.5351(16)	.055(12)
C(E5)	.7112(15)	.1124(10)	.5223(10)	.085(15)
C(E6)	.6817(16)	.1712(10)	.5458(12)	.065(13)
C(F1)	.5592(12)	.2894(23)	.5694(14)	.039(10)
C(F2)	.5665(24)	.3631(24)	.5694(7)	.038(10)
C(F3)	.5245(21)	.3983(10)	.5334(14)	.074(14)
C(F4)	.4753(12)	.3598(23)	.4974(14)	.101(17)
C(F5)	.4680(24)	.2861(24)	.4974(7)	.103(17)
C(F6)	.5099(21)	.2509(10)	.5334(14)	.046(11)
C(G1)	.4600(12)	.3444(19)	.6605(15)	.031(9)
C(G2)	.4859(20)	.4147(24)	.6713(10)	.070(13)
C(G3)	.4535(23)	.4681(14)	.6444(11)	.058(12)
C(G4)	.3950(12)	.4513(19)	.6066(15)	.085(15)
C(G5)	.3691(20)	.3809(24)	.5958(10)	.068(13)
C(G6)	.4016(23)	.3275(14)	.6227(11)	.053(11)
C(H1)	.4321(12)	.2210(11)	.7148(7)	.043(10)
C(H2)	.3551(14)	.2366(17)	.7111(15)	.064(13)
C(H3)	.3065(12)	.1938(20)	.7309(16)	.074(14)
C(H4)	.3348(12)	.1354(11)	.7544(7)	.066(13)
C(H5)	.4118(14)	.1198(17)	.7580(15)	.068(13)
C(H6)	.4604(12)	.1626(20)	.7382(16)	.056(12)
C(I1)	.5858(10)	.2843(11)	.8759(7)	.063(12)
C(I2)	.5555(24)	.3140(15)	.9151(13)	.060(12)
C(I3)	.5092(23)	.2726(13)	.9376(14)	.086(15)
C(I4)	.4933(10)	.2016(11)	.9210(7)	.064(12)
C(I5)	.5236(24)	.1719(15)	.8819(13)	.063(12)
C(I6)	.5698(23)	.2133(13)	.8594(14)	.062(12)
C(J1)	.6592(23)	.4229(14)	.8704(13)	.039(10)
C(J2)	.6306(20)	.4816(22)	.8474(9)	.056(12)
C(J3)	.6408(12)	.5498(17)	.8707(15)	.064(13)
C(J4)	.6795(23)	.5592(14)	.9171(13)	.068(13)
C(J5)	.7080(20)	.5005(22)	.9401(9)	.070(13)
C(J6)	.6978(12)	.4323(17)	.9168(15)	.033(9)
C(K1)	.8621(24)	.3945(15)	.8674(11)	.033(9)
C(K2)	.9135(14)	.3983(16)	.9095(15)	.045(11)
C(K3)	.9356(18)	.4644(23)	.9336(10)	.066(13)
C(K4)	.9062(24)	.5266(15)	.9154(11)	.062(12)
C(K5)	.8548(14)	.5228(16)	.8733(15)	.066(13)
C(K6)	.8327(18)	.4567(23)	.8492(10)	.043(10)
C(L1)	.8811(15)	.2349(20)	.8610(20)	.041(10)
C(L2)	.8698(23)	.2106(11)	.9041(19)	.086(15)
C(L3)	.9123(18)	.1543(18)	.9221(9)	.109(18)
C(L4)	.9661(15)	.1224(20)	.8970(20)	.082(15)
C(L5)	.9773(23)	.1467(11)	.8539(19)	.091(16)
C(L6)	.9348(18)	.2029(18)	.8359(9)	.084(15)
C(M11)	.3083(13)	.4699(22)	.7768(18)	.045(11)
C(M12)	.2562(24)	.5154(20)	.7948(10)	.070(13)
C(M13)	.2060(21)	.5530(11)	.7642(20)	.066(13)
C(M14)	.2078(13)	.5451(22)	.7154(18)	.081(14)
C(M15)	.2598(24)	.4996(20)	.6974(10)	.104(17)
C(M16)	.3101(21)	.4620(11)	.7281(20)	.070(13)
C(M21)	.3063(17)	.3412(14)	.8260(7)	.051(11)
C(M22)	.2283(14)	.3416(19)	.8102(16)	.061(12)
C(M23)	.1820(15)	.2860(24)	.8192(15)	.079(14)
C(M24)	.2138(17)	.2301(14)	.8439(7)	.073(13)
C(M25)	.2918(14)	.2297(19)	.8597(16)	.069(13)
C(M26)	.3380(15)	.2853(24)	.8507(15)	.078(14)
C(M31)	.3912(26)	.4717(23)	.8710(11)	.037(10)
C(M32)	.3781(18)	.4503(14)	.9145(12)	.062(12)
C(M33)	.4020(16)	.4938(18)	.9561(9)	.088(15)
C(M34)	.4389(26)	.5587(23)	.9541(11)	.094(16)
C(M35)	.4520(18)	.5800(14)	.9105(12)	.092(16)
C(M36)	.4281(16)	.5365(18)	.8690(9)	.080(14)

Table 7.3(cont).

	X/A	Y/R	Z/C	U
H(11)	.90170	.42723	.66080	.100
H(12)	.83139	.45483	.68414	.100
H(21)	.55268	.17215	.66242	.100
H(22)	.49174	.20502	.62458	.100
H(31)	.72640	.23903	.85300	.100
H(32)	.74756	.30138	.89345	.100
H(A2)	.90149	.53322	.73229	.100
H(A3)	.97550	.62313	.78146	.100
H(A4)	1.06656	.59211	.84531	.100
H(A5)	1.08369	.47119	.85994	.100
H(A6)	1.00960	.38131	.81082	.100
H(B2)	.92757	.22315	.72237	.100
H(B3)	1.02821	.15459	.69891	.100
H(B4)	1.13383	.21152	.67620	.100
H(B5)	1.13881	.33700	.67691	.100
H(B6)	1.03813	.40558	.70034	.100
H(C2)	.92217	.37962	.58065	.100
H(C3)	1.00210	.30356	.54141	.100
H(C4)	.98951	.17865	.54184	.100
H(C5)	.89695	.12982	.58148	.100
H(C6)	.81693	.20590	.62062	.100
H(D2)	.77322	.35446	.52832	.100
H(D3)	.72294	.43594	.47375	.100
H(D4)	.67196	.54432	.50126	.100
H(D5)	.67130	.57127	.58337	.100
H(D6)	.72159	.48980	.63798	.100
H(E2)	.59680	.08683	.61992	.100
H(E3)	.64673	-.01297	.58025	.100
H(E4)	.71865	.00302	.51897	.100
H(E5)	.74052	.11879	.49727	.100
H(E6)	.69052	.21858	.53693	.100
H(F2)	.60058	.38998	.59432	.100
H(F3)	.52947	.44951	.53319	.100
H(F4)	.44613	.38410	.47219	.100
H(F5)	.43384	.25913	.47237	.100
H(F6)	.50491	.19959	.53352	.100
H(G2)	.52668	.42623	.69756	.100
H(G3)	.47171	.51695	.65193	.100
H(G4)	.37262	.48853	.58786	.100
H(G5)	.32846	.36938	.56951	.100
H(G6)	.38347	.27865	.61514	.100
H(H2)	.33566	.27729	.69479	.100
H(H3)	.25296	.20479	.72828	.100
H(H4)	.30077	.10567	.76812	.100
H(H5)	.43124	.07905	.77442	.100
H(H6)	.51391	.15152	.74091	.100
H(I2)	.56673	.36317	.92662	.100
H(I3)	.48839	.29300	.96485	.100
H(I4)	.46131	.17261	.93666	.100
H(I5)	.51850	.12244	.87018	.100
H(I6)	.59083	.19260	.83194	.100
H(J2)	.60357	.47500	.81511	.100
H(J3)	.62066	.59060	.85462	.100
H(J4)	.68634	.60670	.93321	.100
H(J5)	.73489	.50717	.97230	.100
H(J6)	.71784	.39157	.93275	.100
H(K2)	.93403	.35512	.92236	.100
H(K3)	.97123	.46727	.96302	.100
H(K4)	.92132	.57265	.93214	.100
H(K5)	.83423	.56592	.86061	.100
H(K6)	.79707	.45381	.81994	.100
H(L2)	.83214	.23229	.92139	.100
H(L3)	.90440	.13700	.95193	.100
H(L4)	.99584	.08315	.90944	.100
H(L5)	1.01509	.12458	.83641	.100
H(L6)	.94284	.21988	.80587	.100
H(M12)	.25475	.52082	.82851	.100
H(M13)	.16968	.58461	.77644	.100
H(M14)	.17296	.57122	.69388	.100
H(M15)	.26129	.49403	.66344	.100
H(M16)	.34632	.43024	.71553	.100
H(M22)	.20613	.38048	.79279	.100
H(M23)	.12764	.28636	.80820	.100
H(M24)	.18147	.19156	.85031	.100

Table 7.3(cont).

	X/A	Y/B	Z/C	U
H(M25)	.31376	.19089	.87700	.100
H(M26)	.39222	.28504	.86165	.100
H(M32)	.35192	.40537	.91592	.100
H(M33)	.39265	.47921	.98633	.100
H(M34)	.45569	.58892	.98282	.100
H(M35)	.47803	.62478	.90889	.100
H(M36)	.43735	.55091	.83848	.100

[PF₆]⁻ Anions

	X/A	Y/B	Z/C	U
P(F1)	.3570(6)	.2072(6)	.0294(5)	.109
F(11)	.4403(15)	.2276(24)	.0503(12)	.220(18)
F(12)	.2737(15)	.1868(24)	.0085(12)	.288(25)
F(13)	.3586(8)	.2607(15)	-.0085(11)	.198(17)
F(14)	.3555(8)	.1538(15)	.0674(11)	.243(21)
F(15)	.3293(26)	.2664(10)	.0631(16)	.199(16)
F(16)	.3848(26)	.1480(10)	-.0042(16)	.327(30)
P(F2)	.2530(8)	.1933(6)	.5031(6)	.161
F(21)	.2762(16)	.2703(8)	.4974(11)	.176(14)
F(22)	.2277(16)	.1163(8)	.5087(11)	.160(13)
F(23)	.1729(13)	.2198(14)	.5094(18)	.178(15)
F(24)	.3330(13)	.1668(14)	.4967(18)	.239(21)
F(25)	.2788(23)	.2055(20)	.5573(7)	.296(28)
F(26)	.2271(23)	.1811(20)	.4489(7)	.195(17)

Anisotropic displacement parameters (Å²).

Cation A

	U11	U22	U33	U12	U13	U23
AU	.0392(10)	.0501(10)	.0468(14)	-.0053(8)	.0088(9)	.0103(9)
PT(1)	.0302(9)	.0355(9)	.0330(12)	.0014(7)	.0032(8)	.0048(8)
PT(2)	.0339(9)	.0341(9)	.0304(12)	.0048(7)	.0060(8)	.0042(8)
PT(3)	.0342(9)	.0350(9)	.0334(12)	.0036(7)	.0073(8)	.0042(8)
AG	.0474(20)	.0359(17)	.0683(29)	.0006(15)	.0134(19)	.0042(18)
CL	.136(11)	.050(7)	.118(14)	.002(7)	.075(10)	.014(7)
S	.032(6)	.031(5)	.059(9)	.007(5)	.011(6)	.008(5)
P(1)	.030(6)	.046(6)	.034(8)	.005(5)	.013(6)	.018(6)
P(2)	.047(7)	.041(6)	.032(8)	.001(5)	.010(6)	.002(6)
P(3)	.045(6)	.041(6)	.032(8)	.007(5)	.019(6)	.009(5)
P(4)	.025(6)	.048(6)	.045(9)	-.004(5)	-.002(6)	.006(6)
P(5)	.045(7)	.039(6)	.033(8)	.013(5)	.005(6)	.005(6)
P(6)	.035(6)	.049(6)	.035(8)	-.001(5)	.006(6)	.004(6)
P(7)	.047(7)	.038(6)	.067(10)	-.008(5)	.008(7)	.012(6)

Cation B

	U11	U22	U33	U12	U13	U23
AU	.0386(10)	.0457(10)	.0396(13)	.0075(8)	.0065(9)	-.0008(9)
PT(1)	.0322(9)	.0321(9)	.0365(13)	.0008(7)	.0019(9)	.0017(8)
PT(2)	.0352(9)	.0322(9)	.0308(12)	-.0013(7)	.0047(8)	.0008(8)
PT(3)	.0340(9)	.0324(9)	.0287(12)	.0014(7)	.0053(8)	.0033(8)
AG	.0518(20)	.0327(17)	.0632(28)	.0088(15)	.0096(19)	.0035(17)
CL	.150(12)	.042(7)	.089(12)	.013(7)	.049(10)	.004(7)
S	.038(6)	.043(6)	.030(7)	-.005(5)	.009(5)	.010(5)
P(1)	.024(6)	.048(6)	.034(8)	-.002(5)	-.006(6)	-.006(6)
P(2)	.033(6)	.042(6)	.027(8)	-.015(5)	-.001(6)	-.008(5)
P(3)	.037(6)	.031(6)	.030(8)	.006(5)	-.006(6)	.004(5)
P(4)	.030(6)	.032(6)	.046(8)	.002(5)	.009(6)	.018(6)
P(5)	.048(7)	.030(6)	.037(8)	.002(5)	.008(6)	.006(5)
P(6)	.028(6)	.055(7)	.032(8)	.008(5)	-.001(6)	.005(6)
P(7)	.049(7)	.053(7)	.051(9)	.013(6)	.011(7)	.008(6)

Table 7.3(cont).

[PF₆]⁻ Anions

	U11	U22	U33	U12	U13	U23
P(F1)	.069(10)	.108(11)	.154(18)	.028(8)	.054(11)	.018(11)
P(F2)	.201(21)	.086(11)	.181(26)	-.061(13)	-.054(17)	.052(13)

Table 7.4. Selected bond distances (Å) and angles (°) of cations A and B of complex (3).

	Cation A	Cation B		Cation A	Cation B
Au-Pt(3)	2.577(3)	2.575(3)	Au-P(7)	2.287(11)	2.265(11)
Pt(1)-Pt(2)	2.615(3)	2.619(3)	Pt(1)-P(1)	2.249(10)	2.249(10)
Pt(1)-Ag	2.808(4)	2.843(4)	Pt(2)-P(2)	2.241(10)	2.261(10)
Pt(2)-Ag	2.861(4)	2.832(4)	Pt(1)-P(6)	2.235(12)	2.301(12)
Pt(3)-Ag	2.780(4)	2.768(4)	Pt(2)-P(3)	2.284(10)	2.298(10)
Pt(1)-S	2.278(9)	2.278(9)	Pt(3)-P(4)	2.281(11)	2.272(11)
Pt(2)-S	2.257(11)	2.274(10)	Pt(3)-P(5)	2.294(11)	2.254(12)
Pt(3)-S	2.383(9)	2.394(9)	Ag-Cl	2.413(11)	2.388(10)
Pt(1)...Pt(3)	3.674(2)	3.726(2)	Au...Ag	4.414(3)	4.361(3)
Pt(2)...Pt(3)	3.642(2)	3.631(2)	P-C mean	1.828(4)	
Ag...S	3.255(8)	3.243(9)	P-C range	1.77(5) - 1.87(5)	

	Cation A	Cation B		Cation A	Cation B
Pt(2)-Pt(1)-Ag	63.6(1)	62.3(1)	Pt(1)-Pt(2)-Ag	61.5(1)	62.7(1)
Pt(2)-Pt(1)-S	54.4(3)	54.8(3)	Pt(1)-Pt(2)-S	55.2(3)	55.0(3)
Pt(2)-Pt(1)-P(1)	96.7(3)	96.4(3)	Pt(1)-Pt(2)-P(2)	94.1(3)	94.4(3)
Pt(2)-Pt(1)-P(6)	154.2(3)	152.7(3)	Pt(1)-Pt(2)-P(3)	154.5(3)	154.4(3)
Ag-Pt(1)-S	78.8(3)	77.7(3)	Ag-Pt(2)-S	78.0(3)	78.0(3)
Ag-Pt(1)-P(1)	97.1(3)	94.8(3)	Ag-Pt(2)-P(2)	92.1(3)	92.4(3)
Ag-Pt(1)-P(6)	109.5(3)	106.7(3)	Ag-Pt(2)-P(3)	114.1(3)	114.8(3)
S-Pt(1)-P(1)	149.8(4)	150.6(4)	S-Pt(2)-P(2)	148.8(4)	149.0(4)
S-Pt(1)-P(6)	100.5(4)	99.6(4)	S-Pt(2)-P(3)	99.6(4)	99.5(4)
P(1)-Pt(1)-P(6)	109.0(4)	109.8(4)	P(2)-Pt(2)-P(3)	111.4(4)	111.3(4)
Au-Pt(3)-Ag	110.9(1)	109.4(1)	Ag-Pt(3)-P(4)	101.1(3)	103.9(3)
Au-Pt(3)-S	171.3(3)	172.9(3)	Ag-Pt(3)-P(5)	94.0(3)	93.5(3)
Au-Pt(3)-P(4)	81.7(3)	81.0(3)	S-Pt(3)-P(4)	97.0(4)	95.7(4)
Au-Pt(3)-P(5)	80.7(3)	82.4(3)	S-Pt(3)-P(5)	98.9(4)	99.3(4)
Ag-Pt(3)-S	77.7(3)	77.4(3)	P(4)-Pt(3)-P(5)	160.0(4)	159.1(4)
Pt(1)-Ag-Pt(2)	54.9(1)	55.0(1)	Cl-Ag-Pt(1)	122.2(4)	131.1(4)
Pt(1)-Ag-Pt(3)	82.2(1)	83.2(1)	Cl-Ag-Pt(2)	132.7(4)	134.1(4)
Pt(2)-Ag-Pt(3)	80.4(1)	80.8(1)	Cl-Ag-Pt(3)	145.5(4)	139.3(4)
Pt(3)-S-Pt(1)	104.0(4)	105.8(4)	Pt(3)-S-Pt(2)	103.4(4)	102.1(4)
Pt(1)-S-Pt(2)	70.4(3)	70.3(3)	Pt(3)-Au-P(7)	174.4(3)	174.1(3)
P(1)-C(1)-P(2)	113(2)	115(2)	P(5)-C(3)-P(6)	120(2)	122(2)
P(3)-C(2)-P(4)	118(2)	120(2)			

Table 7.5. Fractional atomic coordinates and displacement parameters (\AA^2) for the crystal structure of complex (7).

	X/A	Y/B	Z/C	U
PT(1)	0.19860(2)	0.00000	0.29380(2)	0.033
PT(2)	0.19204(2)	-0.01483(7)	0.19556(2)	0.033
PT(3)	0.10112(2)	0.01972(6)	0.22621(2)	0.033
SN	0.17442(4)	0.18753(13)	0.23345(5)	0.046
P(1)	0.29188(13)	-0.03269(40)	0.32424(15)	0.041
P(2)	0.28273(15)	-0.07237(38)	0.21121(17)	0.042
P(3)	0.15027(14)	-0.02026(44)	0.10963(15)	0.045
P(4)	0.04779(14)	-0.03892(35)	0.14742(15)	0.037
P(5)	0.05208(14)	0.05657(38)	0.28500(17)	0.042
P(6)	0.15904(14)	-0.01354(40)	0.36159(14)	0.039
P(7)	0.17177(30)	0.38186(53)	0.32239(31)	0.088
P(8)	0.0000	-0.0044(12)	0.5000	0.109
P(9)	1.0000	0.3361(8)	1.0000	0.087
F(11)	0.1693(6)	0.2634(13)	0.3030(6)	0.096(4)
F(12)	0.1439(7)	0.3865(18)	0.3651(8)	0.139(6)
F(13)	0.1528(9)	0.4698(24)	0.2823(9)	0.189(9)
F(14)	0.2306(9)	0.4072(20)	0.3504(9)	0.163(8)
F(21)	0.0640(11)	-0.0021(31)	0.5126(10)	0.225(11)
F(22)	0.0063(12)	0.0679(35)	0.5467(12)	0.261(16)
F(23)	-0.0052(15)	-0.1281(34)	0.5286(13)	0.273(17)
F(31)	1.0182(12)	0.3336(31)	1.0604(12)	0.239(13)
F(32)	0.9623(13)	0.2332(40)	0.9915(16)	0.314(20)
F(33)	0.9618(13)	0.4436(37)	1.0026(14)	0.303(19)
C(1)	0.3262(5)	-0.0152(14)	0.2705(5)	0.036(3)
C(2)	0.0763(5)	0.0153(15)	0.0939(6)	0.045(3)
C(3)	0.0972(6)	0.0836(15)	0.3490(7)	0.046(4)
C(4)	0.2494(10)	0.2630(22)	0.2324(10)	0.082(7)
C(5)	0.1166(8)	0.3019(18)	0.1845(8)	0.064(5)
C(A1)	0.3298(14)	0.0680(23)	0.3699(7)	0.051(4)
C(A2)	0.3031(9)	0.1669(29)	0.3725(11)	0.108(9)
C(A3)	0.3303(11)	0.2487(17)	0.4056(13)	0.118(10)
C(A4)	0.3840(14)	0.2316(23)	0.4361(7)	0.102(8)
C(A5)	0.4107(9)	0.1328(29)	0.4335(11)	0.36(4)
C(A6)	0.3835(11)	0.0510(17)	0.4003(13)	0.123(11)
C(B1)	0.3157(7)	-0.1671(15)	0.3518(9)	0.044(4)
C(B2)	0.2773(6)	-0.2350(20)	0.3653(10)	0.083(7)
C(B3)	0.2938(6)	-0.3360(14)	0.3876(5)	0.108(9)
C(B4)	0.3488(7)	-0.3693(15)	0.3966(9)	0.089(7)
C(B5)	0.3872(6)	-0.3014(20)	0.3831(10)	0.119(10)
C(B6)	0.3707(6)	-0.2004(14)	0.3607(5)	0.072(6)
C(C1)	0.2925(7)	-0.2204(11)	0.2242(8)	0.046(4)
C(C2)	0.2491(6)	-0.2811(16)	0.2334(9)	0.056(5)
C(C3)	0.2561(5)	-0.3922(13)	0.2441(4)	0.091(7)
C(C4)	0.3064(7)	-0.4426(11)	0.2456(8)	0.097(8)
C(C5)	0.3497(6)	-0.3818(16)	0.2364(9)	0.073(6)
C(C6)	0.3428(5)	-0.2707(13)	0.2257(4)	0.065(5)
C(D1)	0.3230(5)	-0.0405(8)	0.1653(5)	0.042(3)
C(D2)	0.3530(8)	0.0564(11)	0.1699(7)	0.063(5)
C(D3)	0.3851(7)	0.0784(12)	0.1364(4)	0.065(5)
C(D4)	0.3871(5)	0.0034(8)	0.0983(5)	0.064(4)
C(D5)	0.3571(8)	-0.0936(11)	0.0937(7)	0.069(5)
C(D6)	0.3250(7)	-0.1155(12)	0.1273(4)	0.055(4)
C(E1)	0.1498(9)	-0.1501(12)	0.0733(7)	0.049(4)
C(E2)	0.1689(7)	-0.2448(19)	0.1007(4)	0.062(5)
C(E3)	0.1623(5)	-0.3456(15)	0.0760(8)	0.076(6)
C(E4)	0.1366(9)	-0.3516(12)	0.0237(7)	0.077(6)
C(E5)	0.1176(7)	-0.2569(19)	-0.0038(4)	0.074(6)
C(E6)	0.1242(5)	-0.1562(15)	0.0210(8)	0.074(6)
C(F1)	0.1782(6)	0.0831(12)	0.0729(4)	0.052(4)
C(F2)	0.1648(8)	0.1933(15)	0.0739(10)	0.086(7)
C(F3)	0.1930(10)	0.2707(13)	0.0530(10)	0.115(10)
C(F4)	0.2347(6)	0.2378(12)	0.0310(4)	0.107(9)
C(F5)	0.2480(8)	0.1276(15)	0.0299(10)	0.107(9)
C(F6)	0.2198(10)	0.0503(13)	0.0509(10)	0.077(6)
C(G1)	0.0495(7)	-0.1894(8)	0.1410(7)	0.038(3)
C(G2)	0.0223(4)	-0.2386(15)	0.0948(6)	0.052(4)
C(G3)	0.0228(8)	-0.3518(16)	0.0900(4)	0.078(6)
C(G4)	0.0505(7)	-0.4158(8)	0.1314(7)	0.070(6)
C(G5)	0.0778(4)	-0.3665(15)	0.1776(6)	0.061(5)
C(G6)	0.0773(8)	-0.2533(16)	0.1824(4)	0.053(4)

Table 7.5(cont).

	X/A	Y/B	Z/C	U
C(H1)	-0.0271(5)	-0.0104(19)	0.1235(8)	0.043(3)
C(H2)	-0.0463(4)	0.0761(11)	0.0902(4)	0.064(5)
C(H3)	-0.1029(4)	0.0989(16)	0.0742(8)	0.072(6)
C(H4)	-0.1402(5)	0.0352(19)	0.0914(8)	0.083(6)
C(H5)	-0.1210(4)	-0.0514(11)	0.1247(4)	0.073(6)
C(H6)	-0.0645(4)	-0.0742(16)	0.1408(8)	0.063(5)
C(I1)	0.0163(6)	0.1895(13)	0.2716(9)	0.043(4)
C(I2)	-0.0030(9)	0.2196(13)	0.2204(8)	0.060(5)
C(I3)	-0.0234(7)	0.3246(18)	0.2076(5)	0.079(6)
C(I4)	-0.0244(6)	0.3994(13)	0.2461(9)	0.072(6)
C(I5)	-0.0051(9)	0.3693(13)	0.2972(8)	0.095(8)
C(I6)	0.0153(7)	0.2643(18)	0.3100(5)	0.081(7)
C(J1)	-0.0002(5)	-0.0426(13)	0.2935(3)	0.044(4)
C(J2)	-0.0348(9)	-0.0175(12)	0.3245(8)	0.065(5)
C(J3)	-0.0733(7)	-0.0939(9)	0.3315(8)	0.078(6)
C(J4)	-0.0772(5)	-0.1954(13)	0.3075(3)	0.085(7)
C(J5)	-0.0427(9)	-0.2204(12)	0.2765(8)	0.080(6)
C(J6)	-0.0042(7)	-0.1440(9)	0.2696(8)	0.061(5)
C(K1)	0.1282(6)	-0.1475(11)	0.3686(5)	0.044(4)
C(K2)	0.1294(6)	-0.2327(9)	0.3349(6)	0.055(4)
C(K3)	0.1037(9)	-0.3316(12)	0.3395(7)	0.081(6)
C(K4)	0.0769(6)	-0.3454(11)	0.3779(5)	0.084(7)
C(K5)	0.0758(6)	-0.2602(9)	0.4116(6)	0.082(7)
C(K6)	0.1013(9)	-0.1612(12)	0.4070(7)	0.063(5)
C(L1)	0.2013(11)	0.0232(34)	0.4237(5)	0.054(4)
C(L2)	0.2327(14)	-0.0562(18)	0.4553(12)	0.128(11)
C(L3)	0.2691(8)	-0.0265(29)	0.5020(12)	0.189(17)
C(L4)	0.2739(11)	0.0826(34)	0.5171(5)	0.095(8)
C(L5)	0.2424(14)	0.1620(18)	0.4855(12)	0.178(17)
C(L6)	0.2061(8)	0.1323(29)	0.4388(12)	0.174(16)
H(11)	.36173	-.05222	.27939	.043
H(12)	.33210	.06174	.26563	.043
H(21)	.05679	-.01757	.06153	.054
H(22)	.07180	.09390	.09115	.054
H(31)	.07688	.07217	.37427	.055
H(32)	.11010	.15849	.35107	.055
H(41)	.26962	.26226	.26697	.098
H(42)	.27145	.21327	.21831	.098
H(43)	.24089	.32789	.21151	.098
H(51)	.09468	.31861	.20798	.077
H(52)	.10828	.36704	.16360	.077
H(53)	.09666	.24995	.15938	.077
H(A2)	.26576	.17884	.35108	.130
H(A3)	.31167	.31753	.40717	.140
H(A4)	.40276	.28855	.45905	.122
H(A5)	.44789	.12088	.45482	.400
H(A6)	.40199	-.01775	.39869	.148
H(B2)	.23901	-.21182	.35898	.099
H(B3)	.26703	-.38324	.39694	.130
H(B4)	.36026	-.43966	.41213	.107
H(B5)	.42546	-.32459	.38936	.143
H(B6)	.39738	-.15317	.35136	.086
H(C2)	.21417	-.24613	.23240	.067
H(C3)	.22595	-.43459	.25046	.109
H(C4)	.31121	-.51993	.25301	.117
H(C5)	.38470	-.41686	.23739	.087
H(C6)	.37293	-.22839	.21939	.078
H(D2)	.35150	.10851	.19635	.076
H(D3)	.40594	.14585	.13954	.078
H(D4)	.40947	.01862	.07498	.076
H(D5)	.35855	-.14561	.06725	.083
H(D6)	.30417	-.18297	.12412	.066
H(E2)	.18671	-.24081	.13718	.074
H(E3)	.17539	-.41150	.09515	.091
H(E4)	.13205	-.42160	.00647	.092
H(E5)	.09983	-.26091	-.04017	.089
H(E6)	.11100	-.09022	.00190	.089
H(F2)	.13585	.21618	.08927	.103
H(F3)	.18376	.34740	.05372	.139
H(F4)	.25435	.29176	.01640	.128
H(F5)	.27698	.10491	.01458	.128
H(F6)	.22907	-.02633	.05008	.092
H(G2)	.00292	-.19414	.06601	.063
H(G3)	.00385	-.38603	.05785	.093
H(G4)	.05089	-.49451	.12807	.084
H(G5)	.09708	-.41100	.20642	.074
H(G6)	.09618	-.21910	.21458	.063

Table 7.5(cont).

	X/A	Y/B	Z/C	U
H(H2)	-.02037	.12042	.07818	.077
H(H3)	-.11624	.15908	.05097	.087
H(H4)	-.17956	.05101	.08023	.099
H(H5)	-.14699	-.09577	.13673	.088
H(H6)	-.05113	-.13439	.16397	.075
H(I2)	-.00229	.16745	.19371	.072
H(I3)	-.03683	.34554	.17197	.095
H(I4)	-.03848	.47252	.23704	.086
H(I5)	-.00567	.42154	.32384	.114
H(I6)	.02877	.24344	.34562	.097
H(J2)	-.03206	.05304	.34112	.078
H(J3)	-.09731	-.07654	.35306	.093
H(J4)	-.10397	-.24857	.31247	.102
H(J5)	-.04541	-.29102	.25995	.096
H(J6)	.01988	-.16150	.24807	.073
H(K2)	.14802	-.22317	.30824	.066
H(K3)	.10454	-.39092	.31611	.097
H(K4)	.05909	-.41418	.38117	.100
H(K5)	.05709	-.26971	.43831	.098
H(K6)	.10050	-.10190	.43039	.076
H(L2)	.22881	-.13236	.44521	.154
H(L3)	.29032	-.08213	.52431	.228
H(L4)	.29919	.10304	.54952	.114
H(L5)	.24635	.23800	.49568	.212
H(L6)	.18485	.18776	.41651	.208

Anisotropic displacement parameters (\AA^2).

	U11	U22	U33	U12	U13	U23
PT(1)	.0326(2)	.0323(3)	.0303(3)	.0007(2)	.0082(2)	-.0004(3)
PT(2)	.0328(2)	.0348(3)	.0302(3)	-.0020(2)	.0092(2)	-.0001(3)
PT(3)	.0311(2)	.0316(3)	.0327(3)	-.0004(2)	.0080(2)	0.0000(3)
SN	.0459(5)	.0303(6)	.0566(8)	-.0043(4)	.0147(5)	.0013(5)
P(1)	.0327(15)	.0506(26)	.0358(20)	-.0004(17)	.0056(14)	-.0023(20)
P(2)	.0321(16)	.0405(23)	.0505(26)	-.0011(16)	.0141(16)	.0016(19)
P(3)	.0409(17)	.0544(26)	.0377(20)	-.0051(20)	.0113(14)	-.0002(22)
P(4)	.0347(15)	.0358(22)	.0377(21)	-.0009(15)	.0075(14)	.0020(17)
P(5)	.0331(17)	.0386(22)	.0512(26)	.0047(16)	.0160(16)	.0033(19)
P(6)	.0422(16)	.0378(21)	.0354(18)	.0036(18)	.0111(14)	.0015(20)
P(7)	.103(5)	.046(3)	.109(6)	.000(3)	.033(4)	-.018(4)
P(8)	.081(5)	.091(8)	.135(9)	.000	-.017(5)	.000
P(9)	.133(8)	.048(5)	.072(6)	.000	.025(6)	.000

Table 7.6. Selected bond distances (\AA) and angles ($^\circ$) of complex (7).

PT(1) - PT(2)	2.609(1)	PT(1) - PT(3)	2.615(1)
PT(1) - SN	2.766(2)	PT(1) - P(1)	2.272(4)
PT(1) - P(6)	2.297(4)	PT(2) - PT(3)	2.635(1)
PT(2) - SN	2.739(2)	PT(2) - P(2)	2.284(4)
PT(2) - P(3)	2.264(4)	PT(3) - SN	2.702(2)
PT(3) - P(4)	2.288(4)	PT(3) - P(5)	2.284(5)
SN - F(11)	2.118(16)	SN - C(4)	2.08(3)
SN - C(5)	2.17(3)	P(1) - C(1)	1.877(13)
P(1) - C(A1)	1.81(3)	P(1) - C(B1)	1.82(2)
P(2) - C(1)	1.804(14)	P(2) - C(C1)	1.833(15)
P(2) - C(D1)	1.826(14)	P(3) - C(2)	1.820(14)
P(3) - C(E1)	1.853(17)	P(3) - C(F1)	1.844(15)
P(4) - C(2)	1.886(16)	P(4) - C(G1)	1.835(11)
P(4) - C(H1)	1.829(13)	P(5) - C(3)	1.812(17)
P(5) - C(I1)	1.829(17)	P(5) - C(J1)	1.829(15)
P(6) - C(3)	1.891(17)	P(6) - C(K1)	1.828(15)
P(6) - C(L1)	1.773(19)	P(7) - F(11)	1.526(17)
P(7) - F(12)	1.49(3)	P(7) - F(13)	1.50(3)
P(7) - F(14)	1.48(3)	P(8) - F(21)	1.53(3)
P(8) - F(22)	1.51(4)	P(8) - F(23)	1.71(5)
P(9) - F(31)	1.56(4)	P(9) - F(32)	1.54(5)
P(9) - F(33)	1.62(5)		

Table 7.6(cont).

PT(2) - PT(1) - PT(3)	60.6(1)	PT(2) - PT(1) - SN	61.2(1)
PT(2) - PT(1) - P(1)	97.6(1)	PT(2) - PT(1) - P(6)	151.1(1)
PT(3) - PT(1) - SN	60.2(1)	PT(3) - PT(1) - P(1)	158.1(2)
PT(3) - PT(1) - P(6)	92.6(1)	SN - PT(1) - P(1)	113.2(2)
SN - PT(1) - P(6)	116.9(2)	P(1) - PT(1) - P(6)	108.2(2)
PT(1) - PT(2) - PT(3)	59.8(1)	PT(1) - PT(2) - SN	62.2(1)
PT(1) - PT(2) - P(2)	92.6(2)	PT(1) - PT(2) - P(3)	157.2(1)
PT(3) - PT(2) - SN	60.3(1)	PT(3) - PT(2) - P(2)	151.0(2)
PT(3) - PT(2) - P(3)	97.4(1)	SN - PT(2) - P(2)	117.0(2)
SN - PT(2) - P(3)	109.3(2)	P(2) - PT(2) - P(3)	109.7(2)
PT(1) - PT(3) - PT(2)	59.6(1)	PT(1) - PT(3) - SN	62.7(1)
PT(1) - PT(3) - P(4)	144.2(1)	PT(1) - PT(3) - P(5)	95.9(2)
PT(2) - PT(3) - SN	61.7(1)	PT(2) - PT(3) - P(4)	89.5(1)
PT(2) - PT(3) - P(5)	155.3(2)	SN - PT(3) - P(4)	120.7(2)
SN - PT(3) - P(5)	106.0(2)	P(4) - PT(3) - P(5)	114.8(2)
PT(1) - SN - PT(2)	56.6(1)	PT(1) - SN - PT(3)	57.1(1)
PT(1) - SN - F(11)	83.9(5)	PT(1) - SN - C(4)	108.5(8)
PT(1) - SN - C(5)	152.5(6)	PT(2) - SN - PT(3)	57.9(1)
PT(2) - SN - F(11)	140.5(5)	PT(2) - SN - C(4)	99.0(8)
PT(2) - SN - C(5)	119.6(6)	PT(3) - SN - F(11)	101.1(4)
PT(3) - SN - C(4)	156.6(8)	PT(3) - SN - C(5)	96.7(6)
F(11) - SN - C(4)	95.0(9)	F(11) - SN - C(5)	94.0(7)
C(4) - SN - C(5)	99.1(9)	PT(1) - P(1) - C(1)	109.4(5)
PT(1) - P(1) - C(A1)	114.8(11)	PT(1) - P(1) - C(B1)	119.3(6)
C(1) - P(1) - C(A1)	100.8(11)	C(1) - P(1) - C(B1)	104.5(9)
C(A1) - P(1) - C(B1)	106.1(11)	PT(2) - P(2) - C(1)	111.7(5)
PT(2) - P(2) - C(C1)	114.1(6)	PT(2) - P(2) - C(D1)	120.1(5)
C(1) - P(2) - C(C1)	101.0(9)	C(1) - P(2) - C(D1)	102.2(6)
C(C1) - P(2) - C(D1)	105.6(7)	PT(2) - P(3) - C(2)	112.7(5)
PT(2) - P(3) - C(E1)	119.9(7)	PT(2) - P(3) - C(F1)	113.1(5)
C(2) - P(3) - C(E1)	102.5(9)	C(2) - P(3) - C(F1)	102.4(8)
C(E1) - P(3) - C(F1)	104.2(8)	PT(3) - P(4) - C(2)	110.9(5)
PT(3) - P(4) - C(G1)	112.0(6)	PT(3) - P(4) - C(H1)	123.1(7)
C(2) - P(4) - C(G1)	104.6(8)	C(2) - P(4) - C(H1)	102.8(8)
C(G1) - P(4) - C(H1)	101.7(10)	PT(3) - P(5) - C(3)	112.6(5)
PT(3) - P(5) - C(I1)	110.9(7)	PT(3) - P(5) - C(J1)	119.1(5)
C(3) - P(5) - C(I1)	99.8(10)	C(3) - P(5) - C(J1)	106.2(7)
C(I1) - P(5) - C(J1)	106.4(7)	PT(1) - P(6) - C(3)	108.2(6)
PT(1) - P(6) - C(K1)	115.2(5)	PT(1) - P(6) - C(L1)	117.4(9)
C(3) - P(6) - C(K1)	103.3(8)	C(3) - P(6) - C(L1)	104.9(13)
C(K1) - P(6) - C(L1)	106.5(13)	F(11) - P(7) - F(12)	108.4(11)
F(11) - P(7) - F(13)	116.7(13)	F(11) - P(7) - F(14)	108.2(12)
F(12) - P(7) - F(13)	113.9(13)	F(12) - P(7) - F(14)	101.3(12)
F(13) - P(7) - F(14)	107.1(14)	F(21) - P(8) - F(21')	177.9(23)
F(21) - P(8) - F(22)	86.4(17)	F(21) - P(8) - F(22')	92.4(17)
F(21) - P(8) - F(23)	96.7(19)	F(21) - P(8) - F(23')	85.2(19)
F(22) - P(8) - F(22')	108.7(23)	F(22) - P(8) - F(23)	98.0(20)
F(22) - P(8) - F(23')	152.3(20)	F(23) - P(8) - F(23')	57.0(18)
F(31) - P(9) - F(31')	177.8(21)	F(31) - P(9) - F(32)	97.7(21)
F(31) - P(9) - F(32')	80.6(21)	F(31) - P(9) - F(33)	88.8(19)
F(31) - P(9) - F(33')	92.9(19)	F(32) - P(9) - F(32')	71.6(21)
F(32) - P(9) - F(33)	108.7(19)	F(32) - P(9) - F(33')	169.3(21)
F(33) - P(9) - F(33')	73.2(19)	SN - F(11) - P(7)	134.8(10)
P(1) - C(1) - P(2)	110.5(7)	P(3) - C(2) - P(4)	107.5(8)
P(5) - C(3) - P(6)	108.7(9)	P(1) - C(A1) - C(A2)	117.0(22)
P(1) - C(A1) - C(A6)	123.0(22)	P(1) - C(B1) - C(B2)	118.0(14)
P(1) - C(B1) - C(B6)	121.9(14)	P(2) - C(C1) - C(C2)	119.2(13)
P(2) - C(C1) - C(C6)	120.7(12)	P(2) - C(D1) - C(D2)	120.0(11)
P(2) - C(D1) - C(D6)	119.8(9)	P(3) - C(E1) - C(E2)	118.3(14)
P(3) - C(E1) - C(E6)	121.2(13)	P(3) - C(F1) - C(F2)	121.4(13)
P(3) - C(F1) - C(F6)	117.8(12)	P(4) - C(G1) - C(G2)	119.8(13)
P(4) - C(G1) - C(G6)	120.2(13)	P(4) - C(H1) - C(H2)	120.8(12)
P(4) - C(H1) - C(H6)	119.2(16)	P(5) - C(I1) - C(I2)	116.7(15)
P(5) - C(I1) - C(I6)	122.6(17)	P(5) - C(J1) - C(J2)	120.5(12)
P(5) - C(J1) - C(J6)	119.5(11)	P(6) - C(K1) - C(K2)	120.9(11)
P(6) - C(K1) - C(K6)	119.0(11)	P(6) - C(L1) - C(L2)	119.8(27)
P(6) - C(L1) - C(L6)	120.0(23)		

Symmetry code:-

' -x y -z

Table 7.7. Fractional atomic positions and displacement parameters (\AA^2) of the crystal structure of complexes (8)/(9).

	X/A	Y/B	Z/C	U
PT	-0.2763(1)	-0.1957(1)	-0.2513(1)	0.031
SN	-0.2925(2)	-0.2925	-0.2925	0.040
CL	-0.18270	-0.18270	-0.18270	0.03(1)
P(1)	-0.3479(6)	-0.1759(5)	-0.2061(5)	0.034
P(2)	-0.2701(6)	-0.1397(6)	-0.3198(6)	0.048
P(3)	-0.50000	0.00000	-0.50000	0.044(9)
F(1)	-0.370(1)	-0.299(1)	-0.282(1)	0.082
F(31)	-0.564(1)	0.000	-0.500	0.06(1)
F(32)	-0.500	0.000	-0.563(1)	0.05(1)
F(33)	-0.500	-0.057(2)	-0.500	0.08(2)
O(2)	-0.18270	-0.18270	-0.18270	0.030
C(1)	-0.370(2)	-0.237(2)	-0.173(2)	0.04(2)
C(2)	-0.20770	-0.20770	-0.20770	0.02(3)
C(A1)	-0.343(4)	-0.132(2)	-0.155(1)	0.10(3)
C(A2)	-0.295(3)	-0.117(1)	-0.138(3)	0.12(3)
C(A3)	-0.291(2)	-0.084(2)	-0.097(2)	0.08(2)
C(A4)	-0.334(3)	-0.065(1)	-0.074(1)	0.09(2)
C(A5)	-0.382(2)	-0.079(2)	-0.092(3)	0.07(2)
C(A6)	-0.386(3)	-0.113(2)	-0.132(2)	0.16(3)
C(B1)	-0.403(1)	-0.150(3)	-0.245(2)	0.07(2)
C(B2)	-0.405(2)	-0.100(3)	-0.259(1)	0.08(2)
C(B3)	-0.446(2)	-0.082(1)	-0.287(2)	0.09(2)
C(B4)	-0.484(1)	-0.115(2)	-0.302(2)	0.06(2)
C(B5)	-0.482(2)	-0.166(2)	-0.288(1)	0.08(2)
C(B6)	-0.442(2)	-0.183(2)	-0.259(2)	0.10(2)
C(C1)	-0.233(2)	-0.083(2)	-0.305(2)	0.06(2)
C(C2)	-0.226(1)	-0.069(1)	-0.255(1)	0.08(2)
C(C3)	-0.200(2)	-0.025(2)	-0.243(1)	0.07(2)
C(C4)	-0.181(1)	0.006(1)	-0.281(1)	0.08(2)
C(C5)	-0.188(1)	-0.008(2)	-0.332(1)	0.04(2)
C(C6)	-0.214(2)	-0.052(2)	-0.343(1)	0.03(2)
C(D1)	-0.329(1)	-0.121(3)	-0.353(2)	0.03(2)
C(D2)	-0.363(2)	-0.157(2)	-0.371(2)	0.12(3)
C(D3)	-0.405(2)	-0.142(2)	-0.399(1)	0.08(2)
C(D4)	-0.412(1)	-0.091(2)	-0.409(1)	0.07(2)
C(D5)	-0.378(2)	-0.055(1)	-0.391(2)	0.11(3)
C(D6)	-0.336(2)	-0.070(3)	-0.363(1)	0.11(3)
H(11)	-0.37339	-0.25893	-0.20299	0.100
H(12)	-0.40193	-0.23475	-0.15616	0.100
H(A2)	-0.265(4)	-0.131(2)	-0.154(4)	0.100
H(A3)	-0.258(3)	-0.075(3)	-0.085(3)	0.100
H(A4)	-0.331(5)	-0.042(2)	-0.046(1)	0.100
H(A5)	-0.412(3)	-0.066(2)	-0.076(4)	0.100
H(A6)	-0.419(3)	-0.122(4)	-0.145(3)	0.100
H(B2)	-0.378(3)	-0.077(4)	-0.249(1)	0.100
H(B3)	-0.447(3)	-0.047(1)	-0.297(2)	0.100
H(B4)	-0.512(1)	-0.103(4)	-0.322(2)	0.100
H(B5)	-0.509(3)	-0.189(3)	-0.298(1)	0.100
H(B6)	-0.440(3)	-0.219(2)	-0.249(3)	0.100
H(C2)	-0.239(1)	-0.090(2)	-0.228(2)	0.100
H(C3)	-0.195(2)	-0.016(3)	-0.208(1)	0.100
H(C4)	-0.163(2)	0.036(2)	-0.273(2)	0.100
H(C5)	-0.175(1)	0.013(2)	-0.358(2)	0.100
H(C6)	-0.219(3)	-0.062(3)	-0.379(1)	0.100
H(D2)	-0.358(3)	-0.193(2)	-0.364(3)	0.100
H(D3)	-0.428(3)	-0.167(3)	-0.412(1)	0.100
H(D4)	-0.441(1)	-0.080(4)	-0.429(2)	0.100
H(D5)	-0.383(3)	-0.019(1)	-0.398(3)	0.100
H(D6)	-0.313(3)	-0.045(3)	-0.350(1)	0.100

Anisotropic displacement parameters (\AA^2).

	U11	U22	U33	U12	U13	U23
PT	0.030(1)	0.027(1)	0.036(1)	0.004(1)	0.004(1)	0.005(2)
SN	0.040(2)	0.040	0.040	-0.002	-0.002	-0.002(2)
P(1)	0.05(1)	0.03(1)	0.02(1)	0.02(1)	0.00(1)	0.01(1)
P(2)	0.06(2)	0.05(1)	0.04(1)	-0.01(1)	0.00(1)	0.01(1)
F(1)	0.12(3)	0.02(2)	0.11(3)	-0.02(2)	0.03(2)	-0.02(2)

Table 7.8. Selected bond distances (Å) and angles (°) of complexes (8)/(9).

Pt - Pt' 2.634(4)
 Pt - Cl 3.045(1)
 Pt - P(2) 2.315(16)
 Sn - F(1) 2.06(3)
 P(1) - C(A1) 1.76(4)
 P(2) - C(1) 1.79(5)
 P(2) - C(D1) 1.82(4)
 P(3) - F(32) 1.65(4)
 O(2) - C(2) 1.129

Pt - Sn 2.773(3)
 Pt - P(1) 2.268(15)
 Pt - C(2) 2.142(3)
 P(1) - C(1) 1.90(5)
 P(1) - C(B1) 1.88(5)
 P(2) - C(C1) 1.81(5)
 P(3) - F(31) 1.66(4)
 P(3) - F(33) 1.50(5)

Pt' - Pt - Pt" 60.0
 Pt' - Pt - P(1) 95.8(4)
 Pt' - Pt - C(2) 52.1(1)
 Pt" - Pt - P(2) 95.8(5)
 Sn - Pt - P(2) 106.6(4)
 P(1) - Pt - P(2) 108.3(6)
 P(2) - Pt - C(2) 116.3(5)
 Pt - Sn - F(1) 100.0(8)
 Pt - Sn - F(1") 105.6(8)
 Pt - P(1) - C(1) 106.5(8)
 Pt - P(1) - C(B1) 115.8(15)
 C(1) - P(1) - C(B1) 108.1(25)
 Pt - P(2) - C(1) 106.4(17)
 Pt - P(2) - C(D1) 118.0(19)
 C(1) - P(2) - C(D1) 101.0(24)
 P(1) - C(1) - P(2) 114.2(26)
 Pt - C(2) - O(2) 134.8(1)

Pt' - Pt - Sn 61.7(1)
 Pt' - Pt - P(2) 155.7(5)
 Pt" - Pt - P(1) 155.7(4)
 Sn - Pt - P(1) 106.5(4)
 Sn - Pt - C(2) 101.5(2)
 P(1) - Pt - C(2) 116.5(4)
 Pt - Sn - Pt' 56.7(1)
 Pt - Sn - F(1') 155.4(8)
 F(1) - Sn - F(1') 93.7(12)
 Pt - P(1) - C(A1) 118.6(34)
 C(1) - P(1) - C(A1) 103.7(22)
 C(A1) - P(1) - C(B1) 103.1(35)
 Pt - P(2) - C(C1) 112.9(16)
 C(1) - P(2) - C(C1) 107.8(22)
 C(C1) - P(2) - C(D1) 109.6(28)
 Pt - C(2) - Pt' 75.9(1)

Symmetry code:-

' z x y
 " y z x

Table 7.9. Fractional atomic coordinates and displacement parameters (\AA^2) of the crystal structure of complex (10).

	X/A	Y/B	Z/C	U
PT(1)	.19606(5)	.11103(5)	.20386(3)	.024
PT(2)	.29846(5)	.13848(5)	.29008(3)	.025
PT(3)	.22125(5)	.27533(5)	.21642(3)	.026
SN(1)	.07564(9)	.19459(9)	.30860(6)	.033
SN(2)	.40535(9)	.14772(9)	.16875(6)	.032
CL(1)	-.2837(7)	.3617(7)	.1717(4)	.124
CL(2)	-.2123(9)	.3359(7)	.0404(5)	.152
P(1)	.2272(3)	-.0458(3)	.2164(2)	.030
P(2)	.3344(3)	-.0142(4)	.3230(2)	.032
P(3)	.3838(4)	.2250(4)	.3367(2)	.034
P(4)	.2865(4)	.3841(3)	.2529(2)	.034
P(5)	.1227(4)	.3491(3)	.1429(2)	.034
P(6)	.0962(4)	.1568(3)	.1266(2)	.031
F(11)	.0280(9)	.2649(8)	.3852(6)	.070
F(12)	-.0556(8)	.2535(9)	.2793(6)	.069
F(13)	.0043(10)	.1009(9)	.3543(6)	.079
F(21)	.4995(8)	.0274(7)	.1625(6)	.058
F(22)	.5247(10)	.2068(9)	.1736(6)	.083
F(23)	.4354(9)	.1720(10)	.0794(5)	.079
C(1)	.3384(13)	-.0853(12)	.2588(8)	.035(4)
C(2)	.4059(14)	.3272(14)	.2837(9)	.044(5)
C(3)	.0319(13)	.2780(12)	.1331(8)	.033(4)
C(4)	-.2762(27)	.3005(25)	.1087(18)	.129(12)
C(A1)	.2722(16)	-.1001(12)	.1421(6)	.038(5)
C(A2)	.3782(12)	-.1149(15)	.1113(7)	.047(5)
C(A3)	.4062(12)	-.1462(8)	.0517(9)	.051(5)
C(A4)	.3281(16)	-.1627(12)	.0229(6)	.061(6)
C(A5)	.2220(12)	-.1479(15)	.0537(7)	.052(6)
C(A6)	.1941(12)	-.1166(8)	.1133(9)	.044(5)
C(B1)	.1262(7)	-.1115(7)	.2578(5)	.029(4)
C(B2)	.0275(14)	-.0677(7)	.2881(9)	.049(5)
C(B3)	-.0484(14)	-.1178(8)	.3205(10)	.056(6)
C(B4)	-.0256(7)	-.2117(7)	.3226(5)	.060(6)
C(B5)	.0731(14)	-.2556(7)	.2923(9)	.058(6)
C(B6)	.1489(14)	-.2055(8)	.2599(10)	.043(5)
C(C1)	.2388(11)	-.0550(9)	.3899(5)	.033(4)
C(C2)	.1988(14)	-.0002(9)	.4394(7)	.050(5)
C(C3)	.1338(19)	-.0312(12)	.4935(8)	.068(7)
C(C4)	.1088(11)	-.1170(9)	.4981(5)	.064(6)
C(C5)	.1488(14)	-.1719(9)	.4485(7)	.068(7)
C(C6)	.2138(19)	-.1409(12)	.3944(8)	.060(6)
C(D1)	.4653(10)	-.0648(15)	.3449(11)	.034(4)
C(D2)	.5572(10)	-.0559(14)	.3005(7)	.046(5)
C(D3)	.6563(11)	-.0957(6)	.3136(7)	.055(6)
C(D4)	.6634(10)	-.1443(15)	.3710(11)	.068(7)
C(D5)	.5714(10)	-.1532(14)	.4154(7)	.062(6)
C(D6)	.4723(11)	-.1134(6)	.4023(7)	.044(5)
C(E1)	.5193(10)	.1755(16)	.3521(11)	.035(4)
C(E2)	.5265(9)	.1133(8)	.4038(7)	.053(6)
C(E3)	.6249(9)	.0786(14)	.4195(9)	.068(7)
C(E4)	.7161(10)	.1061(16)	.3834(11)	.064(6)
C(E5)	.7089(9)	.1683(8)	.3317(7)	.062(6)
C(E6)	.6105(9)	.2031(14)	.3161(9)	.059(6)
C(F1)	.3216(19)	.2709(12)	.4132(8)	.039(5)
C(F2)	.2218(16)	.2547(15)	.4443(5)	.054(6)
C(F3)	.1748(10)	.2905(10)	.5015(8)	.074(7)
C(F4)	.2278(19)	.3425(12)	.5276(8)	.071(7)
C(F5)	.3276(16)	.3587(15)	.4965(5)	.070(7)
C(F6)	.3745(10)	.3229(10)	.4393(8)	.049(5)
C(G1)	.2046(20)	.4523(10)	.3164(10)	.047(5)
C(G2)	.0960(18)	.4514(17)	.3307(5)	.056(6)
C(G3)	.0311(11)	.4978(13)	.3803(10)	.081(8)
C(G4)	.0748(20)	.5450(10)	.4157(10)	.082(8)
C(G5)	.1834(18)	.5459(17)	.4014(5)	.076(7)
C(G6)	.2483(11)	.4995(13)	.3518(10)	.059(6)
C(H1)	.3375(9)	.4682(11)	.1926(6)	.036(4)
C(H2)	.4175(17)	.4391(10)	.1422(11)	.067(7)
C(H3)	.4477(18)	.5003(9)	.0921(10)	.090(8)
C(H4)	.3978(9)	.5906(11)	.0923(6)	.063(6)
C(H5)	.3178(17)	.6197(10)	.1427(11)	.083(8)
C(H6)	.2876(18)	.5585(9)	.1929(10)	.060(6)
C(I1)	.0375(19)	.4566(8)	.1641(11)	.036(5)
C(I2)	-.0595(12)	.4582(8)	.2058(7)	.064(6)
C(I3)	-.1195(14)	.5401(8)	.2260(9)	.077(7)
C(I4)	-.0824(19)	.6203(8)	.2045(11)	.085(8)

Table 7.9(cont).

	X/A	Y/B	Z/C	U
C(I5)	.0146(12)	.6187(8)	.1628(7)	.075(7)
C(I6)	.0746(14)	.5369(8)	.1426(9)	.063(6)
C(J1)	.1858(14)	.3771(12)	.0606(5)	.035(4)
C(J2)	.1228(14)	.4236(11)	.0193(11)	.054(6)
C(J3)	.1670(10)	.4365(16)	-.0442(10)	.063(6)
C(J4)	.2741(14)	.4030(12)	-.0664(5)	.070(7)
C(J5)	.3370(14)	.3565(11)	-.0251(11)	.068(7)
C(J6)	.2929(10)	.3436(16)	.0384(10)	.058(6)
C(K1)	.1686(14)	.1504(10)	.0439(6)	.031(4)
C(K2)	.1273(13)	.2027(10)	-.0048(10)	.048(5)
C(K3)	.1841(9)	.1974(14)	-.0663(8)	.063(6)
C(K4)	.2820(14)	.1398(10)	-.0790(6)	.050(5)
C(K5)	.3232(13)	.0875(10)	-.0303(10)	.048(5)
C(K6)	.2665(9)	.0928(14)	.0311(8)	.047(5)
C(L1)	-.0218(9)	.1047(12)	.1298(6)	.027(4)
C(L2)	-.1003(17)	.1123(15)	.1841(5)	.053(6)
C(L3)	-.1918(14)	.0766(9)	.1889(7)	.056(6)
C(L4)	-.2049(9)	.0334(12)	.1394(6)	.056(6)
C(L5)	-.1264(17)	.0258(15)	.0851(5)	.068(7)
C(L6)	-.0349(14)	.0615(9)	.0803(7)	.053(6)
H(11)	.40676	-.08626	.22940	.075
H(12)	.33699	-.14776	.27699	.075
H(21)	.46573	.31178	.24878	.075
H(22)	.42765	.37090	.30679	.075
H(31)	.00299	.30061	.09530	.075
H(32)	-.02860	.28226	.16870	.075
H(41)	-.35471	.29886	.10438	.075
H(42)	-.25046	.23467	.11852	.075
H(A2)	.43239	-.10291	.13072	.075
H(A3)	.47948	-.15513	.02939	.075
H(A4)	.34707	-.18392	-.01911	.075
H(A5)	.16756	-.16050	.03371	.075
H(A6)	.12046	-.10829	.13507	.075
H(B2)	.01137	-.00240	.28651	.075
H(B3)	-.11749	-.08737	.34113	.075
H(B4)	-.07880	-.24671	.34475	.075
H(B5)	.08876	-.32105	.29368	.075
H(B6)	.21766	-.23607	.23911	.075
H(C2)	.21566	.05965	.43598	.075
H(C3)	.10552	.00720	.52778	.075
H(C4)	.06379	-.13865	.53587	.075
H(C5)	.13216	-.23203	.45212	.075
H(C6)	.24230	-.17961	.36032	.075
H(D2)	.55248	-.02238	.26064	.075
H(D3)	.72050	-.08931	.28288	.075
H(D4)	.73264	-.17198	.38017	.075
H(D5)	.57668	-.18785	.45520	.075
H(D6)	.40863	-.12088	.43296	.075
H(E2)	.46362	.09392	.42880	.075
H(E3)	.63050	.03560	.45559	.075
H(E4)	.78510	.08265	.39450	.075
H(E5)	.77275	.18796	.30674	.075
H(E6)	.60586	.24629	.28001	.075
H(F2)	.18491	.21860	.42670	.075
H(F3)	.10556	.27927	.52377	.075
H(F4)	.19538	.36760	.56787	.075
H(F5)	.36458	.39520	.51497	.075
H(F6)	.44391	.33452	.41791	.075
H(G2)	.06534	.41932	.30596	.075
H(G3)	-.04476	.49770	.39022	.075
H(G4)	.02937	.57737	.45039	.075
H(G5)	.21357	.57866	.42632	.075
H(G6)	.32365	.50033	.34208	.075
H(H2)	.45256	.37696	.14210	.075
H(H3)	.50434	.48106	.05740	.075
H(H4)	.41950	.63407	.05786	.075
H(H5)	.28305	.68303	.14308	.075
H(H6)	.23123	.57892	.22776	.075
H(I2)	-.08623	.40306	.22006	.075
H(I3)	-.18806	.54184	.25424	.075
H(I4)	-.12471	.67778	.21830	.075
H(I5)	.04051	.67496	.14825	.075
H(I6)	.14239	.53618	.11415	.075
H(J2)	.04859	.44686	.03466	.075

Table 7.9(cont).

	X/A	Y/B	Z/C	U
H(J3)	.12328	.46937	-.07295	.075
H(J4)	.30501	.41308	-.11069	.075
H(J5)	.41204	.33431	-.04078	.075
H(J6)	.33732	.31178	.06682	.075
H(K2)	.05922	.24274	.00362	.075
H(K3)	.15577	.23370	-.10050	.075
H(K4)	.32190	.13593	-.12191	.075
H(K5)	.39148	.04732	-.03918	.075
H(K6)	.29492	.05636	.06492	.075
H(L2)	-.09190	.14214	.21874	.075
H(L3)	-.24742	.08235	.22647	.075
H(L4)	-.26947	.00938	.14234	.075
H(L5)	-.13596	-.00380	.05044	.075
H(L6)	.01958	.05599	.04268	.075

	X/A	Y/B	Z/C	U
CLS(11)	.678(2)	.413(2)	.348(1)	.51(2)
CLS(12)	.659(2)	.343(1)	.480(1)	.54(2)
CLS(13)	.760(2)	.409(2)	.447(2)	.38(2)
CLS(14)	.755(3)	.442(2)	.376(2)	.36(2)
CLS(15)	.763(3)	.333(3)	.407(2)	.28(2)
CLS(21)	.547(2)	.351(2)	.660(1)	.51(2)
CLS(22)	.419(2)	.292(2)	.778(1)	.49(2)
CLS(23)	.445(4)	.445(4)	.737(2)	.24(2)
CLS(24)	.512(3)	.350(3)	.745(2)	.33(2)
CLS(25)	.426(3)	.325(3)	.696(2)	.29(2)

Anisotropic displacement parameters (\AA^2).

	U11	U22	U33	U12	U13	U23
PT(1)	.0240(4)	.0233(4)	.0236(4)	-.0049(3)	-.0043(3)	-.0015(3)
PT(2)	.0251(4)	.0264(4)	.0220(4)	-.0029(3)	-.0044(3)	-.0013(3)
PT(3)	.0252(4)	.0231(4)	.0273(4)	-.0053(3)	-.0076(3)	.0002(3)
SN(1)	.0269(6)	.0375(8)	.0313(7)	-.0048(6)	.0001(6)	-.0047(7)
SN(2)	.0277(6)	.0370(8)	.0275(7)	-.0067(6)	.0014(5)	-.0036(6)
CL(1)	.129(7)	.129(8)	.106(7)	-.046(6)	-.035(5)	-.003(6)
CL(2)	.209(10)	.124(8)	.121(8)	-.072(8)	-.054(7)	-.006(6)
P(1)	.028(2)	.029(3)	.029(3)	-.009(2)	-.003(2)	-.003(2)
P(2)	.028(2)	.036(3)	.029(3)	-.001(2)	-.007(2)	.002(2)
P(3)	.032(3)	.040(3)	.028(3)	-.003(2)	-.009(2)	-.009(2)
P(4)	.033(3)	.025(3)	.039(3)	-.005(2)	-.005(2)	0.000(2)
P(5)	.035(3)	.027(3)	.037(3)	-.006(2)	-.011(2)	.003(2)
P(6)	.030(2)	.028(3)	.030(3)	-.006(2)	-.004(2)	.000(2)
F(11)	.059(7)	.059(9)	.090(10)	-.021(7)	-.019(7)	-.024(8)
F(12)	.040(6)	.045(11)	.064(8)	.005(7)	-.014(6)	.004(8)
F(13)	.094(9)	.064(9)	.068(9)	-.040(8)	.028(8)	-.013(7)
F(21)	.049(6)	.034(7)	.077(9)	.007(6)	.011(6)	-.012(6)
F(22)	.075(8)	.083(10)	.091(10)	-.047(8)	.002(8)	-.037(8)
F(23)	.052(7)	.126(13)	.043(7)	.014(8)	-.002(6)	.004(8)

Table 7.10. Selected bond distances (Å) and angles (°) of complex (10).

PT(1) - PT(2)	2.639(1)	PT(1) - PT(3)	2.609(2)
PT(1) - SN(1)	2.746(2)	PT(1) - SN(2)	2.808(2)
PT(1) - P(1)	2.301(5)	PT(1) - P(6)	2.303(5)
PT(2) - PT(3)	2.622(1)	PT(2) - SN(1)	2.801(2)
PT(2) - SN(2)	2.699(2)	PT(2) - P(2)	2.304(6)
PT(2) - P(3)	2.313(5)	PT(3) - SN(1)	2.783(2)
PT(3) - SN(2)	2.830(2)	PT(3) - P(4)	2.278(5)
PT(3) - P(5)	2.308(5)	SN(1) - F(11)	1.995(13)
SN(1) - F(12)	1.957(11)	SN(1) - F(13)	1.901(13)
SN(2) - F(21)	1.969(11)	SN(2) - F(22)	1.965(13)
SN(2) - F(23)	1.901(12)	CL(1) - C(4)	1.72(4)
CL(2) - C(4)	1.63(4)	P(1) - C(1)	1.844(18)
P(1) - C(A1)	1.824(15)	P(1) - C(B1)	1.817(11)
P(2) - C(1)	1.844(18)	P(2) - C(C1)	1.826(14)
P(2) - C(D1)	1.856(16)	P(3) - C(2)	1.836(21)
P(3) - C(E1)	1.854(16)	P(3) - C(F1)	1.831(18)
P(4) - C(2)	1.837(19)	P(4) - C(G1)	1.829(21)
P(4) - C(H1)	1.818(16)	P(5) - C(3)	1.806(18)
P(5) - C(I1)	1.811(17)	P(5) - C(J1)	1.834(14)
P(6) - C(3)	1.852(19)	P(6) - C(K1)	1.840(15)
P(6) - C(L1)	1.844(15)		

PT(2) - PT(1) - PT(3)	59.9(1)	PT(2) - PT(1) - SN(1)	62.6(1)
PT(2) - PT(1) - SN(2)	59.3(1)	PT(2) - PT(1) - P(1)	97.6(2)
PT(2) - PT(1) - P(6)	154.3(2)	PT(3) - PT(1) - SN(1)	62.6(1)
PT(3) - PT(1) - SN(2)	62.9(1)	PT(3) - PT(1) - P(1)	156.8(2)
PT(3) - PT(1) - P(6)	94.5(2)	SN(1) - PT(1) - SN(2)	113.9(1)
SN(1) - PT(1) - P(1)	114.5(2)	SN(1) - PT(1) - P(6)	104.8(2)
SN(2) - PT(1) - P(1)	101.9(2)	SN(2) - PT(1) - P(6)	113.7(2)
P(1) - PT(1) - P(6)	108.2(2)	PT(1) - PT(2) - PT(3)	59.5(1)
PT(1) - PT(2) - SN(1)	60.5(1)	PT(1) - PT(2) - SN(2)	63.5(1)
PT(1) - PT(2) - P(2)	93.3(2)	PT(1) - PT(2) - P(3)	153.4(2)
PT(3) - PT(2) - SN(1)	61.7(1)	PT(3) - PT(2) - SN(2)	64.2(1)
PT(3) - PT(2) - P(2)	152.8(2)	PT(3) - PT(2) - P(3)	94.5(2)
SN(1) - PT(2) - SN(2)	115.7(1)	SN(1) - PT(2) - P(2)	108.3(2)
SN(1) - PT(2) - P(3)	114.3(2)	SN(2) - PT(2) - P(2)	103.9(2)
SN(2) - PT(2) - P(3)	101.9(2)	P(2) - PT(2) - P(3)	112.3(2)
PT(1) - PT(3) - PT(2)	60.6(1)	PT(1) - PT(3) - SN(1)	61.1(1)
PT(1) - PT(3) - SN(2)	62.0(1)	PT(1) - PT(3) - P(4)	156.8(2)
PT(1) - PT(3) - P(5)	96.5(2)	PT(2) - PT(3) - SN(1)	62.3(1)
PT(2) - PT(3) - SN(2)	59.2(1)	PT(2) - PT(3) - P(4)	96.4(2)
PT(2) - PT(3) - P(5)	156.8(2)	SN(1) - PT(3) - SN(2)	112.1(1)
SN(1) - PT(3) - P(4)	113.3(2)	SN(1) - PT(3) - P(5)	104.7(2)
SN(2) - PT(3) - P(4)	104.6(2)	SN(2) - PT(3) - P(5)	115.6(2)
P(4) - PT(3) - P(5)	106.7(2)	PT(1) - SN(1) - PT(2)	56.8(1)
PT(1) - SN(1) - PT(3)	56.3(1)	PT(1) - SN(1) - F(11)	163.6(4)
PT(1) - SN(1) - F(12)	102.1(4)	PT(1) - SN(1) - F(13)	102.3(4)
PT(2) - SN(1) - PT(3)	56.0(1)	PT(2) - SN(1) - F(11)	108.1(4)
PT(2) - SN(1) - F(12)	152.3(4)	PT(2) - SN(1) - F(13)	110.7(4)
PT(3) - SN(1) - F(11)	111.3(4)	PT(3) - SN(1) - F(12)	98.2(4)
PT(3) - SN(1) - F(13)	158.2(4)	F(11) - SN(1) - F(12)	89.8(5)
F(11) - SN(1) - F(13)	88.7(6)	F(12) - SN(1) - F(13)	90.1(6)
PT(1) - SN(2) - PT(2)	57.2(1)	PT(1) - SN(2) - PT(3)	55.1(1)
PT(1) - SN(2) - F(21)	105.2(4)	PT(1) - SN(2) - F(22)	155.6(4)
PT(1) - SN(2) - F(23)	107.4(4)	PT(2) - SN(2) - PT(3)	56.6(1)
PT(2) - SN(2) - F(21)	102.6(4)	PT(2) - SN(2) - F(22)	102.5(4)
PT(2) - SN(2) - F(23)	161.4(4)	PT(3) - SN(2) - F(21)	156.1(4)

Table 7.10(cont).

PT(3) - SN(2) - F(22)	103.5(4)
F(21) - SN(2) - F(22)	91.9(5)
F(22) - SN(2) - F(23)	89.2(6)
PT(1) - P(1) - C(A1)	113.9(6)
C(1) - P(1) - C(A1)	104.4(9)
C(A1) - P(1) - C(B1)	101.0(8)
PT(2) - P(2) - C(C1)	116.6(5)
C(1) - P(2) - C(C1)	104.8(7)
C(C1) - P(2) - C(D1)	102.4(9)
PT(2) - P(3) - C(E1)	118.7(8)
C(2) - P(3) - C(E1)	103.2(9)
C(E1) - P(3) - C(F1)	99.3(11)
PT(3) - P(4) - C(G1)	120.8(8)
C(2) - P(4) - C(G1)	104.5(10)
C(G1) - P(4) - C(H1)	103.9(8)
PT(3) - P(5) - C(I1)	115.7(8)
C(3) - P(5) - C(I1)	104.8(10)
C(I1) - P(5) - C(J1)	101.9(10)
PT(1) - P(6) - C(K1)	117.0(6)
C(3) - P(6) - C(K1)	105.2(8)
C(K1) - P(6) - C(L1)	104.0(8)
P(3) - C(2) - P(4)	112.9(10)
CL(1) - C(4) - CL(2)	116.4(21)

PT(3) - SN(2) - F(23)	106.9(5)
F(21) - SN(2) - F(23)	91.3(6)
PT(1) - P(1) - C(1)	109.0(6)
PT(1) - P(1) - C(B1)	123.4(4)
C(1) - P(1) - C(B1)	103.2(7)
PT(2) - P(2) - C(1)	110.9(6)
PT(2) - P(2) - C(D1)	120.0(7)
C(1) - P(2) - C(D1)	100.0(9)
PT(2) - P(3) - C(2)	108.7(7)
PT(2) - P(3) - C(F1)	121.3(8)
C(2) - P(3) - C(F1)	103.4(9)
PT(3) - P(4) - C(2)	107.6(7)
PT(3) - P(4) - C(H1)	115.1(5)
C(2) - P(4) - C(H1)	103.2(8)
PT(3) - P(5) - C(3)	108.7(6)
PT(3) - P(5) - C(J1)	122.1(6)
C(3) - P(5) - C(J1)	101.6(8)
PT(1) - P(6) - C(3)	109.4(6)
PT(1) - P(6) - C(L1)	118.7(6)
C(3) - P(6) - C(L1)	100.5(8)
P(1) - C(1) - P(2)	113.4(10)
P(5) - C(3) - P(6)	113.0(9)

Table 7.11. Crystallographic and experimental details of the structure analyses of complexes (2), (3), (7), (10) and (8)/(9).

Compound	[Pt ₄ (H)(CO) ₂ (dppm) ₄][PF ₆]	[Pt ₃ (AgCl)(S)(dppm) ₃ (AuPPh ₃)] [PF ₆]
Formula	C ₁₀₂ H ₈₉ F ₆ O ₂ P ₉ Pt ₄	C ₉₃ H ₈₁ AgAuClF ₆ P ₈ Pt ₃ S
Formula Wt.	2519.9	2518.1
Crystal habit	red plate	red plate
Crystal size, mm	0.48x0.48x0.08	0.32x0.36x0.40
Crystal system	orthorhombic	triclinic
Space group	Pcab (No. 61)	P $\bar{1}$
a, Å	21.120(6)	17.639(5)
b, Å	28.962(4)	18.651(5)
c, Å	31.026(4)	28.328(6)
α , °	—	94.83(2)
β , °	—	98.88(2)
γ , °	—	90.33(3)
Obtained from	25 refln, 12< θ <15°	22 refln, 11< θ <14°
V, Å ³	18978(6)	9173(4)
Z	8	4
F(000)	9744	4824
d calc, gcm ⁻³	1.764	1.823
T, K	295	295
μ (Mo-K α), cm ⁻¹	61.5	66.5
Absorption factors on F ²	0.81-1.43	0.76-1.57
Scan width, °	0.70	0.80
Max count time, s	100	100
Total refln measured	14235	15618
Unique refln	13174	15020
R _{INT}	0.050	0.036
Miller indicies	h	-15-15
measured,	k	-16- 9
	l	-24-24
2 θ range, °	4-46	4-36
Unique refln \geq 3 σ (I)	6261	8006
No. of parameters	385	673
R	0.038	0.061
R _w	0.045	0.079
$ \Delta\rho _{\max}$, eÅ ⁻³	0.97	2.03
Δ/σ_{\max}	0.30	0.157

Table 7.11(cont).

Compound	$[\text{Pt}_3(\text{SnMe}_2(\text{PO}_2\text{F}_2))(\text{dppm})_3][\text{PF}_6]$	$[\text{Pt}_3(\text{SnF}_3)_2(\text{dppm})_3]\cdot\text{CH}_2\text{Cl}_2$
Formula	$\text{C}_{77}\text{H}_{72}\text{F}_{80}\text{P}_8\text{Pt}_3\text{Sn}$	$\text{C}_{76}\text{H}_{68}\text{Cl}_2\text{F}_{66}\text{P}_6\text{Pt}_3\text{Sn}_2$
Formula Wt.	2133.2	2174.8
Crystal habit	purple plate	orange plate
Crystal size, mm	0.76x0.20x0.48	0.56x0.40x0.40
Crystal system	monoclinic	triclinic
Space group	C2	P $\bar{1}$
a, Å	24.808(5)	13.002(3)
b, Å	12.136(3)	15.027(3)
c, Å	26.890(6)	21.744(5)
α , °	—	83.05(2)
β , °	105.64(2)	77.29(2)
γ , °	—	78.30(2)
Obtained from	23 refln, $12 < \theta < 14^\circ$	22 refln, $11 < \theta < 16^\circ$
V, Å ³	7796(3)	4045(2)
Z	4	2
F(000)	4104	2072
d calc, gcm ⁻³	1.817	1.786
T, K	295	292
$\mu(\text{Mo-K}\alpha)$, cm ⁻¹	59.7	60.7
Absorption factors on F ²	0.80-1.43	0.69-1.39
Scan width, °	0.90	0.90
Max count time, s	90	60
Total refln measured	11055	12311
Unique refln	10528	11215
R _{INT}	0.051	0.044
Miller indicies	h	-14-14
measured,	k	-16-16
	l	-23- 3
2 θ range, °	4-54	4-46
Unique refln $\geq 3\sigma(I)$	7579	6860
No. of parameters	314	372
R	0.043	0.051
R _w	0.057	0.059
$ \Delta\rho _{\text{max}}$, eÅ ⁻³	1.35	3.25
$\Delta/\sigma_{\text{max}}$	0.12	0.195

Table 7.11(cont).

Compound	[Pt ₃ (dppm) ₃ (SnF ₃)(CO)] _{0.75} [Pt ₃ Cl(SnF ₃)(dppm) ₃] _{0.25} [PF ₆] _{0.75}
Formula	C _{75.75} H ₆₆ Cl _{0.25} F _{7.5} O _{0.75} P _{6.75} Pt ₃ Sn
Formula Wt.	2052.7
Crystal habit	orange plate
Crystal size, mm	0.18x0.32x0.36
Crystal system	cubic
Space group	P43n (No. 218)
a, Å	26.081(5)
b, Å	—
c, Å	—
α, °	—
β, °	—
γ, °	—
Obtained from	22 refln, 10<θ<14°
V, Å ³	17740(6)
Z	8
F(000)	7868
d calc, gcm ⁻³	1.537
T, K	295
μ(Mo-Kα), cm ⁻¹	52.24
Absorption factors on F ²	0.153-0.395
Scan width, °	0.90
Max count time, s	90
Total refln measured	7772
Unique refln	4644
R _{INT}	0.055
Miller indices	h
measured,	k
	l
2θ range, °	4-60
Unique refln ≥ 3σ(I)	944
No. of parameters	102
R	0.058
R _w	0.065
Δρ _{max} , eÅ ⁻³	1.95
Δ/σ _{max}	0.043

CHAPTER 8: SOME COMMENTS ON DPPM CONFORMATIONS IN $M_3(\mu\text{-DPPM})_3$ COMPLEXES

8.1 M_2P_2C Ring Conformations

In previous chapters it has been repeatedly observed that in $M_3(\mu\text{-dppm})_3$ complexes the M_2P_2C rings formed by a bridging dppm ligand attached to a pair of metal atoms show a preference for envelope conformations. The dppm methylene carbon atom defines the flap of the envelope and the M_2P_2 unit is roughly planar. We argue here that this conformation is the lowest in energy compatible with acceptable M-M bonding distances.

A survey of the structures reveals the following means and ranges for the M_2P_2C ring bond lengths and angles.

Table 8.1

		Mean	Range
M-P	Å	2.30	2.24-2.36
P-CH ₂	Å	1.84	1.83-1.85*
M-P-CH ₂	°	109	102-113
P-CH ₂ -P	°	111	108-122

* For the mean P-CH₂ in a given structure.

Although there are some outliers most of the ring angles at the phosphorus and carbon atoms are within 3° of the tetrahedral angle and we conclude that the formation of a M_2P_2C ring in $M_3(\text{dppm})_3$ species does not usually involve significant deformation of bond lengths or angles. If we take the mean distances and angles in the above table as

typical of M_2P_2C rings we can calculate the values of the two M-P-C-P torsion angles, ω_1 and ω_2 , which yield a given M-M distance. The curves in Figure 8(a) give the values of ω_1 and ω_2 which yield M-M = 2.6 Å (inner curve) and 3.0 Å (outer curve). For M-M = 2.6 Å a slightly flattened envelope conformation with $\omega_1 = 47^\circ$, $\omega_2 = -47^\circ$ (or vice versa) gives the smallest deformations from ideal staggering across P-C bonds [$|\omega| = 60^\circ$]. For rings with M-M bonds (i.e. M-M ca. 2.60 Å) the experimental points (marked as crosses) cluster around these values of ω_1 and ω_2 . The lack of any points representing a C_2 conformation with $\omega_1 = \omega_2 = \pm 23^\circ$ on the M-M = 2.6 Å line is striking. Though such a conformation is geometrically possible the deviations from staggering across P-C bonds evidently impose too great a price in conformational energy.

Figure 8(a) also gives some insight into the ability of the $M_3(\mu\text{-dppm})_3$ unit to retain its identity after rupture of two of the M-M bonds. When M-M = 3.0 Å an envelope conformation with $\omega_1 = 60^\circ$, $\omega_2 = -60^\circ$, i.e. with no P-C-P-H torsional strain, is possible. Interestingly, the points corresponding to M_2P_2C rings without M-M bonding (circles in Figure 8(a)) lie close to the M-M = 3.0 Å line although the observed M-M distances are in the range 3.0-3.5 Å. These points typically have $\omega_1 \approx -65^\circ$ and ω_2 in the range $30-50^\circ$ suggesting that rupture of an M-M bond and consequent increase in the M-M separation does not require substantial costs in M-P-C-P torsional strain.

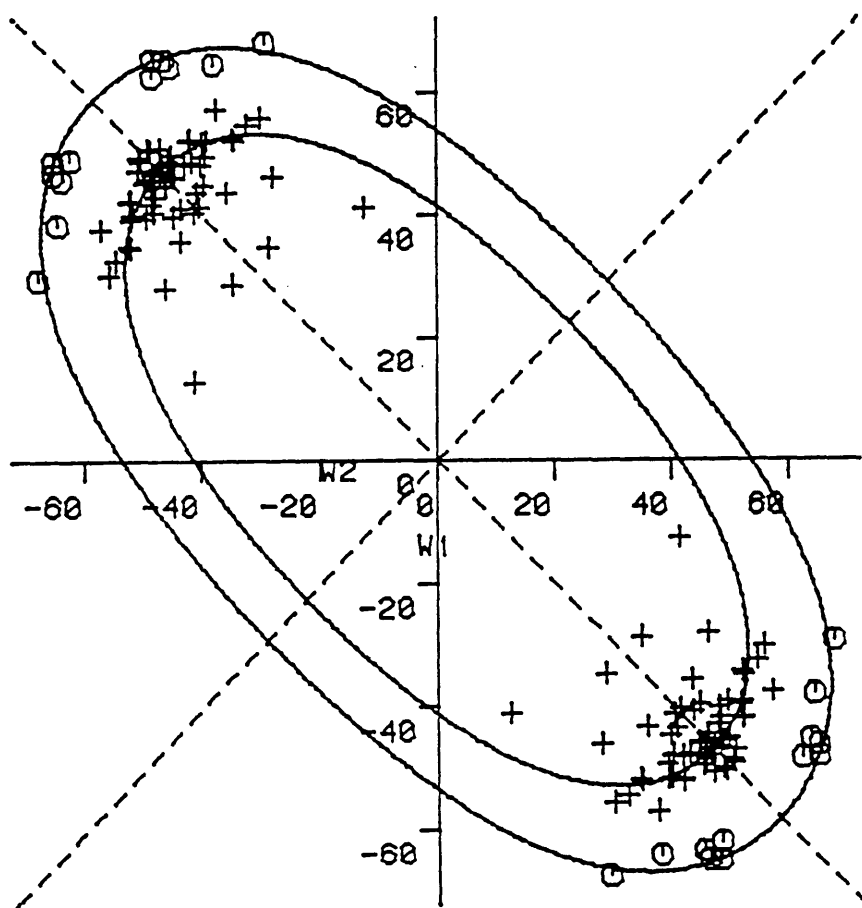


Figure 8(a). Values of P-C-P-M torsion angles($^{\circ}$), ω_1 and ω_2 , in M_2P_2C rings with $M-M = 2.6 \text{ \AA}$ (inner curve) and 3.0 \AA (outer curve). Mean distances in Table 8.1 were used to generate the curves. Crosses represent experimental torsion angles in M_2P_2C rings with M-M bonds and circles values for rings without M-M bonds. Mirror operations across the broken diagonal lines connect energetically equivalent conformations.⁸⁸

8.2 Latitudinal $M_3(\mu\text{-dppm})_3$ Conformations

In the $M_3(\mu\text{-dppm})_3$ complexes described in previous chapters electronic factors require a latitudinal and thus

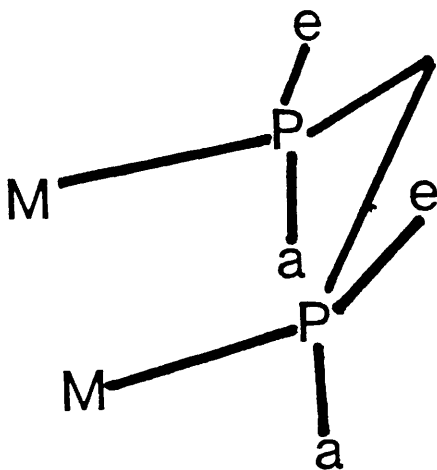


Figure 8(b). The equatorial (e) and axial (a) positions at the phosphorus atoms of the M_2P_2C ring.

roughly planar M_3P_6 core. This, together with the envelope conformation of the M_2P_2C rings, allows the dppm phenyl rings to be classed as axial (P-C bond roughly parallel with the triad axis of the M_3 triangle) or equatorial (P-C bond roughly normal to the triad axis) as shown in Figure 8(b).

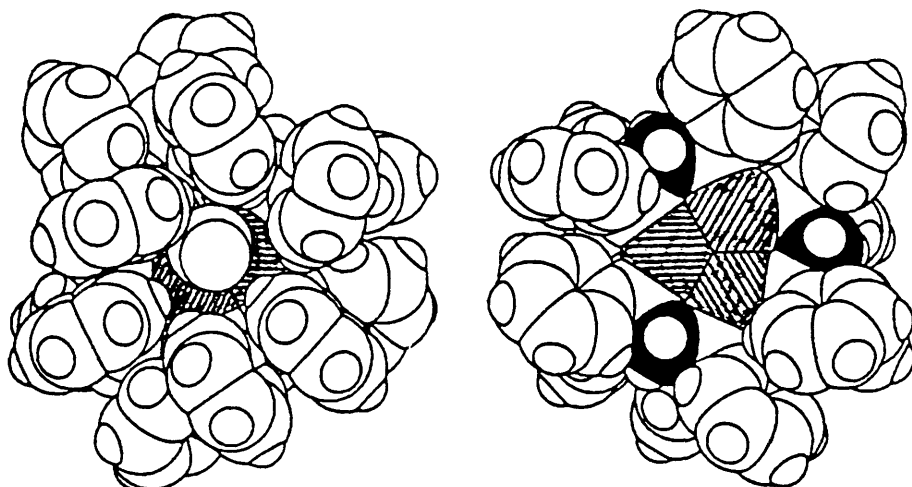


Figure 8(c). Space filled views of the cation $[Pt_3(CO)(P(OPh)_3)(dppm)_3]^{2+}$ (Chapter 5) from either side of the M_3 plane, using van der Waals radii⁵² to calculate the shells. The $P(OPh)_3$ ligand has been removed for clarity.

For the complex as a whole two arrangements are possible. All three methylene carbon atoms can lie on one side of the M_3 plane (Figure 8(c)) as do the six equatorial phenyl rings. The six axial phenyl groups point out from the other side of the plane. Alternatively, two methylene carbon atoms lie on one side of the M_3 plane and one on the other side (Figure 8(d)). In this case four equatorial and two axial phenyl groups lie on one side, with two equatorial and four axial phenyl groups on the other side.

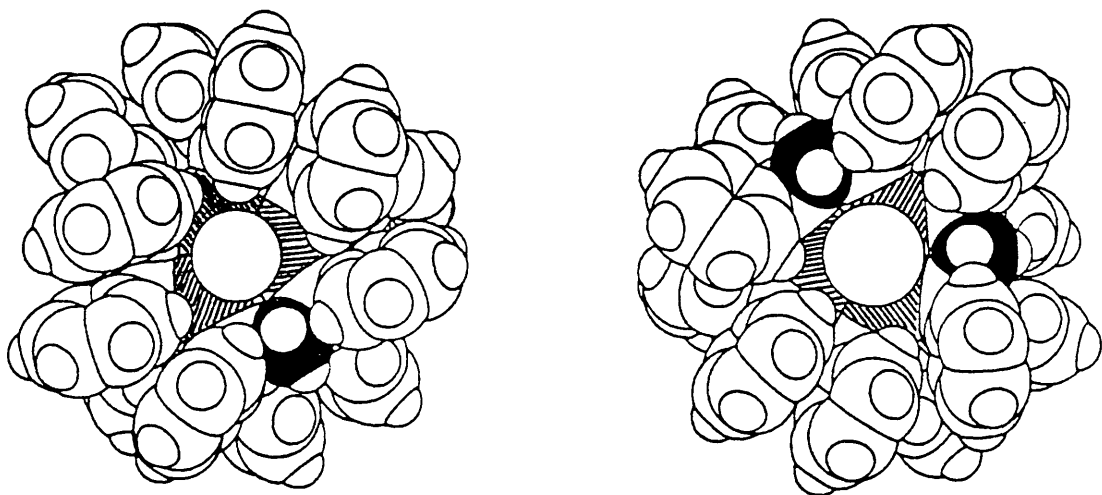


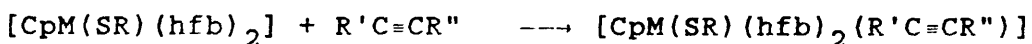
Figure 8(d). Space filled views of $[PdI(CO)(dppm)_3]^+$ ⁴⁹ from either side of the M_3 plane, using van der Waals radii to calculate the shells. ⁵²

In neither case is the same steric environment found above and below the M_3 triangle. Axial phenyl groups tend to shield the M_3 triangle and there is therefore a tendency for faces carrying bulky longitudinal ligands to have mainly or exclusively equatorial phenyl groups projecting towards the ligand.

SECTION II: DERIVATIVES OF $[(C_5H_5)MX(F_3CC_2CF_3)_2]$, M= Mo, W,
AND RELATED COMPOUNDS

CHAPTER 1: INTRODUCTORY COMMENTS ON TRANSITION METAL CATALYSIS

The bulk of the structural work presented in Section II relates to the characterisation of the products of the apparently simple addition reaction:



where $\text{M} = \text{Mo}$ or W , $\text{Cp} = \eta^5\text{-C}_5\text{H}_5$, $\text{R} = \text{alkyl}$ or aryl and $\text{hfb} = \text{F}_3\text{CC}\equiv\text{CCF}_3$, hexafluorobut-2-yne. As we shall see, the tris(alkyne) adduct formed by this reaction can adopt many different structural forms. The structural characterisation of these different species is clearly of interest in understanding transition metal mediated oligomerisation and polymerisation of alkynes. Before considering what was known of the above reaction at the inception of this work some general comments on transition metal catalysis are appropriate.

1.1 Catalysis by Transition Metals

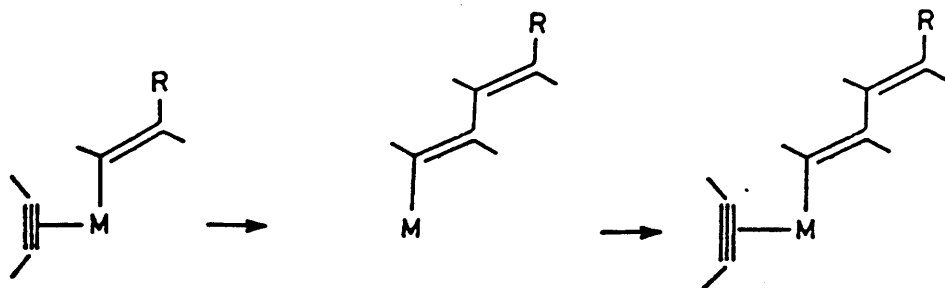
For many years transition metal complexes have been used to catalyse the conversion of unsaturated hydrocarbons obtained from natural gas or petroleum into commercially useful chemicals. Two of the major early discoveries which emphasised the commercial significance of transition metal catalysis were that of low pressure polymerisation of ethylene and propylene, giving polyalkenes and rubbers, by Zeigler and Natta and that of palladium-catalysed oxidation

of alkenes by Smidt.⁸⁹

It is now recognised that complexes which are effective in homogeneous catalysis are usually coordinatively unsaturated, so that one or more vacant sites are available to incoming ligand substrates. Catalysis often involves such processes as oxidative addition (and the reverse process, reductive elimination), insertion into the metal-ligand bonds and attacks on coordinated ligands.^{89,90}

Reactions for which transition metal complexes are effective catalysts include hydrogenation, hydroformylation, alkene isomerisation and metathesis and polymerisation/oligomerisation of alkenes and alkynes. The last of these is particularly relevant to the structural work described in later chapters.

The mechanism of alkene and alkyne polymerisation has been much discussed. For alkynes the Cossee and Arlman mechanism of polymerisation (see Scheme I)⁹¹ has been widely accepted. It involves insertion of the incoming alkyne into a σ -alkenyl-metal bond, thereby lengthening the polyene chain by two carbon atoms.



Scheme I.

Alternative mechanisms, some involving cooperative effects between two metal centres, have also been proposed.⁹²

1.2 Earlier Work on Alkyne Complexes of Molybdenum and Tungsten

Collaboration between the Crystallography Laboratory in Glasgow and Dr.J.L. Davidson of Heriot-Watt University on molybdenum and tungsten organometallics has been of long standing. The starting point of the work described below was the study, in 1979, of the addition of hfb to $[\text{CpW}(\text{SMe})(\text{CO})_3]$.⁹³ X-ray analysis showed that the adduct contained a $\text{C}(\text{CF}_3)\text{C}(\text{CF}_3)\text{C}(\text{O})\text{SMe}$ ligand attached to the metal through the central carbon atoms of the incoming hfb ligand. One W-C bond, of length $2.193(7) \text{ \AA}$, appeared to be a typical σ -W-C bond but the other, of $1.962(8) \text{ \AA}$, was much shorter, being comparable in length to W-C(CO) bonds. The resulting cyclopropene-like W=C-C metallacycle was the first example of an η^2 -vinyl ligand in organometallic chemistry.

Subsequent work⁹⁴⁻⁹⁶ was directed towards understanding the addition of nucleophiles to $[\text{CpM}(\text{hfb})_2\text{X}]$. Formally, this is a coordinatively unsaturated 16-electron species if the alkynes are viewed as two-electron donors. However, it is now well established⁹⁷ that alkynes may use between two and four π -electrons in bonding to metals so that the stability of species like $[\text{CpM}(\text{hfb})_2\text{X}]$ is less surprising. In general it was found that two-electron donors, e.g. PR_3 , added not to the metal but to coordinated

hfb to give η^2 -vinyl $C(CF_3)C(CF_3)L$ ligands. In some cases, e.g. the addition of CO to $[CpMo(hfb)_2X]$, (see also Chapter 5),⁹⁸ a metal-promoted cyclisation was triggered by the incoming ligand. In others, e.g. addition of $CNBU^t$ to $[CpM(hfb)_2X]$,⁹⁶ initial formation of an η^2 -vinyl was followed by cyclisation on addition of a second $CNBU^t$. More unusually the addition gave an η^1 -vinyl, e.g. on addition of $P(OMe)_3$ to $[WBr_2(CO)(hfb)_2]$.⁹⁵ This work established that η^2 -vinyls are often the product of reactions of molybdenum- or tungsten-alkyne complexes.

A particularly significant observation⁹⁹ was that the bis-alkyne thiolato species $[CpM(SR)(hfb)_2]$ could exist either as a bis-alkyne species (2) or as the isomeric η^2 -vinyl (1) in which an hfb has undergone internal nucleophilic attack. In general structure (2) is shown when R is aryl and (1) when R is the more electron releasing alkyl (see Figure 1(a)).

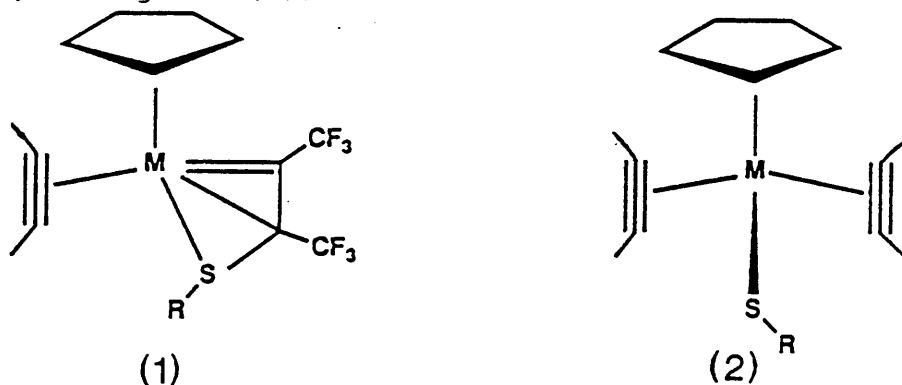


Figure 1(a). The two isomers of the bis-alkyne species $[CpM(SR)(hfb)_2]$

The addition of a third alkyne to species such as (1) or (2) is of obvious relevance to the question of alkyne oligomerisation mechanisms. Preliminary work¹⁰⁰ on the

addition of $\text{MeC}\equiv\text{CMe}$ and $\text{MeO}_2\text{CC}\equiv\text{CCO}_2\text{Me}$ (dmad) to the η^2 -vinyl species $[\text{CpW}(\text{SPr}^i)(\text{hfb})_2]$ established that these reactions are of some complexity. Addition of $\text{MeC}\equiv\text{CMe}$ gave complex (6) whereas dmad gave products (8a), (8b) and (9) ¹⁰⁰ (the numbering scheme is that of Scheme III, in Chapter 3).

Complex $[\text{CpW}(\text{SPr}^i)(\text{hfb})_2(\text{MeC}\equiv\text{CMe})]$, (6), differs from the starting complex (1) in that the incoming ligand has inserted into the $\text{Pr}^i\text{S}-\text{C}(\text{CF}_3)$ bond, and the W-S bond has subsequently ruptured, forming an η^4 -butadienyl ligand. Bond lengths and the planarity of this ligand suggest that it is attached to the tungsten atom through a $\text{W}=\text{C}(\text{carbene})$ bond, an η^2 -alkene-W interaction and a σ -W-C bond.¹⁰⁰ The η^2 -alkyne and the cyclopentadienyl ring complete the 18-electron configuration about the metal atom.

In $[\text{CpW}(\text{SPr}^i)(\text{hfb})_2(\text{dmad})]$, (8), the incoming dmad ligand has linked between the η^2 -alkyne and η^2 -vinyl units to form the chain $\text{Pr}^i\text{SC}(\text{CF}_3)\text{C}(\text{CF}_3)\text{C}(\text{CO}_2\text{Me})\text{C}(\text{CO}_2\text{Me})\text{C}(\text{CF}_3)=\text{C}-\text{CF}_3^-$. This chain is bonded to the central tungsten atom through the sulphur atom, the next four carbon atoms and also the final carbon atom of the C_6 part of the chain. Isomers (8a) and (8b) differ only by the orientation of the alkyl group on the sulphur atom. Relative to the cyclopentadienyl ring they are syn and anti respectively.

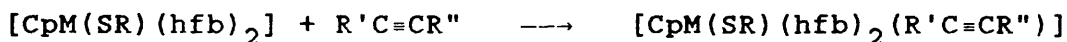
The structure of $[\text{CpW}(\text{SPr}^i)(\text{hfb})_2(\text{dmad})]$, (9), contains a chain similar to that in (8) except there has been a novel transfer of a fluorine atom (from the $-\text{CF}_3$ group attached to the final carbon of the chain) to the tungsten atom. The oligomer is bonded to the tungsten atom through the sulphur atom, a π -allylic interaction with the first

three carbon atoms and a σ -W-C(alkenyl) bond to the final carbon atom.

These structures are all consistent with the reaction of the incoming alkyne with the η^2 -vinyl unit by insertion into the W=C bond followed, in the case of (6), by migration of the thiolate group, and in (8) and (9), by insertion of the η^2 -alkyne initially present.¹⁰⁰

1.3 Strategy of Structural and Chemical Work

The preliminary results discussed above led to a more systematic study of the addition reaction:



In general three factors appear to be of importance in determining the course of this reaction.

(i) The thiolato substituent R: as previously noted, when R = Me, Et, Prⁱ and Bu^t the starting complex is an η^2 -vinyl species, (1), whereas when R = MeC₆H₄ or C₆F₅ it is a bis(alkyne) species, (2).

(ii) The nature of the metal, M = Mo or W.

(iii) The nature of the incoming alkyne: in particular RC=CMe, R = Me, Ph, gives different products from dimethylacetylenedicarboxylate (dmad).

Dr. Davidson has now completed chemical and spectroscopic studies on the products of the addition reaction and the crystal structures of several of these products are described below. The X-ray results have been of critical importance to establishing a sound structural

basis for the interpretation of the chemistry since at least twelve distinct product types can be distinguished. These often differ from each other in rather subtle ways which are difficult to establish from I.R. and N.M.R. spectra. Some of the structural results have already been published.¹⁰¹⁻¹⁰³

In the following chapters the structures of tris(alkyne) adducts are first described. The chemical significance of the structural results is then considered. In addition, results on the structures of other molybdenum and tungsten organometallics related to the tris(alkyne) adducts are also presented.

CHAPTER 2: THERMOLYSIS OF η^2 -VINYLs

2.1 Introduction

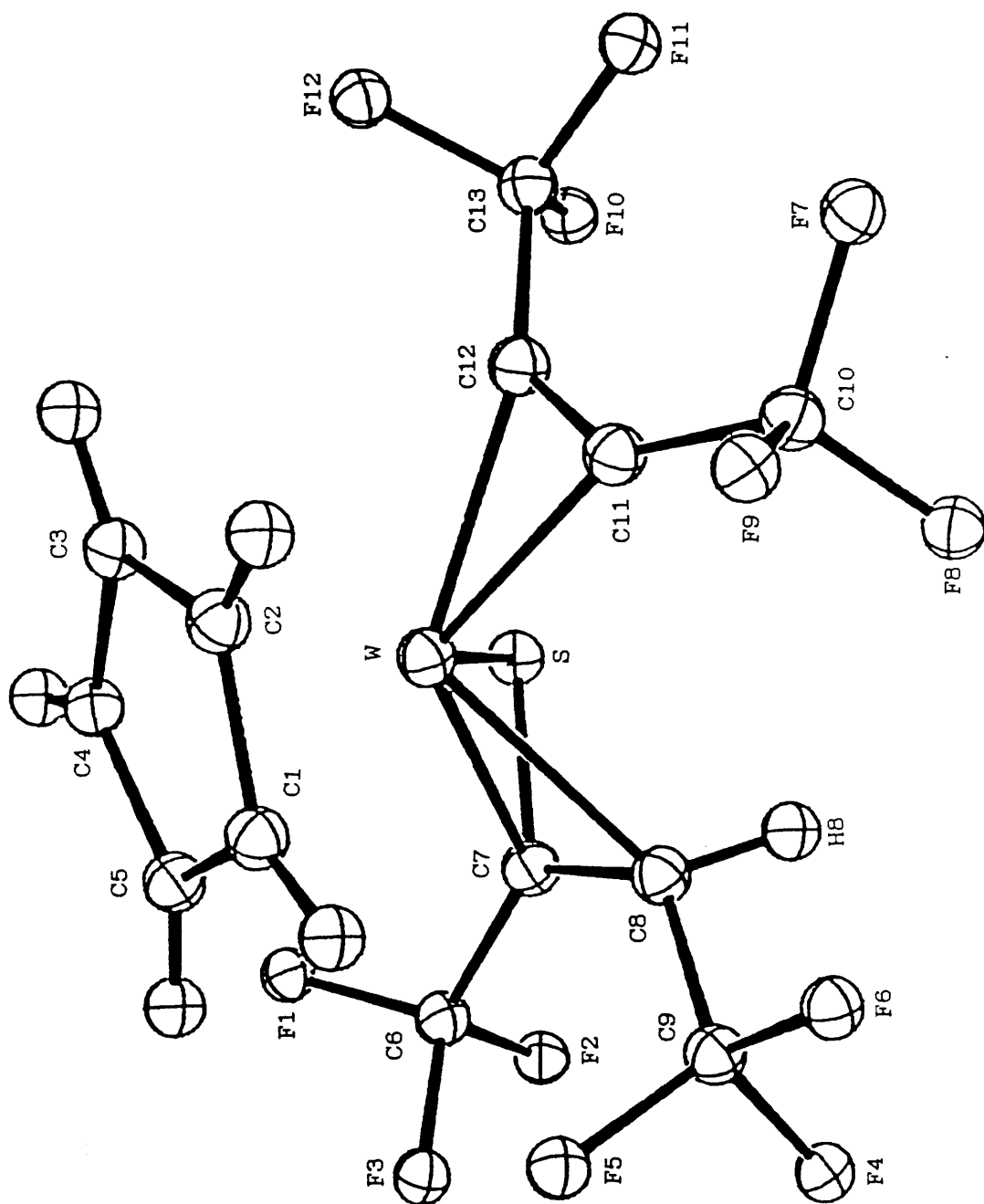
Transition metal complexes are well known for their ability to alter the stereochemistry of alkenes. A number of mechanisms have been proposed for these transformations. For isomerisation of vinyl ligands η^1 -ionic and η^2 -bonded intermediates have been suggested¹⁰⁴ and in isomerisation of butadienyl attached to palladium and rhodium a ring flip process has been proposed.^{105,106} Thermolysis of the η^2 -vinyl complex $[\text{CpM}(\text{SR})(\text{hfb})_2]$, (1),⁹⁹ $\text{M} = \text{W}$ and $\text{R} = \text{Bu}^t$ has now provided evidence of a ring flip mechanism which alters from *cis* to *trans* the arrangement of the trifluoromethyl substituents relative to the C=C double bond.

2.2 Results and Discussion

Thermolysis of (1), $\text{M} = \text{W}$, $\text{R} = \text{Bu}^t$, at 60 °C results in the formation of the complex $[\text{CpW}(\text{hfb})(\text{SC}(\text{CF}_3)=\text{C}(\text{H})\text{CF}_3)]$, (3), and the loss of butane gas. Complex (3) then isomerises to give the thermodynamically more stable complex (4). X-ray diffraction analysis was used to discover the structure of crystals of both (3) and (4). These crystals were found to contain discrete molecules of their respective isomer with contact distances comparable with van der Waal radii.⁵² Structural information for the two isomers is given in Tables 2.1 - 2.5.

The structure of complex (3) (see Figure 2(a)) confirms the rupture of the $\text{S}-\text{Bu}^t$ connection with transfer of a

Figure 2(a). A view of complex (3). All atoms are presented as spheres of arbitrary size.

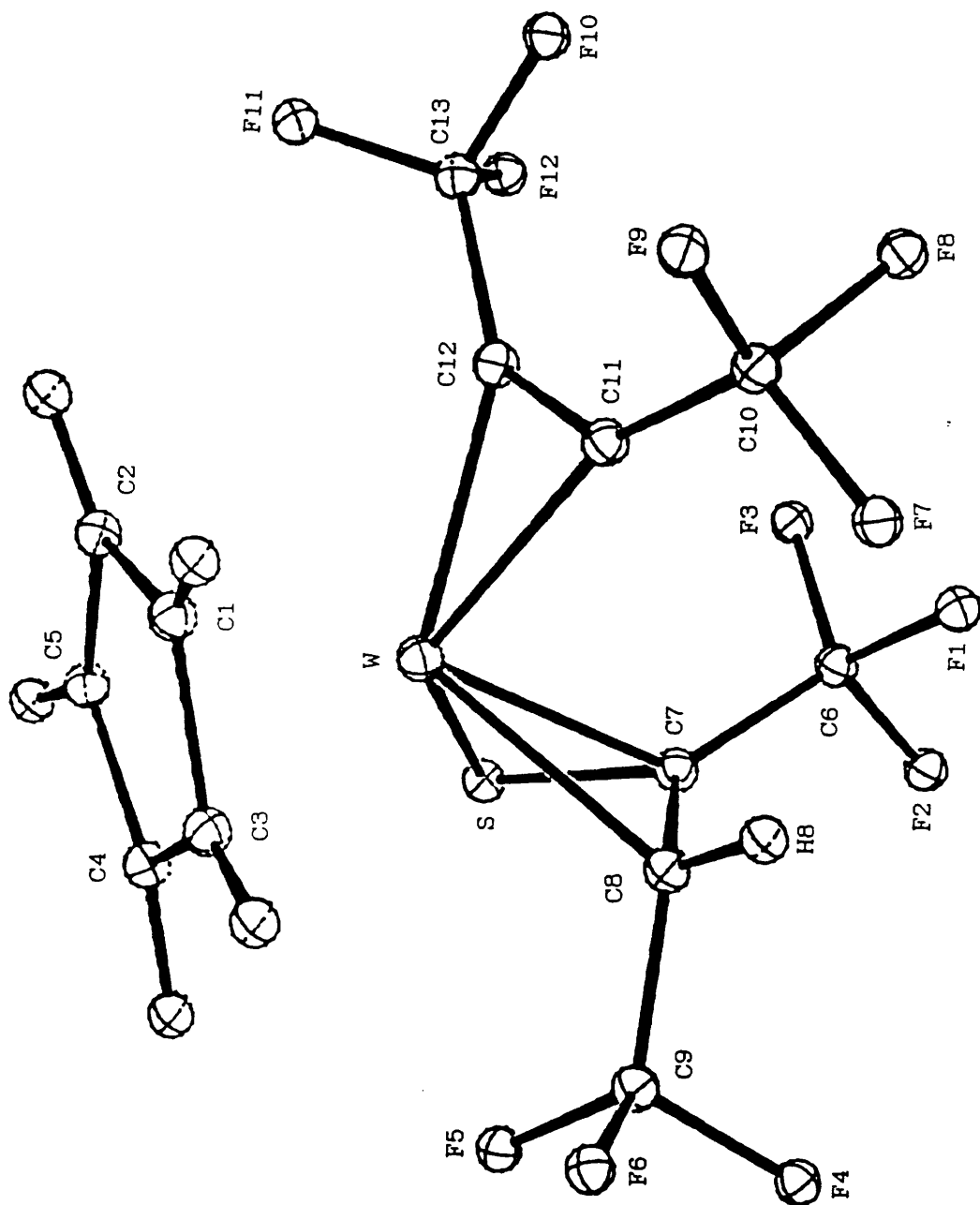


hydrogen atom to the carbon atom β to the sulphur atom of the η^2 -vinyl ligand of (1). The $\text{F}_3\text{CC}(\text{H})=\text{C}(\text{CF}_3)\text{S}^-$ ligand formed has the trifluoromethyl substituents mutually *cis*. The isomerisation to complex (4) (see Figure 2(b)) results in these substituents becoming *trans* [torsion angles $\text{C}(6)-\text{C}(7)-\text{C}(8)-\text{C}(9)$ are $-8(1)$ and $-119(1)^\circ$ respectively]. In each complex the bonding around the tungsten atom is completed by a cyclopentadienyl ring and a hexafluorobut-2-yne ligand. The $\text{CpW}(\eta^2\text{-hfb})$ geometries are almost identical in (3) and (4) (see Table 2.1).

In both isomers $\text{WSC}(7)\text{C}(8)$ forms a butterfly skeleton. However, their orientations relative to the other ligands are distinctly different: e.g. the torsion angles $\text{Cp-W-C}(7)-\text{C}(6)$ are 22 and 180° respectively for (3) and (4) (Cp is the centroid of the cyclopentadienyl ring). These different orientations can be described as the *exo* rotamer for the *cis* isomer and *endo* rotamer for the *trans*, as defined for related $\text{CpM}(\text{allyl})$ complexes.¹⁰⁷ The W-S distances in both (3) and (4) are much shorter than in (1) (see Table 2.1) while C(7) and C(8) are now nearly equidistant from the metal atom, C(8) no longer participating in a W-C multiple bond as it does in (1).

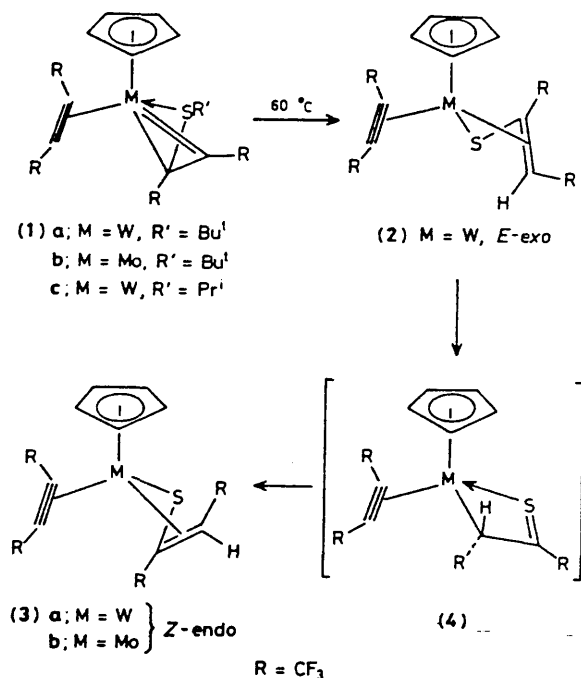
The formal electron count around the tungsten atom in (3) and (4) is 16 if the hfb ligand is considered a two electron donor. It would appear, from the shortness of the W-S distances, that these complexes are stabilised by considerable $\text{S} \rightarrow \text{W}$ π -electron donation rather than by donation from both sets of π -orbitals of the hfb ligand. There is little change in the W-C(alkyne) distances in (1), (3) and (4) [mean distances 2.08 , 2.08 and 2.07 \AA].

Figure 2(b). The structure of complex (4) with all atoms presented as spheres of arbitrary size.



respectively] and this is consistent with the above observation.

Further reactions carried out on (3) and (4) with PPh_3 and with dioxygen have given products which ^{19}F N.M.R. suggested contain an $\eta^1\text{-SC}(\text{CF}_3)=\text{C}(\text{H})\text{CF}_3$ ligand, always with retention of stereochemistry about the double bond.¹⁰¹ These results together with the fact the *exo/endo* and *cis/trans* isomerisations occur simultaneously suggests the ring flip mechanism shown in Scheme II.



Scheme II.

2.3 Experimental

cis-[CpW($\eta^3\text{-SC}(\text{CF}_3)=\text{C}(\text{H})\text{CF}_3$) (hfb)] (3)

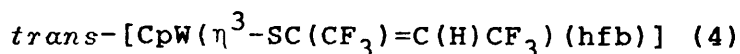
Data Collection:- A red needle of dimensions 0.52x0.16x0.16 mm was mounted on an Enraf-Nonius CAD4F diffractometer equipped with a graphite monochromator. All measurements

were made using Mo-K α X-ray radiation. The cell dimensions were obtained by a least-squares refinement of the setting angles of 23 automatically centred reflections with $13 < \theta < 18^\circ$ (see Table 2.6 for details). The intensities of 6322 reflections with $2 \leq \theta \leq 30^\circ$ were determined by continuous $\omega/2\theta$ scans of 0.80° in ω , increased by 25% at each end to allow for background. After correction for crystal decomposition (about 40% linear reduction of intensity standards), L_p and absorption effects (empirical transmission factors on F 0.87 - 1.09),⁵⁹ symmetrically equivalent reflections were merged to give 3807 unique reflections. Of these 2610 with $I \geq 3\sigma(I)$ were used subsequently.

Structure Analysis:- The structure was solved by Patterson and Fourier methods. Anisotropic displacement parameters were used for all non-hydrogen atoms. All hydrogen atom positions were determined from difference syntheses but they did not refine satisfactorily. Accordingly, the hydrogen positions were idealised with C-H distances 0.96 \AA and the hydrogen atoms were set to ride on their parent carbon atoms. Hydrogen displacement parameters were set to 1.2 times the final isotropic displacement parameters of the parent carbon atom and not refined. Final atomic parameters (see Table 2.2) were obtained from full-matrix least-squares minimisation of $\sum w(|F_o| - |F_c|)^2$ with $w^{-1} = \sigma^2 + 0.0023F^2$ (where σ is derived from counting statistics). Adjustment of 244 parameters converged with $R = 0.025$ and $R_w = 0.044$.

Neutral atom scattering factors and complex anomalous dispersion corrections were taken from ref. 61. All

calculations were performed on a GOULD-3227 computer using the GX program package.⁶⁰



For (4) and later structures in this part of the thesis data collection and structure analysis procedures closely follow those for (3) given above. Accordingly, here and in later experimental sections only variations from procedures used for (3) are explicitly mentioned.

Data Collection :- Cell dimensions and experimental parameters are given in Table 2.6. A decomposition correction was carried out: there was no decomposition in intensity standards for the first 1400 reflection but a 9% linear decrease occurred from that point to the end of data collection.

Structure Analysis:- All hydrogen atom positions were found from difference syntheses and subsequently refined with their isotropic displacement parameters. Final atomic parameters are given in Table 2.4.

Table 2.1 Comparison of bond distances in complexes (1), (3) and (4) in (Å)

Bond		(1)	(3)	(4)
W	- S	2.457(2)	2.303(2)	2.275(3)
W	- C(Cp) min	2.329(5)	2.302(7)	2.308(10)
	max	2.338(5)	2.378(7)	2.368(9)
W	- C(7)	2.178(8)	2.166(6)	2.174(8)
W	- C(8)	1.905(7)	2.211(6)	2.210(8)
W	- C(11)	2.062(9)	2.074(6)	2.082(8)
W	- C(12)	2.103(8)	2.079(7)	2.058(8)
S	- C(7)	1.764(8)	1.757(7)	1.744(9)
C(7)	- C(8)	1.417(10)	1.448(8)	1.440(12)
C(11)	- C(12)	1.300(13)	1.266(9)	1.292(11)

Table 2.2. Fractional atomic coordinates and displacement parameters (\AA^2) of the crystal structure of complex (3).

	X/A	Y/B	Z/C	U
H(1)	.11995(1)	-.09841(3)	.12249(2)	.037
S(1)	.12900(5)	-.30423(24)	.25871(12)	.054
F(1)	.2035(1)	-.0991(7)	.3530(4)	.102
F(2)	.2217(2)	-.3811(7)	.3380(4)	.105
F(3)	.2342(1)	-.1664(8)	.2457(4)	.094
F(4)	.2152(2)	-.4497(7)	.0657(4)	.103
F(5)	.2089(1)	-.1514(7)	.0457(4)	.086
F(6)	.1700(1)	-.3373(7)	-.0604(3)	.096
F(7)	.0239(2)	-.3155(9)	-.1386(4)	.116
F(8)	.0804(2)	-.4359(7)	-.1091(4)	.141
F(9)	.0683(2)	-.1649(8)	-.1718(4)	.124
F(10)	.0323(2)	-.3232(14)	.1854(6)	.210
F(11)	-.0041(2)	-.2706(13)	.0424(5)	.155
F(12)	.0111(2)	-.0612(10)	.1378(7)	.177
C(1)	.1471(3)	.1358(8)	.0412(6)	.072
C(2)	.1071(2)	.1793(8)	.0199(5)	.064
C(3)	.0999(2)	.2178(9)	.1142(5)	.061
C(4)	.1349(2)	.1968(9)	.1938(5)	.065
C(5)	.1644(2)	.1495(9)	.1475(7)	.069
C(6)	.2068(2)	-.2272(12)	.2858(5)	.064
C(7)	.1687(2)	-.2705(8)	.2089(4)	.050
C(8)	.1608(2)	-.3183(8)	.1010(4)	.045
C(9)	.1888(2)	-.3122(11)	.0412(5)	.062
C(10)	.0621(2)	-.2812(11)	-.1026(5)	.065
C(11)	.0741(2)	-.1991(8)	.0015(4)	.046
C(12)	.0608(2)	-.1700(8)	.0770(4)	.050
C(13)	.0259(2)	-.2153(13)	.1108(7)	.075
H(1)	.16025	.10202	-.00875	.113
H(2)	.08806	.18206	-.04703	.088
H(3)	.07510	.25282	.12308	.081
H(4)	.13827	.21243	.26611	.119
H(5)	.19160	.13031	.18415	.071
H(8)	.15057	-.44123	.07828	.057

Anisotropic displacement parameters (\AA^2).

	U11	U22	U33	U12	U13	U23
H(1)	.0368(2)	.0348(1)	.0360(1)	-.0012(1)	.0111(1)	-.0001(1)
S(1)	.058(1)	.058(1)	.043(1)	.000(1)	.019(1)	.010(1)
F(1)	.068(3)	.160(5)	.066(3)	-.005(3)	.003(3)	-.049(3)
F(2)	.081(4)	.135(5)	.081(3)	.025(3)	-.010(3)	.036(3)
F(3)	.045(3)	.162(5)	.068(3)	-.016(3)	.008(2)	.000(3)
F(4)	.085(4)	.128(4)	.091(3)	.063(3)	.033(3)	.013(3)
F(5)	.067(3)	.108(4)	.082(3)	-.005(3)	.036(3)	.006(2)
F(6)	.090(3)	.148(4)	.045(2)	.019(3)	.019(2)	-.013(3)
F(7)	.075(4)	.178(5)	.076(3)	-.030(4)	-.012(3)	-.032(4)
F(8)	.188(7)	.104(4)	.097(4)	.065(4)	-.021(4)	-.054(3)
F(9)	.163(6)	.141(5)	.053(3)	-.040(4)	.024(3)	.008(3)
F(10)	.076(4)	.315(10)	.232(8)	.028(5)	.071(5)	.202(8)
F(11)	.060(3)	.265(8)	.130(5)	-.062(5)	.023(4)	-.013(6)
F(12)	.139(7)	.172(7)	.227(9)	-.003(5)	.133(7)	-.034(6)
C(1)	.092(6)	.044(4)	.078(5)	-.010(3)	.043(5)	.005(3)
C(2)	.075(5)	.040(3)	.066(5)	-.003(3)	.010(4)	.013(3)
C(3)	.072(5)	.043(3)	.064(4)	.004(3)	.023(4)	.001(3)
C(4)	.080(5)	.040(3)	.063(5)	-.004(3)	.010(4)	-.010(3)
C(5)	.058(5)	.045(4)	.090(6)	-.017(3)	-.003(4)	.003(3)
C(6)	.054(4)	.088(5)	.041(4)	.006(4)	.003(3)	.001(4)
C(7)	.047(4)	.048(3)	.049(4)	.006(3)	.014(3)	.004(3)
C(8)	.045(4)	.046(3)	.038(3)	.007(3)	.009(3)	-.003(3)
C(9)	.053(5)	.075(5)	.051(4)	.021(4)	.009(3)	.005(4)
C(10)	.059(5)	.071(4)	.055(4)	0.000(4)	.005(4)	-.002(4)
C(11)	.038(3)	.048(3)	.045(4)	-.001(3)	.005(3)	.001(3)
C(12)	.048(4)	.055(4)	.043(4)	.001(3)	.015(3)	-.002(3)
C(13)	.052(5)	.086(5)	.081(6)	-.004(4)	.026(4)	-.007(5)

Table 2.3. Selected bond distances (Å) and angles (°) of complex (3).

W(1) - S(1)	2.303(2)	W(1) - C(1)	2.356(8)
W(1) - C(2)	2.378(7)	W(1) - C(3)	2.347(7)
W(1) - C(4)	2.302(7)	W(1) - C(5)	2.325(8)
W(1) - C(7)	2.166(6)	W(1) - C(8)	2.211(6)
W(1) - C(11)	2.074(6)	W(1) - C(12)	2.079(7)
S(1) - C(7)	1.757(7)	F(1) - C(6)	1.317(9)
F(2) - C(6)	1.326(10)	F(3) - C(6)	1.325(9)
F(4) - C(9)	1.327(10)	F(5) - C(9)	1.339(10)
F(6) - C(9)	1.357(8)	F(7) - C(10)	1.329(9)
F(8) - C(10)	1.293(10)	F(9) - C(10)	1.316(10)
F(10) - C(13)	1.237(13)	F(11) - C(13)	1.255(11)
F(12) - C(13)	1.312(12)	C(1) - C(2)	1.407(12)
C(1) - C(5)	1.396(12)	C(2) - C(3)	1.405(10)
C(3) - C(4)	1.397(11)	C(4) - C(5)	1.417(11)
C(6) - C(7)	1.480(9)	C(7) - C(8)	1.448(8)
C(8) - C(9)	1.464(9)	C(10) - C(11)	1.471(10)
C(11) - C(12)	1.266(9)	C(12) - C(13)	1.483(10)

S(1) - W(1) - C(7)	46.2(2)	S(1) - W(1) - C(8)	73.5(2)
S(1) - W(1) - C(11)	108.0(2)	S(1) - W(1) - C(12)	88.5(2)
C(7) - W(1) - C(8)	38.6(3)	C(7) - W(1) - C(11)	123.7(3)
C(7) - W(1) - C(12)	127.3(3)	C(8) - W(1) - C(11)	92.8(3)
C(8) - W(1) - C(12)	116.0(3)	C(11) - W(1) - C(12)	35.5(3)
W(1) - S(1) - C(7)	62.8(2)	C(2) - C(1) - C(5)	108.0(8)
C(1) - C(2) - C(3)	107.6(7)	C(2) - C(3) - C(4)	108.9(7)
C(3) - C(4) - C(5)	107.1(7)	C(1) - C(5) - C(4)	108.4(8)
F(1) - C(6) - F(2)	107.5(6)	F(1) - C(6) - F(3)	106.7(7)
F(1) - C(6) - C(7)	112.5(6)	F(2) - C(6) - F(3)	105.0(6)
F(2) - C(6) - C(7)	110.3(7)	F(3) - C(6) - C(7)	114.4(6)
W(1) - C(7) - S(1)	71.0(3)	W(1) - C(7) - C(6)	133.5(5)
W(1) - C(7) - C(8)	72.4(4)	S(1) - C(7) - C(6)	115.7(5)
S(1) - C(7) - C(8)	114.7(5)	C(6) - C(7) - C(8)	128.8(6)
W(1) - C(8) - C(7)	69.0(4)	W(1) - C(8) - C(9)	129.6(5)
C(7) - C(8) - C(9)	126.3(6)	F(4) - C(9) - F(5)	106.6(6)
F(4) - C(9) - F(6)	105.1(6)	F(4) - C(9) - C(8)	112.9(6)
F(5) - C(9) - F(6)	104.8(6)	F(5) - C(9) - C(8)	116.4(7)
F(6) - C(9) - C(8)	110.2(6)	F(7) - C(10) - F(8)	107.7(7)
F(7) - C(10) - F(9)	102.9(7)	F(7) - C(10) - C(11)	113.7(6)
F(8) - C(10) - F(9)	106.8(7)	F(8) - C(10) - C(11)	113.1(7)
F(9) - C(10) - C(11)	112.0(7)	W(1) - C(11) - C(10)	146.6(5)
W(1) - C(11) - C(12)	72.5(4)	C(10) - C(11) - C(12)	140.4(7)
W(1) - C(12) - C(11)	72.0(4)	W(1) - C(12) - C(13)	146.3(5)
C(11) - C(12) - C(13)	139.8(7)	F(10) - C(13) - F(11)	108.9(9)
F(10) - C(13) - F(12)	106.2(9)	F(10) - C(13) - C(12)	114.7(7)
F(11) - C(13) - F(12)	98.2(8)	F(11) - C(13) - C(12)	116.7(7)
F(12) - C(13) - C(12)	110.4(8)		

Table 2.4. Fractional atomic coordinates and displacement parameters (\AA^2) of the crystal structure of complex (4).

	X/A	Y/B	Z/C	U
H(1)	.42999(4)	.33720(2)	.25893(2)	.030
S(1)	.1955(3)	.2745(1)	.1702(2)	.050
F(1)	.2790(9)	.5017(4)	.0746(4)	.094
F(2)	.1018(10)	.4157(4)	-.0010(4)	.096
F(3)	.0555(10)	.4512(5)	.1378(5)	.105
F(4)	.5429(8)	.3004(4)	-.0365(3)	.085
F(5)	.5093(9)	.2089(4)	.0698(4)	.087
F(6)	.7545(7)	.2824(5)	.0725(4)	.092
F(7)	.6491(11)	.5452(4)	.1783(6)	.128
F(8)	.4524(13)	.6106(4)	.2340(7)	.138
F(9)	.6693(16)	.5695(6)	.3222(6)	.194
F(10)	.2443(23)	.5526(6)	.4109(11)	.263
F(11)	.2370(26)	.4474(9)	.4701(7)	.276
F(12)	.0690(14)	.4765(14)	.3725(11)	.268
C(1)	.6792(17)	.3206(8)	.3722(9)	.074
C(2)	.5255(15)	.2953(7)	.4154(6)	.063
C(3)	.4575(15)	.2271(7)	.3676(7)	.064
C(4)	.5634(16)	.2072(7)	.2941(7)	.066
C(5)	.7034(14)	.2673(7)	.2959(7)	.065
C(6)	.1856(14)	.4326(6)	.0835(6)	.058
C(7)	.3006(10)	.3595(5)	.1178(5)	.040
C(8)	.4946(11)	.3545(5)	.1109(5)	.041
C(9)	.5728(12)	.2861(7)	.0564(6)	.057
C(10)	.5630(15)	.5457(7)	.2552(7)	.066
C(11)	.4645(10)	.4679(5)	.2719(5)	.040
C(12)	.3399(10)	.4436(5)	.3249(5)	.040
C(13)	.2294(16)	.4789(8)	.3950(8)	.077
H(1)	.732(10)	.354(4)	.384(5)	.01(2)
H(2)	.497(13)	.320(5)	.476(6)	.07(3)
H(3)	.366(17)	.184(8)	.378(9)	.12(5)
H(4)	.558(16)	.162(7)	.235(7)	.11(5)
H(5)	.803(11)	.274(5)	.261(5)	.04(2)
H(8)	.565(9)	.406(4)	.099(4)	.02(2)

Anisotropic displacement parameters (\AA^2).

	U11	U22	U33	U12	U13	U23
H(1)	.0303(2)	.0300(2)	.0308(2)	.0035(1)	.0029(1)	.0049(1)
S(1)	.0391(10)	.0448(13)	.0641(13)	-.0073(9)	-.0017(10)	-.0026(10)
F(1)	.114(5)	.046(4)	.117(5)	.005(3)	-.040(4)	.029(3)
F(2)	.126(5)	.079(4)	.076(4)	.023(4)	-.062(4)	-.008(3)
F(3)	.096(5)	.115(6)	.103(5)	.068(4)	.004(4)	.018(4)
F(4)	.104(4)	.116(5)	.035(3)	.010(4)	.009(3)	-.011(3)
F(5)	.119(5)	.056(4)	.088(4)	-.002(4)	.042(4)	-.022(3)
F(6)	.061(3)	.145(6)	.072(4)	.022(4)	.013(3)	-.033(4)
F(7)	.171(8)	.080(5)	.138(6)	-.061(5)	.087(6)	-.018(4)
F(8)	.178(8)	.040(4)	.197(9)	.002(5)	.035(7)	.020(5)
F(9)	.271(12)	.193(9)	.109(6)	-.191(9)	-.074(7)	.046(6)
F(10)	.385(20)	.085(7)	.337(17)	-.034(9)	.279(17)	-.096(9)
F(11)	.447(24)	.265(15)	.128(8)	.225(15)	.188(13)	.072(9)
F(12)	.091(7)	.473(27)	.242(15)	.046(11)	.051(8)	-.187(17)
C(1)	.057(7)	.070(9)	.091(9)	.013(6)	-.035(7)	.019(7)
C(2)	.080(7)	.073(8)	.036(5)	.020(6)	.001(5)	.013(5)
C(3)	.076(7)	.055(7)	.059(6)	.015(6)	.008(5)	.025(5)
C(4)	.086(7)	.055(7)	.057(6)	.030(6)	-.003(5)	.007(5)
C(5)	.054(6)	.087(8)	.054(6)	.036(6)	.011(5)	.023(6)
C(6)	.071(6)	.053(6)	.048(5)	.011(5)	-.024(5)	.000(4)
C(7)	.043(4)	.042(5)	.037(4)	-.001(3)	.000(3)	.000(3)
C(8)	.045(4)	.048(5)	.029(4)	.000(4)	.006(3)	.002(3)
C(9)	.055(5)	.074(7)	.042(5)	.002(5)	.006(4)	-.004(5)
C(10)	.084(7)	.047(6)	.067(7)	-.014(5)	.012(6)	-.007(5)
C(11)	.047(4)	.028(4)	.045(4)	-.005(3)	-.003(4)	.002(3)
C(12)	.048(4)	.036(5)	.036(4)	.010(4)	-.001(4)	-.005(3)
C(13)	.071(7)	.082(9)	.077(7)	.017(6)	.024(6)	-.026(6)

Table 2.5. Selected bond distances (Å) and angles (°) of complex (4).

H(1) - S(1)	2.275(3)	H(1) - C(1)	2.349(10)
H(1) - C(2)	2.368(9)	H(1) - C(3)	2.320(10)
H(1) - C(4)	2.308(10)	H(1) - C(5)	2.314(9)
H(1) - C(7)	2.174(8)	H(1) - C(8)	2.210(8)
H(1) - C(11)	2.082(8)	H(1) - C(12)	2.058(8)
S(1) - C(7)	1.744(9)	F(1) - C(6)	1.299(11)
F(2) - C(6)	1.329(10)	F(3) - C(6)	1.311(12)
F(4) - C(9)	1.339(10)	F(5) - C(9)	1.321(11)
F(6) - C(9)	1.337(11)	F(7) - C(10)	1.308(11)
F(8) - C(10)	1.327(13)	F(9) - C(10)	1.238(11)
F(10) - C(13)	1.187(14)	F(11) - C(13)	1.173(14)
F(12) - C(13)	1.196(15)	C(1) - C(2)	1.389(16)
C(1) - C(5)	1.394(16)	C(1) - H(1)	0.67(7)
C(2) - C(3)	1.345(15)	C(2) - H(2)	0.98(9)
C(3) - C(4)	1.389(15)	C(3) - H(3)	0.97(13)
C(4) - C(5)	1.398(15)	C(4) - H(4)	1.10(11)
C(5) - H(5)	0.92(8)	C(6) - C(7)	1.487(12)
C(7) - C(8)	1.440(12)	C(8) - C(9)	1.472(13)
C(8) - H(8)	0.98(7)	C(10) - C(11)	1.454(12)
C(11) - C(12)	1.292(11)	C(12) - C(13)	1.448(12)

S(1) - H(1) - C(7)	46.1(3)	S(1) - H(1) - C(8)	74.9(3)
S(1) - H(1) - C(11)	124.1(2)	S(1) - H(1) - C(12)	110.4(3)
C(7) - H(1) - C(8)	38.3(3)	C(7) - H(1) - C(11)	87.8(3)
C(7) - H(1) - C(12)	99.0(3)	C(8) - H(1) - C(11)	85.9(4)
C(8) - H(1) - C(12)	115.9(4)	C(11) - H(1) - C(12)	36.4(3)
H(1) - S(1) - C(7)	63.9(3)	C(2) - C(1) - C(5)	109.4(12)
C(2) - C(1) - H(1)	126.3(71)	C(5) - C(1) - H(1)	124.3(70)
C(1) - C(2) - C(3)	106.8(11)	C(1) - C(2) - H(2)	121.1(57)
C(3) - C(2) - H(2)	131.2(57)	C(2) - C(3) - C(4)	110.6(11)
C(2) - C(3) - H(3)	135.5(79)	C(4) - C(3) - H(3)	113.1(79)
C(3) - C(4) - C(5)	107.0(10)	C(3) - C(4) - H(4)	137.2(63)
C(5) - C(4) - H(4)	115.5(63)	C(1) - C(5) - C(4)	106.3(11)
C(1) - C(5) - H(5)	120.1(50)	C(4) - C(5) - H(5)	133.4(50)
F(1) - C(6) - F(2)	106.8(8)	F(1) - C(6) - F(3)	106.7(9)
F(1) - C(6) - C(7)	113.0(8)	F(2) - C(6) - F(3)	105.7(8)
F(2) - C(6) - C(7)	110.4(8)	F(3) - C(6) - C(7)	113.7(8)
H(1) - C(7) - S(1)	70.0(3)	H(1) - C(7) - C(6)	128.4(6)
H(1) - C(7) - C(8)	72.2(5)	S(1) - C(7) - C(6)	118.4(7)
S(1) - C(7) - C(8)	117.6(7)	C(6) - C(7) - C(8)	124.0(8)
H(1) - C(8) - C(7)	69.5(5)	H(1) - C(8) - C(9)	122.3(6)
H(1) - C(8) - H(8)	114.9(38)	C(7) - C(8) - C(9)	120.6(8)
C(7) - C(8) - H(8)	120.4(38)	C(9) - C(8) - H(8)	106.2(39)
F(4) - C(9) - F(5)	105.4(8)	F(4) - C(9) - F(6)	104.7(8)
F(4) - C(9) - C(8)	110.9(8)	F(5) - C(9) - F(6)	107.2(9)
F(5) - C(9) - C(8)	116.1(8)	F(6) - C(9) - C(8)	111.8(8)
F(7) - C(10) - F(8)	98.3(9)	F(7) - C(10) - F(9)	109.0(10)
F(7) - C(10) - C(11)	114.5(9)	F(8) - C(10) - F(9)	106.0(10)
F(8) - C(10) - C(11)	112.6(9)	F(9) - C(10) - C(11)	114.9(9)
H(1) - C(11) - C(10)	151.5(7)	H(1) - C(11) - C(12)	70.8(5)
C(10) - C(11) - C(12)	137.1(8)	H(1) - C(12) - C(11)	72.8(5)
H(1) - C(12) - C(13)	147.9(8)	C(11) - C(12) - C(13)	138.9(9)
F(10) - C(13) - F(11)	104.2(14)	F(10) - C(13) - F(12)	98.9(14)
F(10) - C(13) - C(12)	117.3(12)	F(11) - C(13) - F(12)	101.5(16)
F(11) - C(13) - C(12)	118.4(12)	F(12) - C(13) - C(12)	113.6(11)

Table 2.6. Crystallographic and experimental details of the structure analyses of complexes (3) and (4).

Compound	<i>cis</i> -[CpW(SC(CF ₃)CCF ₃)(hfb)] (3)	<i>trans</i> -[CpW(SC(CF ₃)CCF ₃)(hfb)] (4)
Formula	C ₁₃ H ₆ F ₁₂ SW	C ₁₃ H ₆ F ₁₂ SW
Formula Wt.	606.1	606.1
Crystal habit	red needle	red plate
Crystal size, mm	0.52x0.16x0.16	0.20x0.15x0.35
Crystal system	monoclinic	monoclinic
Space group	C 2/c	P 2 ₁ /n
a, Å	35.551(5)	7.353(3)
b, Å	7.072(1)	15.764(12)
c, Å	13.548(2)	14.205(5)
β, °	107.00(1)	94.81(3)
Obtained from	25 refln, 13<θ<18°	25 refln, 11<θ<14°
V, Å ³	3282(1)	1641(2)
Z	8	4
F(000)	2256	1128
d calc, gcm ⁻³	2.453	2.454
T, K	297	295
μ(Mo-Kα), cm ⁻¹	74.3	74.3
Absorption factors on F ²	0.87-1.09	0.81-1.39
Scan width, °	0.95	0.96
Max count time, s	90	120
Total refln measured	6322	5203
Unique refln	3807	2886
R _{INT}	0.024	0.025
Miller indices	h	0-8
measured,	k	-18-18
	l	-16-16
2θ range, °	4-60	4-50
Unique refln ≥ 3σ(I)	2610	2167
No. of parameters	244	268
R	0.025	0.027
R _w	0.031	0.034
Δρ _{max} , eÅ ⁻³	0.86	0.97
Δ/σ _{max}	0.024	0.030

CHAPTER 3: STRUCTURES OF $[\text{CpM}(\text{SR})(\text{hfb})_2(\text{L-L})]$ COMPLEXES

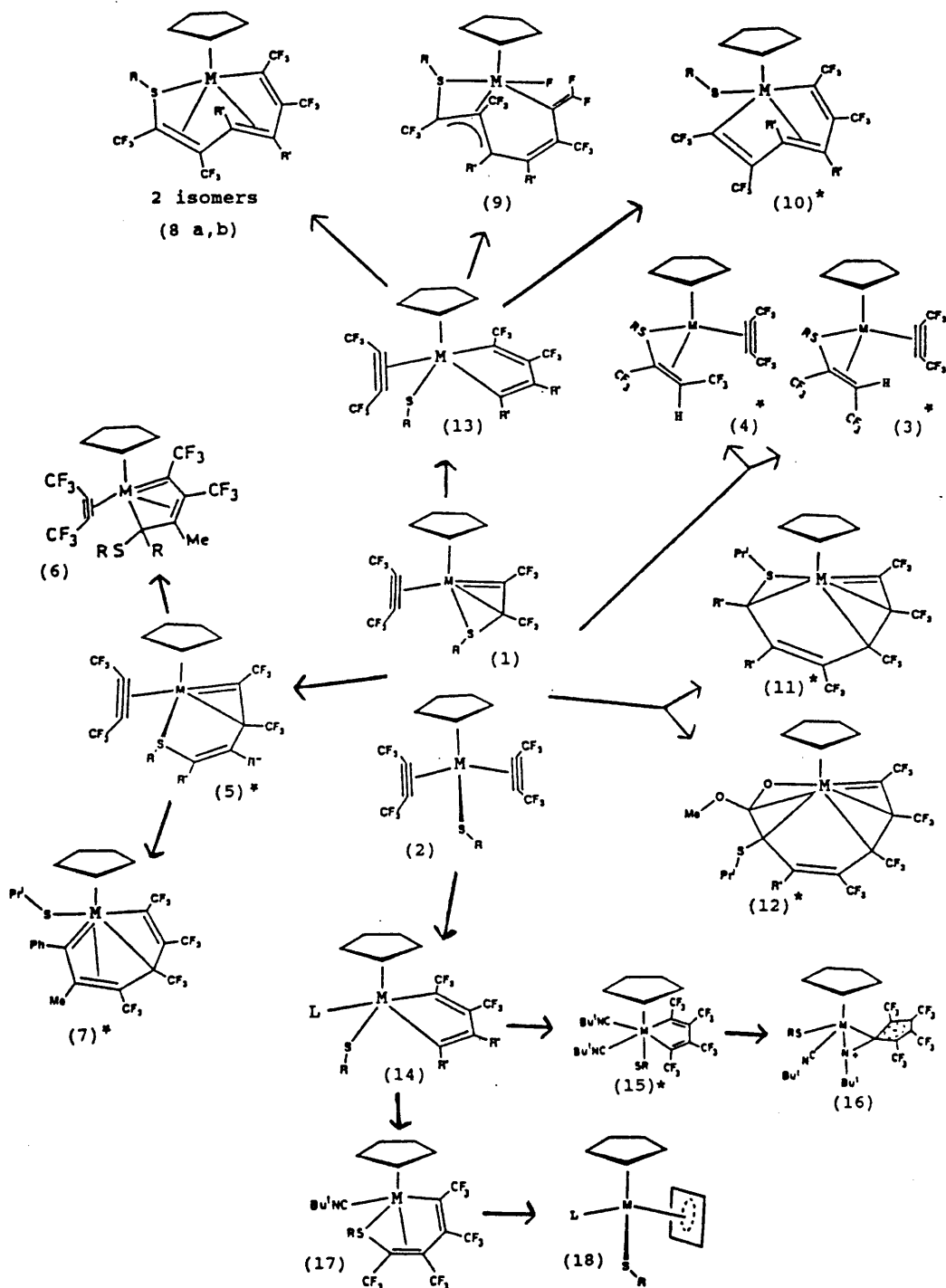
3.1 Introduction

In this chapter the structures are presented of six complexes of general formula $[\text{CpM}(\text{SR})(\text{hfb})_2(\text{L-L})]$, where L-L is either a third alkyne ligand or two isocyanide ligands. The molecules discussed are:

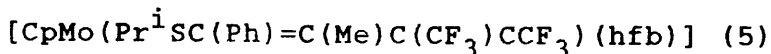
	M	R	L-L
(5)	Mo	Pr^i	$\text{PhC}\equiv\text{CMe}$
(7)	W	Pr^i	$\text{PhC}\equiv\text{CMe}$
(10)	W	$4\text{-MeC}_6\text{H}_4$	dmad
(11)	Mo	Pr^i	dmad
(12)	Mo	Pr^i	dmad
(15)	W	Pr^i	$(4\text{-MeC}_6\text{H}_4\text{NC})_2$

The formulation of each compound, based on its X-ray analysis is shown in Scheme III and the molecular geometries are compared in Table 3.1. The structures are first considered individually (Section 3.2) and the chemical implications of the structural results are then presented (Sections 3.3 and 3.4).

Scheme III.



3.2 Description of Structures

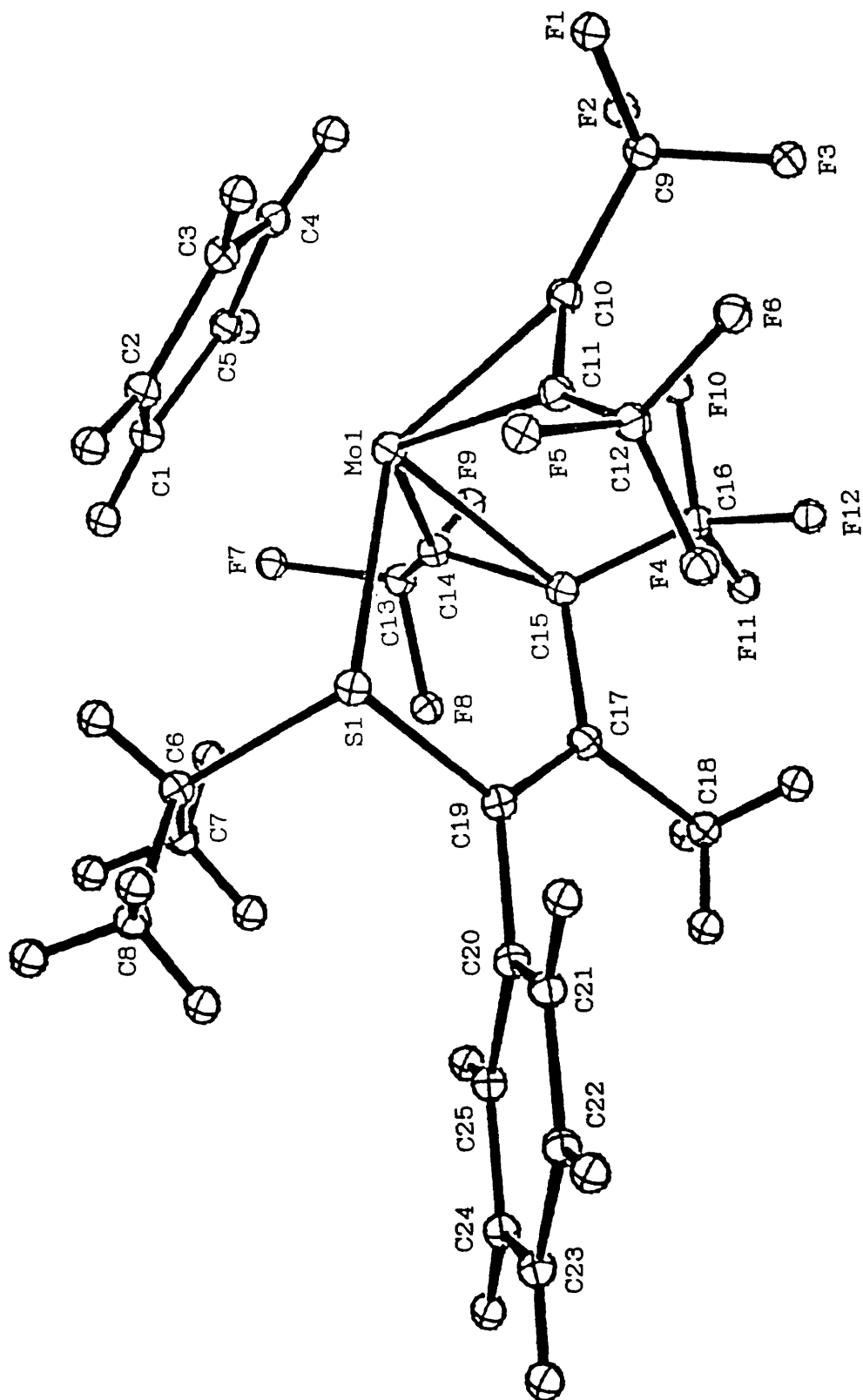


The structure of (5) (see Figure 3(a)) was determined using X-ray diffraction analysis. The crystal used was shown to contain discrete molecules of complex (5) with intermolecular distances consistent with the appropriate sum of van der Waal radii.⁵² Complex (5) differs formally from the starting complex (1) only by the insertion of $\text{PhC}=\text{CMe}$ into the S-C bond of the $\eta^3\text{-Pr}^i\text{SC}(\text{CF}_3)\text{CCF}_3$ ligand of (1).⁹⁹ A comparison of bonding distances between the two structures (see Tables 2.1 and 3.1) shows their remarkable similarity despite the different metal atom in each and the additional alkyne unit in complex (5).

In (5) the Mo-C(14) and Mo-C(15) distances [1.915(5) and 2.239(5) Å respectively] show only minor differences to the corresponding M-C bonds of the η^2 -vinyl ligand in (1), the former clearly showing multiple bonding character and the latter obviously single in nature. The C(14)-C(15) bond length [1.426(7) Å] is short for a single C-C bond,¹⁰⁸ but is similar to that in (1) of 1.417(10) Å. This appears normal for η^2 -vinyllic bonding and is probably due to slight delocalisation of the Mo-C(14) π -bonding. The M-S distances are also closely comparable.

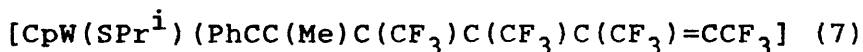
The C(15)-C(17), C(19)-C(17) and C(19)-S bond distances of 1.514(7), 1.338(7) and 1.798(6) Å and a S-C(19)-C(17)-C(15) torsion angle of $-4.1(5)^\circ$ suggest strongly that C(19)-C(17) is a localised double bond connected to atoms C(15) and S by single bonds. Atoms C(19) and C(17) are not

Figure 3(a). A view of complex (5) with all atoms presented as spheres of arbitrary size.



bonded to the molybdenum atom (see Table 3.1).

The Mo-C and C-C bond lengths in the cyclopentadienyl ring agree well with those in other structures in this series (see Table 3.1). The bonding of the η^2 -hfb ligand to the Mo atom is similar to that of other such 2-electron donor ligands.^{94,99} All other bonds and angles also appear unexceptional (see Tables 3.1 and 3.3), the main point of note being the wide range of C-F bond distances [1.25(1)-1.35(1) Å]. The fluorine atoms involved in the shortest of these bonds all have large thermal displacement parameters which could be caused by disorder of their atomic positions.

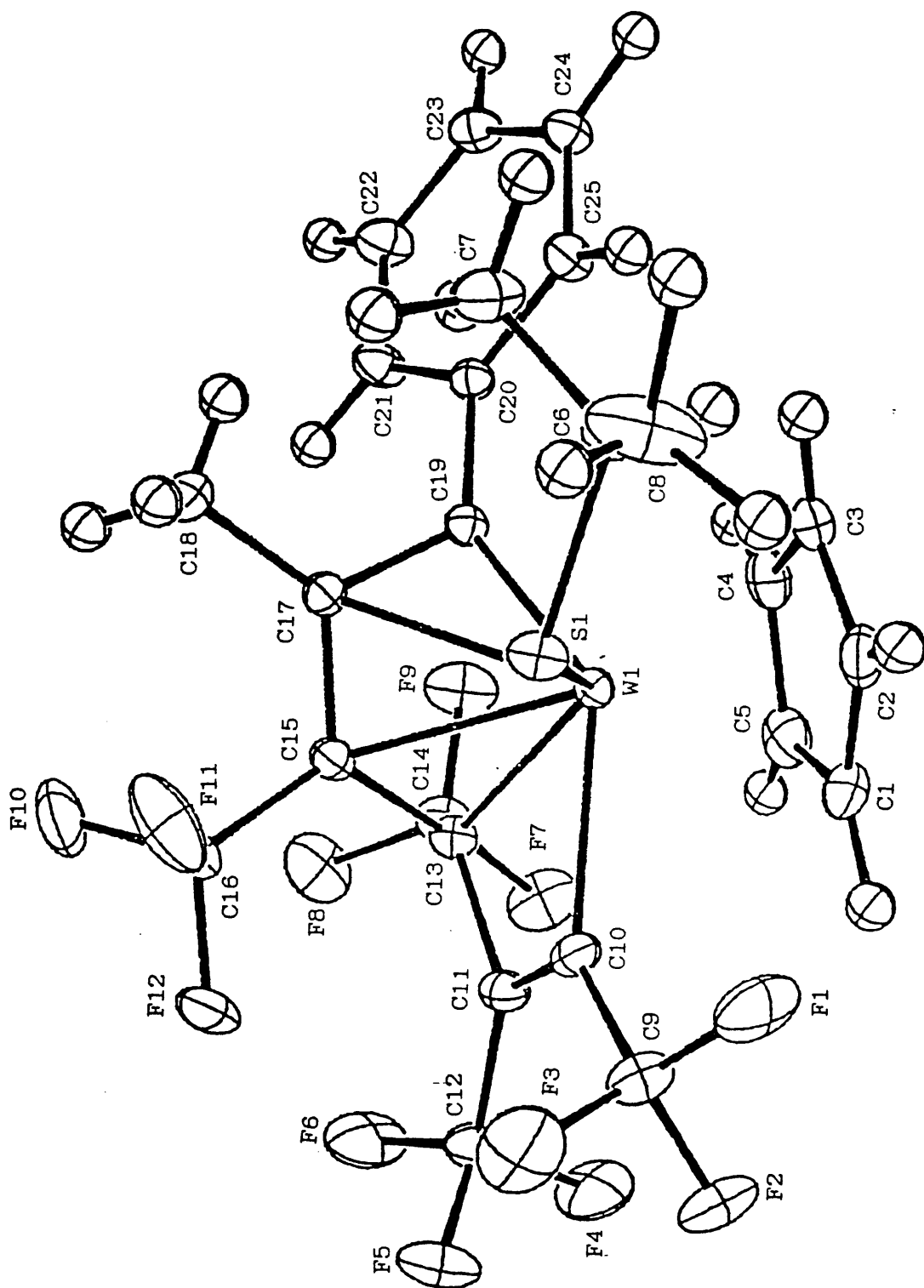


X-ray analysis of a crystal of compound (7) showed it to contain well separated molecules. The tungsten atom in complex (7) has bonded to it a cyclopentadienyl ring, the thiolate group, SPr^i , and a carbon chain consisting of the three alkyne units. The thiolate group is not bonded to this oligomer (see Figure 3(b)).

In complex (7) the M-S distance is clearly shorter than in complex (5). This is probably due to the sulphur being only mono-substituted whereas in complex (5) it was attached to the carbon chain as well as to an alkyl group. There is a wide variation in C-H bond distances. However, this is not unusual when hydrogen atom positions are refined.

The carbon oligomer is bonded to the tungsten atom through five of the six carbon atoms in the chain (see Table 3.1). The C(13)-C(15)-C(17)-C(19) section is a

Figure 3(b). The structure of complex (7). 50% probability ellipsoids are displayed for all atoms.



roughly planar η^4 -butadienyl unit [torsion angle $-22.4(3)^\circ$]. W-C(19) is a multiple bond, length $2.017(3) \text{ \AA}$, which is longer than the similar bond in complex (5) but is clearly much shorter than a M-C single bond. The bonding from atoms C(17) and C(15) to the tungsten atom is similar to that of an η^2 -alkene, as in complexes (3) and (4). However, the C(17)-C(19) and C(15)-C(17) bonds are both much shorter than C-C single bonds, suggesting some delocalisation of charge over atoms C(15), C(17) and C(19). The W-C(13) bond completes the bonding of the η^4 -unit and appears single in nature, as does the C(13)-C(15) bond (see Tables 3.1 and 3.5).

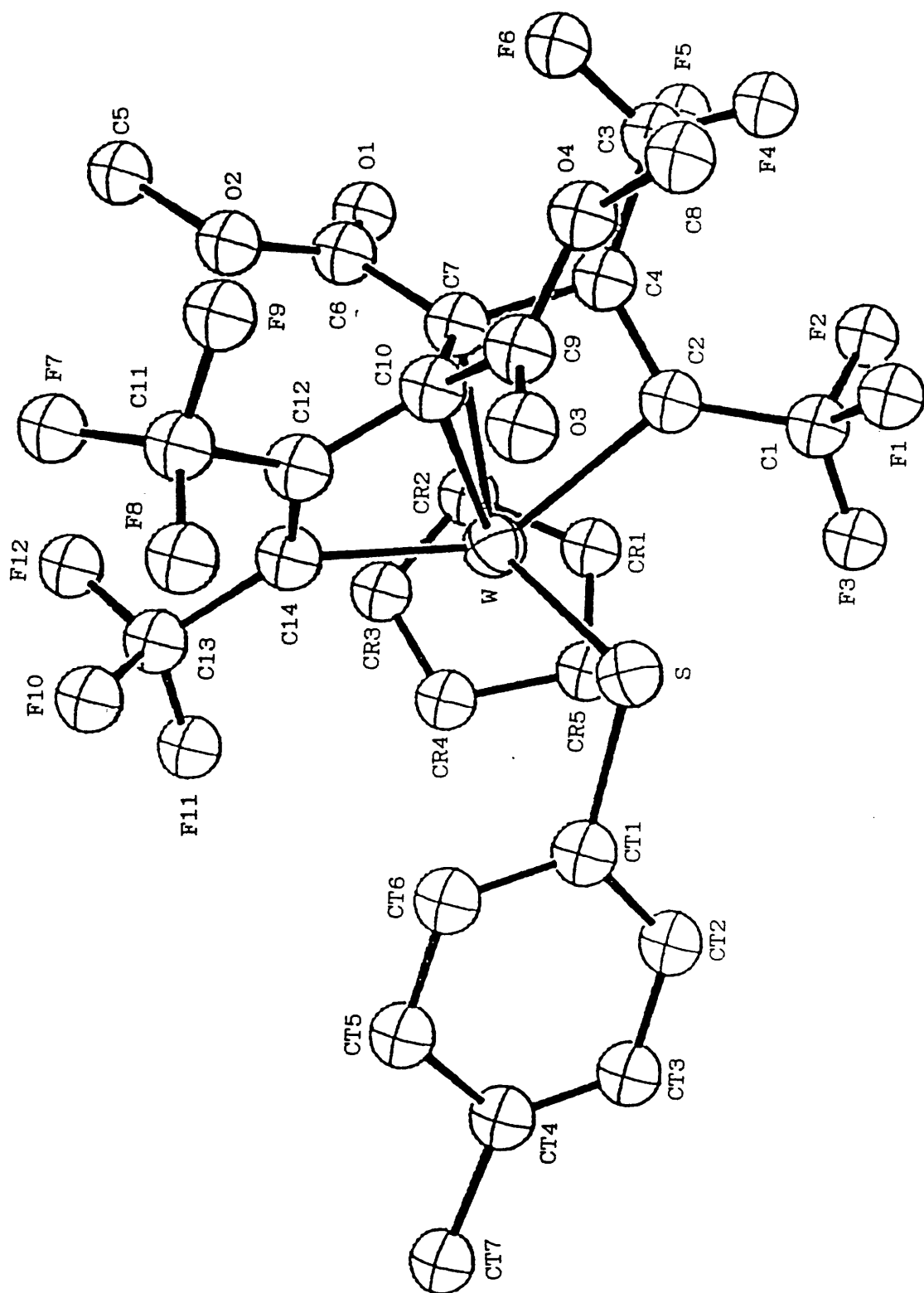
The final C_2 unit of the oligomer is bonded to the η^4 -butadienyl with a C(11)-C(13) single bond and to the metal centre through a W-C(alkenyl) single bond. The torsion angle C(9)-C(10)-C(11)-C(12) of $-4.9(4)^\circ$ and a bond length of $1.321(5) \text{ \AA}$ suggests that C(10)-C(11) is a localised double bond.

[CpW(STol)(F₃CC=C(CF₃)C(CO₂Me)=C(CO₂Me)C(CF₃)=CCF₃)] (10)

X-ray analysis of a crystal of complex (10) showed it to be made up of discrete molecules of separated by distances consistent with those of the appropriate van der Waals radii.⁵²

Complex (10) consists of a cyclopentadienyl ring, a 4-methylbenzenethiolate ligand and a carbon chain constructed from the three alkyne units, with the added dmad again separating the two hfb ligands. These ligands are all attached to the central tungsten atom (see Figure 3(c) and Tables 3.1 and 3.7).

Figure 3(c). The structure of complex (10). Hydrogen atoms are omitted. All atoms shown are displayed as spheres of arbitrary size.

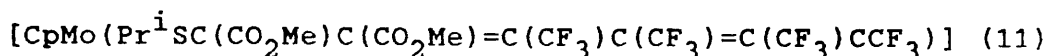


The oligomer appears to be bonded to the tungsten atom through two M-C σ -bonds from the end carbons C(2) and C(14). The bond lengths [2.24(2) and 2.12(2) Å] are consistent with other such bonds (e.g. W-C(10) 2.242(4) Å in complex (7)). The chain is also bonded through an η^2 -alkenyl interaction from atoms C(7) and C(10). The other carbon atoms C(4) and C(12) are clearly not bonded to the tungsten atom [distances are 2.79(2) and 2.67(2) Å respectively].

Bonds C(2)-C(4) and C(12)-C(14) with lengths of 1.35(3) and 1.34(3) Å respectively, appear to be localised double bonds. Distances of 1.50(3) Å for bond C(4)-C(7) and 1.47(2) Å for bond C(10)-C(12) suggest that they are single in nature. The C(7)-C(10) bond length is 1.44(3) Å and is consistent with other η^2 -M-C(alkene) interactions e.g. in complexes (2) and (3) (see Table 2.1). The three C-C-C chain torsion angles [66(2), -166(2) and 51(2)°] seem to confirm this assessment and suggests that the C(7)-C(10) double bond is *trans* in nature. This type of $M(CR)_6$ bonding has previously been observed for $Ni(CCF_3)_6$.¹⁰⁹ It is only the different nature of the chain bonding which distinguishes compound (10) from the formally analogous complex (7).

The major point of note about the geometry of the other ligands is that the W-S distance [2.318(6) Å] is much shorter than the value for the monosubstituted thiolate group in complex (7). This could be due to the more electron withdrawing nature of the 4-MeC₆H₄ substituent compared to isopropyl, but is more likely to be due to stabilisation of the 16-electron tungsten atom by π -electron donation from the sulphur atom. Of the compounds

discussed here (10) is the only one not to contain an 18-electron configuration.

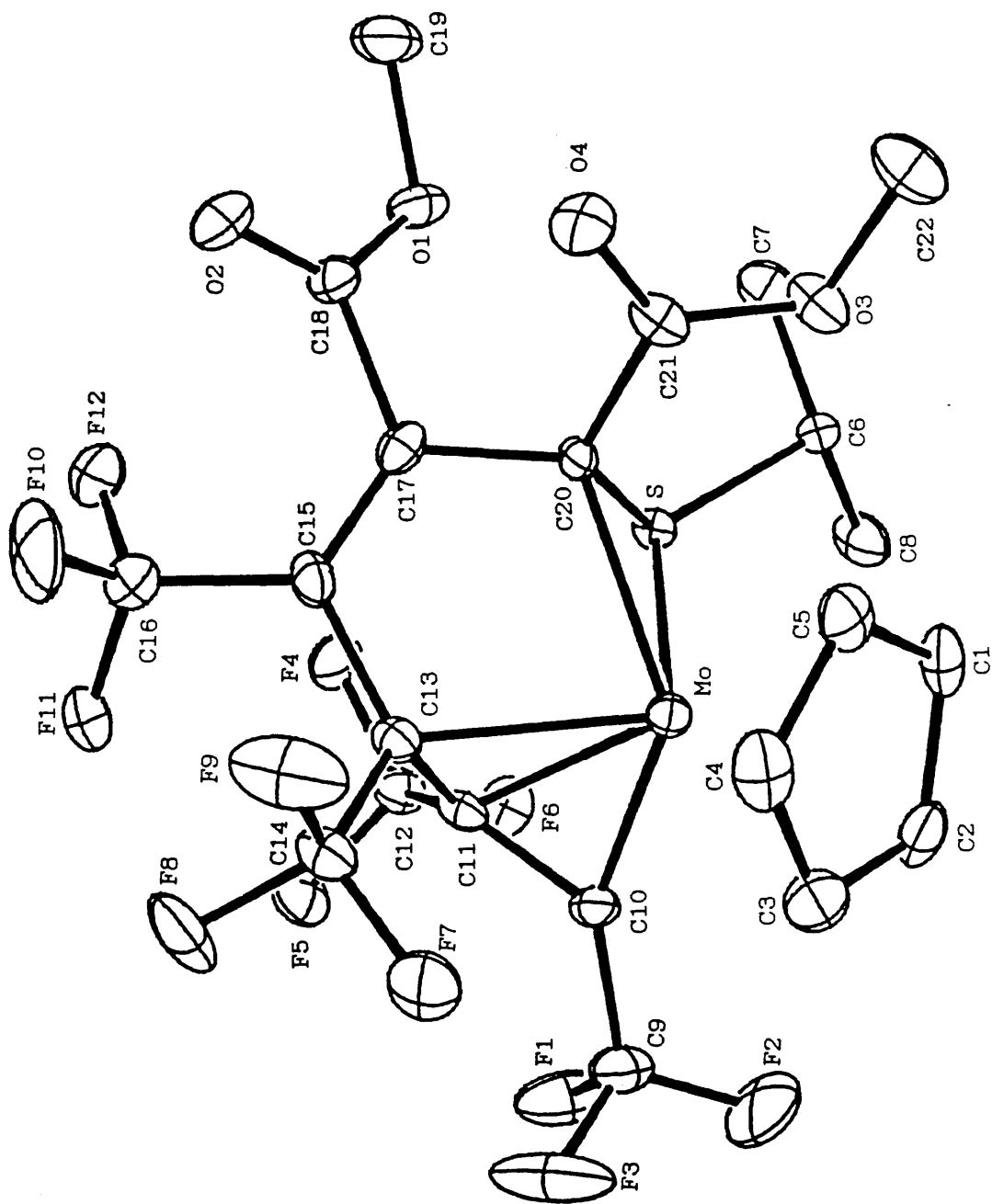


Complex (11) was characterised using X-ray diffraction analysis. This showed the crystal to contain two crystallographically different molecules of (11) separated by distances consistent with van der Waals radii.⁵² While the molecules are crystallographically different, chemically they are identical with only small differences in geometry which can be attributed to crystal packing (see Table 3.9).

In complex (11) there are only two discrete ligands attached to the molybdenum atom: the cyclopentadienyl ring which again has unexceptional geometry and an oligomer containing the three alkyne units and the thiolate group (see Figure 3(d)).

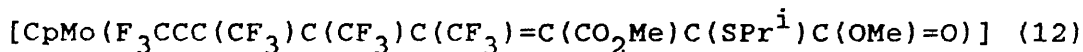
The oligomer is bonded to the molybdenum atom through the C(10), C(11), C(13), C(20) and S atoms. Atoms C(15) and C(17) are clearly non-bonding [Mo...C 3.192(8) and 3.169(8) Å respectively in molecule A and 3.177(8) and 3.162(8) Å in molecule B]. C(10), C(11) and C(13) are bonded in a similar fashion to the first three carbon atoms of the η^4 -butadienyl unit in complex (7) (see C-C and M-C distances for C_a , C_b and C_c in Table 3.1), i.e. Mo-C(10) is a multiple bond and the Mo-C attachment of atoms C(11) and C(13) is a $M-\eta^2$ -alkene interaction but with some delocalisation of π -bonding over all three carbon atoms. The attachment of the molybdenum atom to the oligomer is completed by σ -bond interactions with C(20) and S atoms.

Figure 3(d). A view of complex (11). Hydrogen atoms are omitted and 50% probability ellipsoids are shown for all other atoms.



Atom C(15) is connected to C(13) by a single bond as is C(17) to C(20) [mean bond distances 1.522(10) and 1.483(10) Å]. However, C(15)-C(17) is clearly a localised double bond [mean bond distance 1.318(10) Å and mean C(13)-C(15)-C(17)-C(20) torsion angle 1.9(6)°].

In complex (11) the methyl ester substituents are no longer on the central alkyne unit of the oligomer as in compound (10) but are now on the alkyne unit attached to the thiolate group.

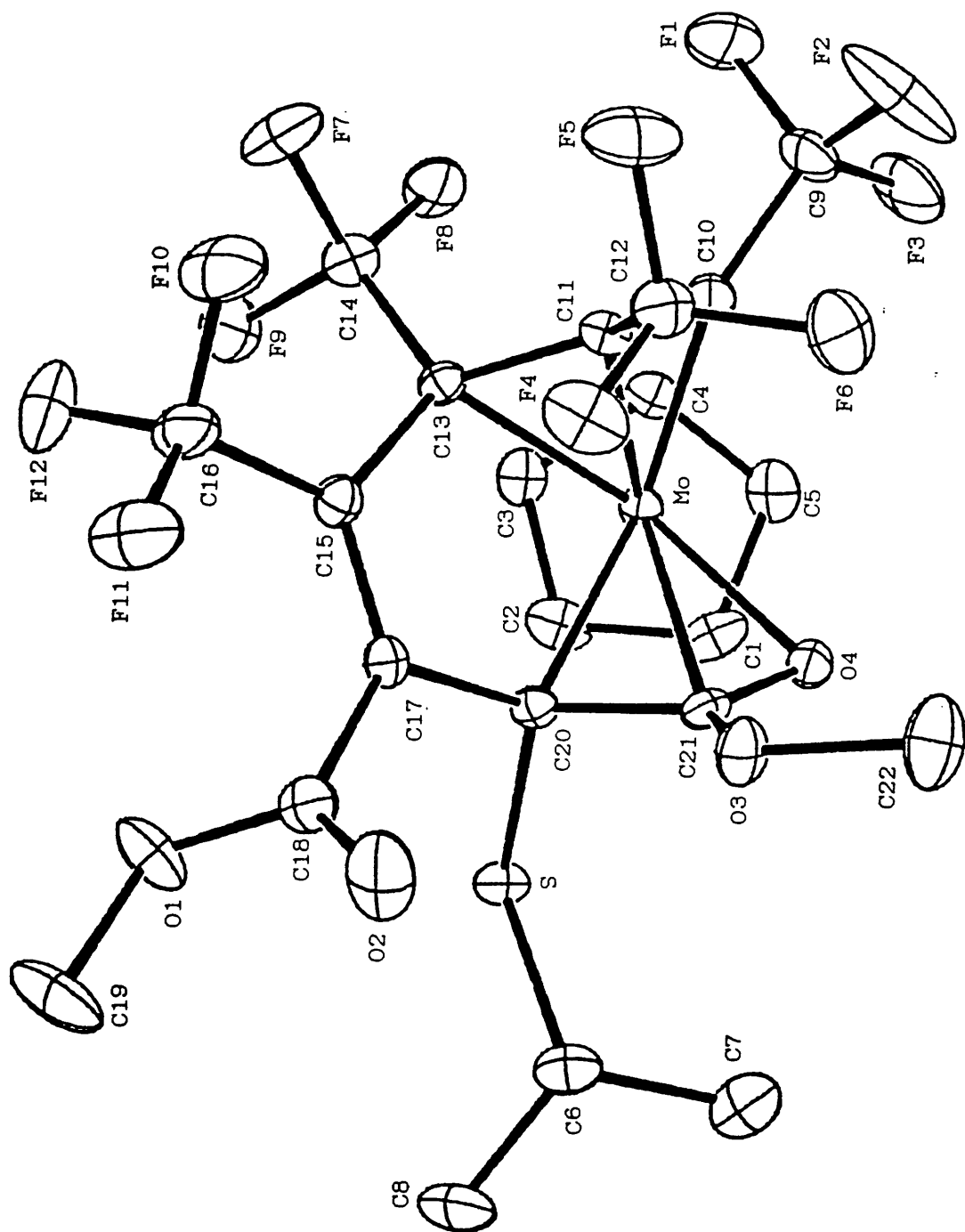


X-ray diffraction analysis of a crystal of complex (12) shows it to be made up of discrete molecules separated by distances consistent with van der Waals radii.⁵²

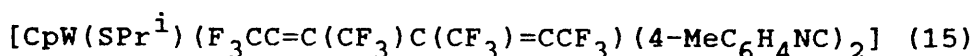
The structure of compound (12) shows it to be an isomer of complex (11), and like (11) it has a cyclopentadienyl ring and the same oligomer, consisting of the three alkyne units and the thiolate group, bonded to the molybdenum atom (see Figure 3(e)). The cyclopentadienyl rings geometry is again consistent with others in this series (see Tables 3.1 and 3.11).

The bonding within the carbon chain and to the molybdenum atom is remarkably similar to that in complex (11) (see Table 3.1 for comparison of M-C and C-C bond distances and C-C-C-C torsion angles), and can be regarded as identical. The structures differ in that there has been a rotation of the substituents attached to C(20) so that in complex (12) the sulphur atom is no longer bonded to the molybdenum atom and Mo-C(21) and Mo-O(4) distances [2.477(6) and 2.125(4) Å respectively] suggest the methyl

Figure 3(e). A view of complex (12). For clarity, hydrogen atoms have been omitted. 50% probability ellipsoids are displayed for all other atoms.



ester group is now attached to the molybdenum atom through the carbonyl group. Comparison of the two methyl ester geometries in complex (12) shows there has been a lengthening of the C=O bond [C(21)-O(4) 1.248(9) Å and C(16)-O(2) 1.178(10) Å] characteristic of η^2 -C=O π -bonding to the metal centre. Evidence that the C(21)-O(4) interaction is no longer a localised double bond is seen in the Mo(1)-O(4)-C(21)-C(20) torsion angle being 32.2(5)°.



X-ray diffraction analysis was used to study the crystals of complex (15). They were found to contain well separated molecules of the compound with all intermolecular contacts consistent with van der waals radii.⁵² Complex (15) contains a $\text{W}(\text{CCF}_3)_4$ metallacycle, with a cyclopentadienyl ring, an isopropylthiolate ligand and two 4-MeC₆H₄NC ligands also bonded to the tungsten atom (see Figure 3(f) and Tables 3.1 and 3.13).

The 18-electron bonding around the tungsten(IV) atom is roughly octahedral with the cyclopentadienyl ring occupying a single vertex *trans* to the thiolato-sulphur atom. The equatorial plane is made up of the two mutually *cis* 4-MeC₆H₄NC ligands and the $\text{C}_4(\text{CF}_3)_4^{2-}$ unit. This means the molecule has an approximate mirror plane which passes through the midpoints of C(1)-C(2) and C(11)-C(13) bonds and atoms W, S and C(4).

The W-C(cyclopentadienyl) bond lengths are comparable with those in the $[\text{CpM}(\text{SR})(\text{hfb})_2(\text{R}'\text{C}=\text{CR}'')]$ complexes (see Table 3.1). However the W-S distance of 2.561(2) Å and the W-CNR distances of 2.081(6) and 2.090(6) Å are longer than

[illegible]

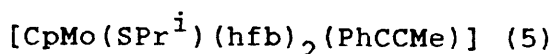
for those in the comparable tungsten(II) structure $[\text{CpW}(\text{SPr}^i)(\text{CNBu}^t)(\eta^2\text{-C}_4(\text{CF}_3)_4\text{CNBu}^t)]$ [W-S 2.374(3) Å and W-CNR 2.017(9) Å].¹¹⁰ The bonding within these ligands in the two structures does not show any major differences.

The geometry of the metallacycle suggests that the W-C_α bonds are single in nature [mean distance 2.210(4) Å] whereas the $\text{C}_\alpha\text{-C}_\beta$ bond lengths are clearly in the range of localised double bonds [1.354(8) and 1.359(8) Å] and the $\text{C}_\beta\text{-C}_\beta$ bond appears to be single [1.472(8) Å]. The internal torsion angles are all 12° or less.

This is clearly different from the WC_4 metallacycle of $[\text{CpW}(\text{C}(\text{CF}_3)\text{C}(\text{CF}_3)\text{C}(\text{CF}_3)\text{C}(\text{CF}_3))(\text{CO})_2\text{Co}(\text{CO})_2]$,¹¹¹ in which the cobalt atom is bonded to the tungsten atom and is roughly equidistant from all four butadiene carbons. The C-C distances in this structure are all in the range 1.42(1)-1.44(1) Å i.e. all effectively the same length. There is one other type of bonding possible for the ring i.e. when the W-C_α and $\text{C}_\beta\text{-C}_\beta$ bonds are localised double bonds and $\text{C}_\alpha\text{-C}_\beta$ bonds are single in nature.

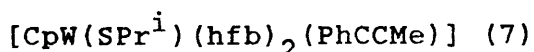
Experimental

The same basic experimental procedure was carried out for the structures in this Chapter as for those in Chapter 2. Therefore only differences from the procedure given in Chapter 2 are mentioned below.



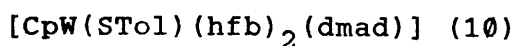
Data Collection. Cell dimensions and experimental parameters are given in Table 3.14. Severe non-linear decomposition (intensity standards fell by 70%) meant even after correction, standards varied by $\pm 10\%$. As a result a Gaussian rather than empirical absorption correction was carried out.⁸⁷

Structure Analysis. All hydrogen atoms save one methyl hydrogen were found in the difference synthesis. Independent refinement of hydrogen parameters was not satisfactory. Accordingly, hydrogen positions were idealised ($\text{C-H} = 0.96\text{\AA}$) and the hydrogen atoms were constrained to ride on their parent carbon atoms with $U(\text{H}) = 1.2 U(\text{C})$. Final atomic parameters were obtained (see Table 3.2) by refinement of blocks of the 351 parameters.



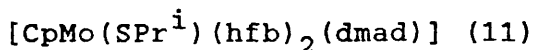
Data Collection. Cell dimensions and experimental parameters are given in Table 3.14. No decomposition correction was necessary.

Structure Analysis. All hydrogen atom positions were found and refined, as were their isotropic displacement parameters. Final atomic parameters for all atoms (Table 3.4) were obtained by refinement of blocks of the 432 parameters. All reflections were used in the final refinement because of a marked improvement in the e.s.d.s of the atomic parameters.



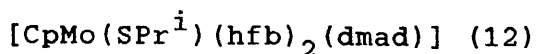
Data Collection. Cell dimensions and experimental parameters are given in Table 3.14. No decomposition correction was necessary.

Structure Analysis. Anisotropic displacement parameters were used for the W and S atoms. All other non-hydrogen atoms used isotropic parameters. Although it proved impossible to locate hydrogen positions in the difference synthesis these were calculated geometrically. Refinement of the chirality factor η from both +1 and -1 gave a convergence to the physically meaningless value of $\eta = -0.10(6)$. Final atomic positions and displacement parameters are given in Table 3.6.



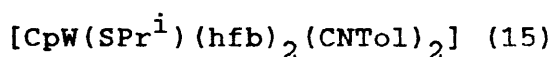
Data Collection. Cell dimensions and experimental parameters are given in Table 3.14. The Type I triclinic cell has two molecules per asymmetric unit. It does not transform to a higher symmetry cell [$A = 118.8$, $B = 259.9$, $C = 305.6$, $D = 124.4$, $E = 55.3$ and $F = 55.9$]. No decomposition correction was necessary.

Structure Analysis. No hydrogen atom positions were found and hence they were calculated geometrically. Final atomic positions and displacement parameters (Table 3.8) were obtained by refinement of blocks of the 721 parameters.



Data Collection. The cell dimensions and experimental parameters are given in Table 3.14. No decomposition correction was applied.

Structure Analysis. Final atomic positions and thermal displacement parameters are given in Table 3.10.



Data Collection. The cell dimensions and experimental parameters are listed in Table 3.14. A decomposition correction was applied to correct a 10% linear decay of the intensity standards.

Structure Analysis. Hydrogen thermal displacement parameters were set at $U(\text{H}) = 0.05 \text{ \AA}^2$ and were not refined. Final atomic positions and displacement parameters were obtained by refinement of blocks of the 433 parameters (see Table 3.12).

3.3 Reaction Scheme and Discussion

Scheme III shows the intermediates and products, in the reaction of alkynes with complexes (1) and (2), elucidated during and prior to this period of research. The * denotes structures which were characterised using X-ray diffraction analysis by the author. The thermolysis products (3) and (4) were referred to in Chapter 2, while all other compounds are discussed below and their geometries have been described more fully in 3.2.

Insertion of $R'C\equiv CMe$

The reactions of the bis(alkyne) species (2), $M = W$, $R = 4-MeC_6H_4$ and the η^2 -vinyl complexes (1), $M = Mo, W$, $R = Pr^i$, with an incoming alkyne of the form $R'C\equiv CMe$, $R' = Me, Ph$ at room temperature, lead initially to the same species (5) irrespective of the metal or incoming alkyne used.

The structure of the 18-electron complex (5), $M = Mo$, $R = Pr^i$, $R' = Ph$ is formally derived from (1) by insertion of the incoming $PhC\equiv CMe$ into the S-C bond of the $\eta^3-C(CF_3)C(CF_3)SPr^i$ ligand with the phenyl substituent on the carbon adjacent to sulphur and involves virtually no change in metal-ligand coordination. Indeed there are no major variations in the metal-ligand bonding between the two complexes (1) and (5) (cf Tables 2.1 and 3.1).

On heating complex (5) can isomerise to either compound (6) or (7), both 18-electron structures, depending on the nature of the metal atom, of the thiolate group or of the incoming alkyne.

Complex (6); $M = W$, $R = Pr^i$, $R' = Me$, was characterised

by X-ray diffraction analysis previously.¹⁰⁰ It differs from complex (5) only in the rearrangement of the SC_4 chain where cleavage of the metal-sulphur bond and attachment of the previously non-bonded chain carbons to form an η^4 -butadienyl ligand has occurred. The η^2 -alkyne and cyclopentadienyl ring remain essentially unaffected.

The structure of complex (7), $M = W$, $R = Pr^i$, $R' = Ph$, (see 3.2 for geometrical description) is derived from complex (5) by incorporation of the η^2 -alkyne into the oligomer and by scission of the $S-C(\text{oligomer})$ bond to form a WC_6 metallacycle and so that the thiolate group remains bonded to the metal only. The oligomer is attached to the metal through a roughly planar η^4 -butadienyl unit, similar to that found in complex (6), and a $\sigma-M-C(\text{alkenyl})$ bond.

Insertion of $MeO_2CC \equiv CCO_2Me$ (dmad)

The reaction of complex (1), $M = W$, $R = Pr^i$, with dmad produced the isomers (8a) and (8b), which differ only in the orientation of the isopropyl group on the sulphur atom relative to the cyclopentadienyl ring, and the novel complex (9), which has undergone transfer of a fluorine atom from a trimethylfluoro group to the metal atom. These structures were all characterised previously using X-ray diffraction analysis,¹⁰⁰ and their bonding is described in 1.2.

The main points of note are that, in both complex (8) and (9), all three alkynes are linked in a continuous chain, with the incoming dmad ligand in the middle position and with the thiolate attached at one end. The differences in structures occur in the mode of bonding of

this oligomer around the metal atom and in the transfer, in complex (9), of the fluorine atom to the tungsten atom.

Reaction of dmad with complex (2), $M = W$, $R = 4\text{-MeC}_6\text{H}_4$, results in a structurally different product (10). In complex (10) (see 3.2 for geometrical description), as in complexes (8) and (9) the three alkynes are connected in a continuous chain around the metal atom with dmad the central unit. However in complex (10) the CO_2Me substituents are clearly *trans* relative to each other whereas in complex (8) they are *cis*. Also in complex (10) there is no bonding between the carbon chain and the thiolate group. The chain is bonded to tungsten atom through the η^2 -alkene type interaction from the central alkyne unit and two $\sigma\text{-M-C(alkenyl)}$ bonds from the outer carbons of the end alkyne units of the C_6 chain. Complex (10) also differs from (8) and (9) in that it is a 16-electron structure rather than having 18-electrons around the tungsten atom.

The reaction of dmad with complex (1), $M = \text{Mo}$, $R = \text{Pr}^i$, produced structurally different complexes from those above. Complexes (11) and (12) are isomers (geometrical details in 3.2). In both complexes the three alkyne units are again connected in a continuous chain. In this case, however, the dmad unit is at one end of the chain and the thiolate group is attached to it.

In complex (11) the oligomer is attached to the molybdenum atom through bonding from the sulphur atom, the carbon atom adjacent to sulphur by a single $\sigma\text{-M-C(alkyl)}$ interaction, and following two clearly non-bonding carbons of the chain through a η^2 -alkene interaction from the next

two carbons and by a M-C(carbene) interaction from the final carbon.

The Mo-C bonding of the oligomer in complex (12) is very similar to that of structure (11). However (12) differs from (11) formally in that there has been a rotation of the groups attached to the first carbon as mentioned above. There has been scission of the Mo-S bond and the methyl-ester substituent is now attached to the molybdenum atom through a η^2 -Mo-C=O interaction.

It has been established by Dr.J.L. Davidson using N.M.R. and I.R. spectroscopy that all the above reactions involving dmad proceed via an intermediate which is unstable above -30°C and it has been predicted that this intermediate has structure (13). Unfortunately this has not as yet been confirmed using X-ray diffraction analysis because suitable crystals have not yet been obtained.

It appears that the 16-electron complex (13) has a metallacyclic ring involving the incoming alkyne and one of the hexafluorobut-2-yne ligands already attached. The thiolate group and the other hfb ligand are attached only to the metal centre.

Addition of R'NC

Attachment of ligands other than alkynes to complexes (1) and (2) can also result in the formation of metallacycles. The reaction of (1) and (2) with R'NC when $\text{R}' = \text{Bu}^t$ is one such example.^{96,112}

Reaction of RNC with (1) led to the formation of complex (14) which is similar in structure to (13) but with the η^2 -hfb ligand replaced by the iso-nitrile. Addition of

a further $R'NC$ ligand produced complex (15). X-ray diffraction analysis of crystals of this compound, $M = W$, $R = Pr^i$, $R' = 4-MeC_6H_4$ has been carried out (see 3.2 for geometrical and experimental details).

Bonded to the tungsten atom are two molecules of $4-MeC_6H_4NC$, a cyclopentadienyl ring, the thiolate group SPr^i and a carbon chain comprising of the two hfb ligands. The last of these ligands creates a metallacycle with the tungsten atom similar to that thought to be present in complexes (13) and (14). The bonding in the metallacycle comprises two $W-C(alkenyl)$ single bonds to the end carbons of the C_4 chain, two localised $C-C$ double bonds with a $C-C$ single bond separating these.

Reaction of the $R'NC$ ligand with complex (2) can also produce (15), however it is also possible, in this case, for the incoming ligand to attach itself to one of the η^2 -alkynes to form an η^2 -vinyl species which on reaction with a further $R'NC$ unit forms complex (15) but not by way of complex (14) as above.¹¹²

On heating complex (15) rearranges to form the novel complex (16), in which a five membered carbon ring has been formed by cyclisation of the carbon chain with a nitrile carbon.⁹⁶

3.4 Conclusions

It is clear from this study that reaction of different alkynes $R'C \equiv CR''$ with the η^2 -vinyl structure (1) and its isomer (2) can create a large number of products, dependant on the nature of the metal atom, the alkylidene groups and

the thiolato substituent R.

Complexes containing η^2 -vinyl species appear to play an important part in these processes and the thiolate group also appears crucial because of its ability not only to bond to the alkyne ligands but also to stabilise 16-electron structures, as in complex (10).

The reactions appear to proceed initially by metallacyclisation of one of the co-ordinated alkyne ligands and the incoming ligand. The incoming ligand seems always to be preferred to the other co-ordinated alkyne ligand. On addition of dmad this intermediate has been detected by N.M.R. and I.R. spectroscopy by Dr.J.L. Davidson, i.e. complex (13), but not as yet when using alkynes of the form $R'C\equiv CMe$ for which complex (5) is the first intermediate detected. However it is likely that complex (5) is formed with attack by $R'C\equiv CMe$ at the $M=C$ bond in (1) with resultant cleavage of the S-C bond and formation of a metallacycle similar to (13). Then immediate attachment of the thiolate group to the new ligand side of the metallacycle forms complex (5).

The products obtained by heating intermediates (5) and (13) also depend on the factors mentioned above. In complex (5) the η^2 -vinyl ligand can isomerise to give the η^4 -butadienyl ligand of complex (6) and cleavage of the M-S bond. Alternatively the η^2 -alkyne can insert into the $M=C$ bond to produce an oligomer comprising three alkyne units, with cleavage of the S-C(oligomer) bond giving complex (7). X-ray diffraction analysis has shown that when $M = W$, $R = Pr^i$ and $R' = Me$ structure (6) is produced¹⁰⁰ and when $M = W$, $R = Pr^i$ and $R' = Ph$ it is structure (7) that is formed

suggesting that substituent R' has an effect on what products are formed. Similarly spectroscopic data ¹⁰² suggest molybdenum complexes are more likely to form complex (6) than (7) compared with tungsten complexes.

When dmad was the incoming ligand used it is clear that there is selective thiolate migration onto either the hfb or dmad metal coordinated carbon of the metallacycle of complex (13) depending on whether the metal centre was a tungsten or molybdenum atom. This means that the product of further insertion of an alkyne into the carbon chain is selective. When M = Mo the sequence of substituents on the carbon chain is $\text{CF}_3\text{CF}_3\text{CF}_3\text{CF}_3\text{CO}_2\text{MeCO}_2\text{MeSR}$ and when M = W the sequence is $\text{CF}_3\text{CF}_3\text{CO}_2\text{MeCO}_2\text{MeCF}_3\text{CF}_3\text{SR}$ as can be seen in structures (11) and (12), and (8), (9) and (10) respectively.

Altering the thiolato substituent, R, of the starting complex also appears to have an effect on the final product, e.g. with M = W and dmad the incoming ligand, when R is the aryl group, C_6H_5 , complex (10) has been formed where the thiolate group is attached only to the metal centre, whereas when the thiolate substituent is the alkyl group, Pr^i , complexes (8) and (9) have been isolated where the sulphur atom is attached to the oligomer as well as the metal atom.

It can be seen that formation of both *cis* and *trans* oligomers is possible. While the first insertion always occurs with *cis* geometry, because of the metallacyclopentadiene intermediate, the final product can keep this *cis* orientation, as in complex (8),¹⁰⁰ or it can change to include a *trans* bond as in complexes (7), (10), (11) and

(12) (see Table 3.1). The ability to change orientation is probably due to the formation of η^4 -butadiene or η^2 -vinyl intermediates whereas an η^1 -bonding mode would lead to an all *cis* structure.

It is also possible for cyclisation reactions to occur if ligands other than alkynes are added to complexes (1) and (2). An example of this is the reaction pathway (14)→(15)→(16) in Scheme III where on addition of two R'NC ligands a five membered ring was formed consisting of the two alkyne ligands and a carbon atom from one of the isonitriles.¹¹² A cyclobutadienyl ring can also be formed, as in complex (18), by cyclisation of the two alkyne ligands after addition of PR'₃ to complex (1). This reaction proceeds via intermediates (14) and (17). A similar cyclobutadienyl complex (19) is described in Chapter 4.

From this work there are at least two important generalisations which can be made. First, it is clear that polymerisation of alkynes by transition metals can involve many different intermediates. Secondly, although the initial step in the oligomerisation reactions discussed here is formation of a metallacyclopentadiene with mutually *cis* double bonds, the final polyene obtained may either retain this configuration if only M-C scission occurs, or it may have a *trans* arrangement of double bonds if further steps involve coordination of several chain atoms to the metal e.g. as η^2 -vinyl or η^4 -butadienyl units.

Table 3.1. A comparison of the geometries of the [CpM(SR)(hfb)₂(L-L)] structures (distances in Å, angles in °).

Complex	(5)	(7)	(10)	(11)	(12)	(15)
M	Mo	W	W	Mo	Mo	W
R	Pr ⁱ	Pr ⁱ	MeC ₄ H ₄	Pr ⁱ	Pr ⁱ	MeC ₄ H ₄
L-L	PhC ₂ Me	PhC ₂ Me	dmad	dmad	dmad	(ToINC) ₂
M-C(Cp)	min 2.342(6)	2.333(5)	2.32(2)	2.310(8)	2.314(8)	2.255(7)
	max 2.410(7)	2.356(5)	2.37(3)	2.378(7)	2.367(7)	2.360(6)
C-C(Cp)	min 1.381(10)	1.396(6)	1.34(3)	1.389(14)	1.383(11)	1.397(9)
	max 1.423(9)	1.419(6)	1.47(3)	1.434(12)	1.416(12)	1.417(9)
C-F	min 1.25(1)	1.294(5)	1.25(4)	1.269(14)	1.273(9)	1.322(9)
	max 1.35(1)	1.341(5)	1.38(4)	1.358(10)	1.345(9)	1.351(9)
M-S	2.468(2)	2.374(1)	2.318(6)	2.400(2)	2.405(2)	-
S-R	1.866(6)	1.840(4)	1.78(2)	1.833(7)	1.846(7)	1.839(8)
S-C(chain)	1.798(6)	-	-	1.784(7)	1.789(6)	1.819(7)
M-C _a	1.915(5)	2.017(3)	2.24(2)	1.915(8)	1.921(6)	1.918(7)
M-C _b	2.239(5)	2.489(3)	[2.79(2)]	2.179(8)	2.176(7)	2.195(7)
M-C _c	[3.216(5)]	2.351(4)	2.21(2)	2.273(8)	2.273(8)	2.266(7)
M-C _d	[3.343(5)]	2.195(4)	2.29(2)	[3.192(8)]	3.177(8)]	[3.187(7)]
M-C _e	[2.133(6)]	[2.766(4)]	[2.67(2)]	[3.169(8)]	3.162(8)]	[3.203(7)]
M-C _f	[2.156(6)]	2.242(4)	2.12(2)	2.267(6)	2.282(7)	2.249(7)
C _a -C _b	1.426(7)	1.400(5)	1.35(3)	1.411(11)	1.410(11)	1.410(10)
C _b -C _c	1.514(7)	1.404(5)	1.50(3)	1.449(10)	1.434(9)	1.431(9)
C _c -C _d	1.338(7)	1.475(5)	1.44(3)	1.533(10)	1.510(11)	1.517(10)
C _d -C _e	-	1.509(5)	1.47(2)	1.316(10)	1.320(10)	1.316(9)
C _e -C _f	1.267(8)	1.321(5)	1.34(3)	1.489(11)	1.477(10)	1.503(10)
C _a -C _b -C _c -C _d	-62.5(6)	-22.4(3)	66(2)	152.0(9)	151.2(9)	152.9(8)
C _b -C _c -C _d -C _e	-	149.0(5)	-166(2)	-78.9(8)	-79.0(8)	-70.7(7)
C _c -C _d -C _e -C _f	-	-64.3(4)	51(2)	1.4(6)	2.4(6)	-3.9(6)

C_a - C_f are the carbon atoms of the oligomer chain, and C_a is the end carbon atom of the chain not attached to the sulphur atom (when applicable).

[] signify the distance is non-bonding.

() signify the carbon atoms are not part of the oligomer.

Table 3.2. Fractional atomic coordinates and displacement parameters (\AA^2) of the crystal structure of complex (5).

	X/A	Y/B	Z/C	U
Mo(1)	.05682(5)	-.35608(2)	.30362(2)	.034
S(1)	.18982(15)	-.23586(7)	.37348(5)	.039
F(1)	.0548(9)	-.3780(3)	.1168(2)	.142
F(2)	.0571(7)	-.5099(3)	.1599(2)	.118
F(3)	.2711(7)	-.4511(5)	.1451(2)	.171
F(4)	.4581(5)	-.2291(4)	.2691(3)	.131
F(5)	.2760(7)	-.1484(3)	.2317(3)	.147
F(6)	.3720(10)	-.2404(5)	.1766(3)	.209
F(7)	-.1012(4)	-.5212(2)	.4120(2)	.082
F(8)	.1355(5)	-.5433(3)	.4513(2)	.086
F(9)	.0402(5)	-.6187(2)	.3709(2)	.089
F(10)	.2297(4)	-.5569(2)	.2701(2)	.065
F(11)	.3912(4)	-.5917(2)	.3491(2)	.071
F(12)	.4539(4)	-.4866(2)	.2861(2)	.068
C(1)	-.1927(7)	-.3321(4)	.3349(3)	.059
C(2)	-.1653(7)	-.2526(4)	.2996(3)	.059
C(3)	-.1579(7)	-.2797(4)	.2400(3)	.061
C(4)	-.1819(8)	-.3785(4)	.2369(3)	.062
C(5)	-.2049(7)	-.4110(4)	.2945(3)	.062
C(6)	.0744(7)	-.2064(3)	.4373(3)	.056
C(7)	.0586(9)	-.2870(4)	.4813(3)	.074
C(8)	.1366(9)	-.1165(4)	.4688(3)	.072
C(9)	.1346(10)	-.4285(5)	.1608(3)	.084
C(10)	.1497(7)	-.3781(3)	.2196(2)	.051
C(11)	.2193(6)	-.3055(3)	.2437(2)	.046
C(12)	.3320(8)	-.2348(4)	.2287(3)	.067
C(13)	.0426(6)	-.5350(3)	.3978(3)	.049
C(14)	.1011(6)	-.4585(3)	.3601(2)	.039
C(15)	.2602(6)	-.4418(3)	.3490(2)	.038
C(16)	.3340(7)	-.5180(3)	.3145(3)	.049
C(17)	.3784(6)	-.3905(3)	.3947(2)	.039
C(18)	.5227(7)	-.4424(4)	.4226(3)	.053
C(19)	.3570(6)	-.3005(3)	.4108(2)	.043
C(20)	.4658(7)	-.2428(3)	.4543(2)	.048
C(21)	.5375(8)	-.1628(4)	.4339(3)	.062
C(22)	.6388(9)	-.1081(4)	.4732(3)	.078
C(23)	.6672(10)	-.1331(5)	.5337(4)	.092
C(24)	.5967(10)	-.2116(5)	.5548(3)	.081
C(25)	.4955(8)	-.2664(4)	.5158(3)	.064
H(1)	-.20187	-.33293	.37780	.035
H(2)	-.15372	-.18878	.31454	.045
H(3)	-.13982	-.23823	.20718	.042
H(4)	-.18203	-.41675	.20099	.050
H(5)	-.22476	-.47538	.30521	.039
H(6)	-.03079	-.18930	.41955	.089
H(7A)	-.01315	-.26940	.50951	.120
H(7B)	.16128	-.29980	.50326	.120
H(7C)	.02002	-.34311	.45973	.120
H(8A)	.15001	-.06669	.43897	.118
H(8B)	.23697	-.12905	.49262	.118
H(8C)	.06479	-.09201	.49505	.118
H(18A)	.58488	-.45999	.39108	.046
H(18B)	.49242	-.49828	.44293	.046
H(18C)	.58537	-.40274	.45168	.046
H(21)	.51530	-.14530	.39166	.073
H(22)	.68870	-.05321	.45854	.094
H(23)	.73614	-.09459	.56143	.106
H(24)	.61887	-.22862	.59716	.104
H(25)	.44619	-.32145	.53054	.093

Table 3.2(cont).

Anisotropic displacement parameters (\AA^2).

	U11	U22	U33	U12	U13	U23
MO(1)	.0305(2)	.0317(2)	.0407(2)	-.0036(2)	.0051(1)	.0028(2)
S(1)	.0443(7)	.0308(4)	.0413(7)	-.0031(4)	.0054(5)	.0019(4)
F(1)	.230(7)	.145(4)	.045(2)	-.001(4)	-.013(3)	.002(2)
F(2)	.169(5)	.098(3)	.081(3)	-.028(3)	-.009(3)	-.035(2)
F(3)	.138(5)	.262(7)	.115(4)	.002(5)	.052(4)	-.106(4)
F(4)	.062(3)	.171(4)	.155(5)	-.048(3)	-.012(3)	.050(4)
F(5)	.125(4)	.076(3)	.239(6)	-.024(3)	.036(4)	.068(3)
F(6)	.285(9)	.227(6)	.126(4)	-.172(6)	.134(5)	-.048(4)
F(7)	.064(2)	.079(2)	.107(3)	-.001(2)	.045(2)	.033(2)
F(8)	.081(3)	.099(2)	.074(2)	-.021(2)	.003(2)	.039(2)
F(9)	.112(3)	.039(1)	.119(3)	-.019(2)	.052(3)	-.002(2)
F(10)	.062(2)	.051(2)	.079(2)	.001(1)	.006(2)	-.025(1)
F(11)	.071(2)	.044(1)	.095(3)	.021(1)	.008(2)	.001(1)
F(12)	.049(2)	.079(2)	.077(2)	.002(2)	.030(2)	-.011(2)
C(1)	.034(3)	.061(3)	.082(4)	.008(2)	.019(3)	.007(3)
C(2)	.040(3)	.046(3)	.089(5)	.007(2)	.006(3)	.006(3)
C(3)	.039(3)	.065(3)	.075(4)	.001(3)	-.009(3)	.018(3)
C(4)	.046(4)	.068(4)	.067(4)	-.004(3)	-.014(3)	-.003(3)
C(5)	.037(3)	.048(3)	.099(5)	-.008(2)	.000(3)	.008(3)
C(6)	.066(4)	.044(2)	.057(4)	.001(2)	.023(3)	-.005(2)
C(7)	.099(6)	.060(3)	.066(4)	-.001(3)	.040(4)	.003(3)
C(8)	.105(6)	.045(3)	.066(4)	.005(3)	.016(4)	-.011(2)
C(9)	.095(6)	.102(5)	.053(4)	-.003(4)	.010(4)	-.008(3)
C(10)	.054(4)	.057(3)	.041(3)	-.004(2)	.008(3)	0.000(2)
C(11)	.042(3)	.051(3)	.046(3)	-.003(2)	.012(3)	.006(2)
C(12)	.064(4)	.078(4)	.059(4)	-.016(3)	.019(3)	.018(3)
C(13)	.041(3)	.043(2)	.064(3)	-.004(2)	.013(3)	.009(2)
C(14)	.040(3)	.034(2)	.044(3)	-.004(2)	.009(2)	0.000(2)
C(15)	.036(3)	.031(2)	.047(3)	.001(2)	.007(2)	.002(2)
C(16)	.043(3)	.045(2)	.060(3)	.002(2)	.012(3)	-.006(2)
C(17)	.036(3)	.041(2)	.040(3)	-.004(2)	.009(2)	.006(2)
C(18)	.041(3)	.055(3)	.060(4)	.005(2)	-.002(3)	.005(2)
C(19)	.046(3)	.040(2)	.042(3)	-.009(2)	-.001(2)	.007(2)
C(20)	.051(4)	.044(2)	.047(3)	-.008(2)	-.005(3)	.001(2)
C(21)	.066(4)	.060(3)	.057(4)	-.017(3)	-.002(3)	.004(2)
C(22)	.085(6)	.064(3)	.080(5)	-.034(4)	-.010(4)	.001(3)
C(23)	.098(6)	.081(4)	.090(5)	-.036(4)	-.032(5)	-.006(4)
C(24)	.094(6)	.088(4)	.054(4)	-.015(4)	-.027(4)	.005(3)
C(25)	.081(5)	.052(3)	.055(4)	-.013(3)	-.008(3)	.008(2)

Table 3.3. Selected bond distances (Å) and angles (°) of complex (5)

MO(1) - S(1)	2.468(2)	MO(1) - C(1)	2.342(6)
MO(1) - C(2)	2.382(6)	MO(1) - C(3)	2.410(7)
MO(1) - C(4)	2.370(7)	MO(1) - C(5)	2.342(6)
MO(1) - C(10)	2.133(6)	MO(1) - C(11)	2.156(6)
MO(1) - C(14)	1.915(5)	MO(1) - C(15)	2.239(5)
S(1) - C(6)	1.866(6)	S(1) - C(19)	1.798(6)
F(1) - C(9)	1.320(9)	F(2) - C(9)	1.324(9)
F(3) - C(9)	1.294(11)	F(4) - C(12)	1.305(9)
F(5) - C(12)	1.315(8)	F(6) - C(12)	1.246(9)
F(7) - C(13)	1.317(7)	F(8) - C(13)	1.339(7)
F(9) - C(13)	1.320(6)	F(10) - C(16)	1.350(7)
F(11) - C(16)	1.345(6)	F(12) - C(16)	1.341(7)
C(1) - C(2)	1.405(8)	C(1) - C(5)	1.423(9)
C(2) - C(3)	1.381(10)	C(3) - C(4)	1.410(8)
C(4) - C(5)	1.390(10)	C(6) - C(7)	1.515(8)
C(6) - C(8)	1.510(8)	C(9) - C(10)	1.472(9)
C(10) - C(11)	1.267(8)	C(11) - C(12)	1.451(9)
C(13) - C(14)	1.489(7)	C(14) - C(15)	1.426(7)
C(15) - C(16)	1.502(7)	C(15) - C(17)	1.514(7)
C(17) - C(18)	1.493(8)	C(17) - C(19)	1.338(7)
C(19) - C(20)	1.488(7)	C(20) - C(21)	1.388(8)
C(20) - C(25)	1.390(8)	C(21) - C(22)	1.378(10)
C(22) - C(23)	1.374(11)	C(23) - C(24)	1.371(11)
C(24) - C(25)	1.374(10)		

S(1) - MO(1) - C(10)	116.9(2)	S(1) - MO(1) - C(11)	82.9(2)
S(1) - MO(1) - C(14)	94.4(2)	S(1) - MO(1) - C(15)	80.0(2)
C(10) - MO(1) - C(11)	34.4(2)	C(10) - MO(1) - C(14)	113.1(2)
C(10) - MO(1) - C(15)	87.6(2)	C(11) - MO(1) - C(14)	124.2(3)
C(11) - MO(1) - C(15)	86.2(2)	C(14) - MO(1) - C(15)	39.2(2)
MO(1) - S(1) - C(6)	112.5(2)	MO(1) - S(1) - C(19)	102.0(2)
C(6) - S(1) - C(19)	103.3(3)	C(2) - C(1) - C(5)	106.3(6)
C(1) - C(2) - C(3)	109.8(5)	C(2) - C(3) - C(4)	107.2(6)
C(3) - C(4) - C(5)	108.6(6)	C(1) - C(5) - C(4)	108.0(5)
S(1) - C(6) - C(7)	114.6(4)	S(1) - C(6) - C(8)	110.4(5)
C(7) - C(6) - C(8)	113.0(5)	F(1) - C(9) - F(2)	104.4(7)
F(1) - C(9) - F(3)	109.2(7)	F(1) - C(9) - C(10)	111.5(6)
F(2) - C(9) - F(3)	104.1(7)	F(2) - C(9) - C(10)	114.9(6)
F(3) - C(9) - C(10)	112.2(7)	MO(1) - C(10) - C(9)	147.0(5)
MO(1) - C(10) - C(11)	73.8(4)	C(9) - C(10) - C(11)	138.2(6)
MO(1) - C(11) - C(10)	71.8(4)	MO(1) - C(11) - C(12)	149.7(4)
C(10) - C(11) - C(12)	138.3(6)	F(4) - C(12) - F(5)	100.4(6)
F(4) - C(12) - F(6)	109.6(7)	F(4) - C(12) - C(11)	113.3(6)
F(5) - C(12) - F(6)	104.4(6)	F(5) - C(12) - C(11)	111.8(6)
F(6) - C(12) - C(11)	115.9(6)	F(7) - C(13) - F(8)	105.2(5)
F(7) - C(13) - F(9)	106.2(5)	F(7) - C(13) - C(14)	114.5(5)
F(8) - C(13) - F(9)	107.0(4)	F(8) - C(13) - C(14)	110.8(5)
F(9) - C(13) - C(14)	112.6(5)	MO(1) - C(14) - C(13)	149.4(4)
MO(1) - C(14) - C(15)	82.8(3)	C(13) - C(14) - C(15)	127.6(5)
MO(1) - C(15) - C(14)	58.0(3)	MO(1) - C(15) - C(16)	120.6(4)
MO(1) - C(15) - C(17)	116.6(3)	C(14) - C(15) - C(16)	116.0(4)
C(14) - C(15) - C(17)	121.8(5)	C(16) - C(15) - C(17)	113.3(5)
F(10) - C(16) - F(11)	104.6(4)	F(10) - C(16) - F(12)	105.3(5)
F(10) - C(16) - C(15)	112.4(5)	F(11) - C(16) - F(12)	106.3(5)
F(11) - C(16) - C(15)	114.2(5)	F(12) - C(16) - C(15)	113.3(4)
C(15) - C(17) - C(18)	118.7(4)	C(15) - C(17) - C(19)	122.0(5)
C(18) - C(17) - C(19)	119.3(5)	S(1) - C(19) - C(17)	119.2(4)
S(1) - C(19) - C(20)	114.0(4)	C(17) - C(19) - C(20)	126.6(5)
C(19) - C(20) - C(21)	119.9(5)	C(19) - C(20) - C(25)	121.3(5)
C(21) - C(20) - C(25)	118.8(5)	C(20) - C(21) - C(22)	121.1(6)
C(21) - C(22) - C(23)	119.1(7)	C(22) - C(23) - C(24)	120.6(7)
C(23) - C(24) - C(25)	120.5(7)	C(20) - C(25) - C(24)	119.9(6)

Table 3.4. Fractional atomic coordinates and displacement parameters (\AA^2) of the crystal structure of complex (7).

	X/A	Y/B	Z/C	U
W(1)	.07416(1)	.14326(1)	.23673(1)	.028
S(1)	.32053(10)	.20816(4)	.23689(6)	.040
F(1)	.3315(4)	.1448(1)	.0481(2)	.090
F(2)	.2954(4)	.0436(1)	.0128(2)	.083
F(3)	.4922(3)	.0696(2)	.0984(2)	.105
F(4)	.0849(4)	-.0691(1)	.0898(2)	.099
F(5)	.3423(4)	-.0640(1)	.1196(2)	.087
F(6)	.1724(4)	-.1010(1)	.2114(2)	.095
F(7)	-.1464(3)	-.0328(1)	.2303(2)	.079
F(8)	-.0151(4)	-.0590(1)	.3449(2)	.084
F(9)	-.1760(3)	.0258(1)	.3434(2)	.072
F(10)	.3773(3)	.0081(2)	.4218(2)	.090
F(11)	.4824(3)	.0810(2)	.3433(2)	.101
F(12)	.4048(4)	-.0122(2)	.2923(2)	.120
C(1)	-.0417(6)	.1496(2)	.0982(3)	.059
C(2)	-.0103(6)	.2150(2)	.1268(3)	.055
C(3)	-.1107(5)	.2280(2)	.1973(2)	.049
C(4)	-.2058(4)	.1697(2)	.2144(3)	.055
C(5)	-.1603(5)	.1211(3)	.1529(3)	.060
C(6)	.2724(5)	.2960(2)	.2640(2)	.043
C(7)	.2483(7)	.3056(3)	.3586(3)	.064
C(8)	.4087(9)	.3390(3)	.2292(6)	.096
C(9)	.3361(5)	.0821(2)	.0795(2)	.057
C(10)	.2225(4)	.0735(2)	.1545(2)	.041
C(11)	.1675(4)	.0158(2)	.1852(2)	.040
C(12)	.1908(6)	-.0547(2)	.1520(3)	.061
C(13)	.0789(4)	.0354(2)	.2662(2)	.036
C(14)	-.0641(5)	-.0066(2)	.2970(3)	.052
C(15)	.1923(4)	.0642(2)	.3309(2)	.035
C(16)	.3662(5)	.0363(2)	.3446(2)	.050
C(17)	.1344(4)	.1130(2)	.3884(2)	.034
C(18)	.2268(6)	.1361(3)	.4659(3)	.056
C(19)	-.0065(4)	.1469(2)	.3586(2)	.033
C(20)	-.1320(4)	.1761(2)	.4171(2)	.035
C(21)	-.1906(5)	.1376(2)	.4842(3)	.048
C(22)	-.3077(5)	.1629(2)	.5394(3)	.054
C(23)	-.3676(4)	.2270(2)	.5293(2)	.047
C(24)	-.3092(4)	.2663(2)	.4649(2)	.047
C(25)	-.1927(5)	.2414(2)	.4089(2)	.045
H(1)	.011(6)	.125(2)	.049(3)	.08(2)
H(2)	.061(6)	.247(3)	.106(3)	.09(2)
H(3)	-.097(5)	.270(2)	.230(3)	.07(1)
H(4)	-.292(6)	.160(3)	.260(3)	.09(2)
H(5)	-.198(5)	.081(2)	.151(2)	.04(1)
H(6)	.175(5)	.308(2)	.239(2)	.05(1)
H(7A)	.222(6)	.351(2)	.373(3)	.08(2)
H(7B)	.359(6)	.296(3)	.385(3)	.08(2)
H(7C)	.157(6)	.272(2)	.388(3)	.08(1)
H(8A)	.498(5)	.328(2)	.241(3)	.04(1)
H(8B)	.428(7)	.399(4)	.247(3)	.14(2)
H(8C)	.422(8)	.342(3)	.158(4)	.11(2)
H(18A)	.319(7)	.152(3)	.459(3)	.08(2)
H(18B)	.268(6)	.103(3)	.504(3)	.08(2)
H(18C)	.158(7)	.167(3)	.486(4)	.11(2)
H(21)	-.151(5)	.091(2)	.492(2)	.06(1)
H(22)	-.354(5)	.133(2)	.586(3)	.05(1)
H(23)	-.443(5)	.246(2)	.562(3)	.06(1)
H(24)	-.352(4)	.316(2)	.461(2)	.04(1)
H(25)	-.140(5)	.273(2)	.370(2)	.06(1)

Table 3.4(cont).

Anisotropic displacement parameters (\AA^2).

	U11	U22	U33	U12	U13	U23
W(1)	.0280(1)	.0277(1)	.0278(1)	-.0001(1)	-.0001(0)	.0003(1)
S(1)	.0354(4)	.0336(4)	.0509(4)	-.0040(3)	.0095(3)	-.0044(3)
F(1)	.139(3)	.075(2)	.055(1)	-.020(2)	.046(2)	-.001(1)
F(2)	.108(2)	.092(2)	.049(1)	-.009(2)	.023(1)	-.029(1)
F(3)	.046(1)	.178(3)	.091(2)	-.006(2)	.022(1)	-.003(2)
F(4)	.118(2)	.081(2)	.097(2)	-.010(2)	-.017(2)	-.055(2)
F(5)	.093(2)	.061(2)	.108(2)	.022(2)	.027(2)	-.029(2)
F(6)	.145(3)	.034(1)	.105(2)	.004(2)	.030(2)	-.006(1)
F(7)	.075(2)	.072(2)	.089(2)	-.039(1)	.001(1)	-.020(1)
F(8)	.092(2)	.053(1)	.108(2)	-.009(1)	.015(2)	.031(1)
F(9)	.058(1)	.058(1)	.098(2)	-.019(1)	.034(1)	-.010(1)
F(10)	.088(2)	.098(2)	.082(2)	.037(2)	-.017(2)	.031(2)
F(11)	.034(1)	.090(2)	.179(3)	.003(1)	-.010(2)	.036(2)
F(12)	.109(2)	.141(3)	.110(2)	.094(2)	-.049(2)	-.072(2)
C(1)	.064(2)	.075(3)	.040(2)	.012(2)	-.016(2)	-.003(2)
C(2)	.063(2)	.059(2)	.043(2)	.007(2)	-.008(2)	.016(2)
C(3)	.053(2)	.047(2)	.046(2)	.013(2)	-.011(2)	.005(2)
C(4)	.034(2)	.068(3)	.063(3)	.005(2)	-.011(2)	.010(2)
C(5)	.052(2)	.057(2)	.070(3)	-.003(2)	-.026(2)	-.003(2)
C(6)	.045(2)	.032(2)	.053(2)	-.006(1)	.006(2)	-.003(1)
C(7)	.079(3)	.050(2)	.064(3)	.000(2)	.007(2)	-.016(2)
C(8)	.088(4)	.044(3)	.156(6)	-.018(3)	.072(4)	-.011(3)
C(9)	.063(2)	.061(3)	.046(2)	-.006(2)	.014(2)	-.011(2)
C(10)	.042(2)	.049(2)	.032(1)	.001(2)	.004(1)	-.009(1)
C(11)	.040(2)	.036(2)	.045(2)	.002(1)	0.000(1)	-.011(1)
C(12)	.075(3)	.040(2)	.068(3)	.003(2)	.006(2)	-.018(2)
C(13)	.034(1)	.033(2)	.042(2)	-.004(1)	.004(1)	-.003(1)
C(14)	.054(2)	.036(2)	.066(2)	-.009(2)	.008(2)	-.001(2)
C(15)	.036(1)	.032(1)	.037(2)	.005(1)	.003(1)	.002(1)
C(16)	.048(2)	.046(2)	.055(2)	.017(2)	-.005(2)	-.002(2)
C(17)	.035(1)	.035(2)	.032(1)	.001(1)	.003(1)	.003(1)
C(18)	.053(2)	.067(3)	.048(2)	.014(2)	-.012(2)	-.017(2)
C(19)	.034(1)	.031(1)	.034(1)	0.000(1)	.003(1)	.000(1)
C(20)	.033(1)	.037(2)	.036(2)	.003(1)	.004(1)	-.002(1)
C(21)	.054(2)	.039(2)	.051(2)	.007(2)	.017(2)	.005(2)
C(22)	.060(2)	.051(2)	.053(2)	.001(2)	.025(2)	.003(2)
C(23)	.035(2)	.055(2)	.052(2)	.000(2)	.012(1)	-.010(2)
C(24)	.047(2)	.043(2)	.053(2)	.010(2)	.005(2)	-.006(2)
C(25)	.046(2)	.041(2)	.047(2)	.004(2)	.010(2)	.005(1)

Table 3.5. Selected bond distances (Å) and angles (°) of complex (7).

W(1) - S(1)	2.374(1)	W(1) - C(1)	2.356(5)
W(1) - C(2)	2.333(5)	W(1) - C(3)	2.335(4)
W(1) - C(4)	2.348(4)	W(1) - C(5)	2.339(5)
W(1) - C(10)	2.242(4)	W(1) - C(13)	2.195(4)
W(1) - C(15)	2.351(4)	W(1) - C(17)	2.489(3)
W(1) - C(19)	2.017(3)	S(1) - C(6)	1.840(4)
F(1) - C(9)	1.341(5)	F(2) - C(9)	1.335(5)
F(3) - C(9)	1.318(5)	F(4) - C(12)	1.322(6)
F(5) - C(12)	1.341(6)	F(6) - C(12)	1.316(5)
F(7) - C(14)	1.339(5)	F(8) - C(14)	1.340(5)
F(9) - C(14)	1.330(5)	F(10) - C(16)	1.332(5)
F(11) - C(16)	1.294(5)	F(12) - C(16)	1.303(6)
C(1) - C(2)	1.400(7)	C(1) - C(5)	1.408(7)
C(2) - C(3)	1.396(6)	C(3) - C(4)	1.419(6)
C(4) - C(5)	1.413(7)	C(6) - C(7)	1.503(7)
C(6) - C(8)	1.500(8)	C(9) - C(10)	1.503(6)
C(10) - C(11)	1.322(5)	C(11) - C(12)	1.509(6)
C(11) - C(13)	1.509(5)	C(13) - C(14)	1.508(5)
C(13) - C(15)	1.476(5)	C(15) - C(16)	1.526(5)
C(15) - C(17)	1.404(5)	C(17) - C(18)	1.491(6)
C(17) - C(19)	1.401(5)	C(19) - C(20)	1.488(5)
C(20) - C(21)	1.384(5)	C(20) - C(25)	1.394(5)
C(21) - C(22)	1.379(6)	C(22) - C(23)	1.372(6)
C(23) - C(24)	1.361(6)	C(24) - C(25)	1.381(6)

S(1) - W(1) - C(10)	83.4(1)	S(1) - W(1) - C(13)	121.2(1)
S(1) - W(1) - C(15)	91.4(1)	S(1) - W(1) - C(17)	88.3(1)
S(1) - W(1) - C(19)	104.8(1)	C(10) - W(1) - C(13)	60.4(2)
C(10) - W(1) - C(15)	74.2(2)	C(10) - W(1) - C(17)	107.0(2)
C(10) - W(1) - C(19)	137.7(2)	C(13) - W(1) - C(15)	37.7(2)
C(13) - W(1) - C(17)	63.9(2)	C(13) - W(1) - C(19)	80.9(2)
C(15) - W(1) - C(17)	33.6(2)	C(15) - W(1) - C(19)	64.3(2)
C(17) - W(1) - C(19)	34.2(2)	W(1) - S(1) - C(6)	109.8(2)
C(2) - C(1) - C(5)	107.7(4)	C(1) - C(2) - C(3)	108.6(4)
C(2) - C(3) - C(4)	108.4(4)	C(3) - C(4) - C(5)	106.8(4)
C(1) - C(5) - C(4)	108.5(5)	S(1) - C(6) - C(7)	112.0(3)
S(1) - C(6) - C(8)	107.6(4)	C(7) - C(6) - C(8)	112.6(5)
F(1) - C(9) - F(2)	104.0(4)	F(1) - C(9) - F(3)	106.5(4)
F(1) - C(9) - C(10)	112.1(4)	F(2) - C(9) - F(3)	107.3(4)
F(2) - C(9) - C(10)	113.2(4)	F(3) - C(9) - C(10)	113.2(4)
W(1) - C(10) - C(9)	135.0(3)	W(1) - C(10) - C(11)	98.5(3)
C(9) - C(10) - C(11)	126.1(4)	C(10) - C(11) - C(12)	129.9(4)
C(10) - C(11) - C(13)	104.0(3)	C(12) - C(11) - C(13)	126.1(4)
F(4) - C(12) - F(5)	106.4(4)	F(4) - C(12) - F(6)	107.0(4)
F(4) - C(12) - C(11)	111.8(4)	F(5) - C(12) - F(6)	106.1(4)
F(5) - C(12) - C(11)	111.9(4)	F(6) - C(12) - C(11)	113.2(4)
W(1) - C(13) - C(11)	94.8(3)	W(1) - C(13) - C(14)	126.6(3)
W(1) - C(13) - C(15)	76.9(2)	C(11) - C(13) - C(14)	119.6(3)
C(11) - C(13) - C(15)	112.2(3)	C(14) - C(13) - C(15)	118.1(3)
F(7) - C(14) - F(8)	106.0(4)	F(7) - C(14) - F(9)	106.0(4)
F(7) - C(14) - C(13)	110.2(4)	F(8) - C(14) - F(9)	105.9(4)
F(8) - C(14) - C(13)	112.6(4)	F(9) - C(14) - C(13)	115.5(4)
W(1) - C(15) - C(13)	65.4(2)	W(1) - C(15) - C(16)	134.6(3)
W(1) - C(15) - C(17)	78.6(2)	C(13) - C(15) - C(16)	121.5(3)
C(13) - C(15) - C(17)	119.9(3)	C(16) - C(15) - C(17)	118.1(3)
F(10) - C(16) - F(11)	105.0(4)	F(10) - C(16) - F(12)	103.9(4)
F(10) - C(16) - C(15)	109.7(3)	F(11) - C(16) - F(12)	108.8(4)
F(11) - C(16) - C(15)	114.6(4)	F(12) - C(16) - C(15)	113.9(4)
W(1) - C(17) - C(15)	67.8(2)	W(1) - C(17) - C(18)	142.3(3)
W(1) - C(17) - C(19)	54.1(2)	C(15) - C(17) - C(18)	124.4(4)
C(15) - C(17) - C(19)	113.2(3)	C(18) - C(17) - C(19)	121.6(4)
W(1) - C(19) - C(20)	91.7(2)	W(1) - C(19) - C(25)	144.9(3)
C(19) - C(20) - C(25)	122.8(3)	C(19) - C(20) - C(21)	119.0(3)
C(19) - C(20) - C(25)	123.3(3)	C(21) - C(20) - C(25)	117.6(4)
C(20) - C(21) - C(22)	120.6(4)	C(21) - C(22) - C(23)	120.7(4)
C(22) - C(23) - C(24)	119.7(4)	C(23) - C(24) - C(25)	120.1(4)
C(20) - C(25) - C(24)	121.1(4)		

Table 3.6. Fractional atomic coordinates and displacement parameters (\AA^2) of the crystal structure of complex (10).

	X/A	Y/B	Z/C	U
H	.34260(8)	.06154(5)	.14820(4)	.036
S	.2244(7)	.1168(4)	.0569(3)	.056
F(1)	.3010(20)	-.0945(12)	-.0200(9)	.119(6)
F(2)	.5028(21)	-.1388(12)	.0107(8)	.117(6)
F(3)	.4701(18)	-.0061(11)	-.0125(8)	.104(5)
F(4)	.2801(16)	-.2498(10)	.0714(8)	.090(5)
F(5)	.4718(17)	-.2492(10)	.1298(7)	.100(5)
F(6)	.2639(16)	-.2552(10)	.1742(7)	.086(4)
F(7)	-.0423(17)	.0629(13)	.3119(8)	.111(5)
F(8)	-.1308(19)	.1156(11)	.2226(8)	.102(5)
F(9)	-.1149(16)	-.0233(9)	.2395(7)	.093(5)
F(10)	.0889(17)	.2178(10)	.2829(7)	.094(5)
F(11)	.3036(16)	.2401(10)	.2616(7)	.088(5)
F(12)	.2503(18)	.1458(11)	.3359(8)	.115(5)
O(1)	.4312(17)	-.1321(10)	.2686(7)	.069(4)
O(2)	.2128(16)	-.0806(11)	.2981(7)	.072(5)
O(3)	-.0641(16)	-.0068(10)	.1048(7)	.065(4)
O(4)	.0424(17)	-.1391(10)	.1161(7)	.067(4)
C(1)	.4081(31)	-.0691(23)	.0155(14)	.091(8)
C(2)	.3634(19)	-.0589(15)	.0856(8)	.049(4)
C(3)	.3441(34)	-.2204(15)	.1261(11)	.075(7)
C(4)	.3287(22)	-.1204(11)	.1305(8)	.041(5)
C(5)	.2448(36)	-.1071(22)	.3668(17)	.123(11)
C(6)	.3241(25)	-.0966(13)	.2569(10)	.055(6)
C(7)	.2748(19)	-.0676(15)	.1874(9)	.046(4)
C(8)	-.0603(28)	-.1774(17)	.0703(13)	.082(7)
C(9)	.0301(20)	-.0541(15)	.1246(9)	.051(5)
C(10)	.1425(20)	-.0218(11)	.1726(8)	.037(4)
C(11)	-.0568(27)	.0549(22)	.2505(13)	.084(7)
C(12)	.1013(18)	.0476(12)	.2192(9)	.039(4)
C(13)	.2173(32)	.1774(21)	.2800(15)	.088(8)
C(14)	.2133(22)	.1020(14)	.2285(11)	.043(5)
C(R1)	.5884(22)	.0452(14)	.1269(10)	.056(6)
C(R2)	.5708(28)	.0322(17)	.1944(13)	.076(8)
C(R3)	.5262(29)	.1128(18)	.2198(12)	.075(7)
C(R4)	.5160(26)	.1725(16)	.1704(11)	.072(7)
C(R5)	.5546(25)	.1275(16)	.1090(11)	.064(6)
C(T1)	.2497(21)	.2327(13)	.0585(10)	.049(5)
C(T2)	.3310(29)	.2753(14)	.0120(11)	.065(6)
C(T3)	.3575(27)	.3651(14)	.0169(10)	.062(6)
C(T4)	.3099(29)	.4100(17)	.0665(14)	.089(9)
C(T5)	.2122(28)	.3713(19)	.1128(13)	.086(8)
C(T6)	.1880(24)	.2801(15)	.1067(11)	.065(7)
C(T7)	.3315(34)	.5116(18)	.0752(13)	.102(9)
H(51)	.16681	-.09511	.39663	.148
H(52)	.32941	-.07590	.38118	.148
H(53)	.26464	-.16915	.36729	.148
H(81)	-.05025	-.23955	.06303	.098
H(82)	-.04663	-.14775	.02924	.098
H(83)	-.15690	-.16631	.08565	.098
H(R1)	.61987	-.00019	.09746	.068
H(R2)	.58989	-.01986	.21961	.092
H(R3)	.50230	.12568	.26454	.090
H(R4)	.48930	.23323	.17526	.086
H(R5)	.55348	.15270	.06588	.077
H(T2)	.37423	.24225	-.02296	.078
H(T3)	.40986	.39477	-.01717	.074
H(T5)	.16341	.40538	.14568	.103
H(T6)	.12855	.24953	.13789	.078
H(T71)	.29552	.54529	.11144	.122
H(T72)	.29050	.53868	.03694	.122
H(T73)	.43431	.52110	.07514	.122

Anisotropic displacement parameters (\AA^2).

	U11	U22	U33	U12	U13	U23
H	.0356(3)	.0366(3)	.0367(4)	-.0018(5)	-.0006(5)	-.0040(5)
S	.067(3)	.044(3)	.057(4)	.000(3)	-.005(3)	.009(3)

Table 3.7. Selected bond distances (Å) and angles (°) of complex (10).

W - S	2.318(7)	W - C(2)	2.240(22)
W - C(7)	2.209(23)	W - C(10)	2.294(18)
W - C(14)	2.119(22)	W - C(R1)	2.324(21)
W - C(R2)	2.351(26)	W - C(R3)	2.370(27)
W - C(R4)	2.368(24)	W - C(R5)	2.340(24)
S - C(T1)	1.775(21)	F(1) - C(1)	1.286(34)
F(2) - C(1)	1.377(38)	F(3) - C(1)	1.253(37)
F(4) - C(3)	1.340(30)	F(5) - C(3)	1.261(35)
F(6) - C(3)	1.339(30)	F(7) - C(11)	1.265(31)
F(8) - C(11)	1.282(34)	F(9) - C(11)	1.322(35)
F(10) - C(13)	1.337(34)	F(11) - C(13)	1.297(35)
F(12) - C(13)	1.275(36)	O(1) - C(6)	1.151(28)
O(2) - C(5)	1.490(38)	O(2) - C(6)	1.351(27)
O(3) - C(9)	1.198(25)	O(4) - C(8)	1.453(30)
O(4) - C(9)	1.307(28)	C(1) - C(2)	1.497(33)
C(2) - C(4)	1.347(27)	C(3) - C(4)	1.528(29)
C(4) - C(7)	1.497(26)	C(6) - C(7)	1.552(28)
C(7) - C(10)	1.438(26)	C(9) - C(10)	1.510(26)
C(10) - C(12)	1.479(25)	C(11) - C(12)	1.598(31)
C(12) - C(14)	1.337(28)	C(13) - C(14)	1.555(38)
C(R1) - C(R2)	1.401(33)	C(R1) - C(R5)	1.339(32)
C(R2) - C(R3)	1.390(38)	C(R3) - C(R4)	1.360(36)
C(R4) - C(R5)	1.471(33)	C(T1) - C(T2)	1.373(31)
C(T1) - C(T6)	1.346(31)	C(T2) - C(T3)	1.389(30)
C(T3) - C(T4)	1.298(36)	C(T4) - C(T5)	1.433(39)
C(T4) - C(T7)	1.565(38)	C(T5) - C(T6)	1.408(37)

S - W - C(2)	82.9(5)	S - W - C(7)	118.7(5)
S - W - C(10)	89.7(5)	S - W - C(14)	104.6(6)
C(2) - W - C(7)	60.4(7)	C(2) - W - C(10)	75.1(7)
C(2) - W - C(14)	136.6(8)	C(7) - W - C(10)	37.2(7)
C(7) - W - C(14)	79.5(8)	C(10) - W - C(14)	62.5(8)
W - S - C(T1)	106.4(8)	C(5) - O(2) - C(6)	112.8(19)
C(8) - O(4) - C(9)	115.1(17)	F(1) - C(1) - F(2)	102.6(25)
F(1) - C(1) - F(3)	108.8(25)	F(1) - C(1) - C(2)	111.0(23)
F(2) - C(1) - F(3)	105.3(23)	F(2) - C(1) - C(2)	108.8(23)
F(3) - C(1) - C(2)	118.9(27)	W - C(2) - C(1)	130.8(19)
W - C(2) - C(4)	99.0(12)	C(1) - C(2) - C(4)	130.1(23)
F(4) - C(3) - F(5)	110.2(21)	F(4) - C(3) - F(6)	103.6(22)
F(4) - C(3) - C(4)	109.9(19)	F(5) - C(3) - F(6)	109.7(20)
F(5) - C(3) - C(4)	115.4(23)	F(6) - C(3) - C(4)	107.4(20)
C(2) - C(4) - C(3)	128.8(18)	C(2) - C(4) - C(7)	103.7(17)
C(3) - C(4) - C(7)	127.5(17)	O(1) - C(6) - O(2)	127.4(20)
O(1) - C(6) - C(7)	125.1(20)	O(2) - C(6) - C(7)	107.2(17)
W - C(7) - C(4)	95.7(12)	W - C(7) - C(6)	120.0(14)
W - C(7) - C(10)	74.6(12)	C(4) - C(7) - C(6)	117.4(18)
C(4) - C(7) - C(10)	112.2(15)	C(6) - C(7) - C(10)	125.3(16)
O(3) - C(9) - O(4)	127.6(19)	O(3) - C(9) - C(10)	121.5(20)
O(4) - C(9) - C(10)	110.3(17)	W - C(10) - C(7)	68.2(11)
W - C(10) - C(9)	126.3(12)	W - C(10) - C(12)	87.4(11)
C(7) - C(10) - C(9)	124.3(17)	C(7) - C(10) - C(12)	115.5(16)
C(9) - C(10) - C(12)	118.4(16)	F(7) - C(11) - F(8)	115.3(26)
F(7) - C(11) - F(9)	107.3(24)	F(7) - C(11) - C(12)	107.8(20)
F(8) - C(11) - F(9)	110.8(22)	F(8) - C(11) - C(12)	111.0(22)
F(9) - C(11) - C(12)	104.0(21)	C(10) - C(12) - C(11)	123.1(18)
C(10) - C(12) - C(14)	109.5(17)	C(11) - C(12) - C(14)	127.4(20)
F(10) - C(13) - F(11)	102.9(24)	F(10) - C(13) - F(12)	110.1(24)
F(10) - C(13) - C(14)	110.3(22)	F(11) - C(13) - F(12)	112.9(25)
F(11) - C(13) - C(14)	111.0(23)	F(12) - C(13) - C(14)	109.5(24)
W - C(14) - C(12)	98.5(14)	W - C(14) - C(13)	136.3(17)
C(12) - C(14) - C(13)	124.7(20)	C(R2) - C(R1) - C(R5)	111.9(21)
C(R1) - C(R2) - C(R3)	106.1(22)	C(R2) - C(R3) - C(R4)	109.3(23)
C(R3) - C(R4) - C(R5)	107.8(22)	C(R1) - C(R5) - C(R4)	104.9(20)
S - C(T1) - C(T2)	121.8(17)	S - C(T1) - C(T6)	119.2(17)
C(T2) - C(T1) - C(T6)	119.0(20)	C(T1) - C(T2) - C(T3)	120.6(21)
C(T2) - C(T3) - C(T4)	120.9(23)	C(T3) - C(T4) - C(T5)	120.8(25)
C(T3) - C(T4) - C(T7)	124.2(25)	C(T5) - C(T4) - C(T7)	114.3(24)
C(T4) - C(T5) - C(T6)	116.5(24)	C(T1) - C(T6) - C(T5)	121.6(22)

Table 3.8. Fractional atomic coordinates and displacement parameters (\AA^2) of the crystal structure of complex (11).

	X/A	Y/B	Z/C	U
MO(A)	.13982(6)	.20293(4)	.49559(4)	.041
MO(B)	.32945(6)	.29559(4)	.00411(4)	.037
S(A)	.23296(17)	.27392(11)	.55408(12)	.040
S(B)	.54513(16)	.22779(11)	-.05649(11)	.037
F(A1)	-.2602(6)	.3203(5)	.5632(5)	.145
F(A2)	-.1849(6)	.3377(8)	.4359(5)	.199
F(A3)	-.2417(6)	.2152(5)	.5236(6)	.166
F(A4)	.0340(5)	.2037(3)	.7350(3)	.070
F(A5)	-.1628(5)	.2100(4)	.7312(3)	.088
F(A6)	-.0702(5)	.3248(3)	.6430(3)	.078
F(A7)	-.0442(6)	.0507(4)	.5592(4)	.103
F(A8)	-.0761(6)	-.0102(4)	.6921(4)	.109
F(A9)	.1054(5)	-.0565(4)	.6228(6)	.149
F(A10)	.1740(7)	-.1002(4)	.7964(4)	.116
F(A11)	.0043(5)	.0100(4)	.8057(4)	.095
F(A12)	.1815(6)	.0090(4)	.8345(3)	.101
F(B1)	.1106(6)	.1751(5)	-.0652(5)	.142
F(B2)	.0854(9)	.1548(7)	.0630(5)	.196
F(B3)	-.0110(5)	.2787(5)	-.0217(6)	.168
F(B4)	.4478(4)	.3014(3)	-.2374(3)	.065
F(B5)	.2550(5)	.2936(4)	-.2339(3)	.083
F(B6)	.3752(4)	.1774(3)	-.1474(3)	.065
F(B7)	.0566(4)	.4450(4)	-.0570(4)	.085
F(B8)	.0920(5)	.5116(4)	-.1917(4)	.102
F(B9)	.1613(5)	.5540(4)	-.1160(5)	.141
F(B10)	.3477(7)	.6040(4)	-.2922(4)	.131
F(B11)	.2984(5)	.4934(4)	-.3053(3)	.086
F(B12)	.4963(6)	.4999(4)	-.3321(3)	.099
O(A1)	.4538(5)	.1134(4)	.6929(4)	.061
O(A2)	.4316(6)	-.0369(4)	.7515(4)	.085
O(A3)	.4845(5)	.1978(3)	.4499(4)	.061
O(A4)	.5138(5)	.0509(4)	.5486(4)	.067
O(B1)	.7378(5)	.3931(4)	-.1932(4)	.063
O(B2)	.6250(6)	.5428(4)	-.2466(4)	.085
O(B3)	.6248(5)	.3017(4)	.0518(4)	.064
O(B4)	.6073(5)	.4498(4)	-.0475(4)	.065
C(A1)	.2518(8)	.2599(6)	.3519(5)	.062
C(A2)	.1158(10)	.2824(6)	.3509(5)	.078
C(A3)	.0720(10)	.1957(7)	.3856(5)	.076
C(A4)	.1825(10)	.1227(6)	.4046(6)	.075
C(A5)	.2938(8)	.1585(5)	.3859(5)	.060
C(A6)	.3297(7)	.3643(5)	.4870(5)	.048
C(A7)	.4296(9)	.3537(6)	.5364(6)	.080
C(A8)	.2261(9)	.4570(5)	.4698(6)	.072
C(A9)	-.1787(10)	.2785(7)	.5133(7)	.086
C(A10)	-.0461(7)	.2357(5)	.5332(5)	.050
C(A11)	-.0049(7)	.1864(5)	.6143(5)	.050
C(A12)	-.0471(8)	.2293(5)	.6809(5)	.055
C(A13)	.0751(7)	.0921(5)	.6261(5)	.045
C(A14)	.0180(8)	.0206(6)	.6253(6)	.061
C(A15)	.1739(7)	.0537(4)	.6873(5)	.045
C(A16)	.1363(9)	-.0075(6)	.7804(6)	.069
C(A17)	.2844(7)	.0834(5)	.6574(5)	.050
C(A18)	.3969(7)	.0451(6)	.7085(5)	.056
C(A19)	.5766(8)	.0860(7)	.7219(7)	.088
C(A20)	.3093(6)	.1539(4)	.5672(4)	.038
C(A21)	.4463(7)	.1286(5)	.5205(5)	.053
C(A22)	.6116(9)	.1743(7)	.3977(7)	.087
C(B1)	.3620(8)	.2375(6)	.1479(5)	.057
C(B2)	.2459(9)	.2143(6)	.1480(5)	.064
C(B3)	.1551(8)	.2992(7)	.1173(6)	.073
C(B4)	.2090(9)	.3757(6)	.0968(5)	.070
C(B5)	.3367(8)	.3383(6)	.1164(5)	.060
C(B6)	.6663(6)	.1370(4)	.0104(4)	.040
C(B7)	.8044(7)	.1476(6)	-.0415(6)	.072
C(B8)	.6378(8)	.0444(5)	.0237(6)	.075
C(B9)	.1016(9)	.2174(7)	-.0131(7)	.077
C(B10)	.2116(6)	.2616(5)	-.0346(5)	.045
C(B11)	.2790(6)	.3146(5)	-.1151(5)	.046
C(B12)	.3407(7)	.2725(5)	-.1841(5)	.055
C(B13)	.2801(7)	.4078(4)	-.1256(5)	.044
C(B14)	.1506(8)	.4776(6)	-.1217(6)	.060
C(B15)	.3964(7)	.4485(5)	-.1851(5)	.049

Table 3.8(cont).

	X/A	Y/B	Z/C	U
C(B16)	.3833(10)	.5134(7)	-.2790(6)	.077
C(B17)	.5076(7)	.4203(5)	-.1558(5)	.045
C(B18)	.6282(8)	.4614(6)	-.2055(5)	.055
C(B19)	.8601(9)	.4253(7)	-.2232(6)	.090
C(B20)	.5188(6)	.3477(4)	-.0678(4)	.035
C(B21)	.5869(7)	.3727(6)	-.0203(5)	.050
C(B22)	.6854(10)	.3248(7)	.1003(7)	.094
H(A1)	.30569	.30472	.33266	.064
H(A2)	.06163	.34496	.32971	.078
H(A3)	-.01615	.18963	.39377	.069
H(A4)	.18237	.05654	.42757	.080
H(A5)	.38074	.12217	.39548	.063
H(A6)	.37581	.35381	.43508	.048
H(A71)	.47228	.40612	.50713	.080
H(A72)	.49479	.29584	.54238	.080
H(A73)	.38359	.35279	.59259	.080
H(A81)	.16216	.46024	.44009	.069
H(A82)	.26467	.51127	.43722	.069
H(A83)	.18242	.46099	.52474	.069
H(A191)	.61843	.13543	.71054	.087
H(A192)	.63688	.04527	.69409	.087
H(A193)	.56327	.05093	.78345	.087
H(A221)	.60371	.13934	.36755	.081
H(A222)	.67694	.13635	.43379	.081
H(A223)	.63779	.23191	.35656	.081
H(B1)	.44170	.19410	.16595	.062
H(B2)	.23390	.15154	.16557	.071
H(B3)	.06778	.30453	.11137	.077
H(B4)	.16736	.44175	.07374	.084
H(B5)	.39727	.37430	.10945	.066
H(B6)	.65456	.14610	.06294	.046
H(B71)	.87097	.09839	-.01126	.072
H(B72)	.81729	.20801	-.05157	.072
H(B73)	.81158	.14446	-.09627	.072
H(B81)	.63591	.04860	-.03248	.068
H(B82)	.55227	.04036	.05888	.068
H(B83)	.70442	-.00916	.04935	.068
H(B191)	.86316	.46159	-.19288	.082
H(B192)	.87124	.46307	-.28421	.082
H(B193)	.93228	.37095	-.21388	.082
H(B221)	.71273	.27418	.15193	.080
H(B222)	.62356	.37446	.11649	.080
H(B223)	.76146	.34805	.06295	.080

Anisotropic displacement parameters (\AA^2).

	U11	U22	U33	U12	U13	U23
MO(A)	.0430(4)	.0349(3)	.0344(4)	-.0132(3)	-.0085(3)	-.0123(3)
MO(B)	.0291(4)	.0344(3)	.0334(4)	-.0041(3)	-.0039(3)	-.0122(3)
S(A)	.037(1)	.031(1)	.039(1)	-.007(1)	-.008(1)	-.015(1)
S(B)	.031(1)	.032(1)	.036(1)	-.006(1)	-.006(1)	-.015(1)
F(A1)	.064(4)	.150(6)	.185(8)	.032(4)	-.060(4)	-.111(6)
F(A2)	.068(4)	.295(12)	.096(6)	-.004(6)	-.043(4)	.048(7)
F(A3)	.081(4)	.122(5)	.259(10)	-.014(4)	-.078(5)	-.090(6)
F(A4)	.069(3)	.069(3)	.052(3)	-.007(2)	-.016(3)	-.032(2)
F(A5)	.059(3)	.101(4)	.074(4)	-.021(3)	.014(3)	-.050(3)
F(A6)	.080(3)	.059(3)	.075(3)	-.005(2)	-.008(3)	-.044(3)
F(A7)	.118(5)	.088(4)	.085(4)	-.063(4)	-.022(4)	-.030(3)
F(A8)	.099(4)	.122(5)	.078(4)	-.082(4)	.017(3)	-.034(4)
F(A9)	.069(4)	.074(4)	.266(10)	-.004(3)	-.026(5)	-.103(5)
F(A10)	.154(6)	.053(3)	.082(4)	-.026(3)	-.004(4)	.017(3)
F(A11)	.067(4)	.106(4)	.068(4)	-.041(3)	.000(3)	-.004(3)
F(A12)	.087(4)	.126(5)	.046(3)	-.032(4)	-.026(3)	.001(3)

Table 3.8(cont).

	U11	U22	U33	U12	U13	U23
F(B1)	.083(4)	.152(6)	.165(7)	-.068(4)	.017(4)	-.102(6)
F(B2)	.197(8)	.228(9)	.106(6)	-.186(8)	-.029(6)	.020(6)
F(B3)	.038(3)	.133(5)	.286(11)	-.013(3)	-.013(5)	-.123(7)
F(B4)	.063(3)	.070(3)	.042(3)	-.022(2)	.004(2)	-.027(2)
F(B5)	.073(3)	.099(4)	.057(3)	-.015(3)	-.025(3)	-.039(3)
F(B6)	.072(3)	.044(3)	.058(3)	-.009(2)	-.003(2)	-.029(2)
F(B7)	.045(3)	.085(3)	.085(4)	.014(3)	-.005(3)	-.031(3)
F(B8)	.064(3)	.119(5)	.080(4)	.027(3)	-.045(3)	-.035(4)
F(B9)	.071(4)	.073(4)	.245(9)	.011(3)	-.046(5)	-.101(5)
F(B10)	.188(7)	.036(3)	.122(6)	-.015(3)	-.089(5)	.016(3)
F(B11)	.083(4)	.092(4)	.051(3)	-.029(3)	-.034(3)	.001(3)
F(B12)	.086(4)	.122(5)	.046(3)	-.055(4)	-.007(3)	.004(3)
U(A1)	.044(3)	.063(3)	.056(3)	-.001(3)	-.018(3)	-.028(3)
U(A2)	.069(4)	.056(4)	.084(5)	.003(3)	-.037(4)	.010(4)
U(A3)	.046(3)	.045(3)	.069(4)	-.014(2)	.015(3)	-.030(3)
U(A4)	.060(4)	.045(3)	.071(4)	.008(3)	-.011(3)	-.030(3)
U(B1)	.041(3)	.072(4)	.056(4)	-.028(3)	.008(3)	-.030(3)
U(B2)	.084(5)	.056(4)	.084(5)	-.043(3)	-.018(4)	.001(4)
U(B3)	.060(4)	.044(3)	.073(4)	.002(3)	-.034(3)	-.028(3)
U(B4)	.072(4)	.042(3)	.067(4)	-.017(3)	-.012(3)	-.028(3)
C(A1)	.065(6)	.060(5)	.044(5)	-.024(4)	-.007(4)	-.019(4)
C(A2)	.102(8)	.063(6)	.046(5)	-.023(5)	-.033(5)	-.009(5)
C(A3)	.083(7)	.083(6)	.046(5)	-.038(6)	-.021(5)	-.025(5)
C(A4)	.086(7)	.067(6)	.055(6)	-.038(6)	-.002(5)	-.030(5)
C(A5)	.072(6)	.057(5)	.036(4)	-.025(4)	.004(4)	-.025(4)
C(A6)	.040(4)	.041(4)	.048(5)	-.011(3)	-.002(4)	-.020(4)
C(A7)	.084(7)	.075(6)	.059(6)	-.044(5)	-.017(5)	-.014(5)
C(A8)	.072(6)	.033(4)	.077(6)	-.008(4)	-.004(5)	-.010(4)
C(A9)	.072(7)	.070(6)	.094(8)	-.015(5)	-.037(6)	-.029(6)
C(A10)	.042(5)	.046(4)	.044(5)	-.011(3)	-.016(4)	-.012(4)
C(A11)	.044(4)	.039(4)	.053(5)	-.019(3)	-.010(4)	-.014(4)
C(A12)	.051(5)	.048(5)	.046(5)	-.005(4)	-.001(4)	-.022(4)
C(A13)	.041(4)	.036(4)	.041(4)	-.018(3)	.004(4)	-.012(4)
C(A14)	.046(5)	.057(5)	.063(6)	-.025(4)	.006(5)	-.028(5)
C(A15)	.044(5)	.026(4)	.048(5)	-.005(3)	-.006(4)	-.008(3)
C(A16)	.066(7)	.058(6)	.047(6)	-.015(5)	-.004(5)	.004(5)
C(A17)	.049(5)	.028(4)	.054(5)	-.003(3)	-.011(4)	-.012(4)
C(A18)	.036(4)	.056(5)	.050(5)	-.002(4)	-.012(4)	-.011(5)
C(A19)	.038(5)	.113(8)	.084(7)	-.011(5)	-.022(5)	-.038(6)
C(A20)	.038(4)	.029(4)	.033(4)	-.005(3)	.001(3)	-.014(3)
C(A21)	.048(5)	.043(5)	.053(5)	-.014(4)	-.005(4)	-.026(4)
C(A22)	.059(6)	.083(6)	.086(7)	-.023(5)	.029(5)	-.049(6)
C(B1)	.062(6)	.054(5)	.029(4)	.009(4)	-.010(4)	-.013(4)
C(B2)	.068(6)	.061(5)	.034(5)	-.022(5)	.005(4)	-.006(4)
C(B3)	.050(5)	.084(7)	.057(6)	-.013(5)	.003(5)	-.029(5)
C(B4)	.066(6)	.066(6)	.049(5)	.016(5)	-.005(5)	-.032(5)
C(B5)	.058(5)	.066(5)	.038(5)	-.007(4)	-.003(4)	-.029(4)
C(B6)	.032(4)	.038(4)	.036(4)	-.011(3)	-.004(3)	-.013(3)
C(B7)	.036(5)	.069(5)	.078(7)	.004(4)	-.017(5)	-.022(5)
C(B8)	.069(6)	.042(4)	.091(7)	-.006(4)	-.041(5)	-.013(5)
C(B9)	.048(6)	.074(6)	.083(7)	-.027(5)	-.009(5)	-.020(6)
C(B10)	.028(4)	.034(4)	.056(5)	-.007(3)	-.009(4)	-.010(4)
C(B11)	.025(4)	.044(4)	.052(5)	.001(3)	-.014(4)	-.019(4)
C(B12)	.045(5)	.054(5)	.051(5)	-.014(4)	-.011(4)	-.024(4)
C(B13)	.038(4)	.032(4)	.049(5)	-.007(3)	-.015(4)	-.015(3)
C(B14)	.041(5)	.056(5)	.063(6)	-.003(4)	-.019(5)	-.023(4)
C(B15)	.055(5)	.032(4)	.044(5)	-.006(3)	-.022(4)	-.008(4)
C(B16)	.074(7)	.065(6)	.065(6)	-.035(5)	-.023(6)	-.001(5)
C(B17)	.053(5)	.032(4)	.038(4)	-.017(3)	.002(4)	-.015(3)
C(B18)	.056(5)	.048(5)	.045(5)	-.027(4)	-.009(4)	-.012(4)
C(B19)	.055(6)	.111(8)	.078(7)	-.041(5)	-.004(5)	-.039(6)
C(B20)	.037(4)	.023(3)	.034(4)	-.008(3)	-.013(3)	-.006(3)
C(B21)	.034(4)	.050(5)	.054(5)	.004(4)	-.011(4)	-.034(4)
C(B22)	.082(7)	.094(7)	.088(7)	-.012(6)	-.035(6)	-.054(6)

Table 3.9. Selected bond distances (Å) and angles (°) of complex (11).

MO - S	2.400(2)	2.405(2)	MO - C(1)	2.353(9)	2.355(7)
MO - C(2)	2.328(9)	2.322(9)	MO - C(3)	2.310(8)	2.319(10)
MO - C(4)	2.329(8)	2.356(9)	MO - C(5)	2.363(9)	2.378(7)
MO - C(10)	1.915(8)	1.921(6)	MO - C(11)	2.179(8)	2.176(7)
MO - C(13)	2.273(8)	2.273(8)	MO - C(20)	2.267(6)	2.282(7)
S - C(6)	1.833(7)	1.846(7)	S - C(20)	1.784(7)	1.789(6)
F(1) - C(9)	1.300(14)	1.324(12)	F(2) - C(9)	1.273(15)	1.269(14)
F(3) - C(9)	1.326(11)	1.308(12)	F(4) - C(12)	1.326(9)	1.334(9)
F(5) - C(12)	1.350(10)	1.326(8)	F(6) - C(12)	1.350(9)	1.345(9)
F(7) - C(14)	1.334(10)	1.323(10)	F(8) - C(14)	1.334(11)	1.358(10)
F(9) - C(14)	1.313(11)	1.323(9)	F(10) - C(16)	1.338(11)	1.316(11)
F(11) - C(16)	1.352(11)	1.319(10)	F(12) - C(16)	1.335(9)	1.330(12)
O(1) - C(18)	1.320(9)	1.337(10)	O(1) - C(19)	1.437(9)	1.454(10)
O(2) - C(18)	1.194(11)	1.180(10)	O(3) - C(21)	1.312(10)	1.340(10)
O(3) - C(22)	1.461(12)	1.433(10)	O(4) - C(21)	1.196(10)	1.189(9)
C(1) - C(2)	1.413(13)	1.428(11)	C(1) - C(5)	1.434(12)	1.426(12)
C(2) - C(3)	1.433(13)	1.391(14)	C(3) - C(4)	1.389(14)	1.394(12)
C(4) - C(5)	1.400(12)	1.407(12)	C(6) - C(7)	1.503(10)	1.533(11)
C(6) - C(8)	1.525(11)	1.524(9)	C(9) - C(10)	1.460(13)	1.460(10)
C(10) - C(11)	1.411(11)	1.410(11)	C(11) - C(12)	1.496(11)	1.526(11)
C(11) - C(13)	1.449(10)	1.434(9)	C(13) - C(14)	1.480(9)	1.504(11)
C(13) - C(15)	1.533(10)	1.510(11)	C(15) - C(16)	1.498(12)	1.527(13)
C(15) - C(17)	1.316(10)	1.320(10)	C(17) - C(18)	1.534(10)	1.517(11)
C(17) - C(20)	1.489(11)	1.477(10)	C(20) - C(21)	1.503(11)	1.498(8)

S - MO - C(10)	104.9(3)	104.4(2)	S - MO - C(11)	83.1(2)	83.3(2)
S - MO - C(13)	94.8(2)	94.5(2)	S - MO - C(20)	44.8(2)	44.8(2)
C(10) - MO - C(11)	39.7(3)	39.6(3)	C(10) - MO - C(13)	69.0(3)	69.2(3)
C(10) - MO - C(20)	130.8(3)	130.5(3)	C(11) - MO - C(13)	37.9(3)	37.5(3)
C(11) - MO - C(20)	92.3(3)	92.2(3)	C(13) - MO - C(20)	75.7(3)	75.5(3)
MO - S - C(6)	123.5(3)	123.1(3)	MO - S - C(20)	63.6(3)	64.0(2)
C(6) - S - C(20)	118.1(3)	117.8(3)	C(18) - O(1) - C(19)	117.3(7)	115.8(7)
C(21) - O(3) - C(22)	117.0(6)	116.1(7)	C(2) - C(1) - C(5)	107.4(7)	106.8(7)
C(1) - C(2) - C(3)	108.5(8)	107.3(8)	C(2) - C(3) - C(4)	106.4(8)	110.2(8)
C(3) - C(4) - C(5)	111.0(8)	107.2(8)	C(1) - C(5) - C(4)	106.7(7)	108.5(7)
S - C(6) - C(7)	108.5(6)	108.2(5)	S - C(6) - C(8)	103.3(5)	102.9(5)
C(7) - C(6) - C(8)	114.5(7)	111.8(6)	F(1) - C(9) - F(2)	106.2(9)	105.3(9)
F(1) - C(9) - F(3)	101.4(8)	103.1(8)	F(1) - C(9) - C(10)	113.8(9)	113.3(8)
F(2) - C(9) - F(3)	104.0(9)	106.1(9)	F(2) - C(9) - C(10)	115.1(9)	115.2(9)
F(3) - C(9) - C(10)	112.8(8)	112.8(8)	MO - C(10) - C(9)	150.1(7)	148.6(7)
MO - C(10) - C(11)	80.3(5)	80.0(4)	C(9) - C(10) - C(11)	129.6(8)	131.3(8)
MO - C(11) - C(10)	60.1(4)	60.4(4)	MO - C(11) - C(12)	135.8(5)	135.5(5)
MO - C(11) - C(13)	74.5(4)	74.9(4)	C(10) - C(11) - C(12)	121.3(7)	119.8(6)
C(10) - C(11) - C(13)	113.4(7)	115.1(7)	C(12) - C(11) - C(13)	125.2(7)	125.1(7)
F(4) - C(12) - F(5)	105.7(7)	106.2(7)	F(4) - C(12) - F(6)	106.0(6)	106.5(6)
F(4) - C(12) - C(11)	117.4(7)	115.7(6)	F(5) - C(12) - F(6)	104.9(6)	106.9(6)
F(5) - C(12) - C(11)	111.0(6)	110.4(6)	F(6) - C(12) - C(11)	110.9(7)	110.8(7)
MO - C(13) - C(11)	67.5(4)	67.5(4)	MO - C(13) - C(14)	116.9(6)	115.7(6)
MO - C(13) - C(15)	112.6(5)	112.7(5)	C(11) - C(13) - C(14)	119.4(6)	118.2(6)
C(11) - C(13) - C(15)	115.8(6)	117.1(6)	C(14) - C(13) - C(15)	115.7(6)	116.3(6)
F(7) - C(14) - F(8)	100.6(6)	102.2(6)	F(7) - C(14) - F(9)	105.6(8)	105.7(7)
F(7) - C(14) - C(13)	114.8(7)	115.7(7)	F(8) - C(14) - F(9)	104.8(8)	104.5(7)
F(8) - C(14) - C(13)	116.2(7)	114.9(7)	F(9) - C(14) - C(13)	113.4(6)	112.6(6)
C(13) - C(15) - C(16)	119.3(7)	118.1(7)	C(13) - C(15) - C(17)	118.1(7)	118.6(7)
C(16) - C(15) - C(17)	122.3(7)	123.0(7)	F(10) - C(16) - F(11)	105.6(7)	108.3(8)
F(10) - C(16) - F(12)	108.6(8)	107.8(8)	F(10) - C(16) - C(15)	113.3(7)	113.4(8)
F(11) - C(16) - F(12)	104.6(7)	104.7(8)	F(11) - C(16) - C(15)	111.0(7)	111.5(8)
F(12) - C(16) - C(15)	113.1(7)	110.8(8)	C(15) - C(17) - C(18)	123.8(7)	124.2(7)
C(15) - C(17) - C(20)	120.1(7)	120.3(7)	C(18) - C(17) - C(20)	116.0(6)	115.4(6)
O(1) - C(18) - O(2)	126.0(7)	125.1(7)	O(1) - C(18) - C(17)	110.7(7)	110.6(7)
O(2) - C(18) - C(17)	123.1(7)	124.3(7)	MO - C(20) - S	71.5(3)	71.2(3)
MO - C(20) - C(17)	113.4(5)	112.8(5)	MO - C(20) - C(21)	118.1(5)	117.9(5)
S - C(20) - C(17)	114.5(5)	115.6(5)	S - C(20) - C(21)	120.8(5)	121.0(5)
C(17) - C(20) - C(21)	112.6(6)	112.3(6)	O(3) - C(21) - O(4)	124.1(7)	124.0(7)
O(3) - C(21) - C(20)	115.1(6)	115.1(6)	O(4) - C(21) - C(20)	120.7(7)	120.9(7)

Table 3.10. Fractional atomic coordinates and displacement parameters (\AA^2) of the crystal structure of complex (12).

	X/A	Y/B	Z/C	U
MO(1)	.21324(5)	.11628(4)	.21204(3)	.032
S(1)	-.06322(16)	.16672(13)	.06463(10)	.042
F(1)	.5628(5)	.1048(4)	.3809(4)	.115
F(2)	.4355(7)	.1035(4)	.4656(3)	.144
F(3)	.4416(5)	.0031(3)	.3814(3)	.100
F(4)	.1576(4)	.3116(3)	.3640(3)	.068
F(5)	.3533(4)	.2915(3)	.4389(3)	.083
F(6)	.2063(5)	.1978(3)	.4367(3)	.079
F(7)	.4986(4)	.3237(3)	.2533(3)	.085
F(8)	.5092(4)	.1952(3)	.2250(3)	.069
F(9)	.4181(4)	.2798(3)	.1242(3)	.083
F(10)	.3201(5)	.4225(3)	.2816(4)	.088
F(11)	.1368(5)	.4591(3)	.2062(4)	.100
F(12)	.2936(6)	.4350(4)	.1440(4)	.116
O(1)	-.0458(5)	.3940(3)	.0699(3)	.065
O(2)	-.1207(6)	.3634(4)	.1867(4)	.086
O(3)	-.0637(5)	.2053(3)	.2874(3)	.051
O(4)	.0409(4)	.0896(3)	.2566(3)	.039
C(1)	.1550(7)	-.0037(5)	.1261(5)	.061
C(2)	.1920(7)	.0566(5)	.0717(4)	.057
C(3)	.3221(7)	.0769(5)	.1040(4)	.050
C(4)	.3693(7)	.0262(5)	.1778(4)	.046
C(5)	.2660(8)	-.0229(5)	.1933(5)	.058
C(6)	-.2324(7)	.1789(5)	.0769(5)	.053
C(7)	-.2809(8)	.0988(7)	.1106(6)	.086
C(8)	-.3141(8)	.2011(7)	-.0125(7)	.107
C(9)	.4441(8)	.0862(6)	.3880(5)	.057
C(10)	.3415(6)	.1273(4)	.3198(4)	.042
C(11)	.2928(6)	.2104(4)	.3123(4)	.036
C(12)	.2508(7)	.2533(5)	.3862(4)	.052
C(13)	.2976(6)	.2482(4)	.2278(4)	.034
C(14)	.4284(7)	.2607(5)	.2085(5)	.053
C(15)	.1968(6)	.3162(4)	.1951(4)	.038
C(16)	.2350(8)	.4081(5)	.2054(6)	.064
C(17)	.0728(6)	.2952(4)	.1683(4)	.033
C(18)	-.0416(7)	.3543(4)	.1450(5)	.046
C(19)	-.1513(8)	.4546(6)	.0466(5)	.084
C(20)	.0416(6)	.2028(4)	.1666(4)	.032
C(21)	.0022(6)	.1635(4)	.2411(4)	.035
C(22)	-.0944(9)	.1629(6)	.3634(5)	.078
H(1)	.06969	-.02823	.11941	.064
H(2)	.13580	.07989	.02026	.066
H(3)	.37123	.11875	.08109	.060
H(4)	.45773	.02476	.21164	.049
H(5)	.26957	-.06214	.24018	.071
H(6)	-.23462	.22302	.11826	.064
H(7A)	-.22439	.08368	.16548	.098
H(7B)	-.27902	.05481	.06910	.098
H(7C)	-.36865	.10586	.11798	.098
H(8A)	-.29932	.15728	-.05120	.099
H(8B)	-.28406	.25360	-.03103	.099
H(8C)	-.40540	.20373	-.01229	.099
H(19A)	-.23263	.43040	.05243	.085
H(19B)	-.16043	.47415	-.01226	.085
H(19C)	-.13125	.50240	.08470	.085
H(22A)	-.14153	.19963	.39388	.078
H(22B)	-.01571	.14456	.40297	.078
H(22C)	-.14639	.11383	.34430	.078

Table 3.10(cont).

Anisotropic displacement parameters (\AA^2).

	U11	U22	U33	U12	U13	U23
MO(1)	.0304(3)	.0302(3)	.0341(3)	.0037(4)	.0085(2)	.0001(3)
S(1)	.0309(9)	.0534(13)	.0379(8)	.0014(9)	.0045(7)	-.0066(9)
F(1)	.048(3)	.133(6)	.147(5)	.001(4)	-.027(3)	.052(5)
F(2)	.178(6)	.197(7)	.041(3)	.125(6)	-.012(3)	.007(4)
F(3)	.119(5)	.065(4)	.100(4)	.023(4)	-.024(3)	.032(3)
F(4)	.071(3)	.075(4)	.054(2)	.028(3)	.015(2)	-.019(2)
F(5)	.069(3)	.114(5)	.057(3)	-.005(3)	-.005(2)	-.039(3)
F(6)	.098(4)	.091(4)	.048(2)	.006(3)	.033(3)	.003(3)
F(7)	.049(3)	.076(4)	.124(4)	-.026(3)	.018(3)	-.027(3)
F(8)	.040(3)	.066(3)	.098(3)	.010(3)	.022(2)	.007(3)
F(9)	.067(3)	.117(5)	.065(3)	-.009(3)	.030(3)	.024(3)
F(10)	.077(4)	.045(3)	.129(4)	-.007(3)	-.015(3)	-.022(3)
F(11)	.069(4)	.032(3)	.180(6)	.007(3)	-.020(4)	-.018(3)
F(12)	.119(5)	.077(5)	.145(5)	-.035(4)	.035(4)	.038(4)
O(1)	.083(4)	.061(4)	.047(3)	.038(3)	.006(3)	.009(3)
O(2)	.079(4)	.066(5)	.117(5)	.038(4)	.058(4)	.029(4)
O(3)	.058(3)	.049(3)	.048(3)	.004(3)	.031(2)	.002(2)
O(4)	.037(3)	.038(3)	.041(2)	.000(2)	.011(2)	.005(2)
C(1)	.045(5)	.051(5)	.085(6)	-.009(4)	.022(4)	-.040(5)
C(2)	.057(5)	.068(6)	.042(4)	.021(5)	.012(3)	-.025(4)
C(3)	.057(5)	.048(5)	.046(4)	.008(4)	.031(4)	-.007(4)
C(4)	.042(4)	.049(5)	.046(4)	.012(4)	.010(3)	-.005(4)
C(5)	.072(6)	.031(4)	.068(5)	.016(4)	.027(4)	-.003(4)
C(6)	.039(4)	.058(6)	.059(4)	.004(4)	.007(3)	-.015(4)
C(7)	.063(6)	.096(8)	.095(7)	-.016(6)	.018(5)	.004(6)
C(8)	.038(5)	.155(11)	.119(8)	.017(6)	-.004(5)	.051(8)
C(9)	.052(5)	.058(6)	.055(4)	.016(5)	-.005(4)	.013(4)
C(10)	.036(4)	.043(5)	.045(3)	.004(4)	.011(3)	.003(3)
C(11)	.033(4)	.038(4)	.033(3)	.004(3)	.003(3)	-.002(3)
C(12)	.049(5)	.061(6)	.046(4)	-.001(4)	.013(4)	-.006(4)
C(13)	.027(4)	.036(4)	.039(3)	.000(3)	.010(3)	-.002(3)
C(14)	.046(5)	.055(5)	.056(4)	-.005(4)	.008(4)	.000(4)
C(15)	.039(4)	.032(4)	.042(4)	-.001(4)	.012(3)	0.000(3)
C(16)	.041(5)	.043(5)	.098(7)	.001(4)	-.002(5)	.003(5)
C(17)	.039(4)	.033(4)	.027(3)	.002(3)	.013(3)	-.001(3)
C(18)	.043(4)	.032(4)	.059(4)	.002(4)	.009(4)	-.005(4)
C(19)	.087(7)	.072(6)	.080(6)	.050(6)	-.020(5)	-.004(5)
C(20)	.026(3)	.031(4)	.037(3)	.005(3)	.008(3)	-.002(3)
C(21)	.022(3)	.042(4)	.039(3)	-.004(4)	.008(3)	-.002(3)
C(22)	.110(7)	.067(6)	.062(5)	-.005(6)	.060(5)	-.005(5)

Table 3.11. Selected bond distances (Å) and angles (°) of complex (12).

MO(1) - O(4)	2.125(5)	MO(1) - C(1)	2.336(9)
MO(1) - C(2)	2.367(7)	MO(1) - C(3)	2.339(7)
MO(1) - C(4)	2.328(8)	MO(1) - C(5)	2.314(8)
MO(1) - C(10)	1.918(7)	MO(1) - C(11)	2.195(7)
MO(1) - C(13)	2.266(7)	MO(1) - C(20)	2.249(7)
MO(1) - C(21)	2.477(6)	S(1) - C(6)	1.839(8)
S(1) - C(20)	1.819(7)	F(1) - C(9)	1.310(10)
F(2) - C(9)	1.273(9)	F(3) - C(9)	1.324(11)
F(4) - C(12)	1.334(10)	F(5) - C(12)	1.345(9)
F(6) - C(12)	1.339(10)	F(7) - C(14)	1.342(10)
F(8) - C(14)	1.331(10)	F(9) - C(14)	1.341(9)
F(10) - C(16)	1.342(11)	F(11) - C(16)	1.313(10)
F(12) - C(16)	1.328(11)	O(1) - C(18)	1.332(9)
O(1) - C(19)	1.450(11)	O(2) - C(18)	1.178(10)
O(3) - C(21)	1.296(8)	O(3) - C(22)	1.470(10)
O(4) - C(21)	1.248(9)	C(1) - C(2)	1.398(12)
C(1) - C(5)	1.416(12)	C(2) - C(3)	1.383(11)
C(3) - C(4)	1.408(10)	C(4) - C(5)	1.401(11)
C(6) - C(7)	1.511(13)	C(6) - C(8)	1.513(13)
C(9) - C(10)	1.487(11)	C(10) - C(11)	1.410(10)
C(11) - C(12)	1.497(10)	C(11) - C(13)	1.471(9)
C(13) - C(14)	1.486(10)	C(13) - C(15)	1.517(10)
C(15) - C(16)	1.513(12)	C(15) - C(17)	1.316(9)
C(17) - C(18)	1.502(10)	C(17) - C(20)	1.503(10)
C(20) - C(21)	1.468(9)		

O(4) - MO(1) - C(10)	101.8(3)	O(4) - MO(1) - C(11)	96.4(2)
O(4) - MO(1) - C(13)	119.3(2)	O(4) - MO(1) - C(20)	63.0(2)
O(4) - MO(1) - C(21)	30.2(2)	C(10) - MO(1) - C(11)	39.4(3)
C(10) - MO(1) - C(13)	69.0(3)	C(10) - MO(1) - C(20)	125.2(3)
C(10) - MO(1) - C(21)	106.7(3)	C(11) - MO(1) - C(13)	38.5(3)
C(11) - MO(1) - C(20)	87.7(3)	C(11) - MO(1) - C(21)	82.6(3)
C(13) - MO(1) - C(20)	74.6(3)	C(13) - MO(1) - C(21)	92.5(3)
C(20) - MO(1) - C(21)	35.8(3)	C(6) - S(1) - C(20)	106.2(4)
C(18) - O(1) - C(19)	114.1(6)	C(21) - O(3) - C(22)	117.9(6)
MO(1) - O(4) - C(21)	90.7(4)	C(2) - C(1) - C(5)	108.0(7)
C(1) - C(2) - C(3)	108.8(7)	C(2) - C(3) - C(4)	107.6(7)
C(3) - C(4) - C(5)	108.7(7)	C(1) - C(5) - C(4)	106.8(7)
S(1) - C(6) - C(7)	111.0(6)	S(1) - C(6) - C(8)	106.5(6)
C(7) - C(6) - C(8)	110.9(7)	F(1) - C(9) - F(2)	108.0(7)
F(1) - C(9) - F(3)	102.7(7)	F(1) - C(9) - C(10)	112.7(7)
F(2) - C(9) - F(3)	106.6(7)	F(2) - C(9) - C(10)	113.6(8)
F(3) - C(9) - C(10)	112.5(7)	MO(1) - C(10) - C(9)	147.7(6)
MO(1) - C(10) - C(11)	81.0(4)	C(9) - C(10) - C(11)	131.3(7)
MO(1) - C(11) - C(10)	59.7(4)	MO(1) - C(11) - C(12)	137.5(5)
MO(1) - C(11) - C(13)	73.4(4)	C(10) - C(11) - C(12)	121.8(6)
C(10) - C(11) - C(13)	111.8(6)	C(12) - C(11) - C(13)	126.3(7)
F(4) - C(12) - F(5)	106.5(7)	F(4) - C(12) - F(6)	106.1(6)
F(4) - C(12) - C(11)	116.2(6)	F(5) - C(12) - F(6)	105.6(6)
F(5) - C(12) - C(11)	110.4(6)	F(6) - C(12) - C(11)	111.4(7)
MO(1) - C(13) - C(11)	68.2(4)	MO(1) - C(13) - C(14)	117.3(5)
MO(1) - C(13) - C(15)	113.3(5)	C(11) - C(13) - C(14)	117.5(6)
C(11) - C(13) - C(15)	115.6(6)	C(14) - C(13) - C(15)	116.2(6)
F(7) - C(14) - F(8)	103.4(6)	F(7) - C(14) - F(9)	105.2(7)
F(7) - C(14) - C(13)	114.6(7)	F(8) - C(14) - F(9)	105.9(7)
F(8) - C(14) - C(13)	115.7(7)	F(9) - C(14) - C(13)	111.2(6)
C(13) - C(15) - C(16)	120.2(6)	C(13) - C(15) - C(17)	119.4(7)
C(16) - C(15) - C(17)	119.8(7)	F(10) - C(16) - F(11)	104.5(8)
F(10) - C(16) - F(12)	105.7(7)	F(10) - C(16) - C(15)	111.5(7)
F(11) - C(16) - F(12)	107.7(8)	F(11) - C(16) - C(15)	114.0(7)
F(12) - C(16) - C(15)	112.8(8)	C(15) - C(17) - C(18)	126.7(7)
C(15) - C(17) - C(20)	116.6(6)	C(18) - C(17) - C(20)	116.5(6)
O(1) - C(18) - O(2)	123.4(7)	O(1) - C(18) - C(17)	112.0(6)
O(2) - C(18) - C(17)	124.5(7)	MO(1) - C(20) - S(1)	111.6(3)
MO(1) - C(20) - C(17)	115.8(4)	MO(1) - C(20) - C(21)	80.6(4)
S(1) - C(20) - C(17)	114.1(5)	S(1) - C(20) - C(21)	110.3(5)
C(17) - C(20) - C(21)	120.1(6)	MO(1) - C(21) - O(3)	150.8(5)
MO(1) - C(21) - O(4)	59.1(4)	MO(1) - C(21) - C(20)	63.6(4)
O(3) - C(21) - O(4)	124.3(6)	O(3) - C(21) - C(20)	121.0(6)
O(4) - C(21) - C(20)	114.7(6)		

Table 3.12. Fractional atomic coordinates and displacement parameters (\AA^2) of the crystal structure of complex (15).

	X/A	Y/B	Z/C	U
W	.13805(2)	.27768(1)	-.33947(1)	.031
S	.36550(15)	.24149(8)	-.26773(9)	.049
F(1)	.1313(4)	.2727(2)	-.1329(2)	.076
F(2)	-.0419(4)	.3074(2)	-.2036(2)	.075
F(3)	-.0432(4)	.2005(2)	-.1504(2)	.078
F(4)	.0286(4)	.0425(2)	-.1711(2)	.086
F(5)	.1675(4)	.1235(2)	-.1186(2)	.086
F(6)	.2365(5)	.0338(2)	-.1795(2)	.089
F(7)	.0885(4)	-.0215(2)	-.2882(2)	.086
F(8)	.1215(5)	-.0038(2)	-.3964(2)	.094
F(9)	.2889(4)	-.0094(2)	-.3102(3)	.105
F(10)	.3027(4)	.0784(2)	-.4355(2)	.073
F(11)	.1055(4)	.1091(3)	-.4869(2)	.082
F(12)	.2624(5)	.1914(2)	-.4738(2)	.086
N(1)	.2115(5)	.4240(3)	-.2342(3)	.048
N(2)	.3629(5)	.3465(3)	-.4274(3)	.054
C(1)	-.0478(6)	.2404(3)	-.4241(3)	.051
C(2)	-.0965(5)	.2692(3)	-.3629(3)	.048
C(3)	-.0580(5)	.3464(3)	-.3547(3)	.047
C(4)	.0166(6)	.3651(3)	-.4106(3)	.054
C(5)	.0210(6)	.2993(4)	-.4533(3)	.052
C(6)	.4796(5)	.3224(3)	-.2427(3)	.044
C(7)	.4935(7)	.3382(4)	-.1613(3)	.066
C(8)	.6121(6)	.3028(4)	-.2648(4)	.066
C(9)	.0868(5)	.2094(3)	-.2476(3)	.039
C(10)	.0356(6)	.2460(4)	-.1841(3)	.053
C(11)	.1189(5)	.1340(3)	-.2473(3)	.040
C(12)	.1347(7)	.0837(4)	-.1799(4)	.062
C(13)	.1589(5)	.1038(3)	-.3147(3)	.042
C(14)	.1666(7)	.0174(3)	-.3283(4)	.066
C(15)	.1747(5)	.1566(3)	-.3665(3)	.039
C(16)	.2116(7)	.1332(3)	-.4387(3)	.056
C(17)	.1851(5)	.3712(3)	-.2705(3)	.040
C(18)	.2839(5)	.3193(3)	-.3972(3)	.042
C(21)	.2522(6)	.4853(3)	-.1868(3)	.043
C(22)	.3424(7)	.5370(4)	-.2041(3)	.057
C(23)	.3869(7)	.5964(4)	-.1574(4)	.062
C(24)	.3415(7)	.6023(3)	-.0918(4)	.057
C(25)	.2491(7)	.5498(4)	-.0744(3)	.063
C(26)	.2052(7)	.4906(3)	-.1215(4)	.058
C(27)	.3884(9)	.6648(5)	-.0385(4)	.098
C(31)	.4624(5)	.3695(3)	-.4667(3)	.045
C(32)	.4590(6)	.3443(4)	-.5372(3)	.055
C(33)	.5598(7)	.3677(4)	-.5751(3)	.061
C(34)	.6605(6)	.4152(3)	-.5444(3)	.052
C(35)	.6618(6)	.4400(4)	-.4728(3)	.054
C(36)	.5604(6)	.4173(4)	-.4349(3)	.055
C(37)	.7700(8)	.4408(4)	-.5831(4)	.082
H(1)	-.05998	.18918	-.44297	.050
H(2)	-.14629	.24105	-.33176	.050
H(3)	-.07940	.38099	-.31794	.050
H(4)	.05741	.41369	-.41773	.050
H(5)	.06388	.29536	-.49568	.050
H(6)	.44460	.36734	-.26889	.050
H(22)	.37610	.53110	-.24923	.050
H(23)	.44870	.63336	-.17097	.050
H(25)	.21625	.55379	-.02878	.050
H(26)	.14129	.45388	-.10981	.050
H(32)	.38057	.31163	-.55990	.050
H(33)	.55931	.35043	-.62425	.050
H(35)	.73317	.47212	-.45003	.050
H(36)	.55858	.43605	-.38648	.050
H(7A)	.55247	.38076	-.14868	.050
H(7B)	.52850	.29325	-.13512	.050
H(7C)	.40748	.34959	-.14876	.050
H(8A)	.67106	.34537	-.25212	.050
H(8B)	.60205	.29325	-.31624	.050
H(8C)	.64722	.25753	-.23917	.050
H(27A)	.35683	.66912	.00745	.050
H(27B)	.36746	.71252	-.06350	.050
H(27C)	.48348	.65957	-.02747	.050
H(37A)	.84054	.47414	-.56155	.050
H(37B)	.72627	.46605	-.62664	.050
H(37C)	.81218	.39467	-.59617	.050

Table 3.12(cont).

Anisotropic displacement parameters (\AA^2).

	U11	U22	U33	U12	U13	U23
H	.0313(1)	.0297(1)	.0326(1)	-.0009(1)	.0050(1)	-.0002(1)
S	.0373(7)	.0383(7)	.0671(10)	-.0014(6)	-.0045(7)	.0022(7)
F(1)	.092(3)	.089(3)	.046(2)	-.016(2)	.010(2)	-.021(2)
F(2)	.091(3)	.066(2)	.070(2)	.025(2)	.037(2)	-.002(2)
F(3)	.087(3)	.079(3)	.071(2)	-.012(2)	.046(2)	.005(2)
F(4)	.101(3)	.070(3)	.088(3)	-.016(2)	.034(3)	.030(2)
F(5)	.107(3)	.095(3)	.051(2)	.005(3)	-.006(2)	.016(2)
F(6)	.094(3)	.083(3)	.087(3)	.029(3)	-.001(3)	.035(2)
F(7)	.110(4)	.044(2)	.103(3)	-.016(2)	.031(3)	.009(2)
F(8)	.128(4)	.053(2)	.103(3)	-.020(2)	.044(3)	-.030(2)
F(9)	.093(3)	.065(3)	.159(5)	.038(2)	.041(3)	.015(3)
F(10)	.067(2)	.074(2)	.079(3)	.012(2)	.031(2)	-.017(2)
F(11)	.073(3)	.110(3)	.060(2)	.003(2)	.003(2)	-.038(2)
F(12)	.136(4)	.068(2)	.059(2)	-.014(3)	.053(2)	-.014(2)
N(1)	.048(3)	.038(2)	.057(3)	.000(2)	.003(3)	-.006(2)
N(2)	.045(3)	.058(3)	.061(3)	-.004(3)	.016(3)	.010(3)
C(1)	.044(3)	.051(3)	.054(3)	.001(3)	-.007(3)	-.009(3)
C(2)	.031(3)	.061(4)	.049(3)	-.005(3)	.002(2)	.001(3)
C(3)	.036(3)	.049(3)	.052(3)	.010(3)	-.002(3)	-.008(3)
C(4)	.049(4)	.049(3)	.060(4)	.003(3)	-.006(3)	.020(3)
C(5)	.048(3)	.070(4)	.035(3)	.000(3)	-.003(3)	.008(3)
C(6)	.036(3)	.044(3)	.050(3)	-.007(2)	-.002(3)	.001(2)
C(7)	.081(5)	.056(4)	.054(4)	-.007(4)	-.010(4)	-.004(3)
C(8)	.042(3)	.067(4)	.086(5)	-.004(3)	.006(4)	.004(4)
C(9)	.034(3)	.049(3)	.033(2)	-.003(2)	.003(2)	.002(2)
C(10)	.060(4)	.058(3)	.043(3)	-.007(3)	.018(3)	-.005(3)
C(11)	.032(3)	.040(3)	.048(3)	-.004(2)	.005(3)	.008(2)
C(12)	.067(4)	.059(4)	.058(4)	.009(3)	.007(3)	.018(3)
C(13)	.038(3)	.039(3)	.049(3)	.001(2)	.002(3)	.008(2)
C(14)	.075(5)	.034(3)	.089(5)	.004(3)	.025(4)	0.000(3)
C(15)	.032(3)	.042(3)	.043(3)	-.003(2)	.009(2)	.003(2)
C(16)	.063(4)	.051(3)	.055(4)	.002(3)	.015(3)	-.014(3)
C(17)	.033(3)	.034(2)	.053(3)	-.003(2)	.013(2)	-.004(2)
C(18)	.041(3)	.040(3)	.046(3)	-.005(2)	.011(3)	0.000(2)
C(21)	.047(3)	.034(2)	.046(3)	-.001(2)	.002(3)	-.004(2)
C(22)	.068(4)	.053(4)	.049(3)	-.011(3)	.014(3)	-.006(3)
C(23)	.066(5)	.053(4)	.063(4)	-.017(3)	.008(4)	.000(3)
C(24)	.063(4)	.045(3)	.056(4)	-.001(3)	-.017(3)	-.008(3)
C(25)	.081(5)	.064(4)	.043(3)	.001(4)	.011(3)	-.007(3)
C(26)	.069(4)	.044(3)	.062(4)	-.009(3)	.017(4)	-.003(3)
C(27)	.099(7)	.095(6)	.089(6)	.004(5)	-.034(5)	-.046(5)
C(31)	.040(3)	.045(3)	.050(3)	0.000(3)	.015(3)	.006(3)
C(32)	.054(4)	.049(3)	.059(4)	-.012(3)	.004(3)	-.007(3)
C(33)	.079(5)	.057(4)	.047(3)	-.002(3)	.027(3)	-.004(3)
C(34)	.050(4)	.048(3)	.057(4)	.002(3)	.019(3)	.012(3)
C(35)	.046(3)	.062(4)	.052(3)	-.015(3)	.003(3)	.002(3)
C(36)	.056(4)	.066(4)	.042(3)	-.005(3)	.014(3)	-.002(3)
C(37)	.075(5)	.083(5)	.092(5)	-.008(4)	.045(5)	.010(4)

Table 3.13. Selected bond distances (Å) and angles (°) of complex (15).

W - S	2.561(2)	W - C(1)	2.353(7)
W - C(2)	2.360(6)	W - C(3)	2.304(6)
W - C(4)	2.255(7)	W - C(5)	2.290(6)
W - C(9)	2.207(6)	W - C(15)	2.213(6)
W - C(17)	2.081(6)	W - C(18)	2.090(6)
S - C(6)	1.839(6)	F(1) - C(10)	1.334(8)
F(2) - C(10)	1.344(8)	F(3) - C(10)	1.348(8)
F(4) - C(12)	1.326(9)	F(5) - C(12)	1.331(8)
F(6) - C(12)	1.350(9)	F(7) - C(14)	1.351(9)
F(8) - C(14)	1.329(9)	F(9) - C(14)	1.322(9)
F(10) - C(16)	1.325(8)	F(11) - C(16)	1.356(8)
F(12) - C(16)	1.350(8)	N(1) - C(17)	1.147(7)
N(1) - C(21)	1.404(8)	N(2) - C(18)	1.150(8)
N(2) - C(31)	1.394(8)	C(1) - C(2)	1.401(9)
C(1) - C(5)	1.397(9)	C(2) - C(3)	1.402(9)
C(3) - C(4)	1.417(9)	C(4) - C(5)	1.398(9)
C(6) - C(7)	1.521(9)	C(6) - C(8)	1.506(9)
C(9) - C(10)	1.503(8)	C(9) - C(11)	1.354(8)
C(11) - C(12)	1.516(9)	C(11) - C(13)	1.472(8)
C(13) - C(14)	1.530(8)	C(13) - C(15)	1.359(8)
C(15) - C(16)	1.504(8)	C(21) - C(22)	1.359(9)
C(21) - C(26)	1.374(9)	C(22) - C(23)	1.380(10)
C(23) - C(24)	1.372(10)	C(24) - C(25)	1.383(10)
C(24) - C(27)	1.499(11)	C(25) - C(26)	1.383(9)
C(31) - C(32)	1.376(9)	C(31) - C(36)	1.361(9)
C(32) - C(33)	1.391(10)	C(33) - C(34)	1.369(10)
C(34) - C(35)	1.396(9)	C(34) - C(37)	1.484(10)
C(35) - C(36)	1.393(9)		

S - W - C(CP)	176.1	S - W - C(9)	76.6(2)
S - W - C(15)	73.4(2)	S - W - C(17)	76.7(2)
S - W - C(18)	72.4(2)	C(CP) - W - C(9)	106.3
C(CP) - W - C(15)	104.7	C(CP) - W - C(17)	105.7
C(CP) - W - C(18)	104.5	C(9) - W - C(15)	74.4(2)
C(9) - W - C(17)	90.3(2)	C(9) - W - C(18)	148.8(2)
C(15) - W - C(17)	148.7(2)	C(15) - W - C(18)	93.2(2)
C(17) - W - C(18)	86.0(3)	W - S - C(6)	115.2(2)
C(17) - N(1) - C(21)	175.5(6)	C(18) - N(2) - C(31)	172.3(6)
S - C(6) - C(7)	110.3(5)	S - C(6) - C(8)	108.2(5)
C(7) - C(6) - C(8)	111.3(6)	W - C(9) - C(10)	122.0(4)
W - C(9) - C(11)	116.1(4)	C(10) - C(9) - C(11)	121.4(5)
F(1) - C(10) - F(3)	104.3(5)	F(1) - C(10) - F(3)	107.0(5)
F(1) - C(10) - C(9)	113.9(6)	F(2) - C(10) - F(3)	103.2(5)
F(2) - C(10) - C(9)	112.7(5)	F(3) - C(10) - C(9)	114.8(6)
C(9) - C(11) - C(12)	124.0(6)	C(9) - C(11) - C(13)	116.2(5)
C(12) - C(11) - C(13)	119.2(5)	F(4) - C(12) - F(5)	106.1(6)
F(4) - C(12) - F(6)	106.7(6)	F(4) - C(12) - C(11)	115.5(6)
F(5) - C(12) - F(6)	104.0(6)	F(5) - C(12) - C(11)	112.7(6)
F(6) - C(12) - C(11)	111.0(6)	C(11) - C(13) - C(14)	121.3(5)
C(11) - C(13) - C(15)	116.1(5)	C(14) - C(13) - C(15)	122.3(6)
F(7) - C(14) - F(8)	103.9(6)	F(7) - C(14) - F(9)	107.0(6)
F(7) - C(14) - C(13)	110.6(6)	F(8) - C(14) - F(9)	108.3(6)
F(8) - C(14) - C(13)	114.3(6)	F(9) - C(14) - C(13)	112.1(6)
W - C(15) - C(13)	116.2(4)	W - C(15) - C(16)	122.3(4)
C(13) - C(15) - C(16)	121.5(5)	F(10) - C(16) - F(11)	106.6(5)
F(10) - C(16) - F(12)	104.2(6)	F(10) - C(16) - C(15)	115.2(5)
F(11) - C(16) - F(12)	103.7(5)	F(11) - C(16) - C(15)	112.9(6)
F(12) - C(16) - C(15)	113.2(5)	W - C(17) - N(1)	178.1(5)
W - C(18) - N(2)	175.9(5)	N(1) - C(21) - C(22)	120.0(6)
N(1) - C(21) - C(26)	119.7(6)	C(22) - C(21) - C(26)	120.2(6)
C(21) - C(22) - C(23)	121.1(6)	C(22) - C(23) - C(24)	119.4(7)
C(23) - C(24) - C(25)	119.5(6)	C(23) - C(24) - C(27)	121.6(7)
C(25) - C(24) - C(27)	118.9(7)	C(24) - C(25) - C(26)	120.6(6)
C(21) - C(26) - C(25)	119.2(6)	N(2) - C(31) - C(32)	119.6(6)
N(2) - C(31) - C(36)	119.3(6)	C(32) - C(31) - C(36)	121.1(6)
C(31) - C(32) - C(33)	118.5(6)	C(32) - C(33) - C(34)	121.7(6)
C(33) - C(34) - C(35)	118.8(6)	C(33) - C(34) - C(37)	123.2(6)
C(35) - C(34) - C(37)	118.1(6)	C(34) - C(35) - C(36)	119.6(6)
C(31) - C(36) - C(35)	120.2(6)		

Table 3.14. Crystallographic and experimental details of the structure analyses of complexes (5), (7), (10), (11), (12) and (15).

Compound	[CpMo(SPr ⁱ)(hfb) ₂ (PhCCMe)] (5)	[CpW(SPr ⁱ)(hfb) ₂ (PhCCMe)] (7)
Formula	C ₂₅ H ₂₀ F ₁₂ MoS	C ₂₅ H ₂₀ F ₁₂ SW
Formula Wt.	676.4	764.3
Crystal habit	yellow needle	dark red needle
Crystal size, mm	0.64x0.20x0.20	0.64x0.20x0.12
Crystal system	monoclinic	monoclinic
Space group	P 2 ₁ /c	P 2 ₁ /c
a, Å	8.509(4)	8.086(1)
b, Å	14.120(1)	19.896(2)
c, Å	22.095(3)	15.614(2)
β, °	97.46(2)	90.33(1)
Obtained from	25 refln, 12<θ<18°	25 refln, 13<θ<18°
V, Å ³	2632(1)	2511.9(6)
Z	4	4
F(000)	1344	1472
d calc, gcm ⁻³	1.707	2.021
T, K	297	295
μ(Mo-Kα), cm ⁻¹	6.6	48.8
Absorption factors on F ²	0.88-0.95	0.87-1.19
Scan width, °	0.75	0.70
Max count time, s	95	100
Total refln measured	8740	9426
Unique refln	6812	7309
R _{INT}	0.058	0.026
Miller indices	h	h
measured,	k	k
	l	l
2θ range, °	4-60	4-60
Unique refln ≥ 3σ(I)	3740	n/a
No. of parameters	352	432
R	0.039	0.054
R _w	0.050	0.031
Δρ _{max} , eÅ ⁻³	0.75	0.89
Δ/σ _{max}	0.26	0.17

Table 3.14(cont).

Compound	[CpW(STol)(hfb) ₂ (dmad)] (10)	[CpMo(SPr ⁱ)(hfb) ₂ (dmad)] (11)
Formula	C ₂₅ H ₁₈ F ₁₂ O ₄ SW	C ₂₂ H ₁₈ F ₁₂ MoO ₄ S
Formula Wt.	838.3	702.4
Crystal habit	black needle	dark red plate
Crystal size, mm	0.48x0.08x0.08	0.52x0.28x0.44
Crystal system	orthorhombic	triclinic
Space group	P2 ₁ 2 ₁ 2 ₁	P $\bar{1}$
a, Å	9.234(3)	10.898(3)
b, Å	15.185(3)	16.120(4)
c, Å	20.420(3)	17.482(4)
α , °	—	63.80(2)
β , °	—	73.11(2)
γ , °	—	71.44(2)
Obtained from	23 refln, 10< θ <15°	23 refln, 12< θ <14°
V, Å ³	2863(1)	2572(1)
Z	4	4
F(000)	1616	1392
d calc, gcm ⁻³	1.945	1.814
T, K	295	297
μ (Mo-K α), cm ⁻¹	43.0	6.9
Absorption factors on F ²	0.79-1.22	0.78-1.17
Scan width, °	0.80	1.00
Max count time, s	180	100
Total refln measured	3895	7171
Unique refln	3868	6268
R _{INT}	—	0.029
Miller indices		
measured, h	0-12	-11-11
k	0-20	-17-17
l	0-26	-18- 2
2 θ range, °	4-56	4-44
Unique refln \geq 3 σ (I)	1676	4602
No. of parameters	188	721
R	0.041	0.044
R _w	0.047	0.060
$ \Delta\rho _{\max}$, eÅ ⁻³	0.72	0.64
Δ/σ_{\max}	0.082	0.09

Table 3.14(cont).

Compound	[CpMo(SPr ⁱ)(hfb) ₂ (dmad)] (12)	[CpW(SPr ⁱ)(hfb) ₂ (CNTol) ₂] (15)
Formula	C ₂₂ H ₁₈ F ₁₂ MoO ₄ S	C ₃₂ H ₂₆ F ₁₂ N ₂ SW
Formula Wt.	702.4	882.5
Crystal habit	red needle	orange plate
Crystal size, mm	0.56x0.24x0.20	0.60x0.25x0.70
Crystal system	monoclinic	monoclinic
Space group	P 2 ₁ /c	P 2 ₁ /c
a, Å	10.496(4)	10.163(1)
b, Å	15.881(3)	17.418(4)
c, Å	15.733(6)	18.577(2)
β, °	103.13(3)	98.934(7)
Obtained from	23 refln, 12<θ<15°	23 refln, 12<θ<14°
V, Å ³	2554(2)	3248.6(7)
Z	4	4
F(000)	1392	1720
d calc, gcm ⁻³	1.827	1.804
T, K	297	295
μ(Mo-Kα), cm ⁻¹	6.9	37.9
Absorption factors on F ²	0.63-1.20	0.65-1.54
Scan width, °	0.80	0.80
Max count time, s	100	100
Total refln measured	6618	12075
Unique refln	4470	9490
R _{INT}	0.048	0.059
Miller indices	h	0-12
measured,	k	-18- 7
	l	-18-18
2θ range, °	4-50	4-60
Unique refln ≥ 3σ(I)	2881	5735
No. of parameters	361	433
R	0.046	0.037
R _w	0.054	0.044
Δρ _{max} , eÅ ⁻³	1.12	2.40
Δ/c _{max}	0.051	0.154

CHAPTER 4: OTHER ORGANOMETALLIC COMPLEXES OF Mo AND W

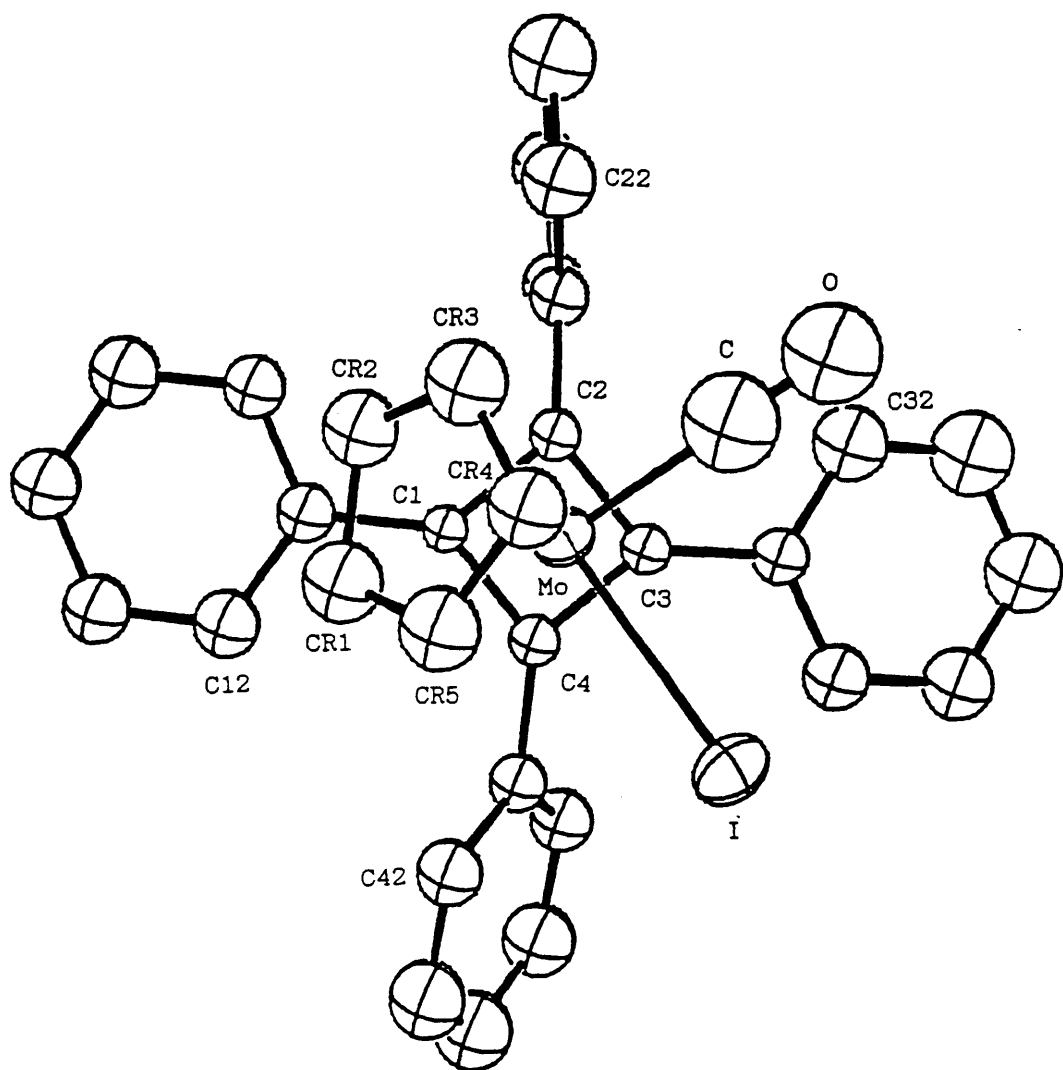
4.1 The Structure of $[\text{CpMoI}(\text{CO})(\text{C}_4\text{Ph}_4)]\cdot\text{CH}_2\text{Cl}_2$ Black Isomer

The complex $[\text{CpMoI}(\text{CO})(\text{C}_4\text{Ph}_4)]\cdot\text{CH}_2\text{Cl}_2$, (19) exists in two isomeric forms. Recrystallisation from hexane gives yellow crystals. On dissolving the yellow crystals in dichloromethane a yellow solution is obtained. This slowly turns black and deposits isomeric black crystals. To determine the nature of the isomerism it was decided to carry out crystal structure analyses of the yellow and black forms of (19). Unfortunately, Dr.J.L. Davidson was able to obtain diffraction quality crystals only of the black isomer whose structure is described below. Work on obtaining crystals of the yellow isomer is continuing at Heriot-Watt.

The black crystals of complex (19) contain well separated molecules of the type shown in Figure 4(a) and of dichloromethane, the solvent of crystallisation, in equal numbers (see Table 4.2 for bond distances and angles). The structure is of the form (18) shown in Scheme III in Chapter 3. The molybdenum atom has bonded to it in a tetrahedral arrangement a cyclopentadienyl ring, a cyclobutadienyl ring with phenyl substituents and disordered molecules of iodine and carbonyl.

In Figure 4(a) the site marked I is a superposition of $\alpha\text{I} + (1-\alpha)\text{CO}$ where $\alpha = 0.754(3)$. The second site I' contains $\alpha\text{CO} + (1-\alpha)\text{I}$. During the refinement the Mo-C-O and Mo-I vectors were constrained to coincide and the C-O distances was fixed at 1.15 \AA . The Mo-I distances are

Figure 4(a). The structure of the major isomer of complex (19). For clarity hydrogen atoms are omitted. 50% probability ellipsoids are displayed for all atoms shown.



2.762(2) and 2.739(7) Å.

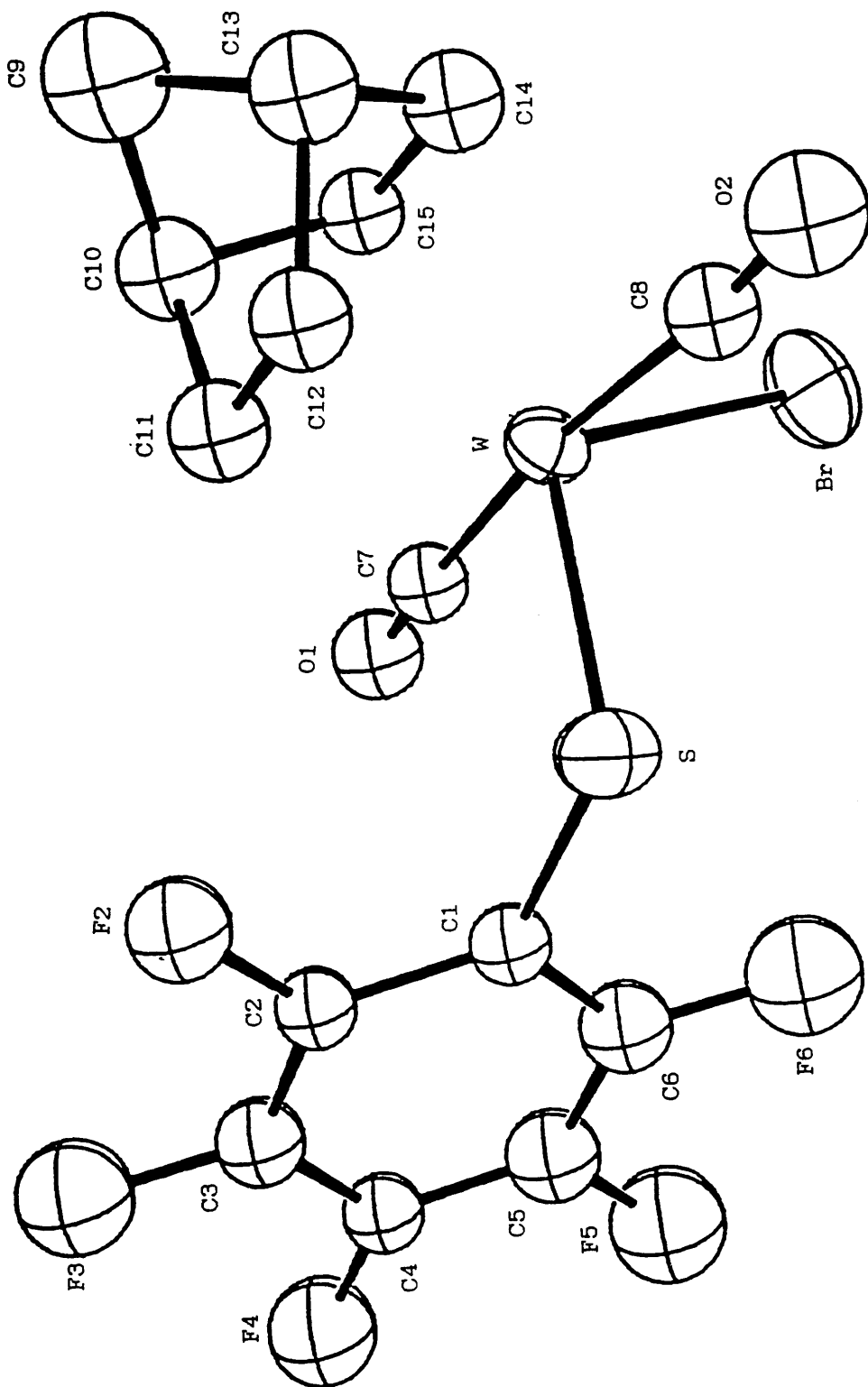
The cyclopentadienyl ring geometry is similar to those described in Chapter 3. Mo-C distances range from 2.307(17) to 2.339(18) Å and C-C distances from 1.36(3) to 1.42(3) Å. The Mo-C distances involving the cyclobutadienyl ligand are significantly shorter than those involving the five membered ring [2.231(14)-2.302(14) Å]. The ring C-C distances range from 1.450(20) to 1.478(20) Å and suggest that the π -bonding within the ring is delocalised over all four bonds. The phenyl substituents are staggered to reduce mutual steric interaction, resulting in two rings being roughly coplanar with the cyclobutadienyl ring [torsion angles C(2)-C(3)-C(31)-C(32) 4.5(15)° and C(4)-C(1)-C(11)-C(12) 11.6(15)°] and the other two clearly out of the plane [C(3)-C(4)-C(41)-C(42) 138.8(23)° and C(1)-C(2)-C(21)-C(22) 102.8(21)°].

It is difficult from the structure of the black form of (19) alone to suggest how the isomerism shown by (19) arises.

4.2 The Structure of $[\text{WBr}(\text{CO})_2(\text{SC}_6\text{F}_5)(\text{C}_7\text{H}_8)]$

The complex $[\text{WBr}(\text{CO})_2(\text{SC}_6\text{F}_5)(\text{C}_7\text{H}_8)]$, (20), shows unexpected fluxionality in that the six norbornadiene C-H groups are inequivalent at low temperature in the ^1H N.M.R. spectrum whereas at higher temperatures only three C-H resonances are observed. Restricted rotation about the W-S bond is one possible explanation for these observations. Alternatively, monomer-dimer equilibrium might explain them. A crystal structure analysis of (20) has accordingly

Figure 4(b). A view of complex (20). Hydrogen atoms are omitted and 50% ellipsoids are displayed for all other atoms.



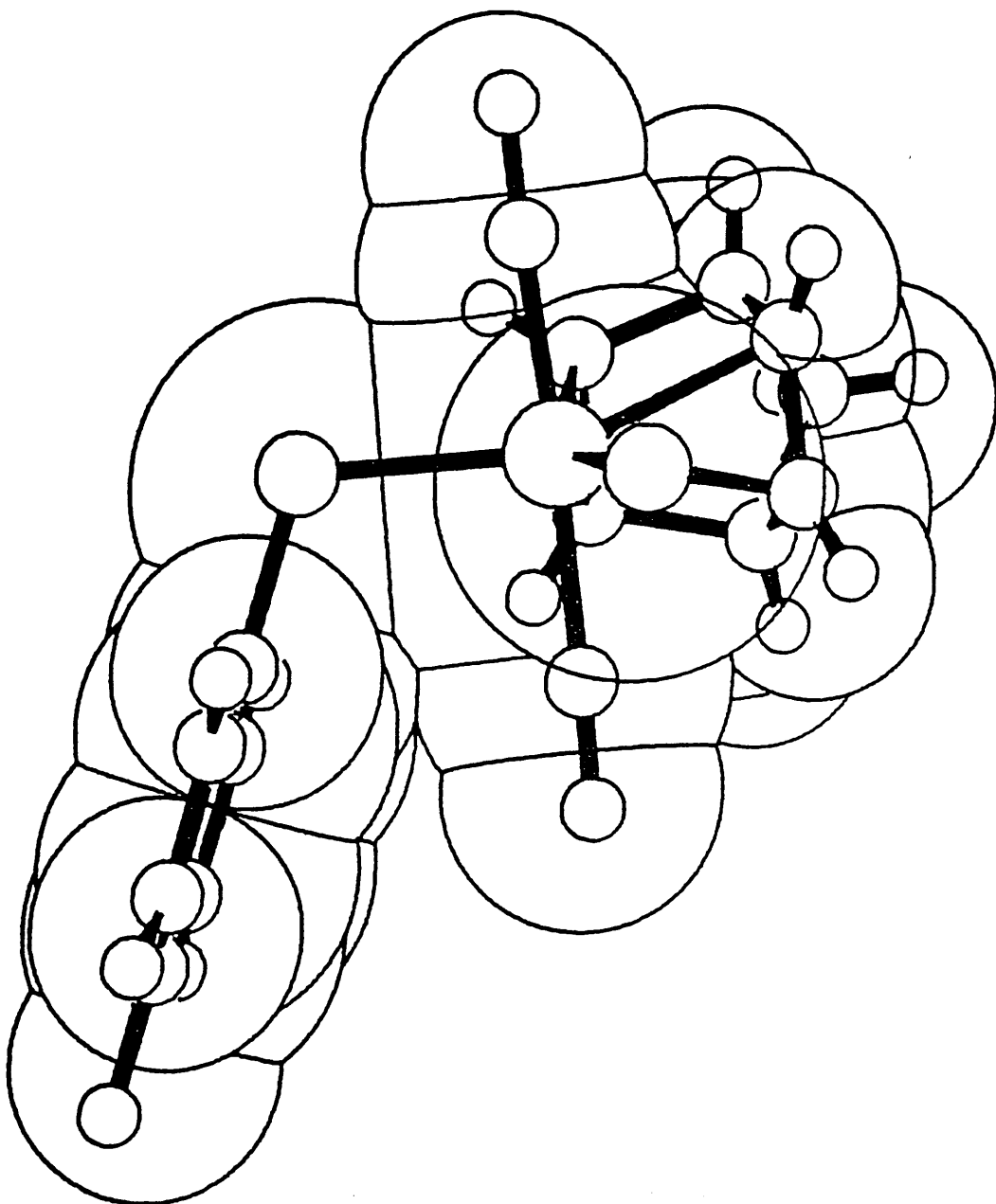
been carried out to see if the solid state structure at ambient temperature offers any explanation of the fluxionality.

Crystals of (20) are built of the well separated mononuclear molecules shown in Figure 4(b). The 16-electron divalent tungsten atom displays distorted octahedral coordination with mutually *trans* carbonyl ligands normal to a basal plane containing *cis* S and Br atoms respectively *trans* to the centroids of the C(14)-C(15) and C(11)-C(12) norbornadiene double bonds (see Table 4.4). The geometries of the penta-fluorophenylthiolato and norbornadiene ligands present no unusual features. Two features of the metal coordination merit comment: (i) the rather short W-S distance of 2.322(5) Å which lies at the lower end of the range of M(II)-S (M = Mo, W) distances observed in compounds discussed in this section of the thesis and (ii) the strong *trans*-influence exerted by the sulphur atom evident in the W-C(norbornadiene) distances of 2.38(2) and 2.40(2) Å *trans* to S compared with W-C distances of 2.23(3) and 2.28(2) Å *trans* to Br.

Inspection of interligand non-bonded distances confirms the conclusion, evident from Figure 4(c) which shows a space-filling view of the molecule, that steric effects leading to restricted rotation about the W-S bond do not provide a plausible explanation for the fluxionality.

However, the structural results are at least consistent with restricted rotation about the W-S bond arising from the electronic nature of the molecule. Only two of the t_{2g} orbitals can be filled and the strong π -acidity of carbonyl suggests that the vacant t_{2g} orbital will lie in the basal

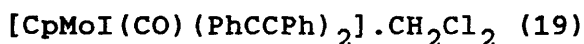
Figure 4(c). Space filled view of complex (20) using van der Waal radii ⁵² to calculate shells.



plane containing the S and Br atoms. S→W π -donation is thereby favoured and is consistent with the short W-S bond length. On the assumption that the sulphur hybridisation is closer to sp^2 than sp^3 the multiple S-W bonding could thereby lead to locking of the thiolato conformation at low temperature. The flow of charge from S to W is also consistent with the high thiolato *trans*-influence.

4.3 Experimental

A fuller description of the experimental procedure is given in Chapter 2. Only deviations from the method described there are noted below.



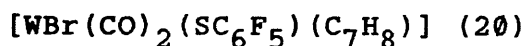
Data Collection. The cell dimensions and experimental parameters are listed in Table 4.5. A decomposition correction was applied which allowed for the 45% reduction in the mean intensity of the standard reflections.

Structure Analysis. The phenyl ring carbon atom positions were idealised [C-C distances 1.38 Å, C-C-C angles 120°] and refined as groups. All hydrogen atom positions were geometrically calculated. Phenyl hydrogens were included in the group refinement. Other hydrogen atoms were set to ride on their parent carbon atoms.

The disordered I and CO groups were treated thus: Mo-C distances were set at 1.97 Å and C-O distances at 1.15 Å along the Mo-I vectors. Atoms C' and O', and atoms C and O were set to ride on the I and I' atoms respectively. The

occupancy, α , of I was refined and those of C and O constrained to equal its value. Occupancies of C', O' and I' were constrained to $1-\alpha$.

Anisotropic displacement parameters were used for the Cl, Mo and I atoms. Atoms C, O, C' and O' had their displacement parameters set to and constrained to the I' isotropic displacement parameter. All other atoms were given isotropic displacement parameters and those of the hydrogen atoms were constrained to equal $1.2U(C)$. Final atomic positions and displacement parameters are listed in Table 4.1



Data Collection. The cell dimensions and experimental parameters are given in Table 4.5. No decomposition correction was applied. Two scale factors were used because it was necessary to realign the crystal after 1665 reflections.

The horizontal diffracted beam aperture was set at its minimum, 1.3mm, by mistake. The refinement, final individual ΔF values and the final difference synthesis do not suggest that major systematic errors in results have arisen in consequence.

Structure Analysis. Anisotropic displacement parameters were used for the W, Br, and S atoms. Isotropic parameters were used for all other atoms. The hydrogen atom positions were calculated geometrically and constrained to ride on their carbon atoms. $U(H)$ were constrained to $1.2U(C)$. Final atomic positions and displacement parameters are given in

Table 4.3.

Table 4.1. Fractional atomic coordinates and displacement parameters (\AA^2) of the crystal structure of complex (19).

	X/A	Y/B	Z/C	U
I	.08717(13)	-.29229(8)	.04775(5)	.047
I'	-.2040(5)	-.2555(4)	.0965(2)	.091(2)
MO	-.02139(11)	-.17309(7)	.07579(5)	.034
CL(1)	-.4710(7)	-.0081(4)	.0839(4)	.201
CL(2)	-.2767(9)	.0671(5)	.0628(6)	.315
O'	.10121	-.30772	.04411	.091
O	-.22790	-.26623	.09919	.091
C'	.05602	-.25807	.05583	.091
C	-.15117	-.23162	.09050	.091
C(1)	.0602(11)	-.0997(7)	.1317(5)	.023(4)
C(2)	-.0370(12)	-.1358(7)	.1529(5)	.027(4)
C(3)	.0311(11)	-.1994(7)	.1542(5)	.027(4)
C(4)	.1238(11)	-.1632(8)	.1298(5)	.028(4)
C(11)	.0894(13)	-.0269(4)	.1273(5)	.035(4)
C(12)	.1960(13)	-.0079(5)	.1108(6)	.048(5)
C(13)	.2233(7)	.0605(6)	.1063(4)	.049(5)
C(14)	.1441(11)	.1100(4)	.1183(4)	.050(5)
C(15)	.0375(11)	.0910(5)	.1349(6)	.053(5)
C(16)	.0101(7)	.0226(6)	.1394(4)	.041(5)
C(21)	-.1303(15)	-.1081(4)	.1848(7)	.037(4)
C(22)	-.2379(16)	-.0942(8)	.1669(3)	.057(5)
C(23)	-.3183(9)	-.0639(9)	.1969(6)	.080(7)
C(24)	-.2911(13)	-.0475(4)	.2448(6)	.060(5)
C(25)	-.1834(14)	-.0613(8)	.2627(3)	.063(6)
C(26)	-.1031(6)	-.0917(9)	.2327(7)	.047(5)
C(31)	.0143(9)	-.2644(6)	.1802(5)	.034(4)
C(32)	-.0831(12)	-.2761(7)	.2074(7)	.061(5)
C(33)	-.0953(9)	-.3364(5)	.2333(5)	.079(6)
C(34)	-.0101(9)	-.3851(5)	.2320(4)	.071(6)
C(35)	.0873(11)	-.3734(7)	.2048(6)	.057(5)
C(36)	.0995(9)	-.3131(5)	.1789(4)	.047(5)
C(41)	.2485(7)	-.1790(8)	.1248(6)	.035(4)
C(42)	.3032(9)	-.1682(10)	.0806(5)	.046(4)
C(43)	.4194(10)	-.1779(5)	.0770(3)	.064(5)
C(44)	.4810(7)	-.1984(7)	.1176(5)	.070(6)
C(45)	.4262(9)	-.2093(8)	.1618(5)	.064(5)
C(46)	.3100(11)	-.1996(4)	.1654(4)	.045(5)
C(R1)	.0308(15)	-.0894(9)	.0186(6)	.059(5)
C(R2)	-.0746(15)	-.0700(9)	.0401(6)	.052(5)
C(R3)	-.1535(15)	-.1176(10)	.0265(6)	.061(6)
C(R4)	-.1004(14)	-.1699(9)	-.0017(6)	.053(5)
C(R5)	.0116(15)	-.1488(9)	-.0078(6)	.060(5)
C(S)	-.391(3)	.066(2)	.092(1)	.159(11)
H(12)	.25111	-.04230	.10238	.057
H(13)	.29748	.07376	.09473	.058
H(14)	.16309	.15759	.11518	.060
H(15)	-.01770	.12540	.14323	.063
H(16)	-.06408	.00934	.15089	.049
H(22)	-.25689	-.10566	.13356	.068
H(23)	-.39317	-.05426	.18441	.096
H(24)	-.34697	-.02637	.26563	.072
H(25)	-.16450	-.04993	.29605	.075
H(26)	-.02824	-.10130	.24520	.056
H(32)	-.14238	-.24218	.20836	.073
H(33)	-.16301	-.34444	.25228	.095
H(34)	-.01854	-.42703	.25003	.085
H(35)	.14660	-.40731	.20392	.068
H(36)	.16725	-.30504	.16001	.056
H(42)	.26038	-.15384	.05232	.056
H(43)	.45745	-.17029	.04620	.077
H(44)	.56185	-.20512	.11511	.084
H(45)	.46904	-.22363	.19009	.077
H(46)	.27197	-.20708	.19620	.054
H(R1)	.10220	-.06567	.02122	.070
H(R2)	-.08718	-.03086	.06081	.062
H(R3)	-.23305	-.11624	.03476	.073
H(R4)	-.13534	-.21073	-.01434	.063
H(R5)	.06742	-.17280	-.02710	.072
H(S1)	-.37310	.06974	.12657	.191
H(S2)	-.43607	.10469	.08265	.191

Table 4.1(cont).

Anisotropic displacement parameters (\AA^2).

	U11	U22	U33	U12	U13	U23
I	.0505(10)	.0441(11)	.0466(10)	.0081(10)	-.0070(9)	-.0097(10)
MO	.0331(8)	.0352(9)	.0333(8)	.0039(9)	-.0043(8)	-.0015(9)
CL(1)	.140(7)	.161(8)	.302(12)	-.005(6)	-.074(8)	.022(9)
CL(2)	.197(10)	.168(9)	.579(22)	.080(8)	.234(13)	.124(12)

Table 4.2. Selected bond distances (\AA) and angles ($^\circ$) of complex (19).

I - MO	2.762(3)	I' - MO	2.739(7)
MO - C'	1.969	MO - C	1.946
MO - C(1)	2.302(14)	MO - C(2)	2.231(14)
MO - C(3)	2.282(13)	MO - C(4)	2.259(13)
MO - C(R1)	2.339(18)	MO - C(R2)	2.322(17)
MO - C(R3)	2.319(18)	MO - C(R4)	2.307(17)
MO - C(R5)	2.359(17)	MO - C(BU)	2.020
MO - C(CP)	2.004	CL(1) - C(S)	1.74(4)
CL(2) - C(S)	1.57(4)	C(1) - C(2)	1.461(19)
C(1) - C(4)	1.450(20)	C(1) - C(11)	1.467(17)
C(2) - C(3)	1.478(20)	C(2) - C(21)	1.499(22)
C(3) - C(4)	1.458(19)	C(3) - C(31)	1.468(19)
C(4) - C(41)	1.502(16)	C(R1) - C(R2)	1.42(3)
C(R1) - C(R5)	1.38(3)	C(R2) - C(R3)	1.36(3)
C(R3) - C(R4)	1.42(3)	C(R4) - C(R5)	1.39(3)

I - MO - I'	85.6(2)	I - MO - C(BU)	105.2
I - MO - C(CP)	108.2	I' - MO - C(BU)	104.4
I' - MO - C(CP)	108.3	C(BU) - MO - C(CP)	134.3
C(2) - C(1) - C(4)	90.1(11)	C(2) - C(1) - C(11)	133.1(13)
C(4) - C(1) - C(11)	135.0(12)	C(1) - C(2) - C(3)	89.6(11)
C(1) - C(2) - C(21)	128.6(12)	C(3) - C(2) - C(21)	133.3(12)
C(2) - C(3) - C(4)	89.2(11)	C(2) - C(3) - C(31)	131.8(12)
C(4) - C(3) - C(31)	137.7(13)	C(1) - C(4) - C(3)	90.8(11)
C(1) - C(4) - C(41)	132.9(13)	C(3) - C(4) - C(41)	132.0(13)
C(1) - C(11) - C(12)	119.9(11)	C(1) - C(11) - C(16)	120.1(13)
C(2) - C(21) - C(22)	122.3(15)	C(2) - C(21) - C(26)	117.5(14)
C(3) - C(31) - C(32)	120.8(12)	C(3) - C(31) - C(36)	119.1(11)
C(4) - C(41) - C(42)	120.0(13)	C(4) - C(41) - C(46)	119.8(13)
C(R2) - C(R1) - C(R5)	107.3(16)	C(R1) - C(R2) - C(R3)	107.3(16)
C(R2) - C(R3) - C(R4)	109.9(16)	C(R3) - C(R4) - C(R5)	105.5(16)
C(R1) - C(R5) - C(R4)	109.9(16)	CL(1) - C(S) - CL(2)	114.1(19)

Table 4.3. Fractional atomic coordinates and displacement parameters (\AA^2) of the crystal structure of complex (20).

	X/A	Y/B	Z/C	U
W	-.19348(7)	.23035(6)	-.24149(5)	.031
BR	-.0174(2)	.3613(2)	-.2044(2)	.057
S	-.3496(4)	.2407(4)	-.1098(3)	.042
F(2)	-.3934(9)	.0198(8)	-.0730(8)	.050(3)
F(3)	-.3180(11)	-.0885(10)	.0966(9)	.070(3)
F(4)	-.1881(11)	-.0041(9)	.2623(9)	.069(3)
F(5)	-.1298(12)	.1933(10)	.2568(9)	.079(4)
F(6)	-.2000(12)	.3027(10)	.0873(9)	.074(4)
O(1)	-.0055(12)	.0928(10)	-.0955(9)	.047(3)
O(2)	-.3714(13)	.3863(11)	-.3672(10)	.060(4)
C(1)	-.3004(16)	.1662(14)	.0014(12)	.032(4)
C(2)	-.3275(16)	.0646(14)	.0077(12)	.033(4)
C(3)	-.2895(17)	.0103(15)	.0942(14)	.042(5)
C(4)	-.2243(16)	.0534(15)	.1776(13)	.035(4)
C(5)	-.1941(18)	.1474(16)	.1728(15)	.049(5)
C(6)	-.2323(17)	.2064(15)	.0866(13)	.044(5)
C(7)	-.0739(16)	.1394(14)	-.1498(13)	.035(4)
C(8)	-.3044(17)	.3317(15)	-.3166(13)	.038(4)
C(9)	-.1958(22)	.0571(18)	-.4821(17)	.071(7)
C(10)	-.1287(19)	.0515(16)	-.3674(15)	.052(5)
C(11)	-.2483(19)	.0768(16)	-.2971(13)	.047(5)
C(12)	-.3243(18)	.1418(15)	-.3592(14)	.044(5)
C(13)	-.2526(19)	.1596(16)	-.4621(14)	.053(5)
C(14)	-.1333(18)	.2171(15)	-.4238(14)	.048(5)
C(15)	-.0549(16)	.1485(14)	-.3644(13)	.037(4)
H(9A)	-.26245	.00674	-.49518	.085
H(9B)	-.13353	.05553	-.53843	.085
H(10)	-.09493	-.01413	-.35154	.062
H(11)	-.25495	.03157	-.23829	.057
H(12)	-.41891	.14256	-.35029	.053
H(13)	-.31410	.17987	-.51806	.063
H(14)	-.12667	.28113	-.45771	.057
H(15)	.03496	.14140	-.33698	.044

Anisotropic displacement parameters (\AA^2).

	U11	U22	U33	U12	U13	U23
W	.0309(4)	.0311(5)	.0298(4)	.0032(5)	.0008(3)	.0059(4)
BR	.046(1)	.046(2)	.077(1)	-.011(1)	-.007(1)	.009(1)
S	.041(2)	.047(4)	.038(2)	.013(3)	.005(2)	.006(2)

Table 4.4. Selected bond distances (Å) and angles (°) of complex (20).

H - BR	2.521(3)	H - S	2.322(5)
H - C(7)	2.040(18)	H - C(8)	1.975(19)
H - C(11)	2.233(21)	H - C(12)	2.279(19)
H - C(14)	2.397(18)	H - C(15)	2.380(17)
H - C(A)	2.148	H - C(B)	2.282
S - C(1)	1.774(17)	F(2) - C(2)	1.34(2)
F(3) - C(3)	1.35(3)	F(4) - C(4)	1.35(3)
F(5) - C(5)	1.37(3)	F(6) - C(6)	1.33(3)
O(1) - C(7)	1.14(3)	O(2) - C(8)	1.17(3)
C(1) - C(2)	1.37(3)	C(1) - C(6)	1.36(3)
C(2) - C(3)	1.35(3)	C(3) - C(4)	1.35(3)
C(4) - C(5)	1.30(3)	C(5) - C(6)	1.38(3)
C(9) - C(10)	1.57(3)	C(9) - C(13)	1.51(4)
C(10) - C(11)	1.55(3)	C(10) - C(15)	1.49(3)
C(11) - C(12)	1.38(3)	C(12) - C(13)	1.52(3)
C(13) - C(14)	1.49(3)	C(14) - C(15)	1.41(3)

BR - H - S	108.2(2)	BR - H - C(7)	85.1(6)
BR - H - C(8)	89.6(6)	BR - H - C(A)	155.2
BR - H - C(B)	91.5	S - H - C(7)	91.8(5)
S - H - C(8)	85.2(5)	S - H - C(A)	96.6
S - H - C(B)	160.1	C(7) - H - C(8)	172.7(8)
C(7) - H - C(A)	94.3	C(7) - H - C(B)	92.5
C(8) - H - C(A)	92.7	C(8) - H - C(B)	92.6
C(A) - H - C(B)	63.7	H - S - C(1)	110.6(6)
S - C(1) - C(2)	123.0(12)	S - C(1) - C(6)	121.2(15)
C(2) - C(1) - C(6)	115.8(16)	F(2) - C(2) - C(1)	119.3(15)
F(2) - C(2) - C(3)	119.3(17)	C(1) - C(2) - C(3)	121.4(16)
F(3) - C(3) - C(2)	119.3(16)	F(3) - C(3) - C(4)	119.9(17)
C(2) - C(3) - C(4)	120.9(19)	F(4) - C(4) - C(3)	118.7(18)
F(4) - C(4) - C(5)	122.1(16)	C(3) - C(4) - C(5)	119.1(18)
F(5) - C(5) - C(4)	120.3(18)	F(5) - C(5) - C(6)	117.5(19)
C(4) - C(5) - C(6)	122.0(18)	F(6) - C(6) - C(1)	120.3(16)
F(6) - C(6) - C(5)	119.1(16)	C(1) - C(6) - C(5)	120.7(19)
H - C(7) - O(1)	176.5(16)	H - C(8) - O(2)	174.8(16)
C(10) - C(9) - C(13)	92.5(16)	C(9) - C(10) - C(11)	101.3(15)
C(9) - C(10) - C(15)	100.3(16)	C(11) - C(10) - C(15)	100.9(16)
C(10) - C(11) - C(12)	103.8(15)	C(11) - C(12) - C(13)	108.1(16)
C(9) - C(13) - C(12)	101.0(17)	C(9) - C(13) - C(14)	102.6(17)
C(12) - C(13) - C(14)	101.8(15)	C(13) - C(14) - C(15)	105.3(17)
C(10) - C(15) - C(14)	106.5(15)		

Table 4.5. Crystallographic and experimental details of the structure analyses of complexes (19) and (20).

Compound	[CpMoI(CO)(PhCCPh) ₂].CH ₂ Cl ₂	[WBr(CO) ₂ (SC ₆ F ₅)(C ₇ H ₈)]
Formula	C ₃₅ H ₂₇ Cl ₂ MoOI	C ₁₅ H ₈ BrF ₅ O ₂ SW
Formula Wt.	757.4	611.0
Crystal habit	dark purple plate	purple plate
Crystal size, mm	0.30x0.20x0.32	0.36x0.16x0.02
Crystal system	orthorhombic	monoclinic
Space group	Pbca (No.61)	P 2 ₁ /n
a, Å	11.736(4)	10.040(4)
b, Å	19.538(4)	13.379(4)
c, Å	27.238(4)	12.575(3)
β, °	—	91.88(3)
Obtained from	23 refln, 9<θ<14°	23 refln, 11<θ<15°
V, Å ³	6246(3)	1688.2(8)
Z	8	4
F(000)	2992	1136
d calc, gcm ⁻³	1.611	2.404
T, K	296	293
μ(Mo-Kα), cm ⁻¹	15.9	95.1
Absorption factors on F ²	0.91-1.10	0.80-1.32
Scan width, °	0.80	0.80
Max count time, s	100	90
Total refln measured	4284	3936
Unique refln	3806	2640
R _{INT}	—	0.063
Miller indicies	h	0-11
measured,	k	-15- 4
	l	-14-14
2θ range, °	4-44	4-48
Unique refln ≥ 3σ(I)	1385	1587
No. of parameters	130	116
R	0.047	0.052
R _w	0.053	0.056
Δρ _{max} , eÅ ⁻³	0.70	2.3
Δ/σ _{max}	0.04	0.026

SECTION III: MISCELLANEOUS STRUCTURES

CHAPTER 1: A SULPHIMIDE COMPLEX OF PALLADIUM

The succesful clinical exploitation of the complex Cisplatin, *cis*-[PtCl₂(NH₃)₂],¹¹³ has led to an expansion in research to find structural analogues of the complex which would have similar chemotherapeutic properties but reduced toxicity.¹¹⁴ One such group of complexes is comprised of [PtCl₂L₂] species where L is a heterocyclic ligand.

A number of palladium and platinum complexes containing heterocyclic sulphimide ligands were synthesised by Dr J.L. Davidson and his group but were found to form dimers, with bridging chlorine atoms. These compounds were insoluble in most solvents and therefore could not be purified. However, it was possible to split these dimers, particularly the palladium complexes, by reaction with phosphines such as L = PEt₃, PMe₂Ph, PMePh₂ and PPh₃, to produce complexes of the form [PdCl₂L(Me₂SNR)].

When R = 2-pyridyl or 2-pyrimidinyl, so that there is at least one heteroaromatic ring nitrogen as well as the ylide nitrogen present in the ligand, four possible isomers of [PdCl₂L(Me₂SNR)] can exist. The chloride ligands can be either *cis* or *trans* and the sulphimide ligand can bond either through the ylide nitrogen or the ring nitrogen atom.

Since N.M.R. could not conclusively distinguish which structures were formed, X-ray diffraction analysis was used to determine the precise nature of the metal coordination. When L = PEt₃ and R = 2-pyridyl, (1), it was found that the complex was *trans*-square planar and that the sulphimide ligand bonded via the ylide nitrogen with the R-N-S plane

approximately normal to the square plane of the metal.¹¹⁵

The N.M.R. data for the complex with $L = PPh_3$ and $R = 2\text{-pyrimidinyl}$, (2), differed from (1). The most notable difference was that in complex (1) the methyl groups attached to the sulphur atom were equivalent whereas in (2) they clearly differ. Consequently an X-ray analysis was carried out to determine its structure.

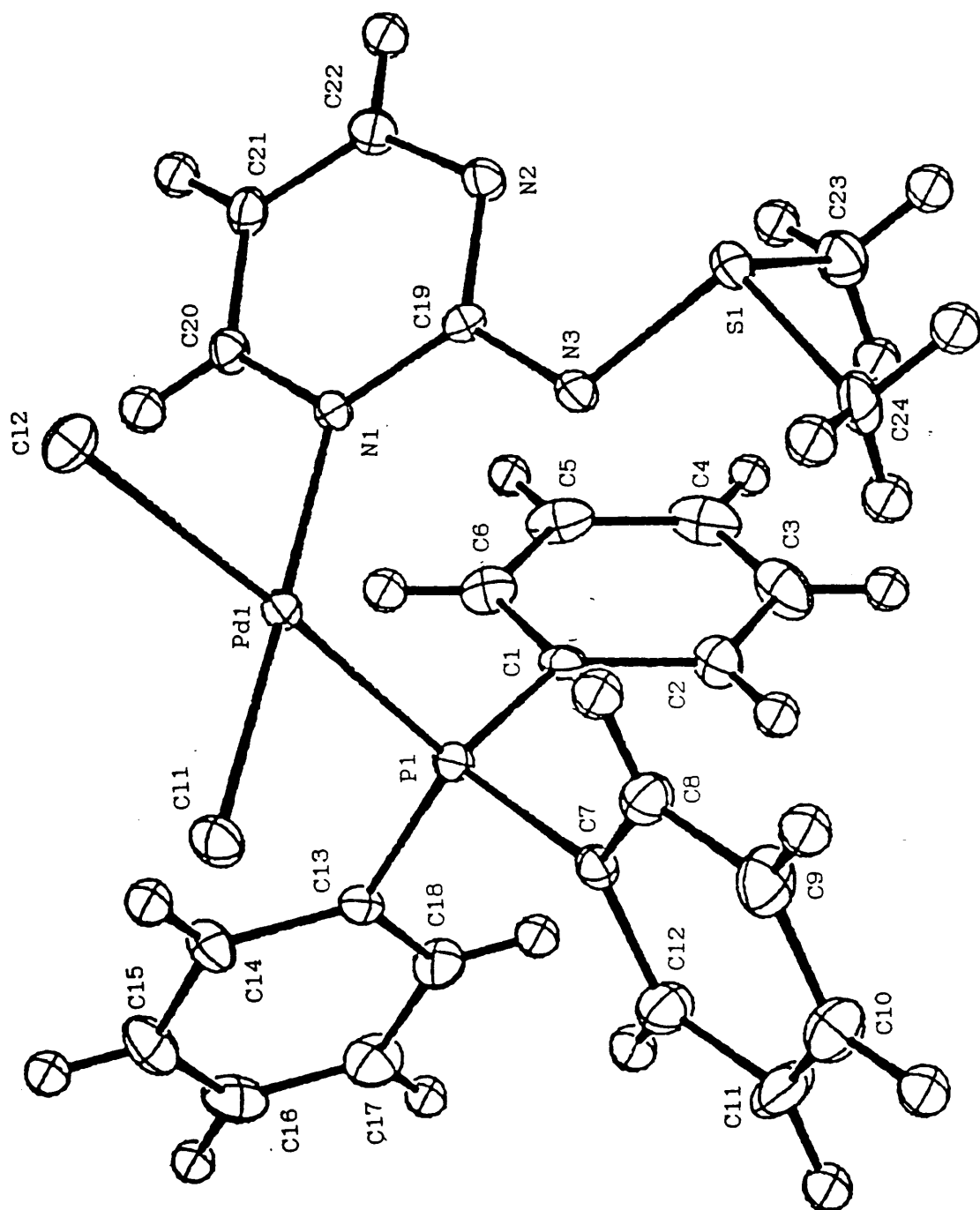
The crystals were shown to contain distinct molecules of complex (2) and of dichloromethane in equal numbers. Intermolecular contacts were consistent with appropriate van der Waals radii.⁵²

Complex (2) (see Figure 1(a) and Tables 1.1 and 1.2)¹¹⁶ consists of a central palladium atom with two chlorine atoms, a triphenylphosphine, and the $Me_2SNC_4H_3N_2$ sulphimide ligand attached to it in a square planar arrangement with the chloro ligands mutually *cis*. The sulphimide ligand is bonded to the palladium via a nitrogen atom of the pyrimidine ring. The Pd-N bond length of 2.022(5) Å is considerably shorter than that in *cis*-[PdCl₂-(Bu^tPCH₂C₁₀H₈N)], of 2.078(4) Å, the only other *cis*-PdCl₂PN sp² structure characterized.¹¹⁷ However this could be due to intramolecular overcrowding in the latter complex.

A comparison of the geometry of the sulphimide ligand of (2) with free sulphimides show no major differences.^{115,118} The short N(3)-C(19) bond length in (2) [1.332(7) Å compared to 1.366(6)-1.395(6) Å] and the C-N-S-Me torsion angles [-77.4(5) and -179.6(5)° compared to 92-100 and 157-164°] are the most obvious differences.

The PPh_3 ligand is attached to the palladium atom by a Pd-P single bond, length 2.246(2) Å, which is consistent

Figure 1(a). A view of complex (2). 50% probability ellipsoids are displayed for all atoms shown.



with other *cis*-PdCl₂PN species.^{117,119} The internal geometry of the PPh₃ ligand appears unaffected by coordination to the palladium atom, with the Pd-P-C-C torsion angles [16.0(5), 28.1(5) and 59.0(5)°] similar to those for solid PPh₃ [l.p.-P-C-C 21.7, 32.7 and 56.2°, l.p. = lone pair].¹²⁰

The Pd-Cl(1) bond distance is 0.085(3) Å shorter than that for Pd-Cl(2). This is probably due to the higher *trans*-influence of the phosphine compared with the nitrogen atom. This effect is confirmed in other *cis*-PdCl₂PN structures characterised.^{117,119}

The bonding of the sulphimide ligand to the palladium atom through atom N(1) differs from *trans*-[PdCl₂(PEt₃)-(Me₂SNC₅H₄N)], (1), where bonding is via the sulphimide nitrogen atom rather than that of the pyridine ring.¹¹⁵ This does not appear to be due to steric effects. A plot of the conformation energy¹²¹ of complex (2) as a function of the torsion angles $\omega_1 = \text{N}(1)\text{-Pd}(1)\text{-P}(1)\text{-C}(1)$ and $\omega_2 = \text{P}(1)\text{-Pd}(1)\text{-N}(1)\text{-C}(19)$ had its lowest local minimum close to the conformation found in the crystal structure of complex (2). When a similar calculation was carried out on the hypothetical molecule with the metal atom attached to atom N(3), rather than atom N(1), a local minimum of comparable energy to that above was found. An inspection of this conformation showed there to be no abnormally short non-bonded distances.

This suggests that electronic factors probably determine the mode of attachment of the sulphimide ligand in structures (1) and (2). However, the effect of a change in *trans* ligand together with a possible change in the

donor properties of the sulphimide nitrogen relative to the differing heteroaromatic ring nitrogens makes the exact reason difficult to predict. It should be noted however that the pyridine nitrogen atom is considerably more basic than the pyrimidine nitrogen atom [$pK_a = 5.6$, $pK_{a1} = 1.3$ respectively].¹²²

Experimental

The experimental procedure used for this structure is broadly similar to that described more fully in Chapter 2 of Section II. Only differences between the two methods are mentioned below.

Data Collection. The cell dimensions and experimental parameters are listed in Table 1.3. No decomposition correction was applied.

Structure Analysis. Anisotropic displacement parameters were used for all non-hydrogen atoms. All the hydrogen atoms were calculated geometrically (C-H 0.96 Å) and constrained to ride on their parent carbon atom. Isotropic displacement parameters were used for the hydrogen atoms and set at $1.2U_{ISO}(C)$ but not refined. Conformational energy calculations were made on a Uman computer using the Chemmod package.¹²¹

Table 1.1. Fractional atomic coordinates and displacement parameters (\AA^2) of the crystal structure of complex (2).

	X/A	Y/B	Z/C	U
PD(1)	.13484(2)	-.02586(5)	.06308(2)	.028
CL(1)	.16331(9)	-.26864(16)	.05638(9)	.044
CL(2)	.02778(9)	-.07232(18)	.13044(8)	.047
CL(3)	.55382(15)	-.88938(31)	.22358(19)	.135
CL(4)	.53335(18)	-.60839(33)	.15375(15)	.127
S(1)	.26966(9)	.31325(17)	.22255(8)	.037
P(1)	.23098(8)	.01252(17)	-.00784(7)	.032
N(1)	.1086(3)	.1862(5)	.0768(2)	.031
N(2)	.1303(3)	.4006(5)	.1516(3)	.036
N(3)	.2228(3)	.2130(5)	.1548(2)	.033
C(1)	.2538(3)	.2025(6)	-.0237(3)	.036
C(2)	.3274(4)	.2610(7)	-.0035(4)	.050
C(3)	.3411(5)	.4072(9)	-.0174(4)	.074
C(4)	.2826(6)	.4941(8)	-.0497(4)	.076
C(5)	.2097(5)	.4365(8)	-.0702(4)	.063
C(6)	.1960(4)	.2919(7)	-.0580(4)	.052
C(7)	.3218(3)	-.0721(6)	.0315(3)	.036
C(8)	.3353(3)	-.0908(7)	.1091(3)	.047
C(9)	.4023(4)	-.1555(9)	.1412(4)	.062
C(10)	.4572(4)	-.2003(9)	.0986(5)	.073
C(11)	.4461(4)	-.1807(10)	.0225(4)	.074
C(12)	.3778(4)	-.1194(8)	-.0101(3)	.056
C(13)	.2095(3)	-.0543(6)	-.1045(3)	.036
C(14)	.1494(4)	-.1514(7)	-.1238(3)	.049
C(15)	.1335(4)	-.1982(8)	-.1972(4)	.063
C(16)	.1768(5)	-.1485(9)	-.2512(4)	.066
C(17)	.2360(4)	-.0541(9)	-.2329(4)	.065
C(18)	.2522(3)	-.0048(7)	-.1592(3)	.052
C(19)	.1537(3)	.2692(6)	.1288(3)	.029
C(20)	.0401(3)	.2404(7)	.0473(3)	.035
C(21)	.0140(3)	.3736(7)	.0677(3)	.040
C(22)	.0610(4)	.4482(6)	.1204(3)	.043
C(23)	.3104(4)	.4680(7)	.1816(3)	.054
C(24)	.3544(3)	.2050(7)	.2393(3)	.048
C(25)	.5805(5)	-.7084(11)	.2282(5)	.096
H(2)	.36847	.19985	.01972	.059
H(3)	.39207	.44703	-.00316	.087
H(4)	.29245	.59510	-.05868	.088
H(5)	.16857	.49768	-.09341	.076
H(6)	.14508	.25207	-.07228	.062
H(8)	.29650	-.05890	.13934	.057
H(9)	.41057	-.16823	.19468	.075
H(10)	.50429	-.24493	.12184	.087
H(11)	.48494	-.21263	-.00771	.084
H(12)	.36948	-.10675	-.06360	.064
H(14)	.11896	-.18590	-.08624	.055
H(15)	.09171	-.26567	-.21050	.076
H(16)	.16546	-.18155	-.30205	.077
H(17)	.26637	-.01964	-.27048	.075
H(18)	.29401	.06272	-.14589	.059
H(20)	.00836	.18359	.01039	.041
H(21)	-.03508	.41221	.04547	.048
H(22)	.04361	.54100	.13645	.049
H(23A)	.33781	.52588	.22083	.067
H(23B)	.26959	.52476	.15443	.067
H(23C)	.34568	.43674	.14754	.067
H(24A)	.38187	.26281	.27855	.055
H(24B)	.38192	.19801	.19614	.055
H(24C)	.34352	.10859	.25586	.055
H(25A)	.63571	-.70174	.22697	.113
H(25B)	.56812	-.66742	.27445	.113

Table 1.1(cont).

Anisotropic displacement parameters (\AA^2).

	U11	U22	U33	U12	U13	U23
PD(1)	.0281(2)	.0282(2)	.0274(2)	-.0015(3)	.0028(2)	-.0017(2)
CL(1)	.0479(10)	.0289(9)	.0530(10)	-.0011(7)	.0046(8)	-.0003(7)
CL(2)	.0406(9)	.0559(11)	.0457(9)	.0006(8)	.0143(7)	.0093(8)
CL(3)	.083(2)	.101(2)	.223(3)	-.005(2)	.061(2)	-.007(2)
CL(4)	.142(2)	.141(3)	.095(2)	-.030(2)	.020(2)	-.001(2)
S(1)	.0359(9)	.0416(10)	.0326(8)	-.0018(7)	-.0027(7)	-.0065(7)
P(1)	.0327(8)	.0298(9)	.0317(8)	.0001(7)	.0053(6)	-.0021(7)
N(1)	.027(3)	.037(3)	.027(3)	.000(2)	0.000(2)	-.006(2)
N(2)	.035(3)	.029(3)	.043(3)	.004(2)	-.001(2)	-.005(2)
N(3)	.033(3)	.034(3)	.030(3)	0.000(2)	-.001(2)	-.006(2)
C(1)	.043(4)	.033(4)	.033(3)	.000(3)	.016(3)	-.005(3)
C(2)	.054(5)	.044(4)	.051(4)	-.011(4)	.014(3)	.005(3)
C(3)	.098(6)	.058(5)	.066(5)	-.037(5)	.025(5)	-.006(4)
C(4)	.125(7)	.040(5)	.064(5)	-.009(5)	.045(5)	.004(4)
C(5)	.085(6)	.044(5)	.062(5)	.019(4)	.033(4)	.017(4)
C(6)	.062(5)	.048(5)	.047(4)	.006(4)	.019(4)	.009(4)
C(7)	.036(3)	.033(4)	.038(3)	-.008(3)	.008(3)	-.006(3)
C(8)	.040(4)	.053(4)	.047(4)	.007(3)	.003(3)	-.003(3)
C(9)	.052(5)	.077(5)	.054(4)	.005(4)	-.011(4)	.005(4)
C(10)	.048(5)	.072(6)	.095(7)	.023(4)	-.016(5)	-.007(5)
C(11)	.042(5)	.104(7)	.076(6)	.025(4)	.015(4)	-.015(5)
C(12)	.049(4)	.074(5)	.045(4)	.012(4)	.001(4)	-.012(4)
C(13)	.039(3)	.032(4)	.035(3)	.005(3)	.004(3)	-.003(3)
C(14)	.059(5)	.048(4)	.039(4)	-.012(4)	.002(3)	-.006(3)
C(15)	.077(5)	.060(5)	.049(5)	-.009(4)	-.014(4)	-.011(4)
C(16)	.089(6)	.072(5)	.035(4)	.012(5)	.003(4)	-.016(4)
C(17)	.069(5)	.088(6)	.039(4)	.006(5)	.012(3)	-.004(4)
C(18)	.051(4)	.066(5)	.038(4)	-.003(4)	.016(3)	-.001(4)
C(19)	.029(3)	.031(4)	.028(3)	.004(3)	.007(3)	.003(3)
C(20)	.029(4)	.045(4)	.030(3)	-.002(3)	-.003(3)	-.002(3)
C(21)	.027(3)	.043(4)	.049(4)	0.000(3)	.000(3)	.001(3)
C(22)	.042(4)	.027(4)	.058(4)	.005(3)	.006(3)	-.001(3)
C(23)	.054(4)	.044(4)	.063(4)	-.003(4)	.001(3)	-.003(4)
C(24)	.040(4)	.047(4)	.053(4)	-.003(3)	-.019(3)	-.002(3)
C(25)	.090(7)	.106(8)	.090(6)	-.013(6)	.022(5)	-.021(6)

Table 1.2. Selected bond distances (Å) and angles (°) of complex (2).

PD(1) - CL(1)	2.289(2)	PD(1) - CL(2)	2.374(2)
PD(1) - P(1)	2.246(2)	PD(1) - N(1)	2.022(5)
CL(3) - C(25)	1.724(10)	CL(4) - C(25)	1.740(10)
S(1) - N(3)	1.656(5)	S(1) - C(23)	1.786(7)
S(1) - C(24)	1.767(7)	P(1) - C(1)	1.819(6)
P(1) - C(7)	1.815(6)	P(1) - C(13)	1.836(6)
N(1) - C(19)	1.372(7)	N(1) - C(20)	1.332(8)
N(2) - C(19)	1.352(8)	N(2) - C(22)	1.333(8)
N(3) - C(19)	1.332(7)	C(1) - C(2)	1.387(9)
C(1) - C(6)	1.379(9)	C(2) - C(3)	1.391(11)
C(3) - C(4)	1.361(13)	C(4) - C(5)	1.375(13)
C(5) - C(6)	1.371(10)	C(7) - C(8)	1.396(9)
C(7) - C(12)	1.367(9)	C(8) - C(9)	1.365(10)
C(9) - C(10)	1.356(11)	C(10) - C(11)	1.368(12)
C(11) - C(12)	1.374(11)	C(13) - C(14)	1.382(9)
C(13) - C(18)	1.377(8)	C(14) - C(15)	1.381(10)
C(15) - C(16)	1.375(11)	C(16) - C(17)	1.350(12)
C(17) - C(18)	1.394(9)	C(20) - C(21)	1.370(9)
C(21) - C(22)	1.355(9)		

CL(1) - PD(1) - CL(2)	92.2(1)	CL(1) - PD(1) - P(1)	86.8(1)
CL(1) - PD(1) - N(1)	175.9(2)	CL(2) - PD(1) - P(1)	176.0(1)
CL(2) - PD(1) - N(1)	84.8(2)	P(1) - PD(1) - N(1)	96.4(2)
N(3) - S(1) - C(23)	108.8(3)	N(3) - S(1) - C(24)	97.6(3)
C(23) - S(1) - C(24)	99.0(3)	PD(1) - P(1) - C(1)	115.5(2)
PD(1) - P(1) - C(7)	112.3(2)	PD(1) - P(1) - C(13)	114.0(2)
C(1) - P(1) - C(7)	106.1(3)	C(1) - P(1) - C(13)	101.2(3)
C(7) - P(1) - C(13)	106.8(3)	PD(1) - N(1) - C(19)	120.0(4)
PD(1) - N(1) - C(20)	120.9(4)	C(19) - N(1) - C(20)	117.8(5)
C(19) - N(2) - C(22)	116.9(5)	S(1) - N(3) - C(19)	112.2(4)
P(1) - C(1) - C(2)	122.8(5)	P(1) - C(1) - C(6)	118.8(5)
C(2) - C(1) - C(6)	118.5(6)	C(1) - C(2) - C(3)	119.9(7)
C(2) - C(3) - C(4)	120.5(8)	C(3) - C(4) - C(5)	119.8(8)
C(4) - C(5) - C(6)	120.0(7)	C(1) - C(6) - C(5)	121.2(7)
P(1) - C(7) - C(8)	118.1(5)	P(1) - C(7) - C(12)	124.2(5)
C(8) - C(7) - C(12)	117.7(6)	C(7) - C(8) - C(9)	120.2(6)
C(8) - C(9) - C(10)	120.8(7)	C(9) - C(10) - C(11)	120.2(8)
C(10) - C(11) - C(12)	119.1(7)	C(7) - C(12) - C(11)	121.9(6)
P(1) - C(13) - C(14)	120.7(5)	P(1) - C(13) - C(18)	120.0(5)
C(14) - C(13) - C(18)	119.2(6)	C(13) - C(14) - C(15)	119.6(6)
C(14) - C(15) - C(16)	120.6(7)	C(15) - C(16) - C(17)	120.3(7)
C(16) - C(17) - C(18)	119.8(7)	C(13) - C(18) - C(17)	120.4(6)
N(1) - C(19) - N(2)	122.3(5)	N(1) - C(19) - N(3)	115.7(5)
N(2) - C(19) - N(3)	122.0(5)	N(1) - C(20) - C(21)	122.2(6)
C(20) - C(21) - C(22)	116.7(6)	N(2) - C(22) - C(21)	124.0(6)
CL(3) - C(25) - CL(4)	112.0(6)		

Table 1.3. Crystallographic and experimental details of the structure analysis of complex (2).

Compound	$\text{cis-[PdCl}_2(\text{PPh}_3)(\text{Me}_2\text{SNC}_4\text{H}_3\text{N}_2)] \cdot \text{CH}_2\text{Cl}_2$
Formula	$\text{C}_{25}\text{H}_{26}\text{Cl}_4\text{N}_3\text{PPdS}$
Formula Wt.	679.8
Crystal habit	yellow plate
Crystal size, mm	$0.28 \times 0.10 \times 0.05$
Crystal system	monoclinic
Space group	$\text{P } 2_1/\text{n}$
a, Å	17.307(6)
b, Å	9.182(2)
c, Å	17.960(2)
β , °	97.15(2)
Obtained from	21 refln, $10^\circ < \theta < 15^\circ$
V, Å ³	2832(1)
Z	4
F(000)	1368
d calc, gcm ⁻³	1.594
T, K	295
$\mu(\text{Mo-K}\alpha)$, cm ⁻¹	11.74
Absorption factors on F	0.95-1.05
Scan width, °	0.70
Max count time, s	150
Total refln measured	4414
Unique refln	3463
R _{INT}	0.024
Miller indicies	h
measured,	k
	l
2 θ range, °	4-44
Unique refln $\geq 3\sigma(\text{I})$	1999
No. of parameters	316
R	0.024
R _w	0.028
$ \Delta\rho _{\text{max}}$, eÅ ⁻³	0.30
$\Delta/\sigma_{\text{max}}$	0.016

CHAPTER 2: STRUCTURE OF HEXAKIS(THIOUREA)RUTHENIUM(II)
TRIFLUOROMETHYLSULPHONATE.

During kinetic studies of the replacement of coordinated water in $[\text{Ru}(\text{H}_2\text{O})_6]^{2+}$, purple crystals of $[\text{Ru}(\text{SC}(\text{NH}_2)_2)_6][\text{CF}_3\text{SO}_3]_2$ were isolated by Dr.D. Richens of St. Andrews University. The new compound was the first thiourea complex of ruthenium(II) to be isolated and the mode of attachment (through N or S) of thiourea to the metal atom could not be unambiguously determined from spectroscopic data. Moreover, oxidation of the complex in solution gave a species, $[\text{Ru}(\text{SC}(\text{NH})\text{NH}_2)_3]$, in which deprotonated thiourea appears to behave as a N,S-bonded chelate ligand.

X-ray diffraction analysis showed the crystal structure to contain layers of discrete cations of $[\text{Ru}(\text{SC}(\text{NH}_2)_2)_6]^{2+}$ (Figure 2(a)), the metal atoms lying exactly on the 001 planes, with $[\text{CF}_3\text{SO}_3]^-$ anions (Figure 2(b)) lying between these layers. Each anion is hydrogen-bonded to four cations, three in the same layer and one in an adjacent layer (Figure 2(c)). Two N-H...S hydrogen bonds also connect adjacent cations (see Table 2.2 for bond distances and angles).

The centrosymmetric cations contain planar thiourea ligands coordinated to the metal atom through sulphur as is usual for homoleptic $\text{M}(\text{SC}(\text{NH}_2)_2)_6$ species.¹²³⁻¹²⁵ The RuS_6 coordination is D_{3d} trigonal antiprismatic rather than octahedral, with elongation along an axis through the metal atom and the centroid of the triangle defined by the three unique sulphur atoms so that the S-Ru-S angles are in the

Figure 2(a). Structure of the $[\text{Ru}(\text{SC}(\text{NH}_2)_2)_6]^{2+}$ cation. Hydrogen atoms are presented as spheres of arbitrary size and 50% probability ellipsoids are displayed for all other atoms.

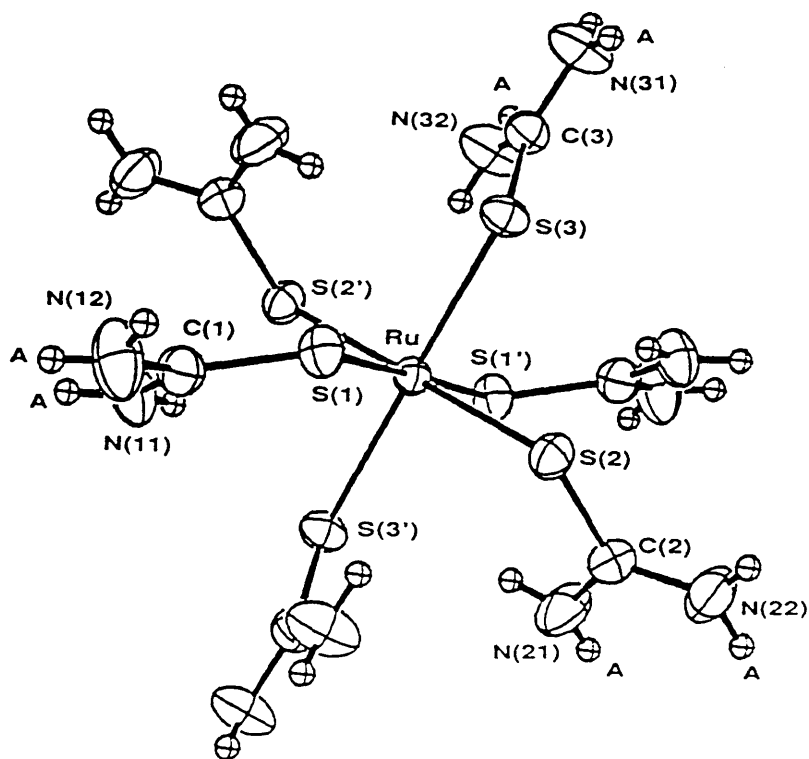
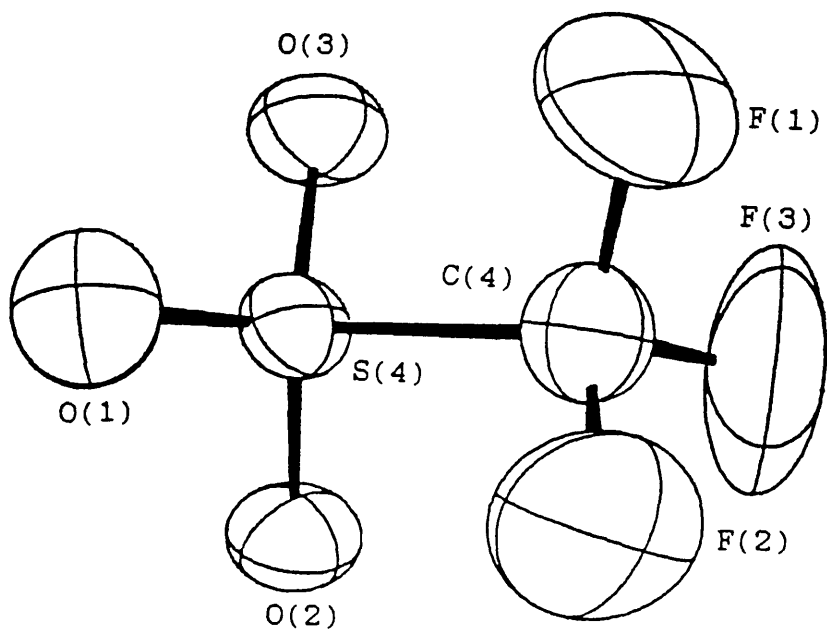


Figure 2(b). Structure of the $[\text{CF}_3\text{SO}_3]^-$ anion. 50% probability ellipsoids are displayed for all atoms.



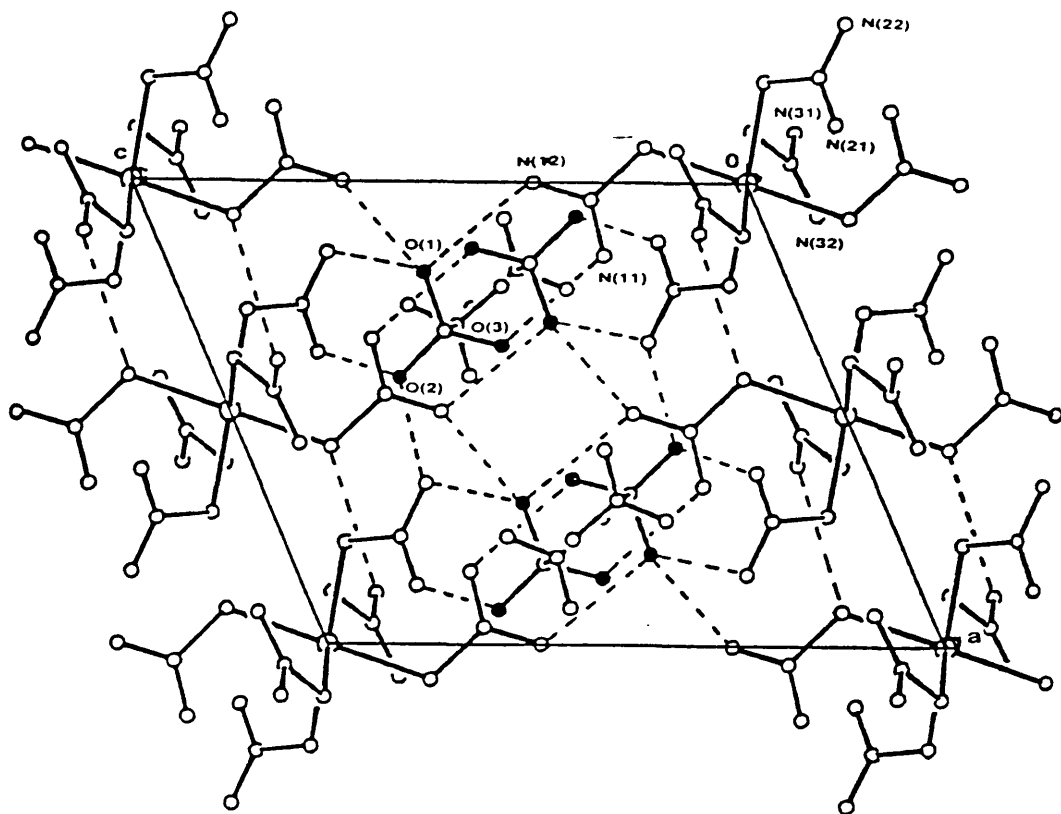


Figure 2(c). Crystal packing diagram showing the hydrogen bonding between adjacent ions.

range $81.4(1)$ – $82.9(1)^\circ$. Similar distortions occur in several other metal ion complexes of thiourea. This suggests that it is unlikely to be due to electronic factors such as the Jahn-Teller effect.

The mean Ru-S bond distance [$2.421(3)$ Å] is longer than values reported for other Ru(II)-S bonds¹²⁶ but comparable with that for the Tc^{3+} complex of thiourea¹²⁴ [$2.428(16)$ Å]. The mean S-C bond length is similar to values for X-ray studies of free thiourea^{127,128} but the mean N-C distance is slightly shorter than those in the free ligand [$1.316(2)$ compared to $1.340(6)$ and $1.333(3)$ Å]. The Ru-S-C-N torsion angles show the ruthenium atom to be close to the plane of the thiourea ligands as previously noted in the

Tc³⁺ complex.¹²⁴ The S-Ru-S-C tilt angles¹²⁹ indicate that these Ru-thiourea planes bisect S-Ru-S angles. This probably minimises thiourea-thiourea ligand repulsion.

The [CF₃SO₃]⁻ anions adopt a staggered C_{3v} conformation and their bond distances and angles are unexceptional.¹³⁰

Experimental

The procedure used for this crystal structure analysis was similar to that described in Section 2, Chapter 2. Only differences in the methods are mentioned below.

Data Collection. The cell dimensions and experimental parameters are listed in Table 2.3. No decomposition correction was required.

Application of the transformation (0 $\bar{1}$ 0/100/102) gave an A-centered cell with $a = 11.06$, $b = 11.17$ and $c = 25.06$ Å and $\alpha = \beta = \gamma = 90^\circ$ to within experimental error. However, transformation of the intensity data and subsequent merging assuming Laue symmetry mmm gave $R_{\text{INT}} = 0.39$ confirming the monoclinic symmetry.

Structure Analysis. All hydrogen atom positions were found in difference syntheses. The positional and the isotropic displacement parameters of the hydrogen atoms were refined.

The ruthenium atom lies on a centre of symmetry. Consequently, each asymmetric unit comprises half a cation and one anion in each asymmetric unit.

Final atomic positions and displacement parameters are given in Table 2.1.

Table 2.1. Fractional atomic coordinates and displacement parameters (\AA^2) of $[\text{Ru}(\text{SC}(\text{NH}_2)_2)_6][\text{CF}_3\text{SO}_3]_2$.

	X/A	Y/B	Z/C	U
RU	.00000	.00000	.00000	.026
S(1)	-.07517(5)	-.00101(5)	.14201(4)	.037
S(2)	-.22020(4)	.07580(5)	-.09936(4)	.034
S(3)	-.11327(4)	-.19199(5)	-.04361(4)	.035
S(4)	-.17574(6)	-.53665(6)	-.40784(4)	.043
N(11)	.1588(2)	.0688(2)	.2833(2)	.048
N(12)	.0016(3)	.0432(3)	.3459(2)	.063
N(21)	-.1286(2)	.2461(2)	-.1865(2)	.058
N(22)	-.3459(2)	.2061(2)	-.2714(2)	.054
N(31)	-.1111(2)	-.4112(2)	-.1140(2)	.058
N(32)	.0689(2)	-.2994(3)	-.0925(2)	.064
O(1)	-.3038(2)	-.5707(2)	-.4149(2)	.065
O(2)	-.0753(2)	-.5497(2)	-.3020(1)	.057
O(3)	-.1446(2)	-.5847(2)	-.4908(1)	.065
F(1)	-.2734(3)	-.3523(2)	-.5312(2)	.136
F(2)	-.2319(3)	-.3215(2)	-.3667(2)	.136
F(3)	-.0800(3)	-.3273(3)	-.4177(3)	.163
C(1)	.0382(2)	.0404(2)	.2660(2)	.041
C(2)	-.2313(2)	.1834(2)	-.1929(2)	.040
C(3)	-.0458(2)	-.3087(2)	-.0866(2)	.037
C(4)	-.1891(4)	-.3759(3)	-.4315(2)	.085
H(11A)	.217(2)	.075(3)	.344(2)	.049(7)
H(11B)	.178(2)	.070(3)	.231(2)	.052(7)
H(12A)	.055(3)	.052(3)	.403(2)	.049(7)
H(12B)	-.062(3)	.005(3)	.340(3)	.064(10)
H(21A)	-.134(3)	.290(3)	-.225(2)	.049(8)
H(21B)	-.051(3)	.220(3)	-.145(2)	.054(8)
H(22A)	-.354(3)	.260(3)	-.320(3)	.078(10)
H(22B)	-.407(3)	.171(3)	-.283(3)	.073(11)
H(31A)	-.188(3)	-.415(3)	-.116(2)	.073(9)
H(31B)	-.078(4)	-.452(4)	-.139(3)	.080(12)
H(32A)	.087(3)	-.350(3)	-.117(3)	.080(12)
H(32B)	.111(3)	-.234(3)	-.075(3)	.075(10)

Anisotropic displacement parameters (\AA^2).

	U11	U22	U33	U12	U13	U23
RU	.0234(1)	.0251(1)	.0241(1)	-.0018(1)	.0097(1)	.0002(1)
S(1)	.0332(2)	.0426(3)	.0299(2)	-.0038(2)	.0153(2)	-.0015(2)
S(2)	.0282(2)	.0353(3)	.0331(2)	.0015(2)	.0108(2)	.0045(2)
S(3)	.0289(2)	.0291(2)	.0425(2)	-.0044(2)	.0170(2)	-.0033(2)
S(4)	.0522(3)	.0325(3)	.0354(2)	.0041(3)	.0158(2)	.0047(2)
N(11)	.0454(9)	.0600(14)	.0314(8)	-.0107(10)	.0158(7)	-.0093(9)
N(12)	.0552(12)	.0943(21)	.0342(9)	-.0165(14)	.0231(9)	-.0133(12)
N(21)	.0512(11)	.0534(14)	.0552(12)	-.0024(11)	.0139(10)	.0259(11)
N(22)	.0451(10)	.0543(14)	.0477(11)	.0042(11)	.0054(9)	.0177(10)
N(31)	.0510(11)	.0363(12)	.0782(15)	-.0108(10)	.0350(11)	-.0207(11)
N(32)	.0451(10)	.0437(14)	.0947(18)	-.0070(10)	.0405(12)	-.0191(13)
O(1)	.0533(9)	.0615(13)	.0684(11)	.0035(10)	.0221(8)	.0059(10)
O(2)	.0571(9)	.0591(12)	.0393(8)	.0052(9)	.0088(7)	.0043(8)
O(3)	.0783(11)	.0570(13)	.0477(9)	.0049(11)	.0290(8)	-.0071(9)
F(1)	.231(3)	.056(1)	.073(1)	.030(2)	.027(2)	.026(1)
F(2)	.226(3)	.049(1)	.103(2)	.037(2)	.073(2)	-.003(1)
F(3)	.183(3)	.070(2)	.207(3)	-.053(2)	.097(2)	.001(2)
C(1)	.0454(10)	.0375(11)	.0325(8)	-.0015(10)	.0185(8)	-.0009(9)
C(2)	.0401(9)	.0330(11)	.0374(9)	.0040(9)	.0121(8)	.0028(8)
C(3)	.0334(8)	.0332(10)	.0374(9)	-.0019(8)	.0137(7)	-.0020(8)
C(4)	.137(3)	.037(2)	.059(2)	-.007(2)	.036(2)	.001(1)

Table 2.2. Selected bond distances (Å) and angles (°) of
[Ru(SC(NH₂)₂)₆][CF₃SO₃]₂.

RU - S(1)	2.415(1)	RU - S(2)	2.427(1)
RU - S(3)	2.420(1)	S(1) - C(1)	1.720(3)
S(2) - C(2)	1.717(3)	S(3) - C(3)	1.716(3)
S(4) - O(1)	1.445(2)	S(4) - O(2)	1.437(2)
S(4) - O(3)	1.422(2)	S(4) - C(4)	1.804(4)
N(11) - C(1)	1.307(3)	N(12) - C(1)	1.317(3)
N(21) - C(2)	1.312(4)	N(22) - C(2)	1.320(3)
N(31) - C(3)	1.317(4)	N(32) - C(3)	1.321(3)
F(1) - C(4)	1.335(5)	F(2) - C(4)	1.313(5)
F(3) - C(4)	1.275(6)	N - H	0.70(3)-0.87(4)

S(1) - RU - S(2)	81.3(1)	S(1) - RU - S(3)	82.9(1)
S(2) - RU - S(3)	81.7(1)	RU - S(1) - C(1)	116.0(1)
RU - S(2) - C(2)	114.1(1)	RU - S(3) - C(3)	119.1(1)
O(1) - S(4) - O(2)	112.8(2)	O(1) - S(4) - O(3)	114.0(2)
O(1) - S(4) - C(4)	103.8(2)	O(2) - S(4) - O(3)	115.3(2)
O(2) - S(4) - C(4)	104.7(2)	O(3) - S(4) - C(4)	104.7(2)
S(1) - C(1) - N(11)	122.3(2)	S(1) - C(1) - N(12)	118.4(2)
N(11) - C(1) - N(12)	119.3(2)	S(2) - C(2) - N(21)	121.4(2)
S(2) - C(2) - N(22)	119.6(2)	N(21) - C(2) - N(22)	119.0(3)
S(3) - C(3) - N(31)	118.5(2)	S(3) - C(3) - N(32)	122.3(2)
N(31) - C(3) - N(32)	119.2(3)	S(4) - C(4) - F(1)	110.4(3)
S(4) - C(4) - F(2)	110.9(3)	S(4) - C(4) - F(3)	112.6(3)
F(1) - C(4) - F(2)	107.8(4)	F(1) - C(4) - F(3)	107.7(4)
F(2) - C(4) - F(3)	107.2(3)		

Torsion Angles

S(3)-Ru-S(1)-C(1)	135.3(1)	Ru-S(1)-C(1)-N(11)	-1.2(2)
S(1)-Ru-S(2)-C(2)	143.4(1)	Ru-S(2)-C(2)-N(21)	-21.3(2)
S(2)-Ru-S(3)-C(3)	136.2(1)	Ru-S(3)-C(3)-N(32)	2.9(2)

Hydrogen Bonds

A-H---B	A-H	H...B	A...B	A-H---B
N(11)-H(11A)...O(3 ⁱ)	0.83(3)	2.16(3)	2.989(3)	174(2)
N(12)-H(12A)...O(1 ⁱ)	0.77(3)	2.36(3)	3.131(3)	172(3)
N(12)-H(12B)...O(1 ⁱⁱ)	0.80(3)	2.29(3)	3.012(3)	150(3)
N(21)-H(21A)...O(2 ⁱⁱⁱ)	0.70(3)	2.30(3)	2.958(3)	158(3)
N(22)-H(22A)...O(1 ⁱⁱⁱ)	0.87(3)	2.47(3)	3.308(3)	161(3)
N(22)-H(22B)...O(2 ^{iv})	0.74(3)	2.24(3)	2.978(3)	172(3)
N(31)-H(31A)...S(1 ^v)	0.85(3)	2.70(3)	3.512(2)	161(3)

Symmetry Code

i	$\frac{1}{2}+x, -\frac{1}{2}-y, 1+z$
ii	$-\frac{1}{2}-x, \frac{1}{2}+y, -z$
iii	$x, 1+y, z$
iv	$-\frac{1}{2}+x, -\frac{1}{2}-y, z$
v	$-\frac{1}{2}-x, -\frac{1}{2}+y, -z$

Table 2.3. Crystallographic and experimental details of the structure analysis of $[\text{Ru}(\text{SC}(\text{NH}_2)_2)_6][\text{CF}_3\text{SO}_3]_2$.

Compound	$[\text{Ru}(\text{SC}(\text{NH}_2)_2)_6][\text{CF}_3\text{SO}_3]_2$
Formula	$\text{C}_8\text{H}_{24}\text{F}_6\text{N}_{12}\text{O}_6\text{RuS}_8$
Formula Wt.	855.9
Crystal habit	purple plate
Crystal size, mm	0.48x0.40x0.36
Crystal system	monoclinic
Space group	$P 2_1/a$
a, Å	11.173(1)
b, Å	11.064(1)
c, Å	13.722(1)
β , °	113.96(1)
Obtained from	22 refln, $11^\circ < \theta < 15^\circ$
V, Å ³	1550.1(3)
Z	2
F(000)	860
d calc, gcm ⁻³	1.834
T, K	295
$\mu(\text{Mo-K}\alpha)$, cm ⁻¹	11.0
Absorption factors on F	0.86-1.10
Scan width, °	0.80
Max count time, s	90
Total refln measured	6292
Unique refln	4506
R_{INT}	0.025
Miller indices	h
measured,	k
	l
2θ range, °	4-60
Unique refln $\geq 3\sigma(I)$	3609
No. of parameters	236
R	0.026
R_w	0.037
$ \Delta\rho _{\text{max}}$, eÅ ⁻³	0.43
$\Delta/\sigma_{\text{max}}$	0.034

CHAPTER 3: THE STRUCTURE OF TETRACHLOROTHIOPHENE 1,1-DIOXIDE AT 295 AND 150K

3.1 Introduction

The main objectives of most modern single crystal X-ray diffraction studies are the identification of the chemical nature of each atom in the asymmetric unit and the derivation of reasonably accurate atomic positions. The analyses described in earlier parts of this thesis are typical of this approach. The objectives can be met by employing relatively standard experimental and computational techniques which incorporate the assumption that the individual atomic electron densities are spherically symmetric before allowance is made for thermal smearing.

During the last decade great progress has been made in developing methods of X-ray analysis which lead to accurate determination of the true molecular electron density. These methods allow deformation of the valence electron density from spherical symmetry to be studied and thus reveal the effects of bond formation and lone pair localisation. They also permit atomic charges to be determined experimentally.¹³¹⁻¹³³

To achieve experimental electron density determinations which show these features it is necessary to work at low temperature so that thermal smearing is lessened and the intensities of high angle ($\sin\theta/\lambda$ up to 1.2 \AA^{-1}) reflections can be measured. Careful attention to

processing the raw intensity measurements is required and least-squares refinement must allow for the asphericity of the atomic electron densities. The recent study of an iron(II)-porphyrin illustrates this approach.¹³⁴

The work described below was carried out as a preliminary exploration of the experimental and computational procedures required in studies of accurate electron densities. Specifically (i) an attempt was made to collect high-angle low temperature data, (ii) Blessing's data processing package¹³⁵ was installed on the Gould 3227 computer by the present author and used to process the data and (iii) the resulting structure was refined by Hansen's LSMOL multipole charge density refinement program as modified by Dr.P.R. Mallinson.¹³⁶

3.2 Structure Analysis Of Tetrachlorothiophene

1,1-Dioxide At 295K

The chlorinated thiophene 1,1-dioxide, (1) was prepared in the du Pont Laboratories in 1980 and is useful as a Diels-Alder reagent since it is stable, yet reacts readily with double bonds with loss of SO₂ to add a tetrachlorobutadienediyl unit.¹³⁷ The small size and symmetry of the molecule and the stability of its crystals make it a suitable choice of test crystal. The details of the data collection at 295K are given in Table 3.5. The structure was readily solved by direct methods and refined by full-matrix least-squares with all atoms anisotropic. Final coordinates and temperature factors are given in Table 3.1. The refinement gave normal bond lengths and angles (Table

3.3) and the structure proved to be free of disorder, although the chlorine atoms showed fairly large thermal anisotropy (U_{ii} 0.04-0.14 Å²).

3.3 Structure Analysis At 150K

Accordingly an attempt was made to collect a high resolution data set at 150K (Table 3.5). Initially it was intended to collect to higher angle and also to collect several asymmetric units of data but the crystal was lost during replenishment of the liquid nitrogen supply after an asymmetric unit of data to $\theta = 35^\circ$ had been measured. For the data set thus collected reflection profiles were recorded and the data were processed using Blessing's program.¹³⁵ This determines the variable reflection width as a function of a unit vector normal to the incident and diffracted beam and also makes a careful allowance for all factors in deriving $\sigma(F)$. (It also has a very flexible procedure for scaling data according to the intensity standards but the small variation in the standards during this experiment made this feature superfluous.) A Gaussian absorption correction was also applied to the 150K data.

The data so obtained were used to refine the structure in the same way as for the 295K data set. Coordinates and temperature factors are presented in Table 3.2 and bond lengths and angles for both refinements in Table 3.3. A drawing of the molecule based on the 150K refinement is given in Figure 3(a).

As expected the transition from 295 to 150K greatly lowers the thermal motion (U_{ii} 0.02-0.07 Å² for Cl, see

Figure 3(a). Structure of $C_4Cl_4O_2S$ at 150K. 50% probability ellipsoids are displayed for all atoms.

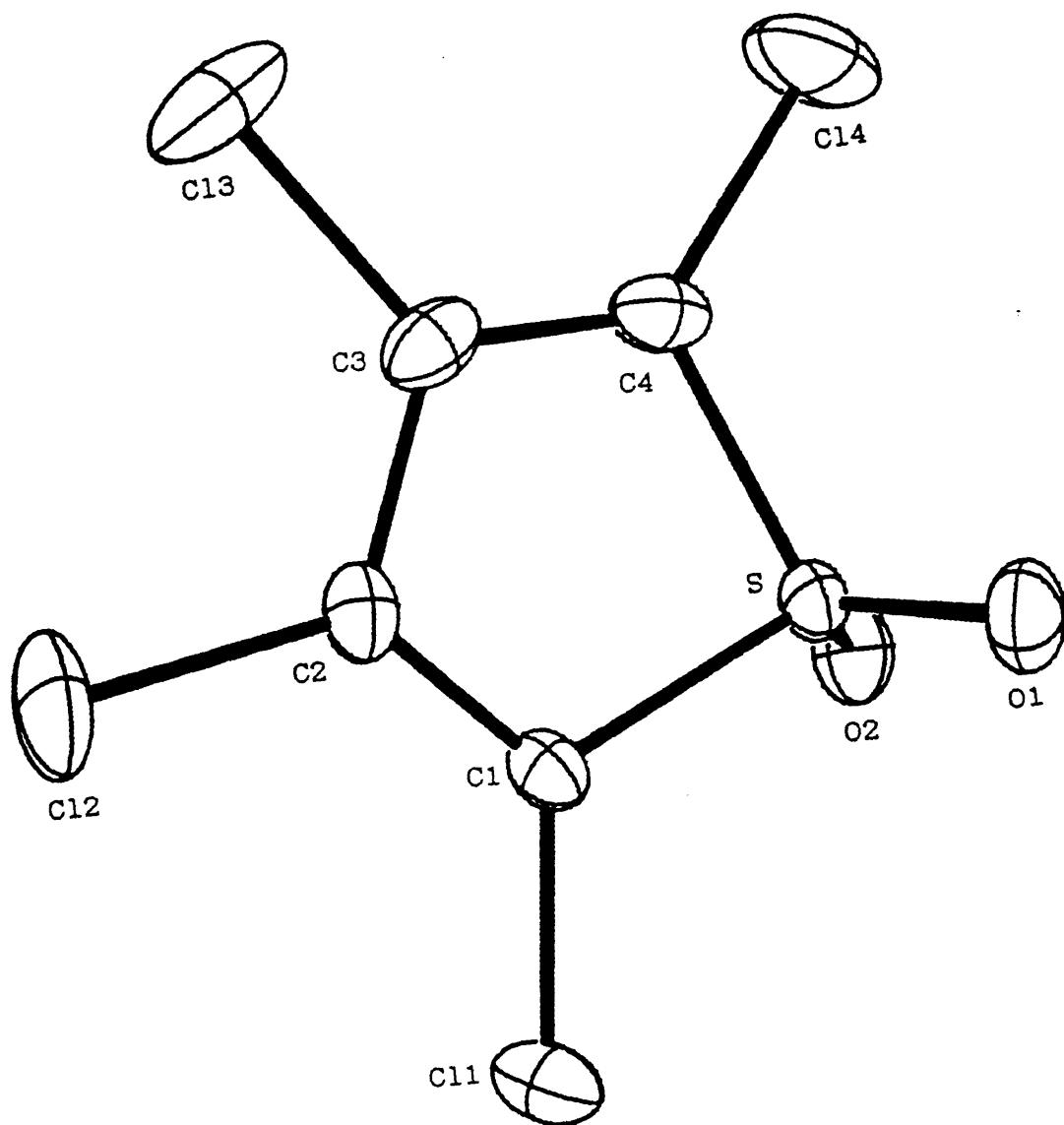


Table 3.2) and bond lengths tend to be slightly longer at 150K (see Table 3.3). The molecular structure at the two temperatures is, however, basically the same and approximates to C_{2v} symmetry. Since they are more accurate only the results for the 150K structure will be discussed. The butadiene unit, C(1)-C(4), is coplanar to ± 0.06 Å with S displaced 0.09 Å from this plane. Atoms Cl(1) and Cl(2) are displaced in the same direction by 0.13 and 0.08 Å from this plane whereas Cl(3) and Cl(4) show smaller displacements of 0.07 and 0.04 Å in the opposite direction. The C-Cl bond lengths show significant variation, from 1.682(3) to 1.696(3) Å, for reasons which are not obvious.

As a check on the profile analysis (and of Blessing's program in general) the 150K data were processed normally (i.e. assuming that the peak was centred in the middle 2/3rds of the scan). This yielded fewer unique reflections with $I \geq 3\sigma(I)$ (2674 compared with 2721) showing the improved $\sigma(I)/I$ ratios for weaker reflections obtained by profile analysis. Refinement of this data set gave $R = 0.043$, $R_w = 0.067$ and distances and angles little different from those in Table 3.3. The maximum change in a bond length between the two low temperature refinements was 0.003 Å, for example.

3.4 Multipole Refinement

As previously noted, the 150K data collection was prematurely terminated. Nevertheless, it was considered worthwhile to attempt a multipole refinement using Dr.P.R. Mallinson's adaption of LSMOL. this expands the rest atomic

density, $\rho(r)$, in spherical harmonic functions of up to order 4:

$$\rho(r) = P_C \rho_C(r) + P_V k'^3 \rho_V(k'r) + \sum_{l=0}^4 k''^3 R_l(k''r) \sum_{m=-l}^{+l} P_{lm} y_{lm}$$

where $\rho_C(r)$ is a spherical core electron density normalised to unity, $\rho_V(k'r)$ is a spherical valence electron density normalised to unity, $R_l(k''r)$ defines the radial dependence of the multipole density function and y_{lm} defines its angular dependence.

P_C , P_V , P_{lm} are refinable population coefficients and the total number of valence electrons associated with the atom, Z , can be obtained from $P_V + P_{00}$ since the $l \neq 0$ functions integrate to zero. k' and k'' are refinable radial parameters. Analytical forms for R_l and y_{lm} are given in reference 136: $l = 0, 1$ and 2 correspond to s , p and d type electron densities. For first row atoms $l = 0$ to 3 is considered a sufficient multipole expansion, whereas $l = 4$ functions are required for second and higher row elements.

The multipole refinement of the 150K data set by LSMOL was based on the 2792 reflections with $I \geq 3\sigma(I)$ from the total of 3152 for which intensity measurements were made. Inclusion of the multipole expansion increased the number of refined parameters from 101 to 330. This refinement converged with $R = 0.028$ and $R_w = 0.044$ indicating a highly significant improvement in structure factor agreement. However, the derived atomic charges (Table 3.4) are poorly determined and some are implausible, notably the value for the sulphur atom and the differences between the equivalent pairs C(1) and C(4), C(2) and C(3). In addition, although the deformation density in the mean ring plane

(Figure 3(b)) contains maxima which could be attributed to bonding electron density, notably at the C-C bond centroids, it is dominated by features around Cl(2) and Cl(3) rising to nearly $2\text{e}\text{\AA}^{-3}$ which can only be attributed to the effects of uncorrected absorption or extinction.

3.5 Conclusions

The work just described led to succesful structure analyses of (1) at 295 and 150K in that the atomic positions were accurately located. In addition Blessing's data reduction program was succesfully installed and experience was gained using LSMOL. However, the quality of the 150K data is insufficient for accurate determination of the features of the bonding electron density.

Figure 3(b). Deformation density in plane of C_4S ring.

The coefficients used to compute the difference density function were of the form $F^m - F^s$ where F^m is the calculated structure factor from the multipole refinement and F^s is the structure factor calculated assuming spherical atomic densities centred on the positions derived from the multipole refinement. Contour levels are at $0.1e\text{\AA}^{-3}$ intervals starting at $-0.7e\text{\AA}^{-3}$ and are numbered 1 ($-0.7e\text{\AA}^{-3}$) to 13 ($+1.0e\text{\AA}^{-3}$).

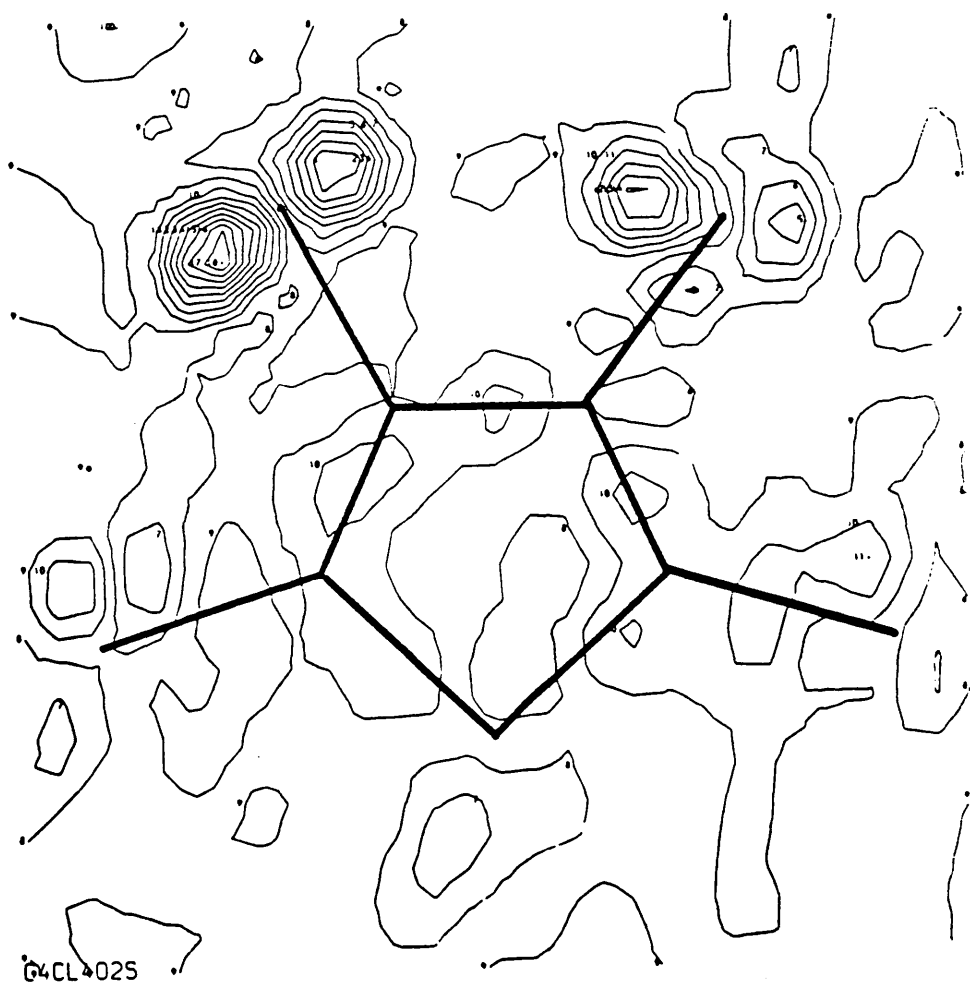


Table 3.1 Atomic positional and anisotropic displacement parameters (\AA^2) of (1) at 295K

	X/A	Y/B	Z/C	U
CL(1)	.27352(6)	.43155(14)	.72829(9)	.077
CL(2)	.34478(8)	.72255(13)	.55626(13)	.094
CL(3)	.44515(7)	.55630(20)	.36593(10)	.095
CL(4)	.44699(8)	.12632(19)	.38339(11)	.106
S	.35739(5)	.19408(10)	.58474(7)	.048
O(1)	.40873(16)	.13019(36)	.66518(22)	.072
O(2)	.29582(16)	.09288(29)	.55500(25)	.068
C(1)	.33004(18)	.40217(37)	.62011(28)	.048
C(2)	.35825(18)	.51195(42)	.55072(31)	.052
C(3)	.40163(19)	.43363(48)	.46092(28)	.055
C(4)	.40450(20)	.26743(51)	.46513(26)	.054

	U11	U22	U33	U12	U13	U23
CL(1)	.0724(7)	.0806(7)	.0776(7)	.0071(5)	.0309(5)	-.0145(5)
CL(2)	.1029(10)	.0359(5)	.1439(11)	-.0009(5)	.0090(8)	.0069(5)
CL(3)	.0730(7)	.1364(12)	.0767(7)	-.0256(7)	.0065(5)	.0489(7)
CL(4)	.1165(11)	.1216(12)	.0792(8)	.0349(9)	.0348(7)	-.0267(7)
S	.0580(5)	.0357(4)	.0512(4)	.0047(3)	.0078(3)	.0006(3)
O(1)	.087(2)	.062(2)	.066(2)	.019(1)	-.004(1)	.013(1)
O(2)	.074(2)	.041(1)	.090(2)	-.011(1)	.005(1)	.000(1)
C(1)	.049(2)	.039(1)	.057(2)	.001(1)	.010(1)	-.005(1)
C(2)	.048(2)	.040(2)	.069(2)	-.001(1)	-.004(2)	.004(1)
C(3)	.044(2)	.071(2)	.051(2)	-.005(2)	-.002(1)	.014(2)
C(4)	.052(2)	.067(2)	.044(2)	.005(2)	.010(1)	-.002(1)

Table 3.2 Atomic positional and anisotropic displacement parameters (\AA^2) of (1) at 150K

	X/A	Y/B	Z/C	U
CL(1)	.27236(3)	.43417(7)	.72966(4)	.031
CL(2)	.34378(4)	.72891(7)	.55617(7)	.041
CL(3)	.44528(3)	.56237(11)	.36506(5)	.044
CL(4)	.44822(4)	.12835(10)	.38269(5)	.048
S	.35837(2)	.19584(6)	.58590(3)	.019
O(1)	.41166(9)	.13183(21)	.66767(13)	.030
O(2)	.29570(9)	.09306(19)	.55521(14)	.029
C(1)	.33035(10)	.40449(23)	.62179(15)	.021
C(2)	.35816(10)	.51668(25)	.55064(16)	.023
C(3)	.40177(10)	.43896(29)	.46038(15)	.026
C(4)	.40437(11)	.26929(29)	.46424(16)	.026

	U11	U22	U33	U12	U13	U23
CL(1)	.0274(3)	.0354(3)	.0317(3)	.0039(2)	.0089(2)	-.0065(2)
CL(2)	.0424(3)	.0152(3)	.0659(4)	-.0014(2)	-.0021(3)	.0046(2)
CL(3)	.0297(3)	.0698(5)	.0335(3)	-.0159(3)	-.0023(2)	.0262(3)
CL(4)	.0493(4)	.0602(5)	.0346(3)	.0157(3)	.0124(3)	-.0156(3)
S	.0228(2)	.0153(3)	.0203(2)	.0021(2)	.0008(2)	-.0001(1)
O(1)	.0351(8)	.0262(7)	.0280(7)	.0070(6)	-.0055(6)	.0044(5)
O(2)	.0323(7)	.0178(7)	.0359(7)	-.0041(6)	-.0019(6)	-.0003(5)
C(1)	.0205(7)	.0186(8)	.0232(7)	.0008(6)	.0018(6)	-.0040(6)
C(2)	.0217(7)	.0168(8)	.0293(8)	-.0012(6)	-.0016(6)	.0022(7)
C(3)	.0199(7)	.0361(11)	.0214(7)	-.0041(7)	-.0011(6)	.0081(7)
C(4)	.0230(8)	.0365(11)	.0196(7)	.0045(7)	.0007(6)	-.0022(7)

Table 3.3 Bond lengths (Å) and angles (°) in (1) at 295 and 150K

Cl(1) - C(1)	1.683(4)	1.691(2)	Cl(2) - C(2)	1.687(4)	1.689(3)
Cl(3) - C(3)	1.699(4)	1.696(3)	Cl(4) - C(4)	1.681(4)	1.682(3)
S - O(1)	1.432(3)	1.435(2)	S - O(2)	1.430(3)	1.435(2)
S - C(1)	1.775(4)	1.771(2)	S - C(4)	1.775(4)	1.775(2)
C(1) - C(2)	1.311(5)	1.326(3)	C(2) - C(3)	1.479(5)	1.477(3)
C(3) - C(4)	1.318(6)	1.334(4)			

O(1) - S - O(2)	119.0(2)	119.1(1)	O(1) - S - C(1)	110.8(2)	111.0(1)
O(1) - S - C(4)	109.2(2)	109.1(1)	O(2) - S - C(1)	110.7(2)	110.4(1)
O(2) - S - C(4)	112.1(2)	111.8(1)	C(1) - S - C(4)	91.7(2)	92.1(1)
Cl(1) - C(1) - S	119.3(2)	119.8(2)	Cl(1) - C(1) - C(2)	130.2(3)	129.8(2)
S - C(1) - C(2)	110.5(3)	110.3(2)	Cl(2) - C(2) - C(1)	124.8(3)	124.7(2)
Cl(2) - C(2) - C(3)	121.6(3)	121.5(2)	C(1) - C(2) - C(3)	113.6(4)	113.7(2)
Cl(3) - C(3) - C(2)	120.4(3)	120.7(2)	Cl(3) - C(3) - C(4)	125.2(3)	125.2(2)
C(2) - C(3) - C(4)	114.4(4)	114.1(2)	Cl(4) - C(4) - S	118.9(3)	119.3(2)
Cl(4) - C(4) - C(3)	131.3(3)	130.9(2)	S - C(4) - C(3)	109.7(3)	109.6(2)

Table 3.4 Refined number of valence electrons, Z, and atomic charge, A-Z, where A is the number of valence electrons of the isolated atom

	Z	A-Z		Z	A-Z
Cl(1)	6.9(1)	+0.1(1)	O(1)	6.2(1)	-0.2(1)
Cl(2)	7.0(1)	0.0(1)	O(2)	6.3(1)	-0.3(1)
Cl(3)	7.1(1)	-0.1(1)	C(1)	4.6(2)	-0.6(2)
Cl(4)	7.1(1)	-0.1(1)	C(2)	3.7(2)	+0.3(2)
S	4.8(3)	+1.2(3)	C(3)	4.2(2)	-0.2(2)
			C(4)	4.1(2)	-0.1(2)

Table 3.5: Summary Of Experimental Measurements On
Tetrachlorothiophene 1,1-Dioxide

Formula $C_4Cl_4O_2S$, F.Wt. = 253.9, monoclinic, space group
C 2/c, Z = 8, F(000) = 992.

T, K	295	150
Crystal size, mm	0.52x1.04x1.04	0.76x0.68x0.34
a, Å	18.424(4)	18.204(5)
b, Å	7.915(1)	7.855(2)
c, Å	11.927(5)	11.751(3)
β , °	90.48(3)	91.77(2)
V, Å ³	1379.2(8)	1679.6(6)
d(calc), gcm ⁻³	1.939	2.008
Refln for cell determination	21	23
Their θ range, °	14-18	12-17
μ (Mo-K α), cm ⁻¹	15.5	16.0
Transmission factors on F ²	-	0.37-0.58
Scan width in ω , °	1.2	1.2+0.53tan θ
Max. counting time, s	90	100
2 θ range	4-60	4-70
Range of h	0-25	0-29
k	$\bar{1}\bar{1}$ -11	0-12
l	$\bar{1}\bar{6}$ -16	$\bar{1}\bar{9}$ -19
Total refln measured	3370	3152
Unique refln	2073	3078
R _{int}	0.058	0.040
Unique refln > 3 σ (I)	1843	2721
Refined parameters*	101	101
R	0.054	0.043
R _w	0.081	0.080
(Δ/σ)max	0.04	0.01
$ \Delta\rho _{\max}$, eÅ ⁻³	0.76	1.22

* 99 positional and U_{ij} parameters plus the scale and the isotropic extinction parameters.

Appendix 1: Single Crystal X-Ray Diffraction Analysis

The main experimental techniques and equations used for the structure analyses discussed earlier are briefly summarised.

Crystal Structure

A crystal is built by translational repetition in three dimensions of the structural motif contained in the unit cell. The geometry of the unit cell is defined by three lengths a , b , and c , which are the axes of the cell, and the three interaxial angles α , β , and γ . A unit cell belongs to one of the seven different crystal systems. This classification is based on the symmetry displayed both by the six cell parameters and by the contents of the cell.

The contents of the unit cell can be generated from the asymmetric unit by applying a series of symmetry operations. This set of symmetry operations defines the space group. The operations are most conveniently specified in terms of the equivalent positions: the fractional coordinates of points symmetrically equivalent to an arbitrary general position, x , y , z , within the cell.

Bragg's Law

The conditions under which all unit cells in a crystal scatter in phase when the crystal is bathed in a uniform beam of monochromatic X-rays were first derived by Laue. Bragg suggested an alternative and conceptually simpler

approach in which the scattering is regarded as arising from stacks of symmetrically equivalent planes defined by integral Miller indices, h , k , l , where the intercepts of any plane on the axes of a reference unit cell are a/h , b/k , and c/l . It can be easily shown that for symmetrical scattering from such a stack of planes the interference condition is

$$2d\sin\theta = \lambda$$

where d is the interplanar spacing, θ the angle of incidence and λ the wavelength of the X-ray beam. Higher orders of diffraction are conventionally regarded as first order diffraction from planes of spacing d/n where n is the order of diffraction for planes of spacing d .

Measurements on Diffractometers

All the structure analyses described in this thesis were performed on an Enraf-Nonius CAD4 diffractometer. This, like other modern diffractometers, allows the crystal to be oriented by rotation about three independent axes and the counter to be swung about the crystal specimen so that a high proportion of all accessible Bragg reflections can be observed.

In a typical experiment 20-25 Bragg reflections were located and their setting angles carefully measured. The unit cell and its orientation were then deduced from the setting angles. If necessary, the reduced cell was derived and then the conventional cell by methods of International Tables for Crystallography, Volume A, Section 9.

A preliminary survey of intensities, out to $\theta = 8^\circ$

typically with molybdenum radiation, was then made. Usually this would cover at least a hemisphere of reciprocal space, include lattice absences and sometimes might be based on doubled unit cell axes. Sorting and averaging these measurements would allow final decisions to be made concerning the Laue group and lattice type.

Final intensity measurements would then be made out to as high an angle as possible using the $\omega/2\theta$ scan method and an ω scan width of $0.8-1.0^\circ$, increased by 25% at each end to allow for background. This set of measurements would normally comprise more than an asymmetric unit of data.

Data Processing

Structure amplitudes were derived from the expression

$$I(hkl) = kLpAF^2(hkl)$$

where k is a scale factor for the experiment, L and p are the Lorentz and polarisation factors allowing respectively for the rapidity with which each reflection passes through the Bragg condition and for angle-dependant polarisation on scattering from the crystal monochromator and specimen. A allows for X-ray absorption and was estimated either by Gaussian quadrature⁸⁷ or by the empirical method of Stuart and Walker.⁵⁹ In cases where crystal decomposition was a problem the mean of the intensity standards was used to apply a linear decay correction to segments of data.

The scale factor, k , and any allowance for extinction were derived by least-squares refinement (see below).

The final structure amplitudes were then sorted and averaged according to the point (rather than Laue) group to

give a unique data set. Unobserved reflections with $I < K\sigma(I)$, with $K = 2-3$, were not subsequently used.

Structure Analysis

The wave scattered from a single unit cell during Bragg reflection from the hkl planes is defined by the structure factor

$$F(hkl) = \sum_{j=1}^N f_j \exp 2\pi i (hx_j + ky_j + lz_j) = |F(hkl)| \exp(i\alpha_{hkl})$$

where the summation is over the N atoms in the unit cell, f_j , x_j , y_j , z_j , are respectively the scattering factor and fractional coordinates of the j th atom and α_{hkl} is the phase of the reflection.

For atoms or ions at rest the scattering factor, f_j , can be calculated from suitable electronic wave functions which assume a spherical distribution of electron density. The scattering factor is therefore real and a function of $\sin\theta/\lambda$ only.

Such calculations ignore the perturbing effect of the atomic nucleus on the scattering process. For heavier atoms this approximation is not acceptable and, when allowance is made for anomalous dispersion, the scattering factor is of the form $f + f' + if''$ where f' and f'' correct the scattering factor, f , calculated for a spherical distribution of classical electrons. In non-centrosymmetric space groups anomalous dispersion leads to inequivalence of $F(hkl)$ and $F(\bar{h}\bar{k}\bar{l})$ and thus implies that the chirality of the crystal structure can be determined.

The scattering factor also contains an allowance for

thermal vibration of the form $\exp(-8\pi^2 U \sin^2 \theta / \lambda^2)$ where U is the mean-square amplitude of vibration normal to the Bragg planes. For anisotropic vibration this temperature factor is defined by six independent U_{ij} parameters in the expression:

$$\exp - 2\pi^2 (h^2 a^{*2} U_{11} + k^2 b^{*2} U_{22} + \dots + 2hka^*b^*U_{12} \dots)$$

where a^* , b^* and c^* are the reciprocal cell parameters.

The electron density at the point with fractional coordinates x , y and z is given by

$$\rho(xyz) = 1/V \sum_{hkl} F(hkl) \exp - 2\pi i (hx + ky + lz)$$

where V is the cell volume. The summation, in principle triply infinite, is over all the Bragg reflections.

Phase Problem

In order to calculate the electron density function both the structure amplitude and the phase angle are required. However, only the structure amplitude is experimentally available. Without phase information it is impossible to work out the relative phases of the individual component waves from each atom. However, if atomic positions are known it is possible to work back and calculate phase angles.

There are two main methods used to overcome the phase problem:

(1) the heavy atom method

and (2) direct methods

(1) Heavy Atom Method

The Patterson function is defined by

$$P(uvw) = \sum_{hkl} |F_{hkl}|^2 \cos 2\pi(hu + kv + lw)$$

where $P(uvw)$ is the value of the function at the position with fractional coordinates u , v and w , the summation is over all Bragg reflections and $|F_{hkl}|$ is the measurable structure amplitude. Atoms at positions (x_1, y_1, z_1) and (x_2, y_2, z_2) give rise to a peak in $P(uvw)$ at $u = x_2 - x_1$, $v = y_2 - y_1$, $w = z_2 - z_1$, and the height of the peak is roughly proportional to $\rho(x_1 y_1 z_1) \times \rho(x_2 y_2 z_2)$. Accordingly, heavy atom positions are easily calculated if the equivalent positions of the unit cell are known. These atomic positions can then be used to calculate approximate phase angles to be used with the observed structure amplitudes, thereby allowing an approximate electron density distribution to be obtained.

(2) Direct Methods

These methods use probabilistic calculations to derive a set of phases from the structure amplitudes using criteria such as the non-negative nature of electron density. Usually it is sufficient to obtain phases for only a few hundred critically selected reflections to obtain an interpretable electron density.

In general, Patterson rather than direct methods were used in the structure analyses described earlier. Sometimes Patterson methods were used to assess solutions suggested by Direct Methods.

Refinement

Once an approximate electron density had been derived the positions of at least some atoms were known. Approximate phases for all structure amplitudes would then be calculated and an improved electron density distribution. This process was repeated till all atoms had been located.

Final parameters (scale, atomic coordinates and U_{ij} values) were obtained by minimising

$$M = \sum_{hkl} \omega (|F^O| - |F^C|)^2$$

by the method of least-squares using weights $\omega = 1/\sigma^2(F^O)$. When allowance for extinction was deemed necessary $|F^C|$ was modified by the factor $[1 + gF^{C2}f(\theta)]^{-1/2}$ where g is the refined extinction parameter and $f(\theta)$ a modified L_p factor.¹³⁸ Refinement was continued till the maximum shift/e.s.d. ratio was typically 0.5 or less. Structure factor agreement was assessed from

$$R = (\sum ||F^O| - |F^C||) / \sum |F^O|$$

and

$$R_w = \{\sum \omega (|F^O| - |F^C|)^2 / \sum \omega F^{O2}\}^{1/2}$$

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X-RAY DIFFRACTION ANALYSIS OF ORGANOMETALLIC
COMPOUNDS WITH CATALYTIC PROPERTIES
VOLUME II

Submitted to the University of Glasgow in partial
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Doctor of Philosophy in the Faculty of Science

by

Graeme Douglas

Chemistry Department
September 1990
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APPENDIX 2

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CALCULATED AND OBSERVED STRUCTURE FACTORS OF	PAGE
1. $[\text{Pt}_3(\text{CO})(\text{CNC}_6\text{H}_{11})(\text{dppm})_3][\text{PF}_6]_2$	1
2. $[\text{Pt}_3(\text{CO})(\text{S}_2\text{CNMe}_2)(\text{dppm})_3][\text{PF}_6]$	29
3. $[\text{Pt}_3(\text{CO})(\text{SnF}_3)(\text{dppm})_3]_{0.75}$ $[\text{Pt}_3\text{Cl}(\text{SnF}_3)(\text{dppm})_3]_{0.25}[\text{PF}_6]_{0.25}$	60
4. $[\text{Pt}_3(\text{SnF}_3)_2(\text{dppm})_3]$	66
5. $[\text{CpMo}(\text{Pr}^i\text{SC}(\text{CO}_2\text{Me})\text{C}(\text{CO}_2\text{Me})=\text{C}(\text{CF}_3)\text{C}(\text{CF}_3)=\text{C}(\text{CF}_3)\text{CCF}_3)]$	99
6. $[\text{CpMo}(\text{F}_3\text{CCC}(\text{CF}_3)\text{C}(\text{CF}_3)=\text{C}(\text{CO}_2\text{Me})\text{C}(\text{SPr}^i)\text{C}(\text{OMe})=\text{O})]$	121
7. $[\text{CpMoI}(\text{CO})(\text{C}_4\text{Ph}_4)] \cdot \text{CH}_2\text{Cl}_2$	135
8. $[\text{WBr}(\text{CO})_2(\text{SC}_6\text{F}_5)(\text{C}_7\text{H}_8)]$	142
9. $[\text{Ru}(\text{SC}(\text{NH}_2)_2)_6][\text{CF}_3\text{SO}_3]_2$	150
10. $\text{C}_4\text{Cl}_4\text{O}_2\text{S}$ at 295K	167
11. $\text{C}_4\text{Cl}_4\text{O}_2\text{S}$ at 150K	176

CALCULATED AND OBSERVED STRUCTURE FACTORS OF
 $[\text{Pt}_3(\text{CO})(\text{CNC}_6\text{H}_{11})(\text{dppm})_3][\text{PF}_6]_2$

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5	2152	1714	15	738	959	-25	908	994	-19	560	614	-23	774	-22	659
6	2121	1902	16	779	987	-23	814	959	-14	637	583	-22	882	-21	761
10	548	799	17	1259	1157	-16	1265	1160	-13	1184	1123	-21	995	-20	785
12	1619	1617	19	1358	1262	-15	1108	1175	-12	1068	1130	-20	790	-19	643
14	1531	1615	20	864	901	-15	1108	1175	-11	1382	1351	-19	619	-18	621
16	2392	2263	21	613	537	-14	1668	1483	-10	1263	1478	-18	701	-17	636
17	661	1001	O, 2,			-13	1747	1641	-9	1177	1411	-17	790	-16	560
18	1488	1588	-19	776	878	-12	1601	1462	-8	1692	1538	-12	649	-15	502
19	931	1041	-17	1566	1495	-11	2033	2099	-7	1125	1086	-11	1003	-14	475
20	840	900	-16	1303	1290	-10	731	826	-6	1884	1973	-10	811	-13	399
21	364	805	-15	1833	1757	-9	2404	2552	-5	639	468	-9	1468	-12	569
O, 1,			-14	2435	2107	-7	2370	2703	-4	810	978	-8	940	-10	510
-20	1065	1213	-13	1448	1525	-6	810	460	-3	659	705	-7	1286	-8	904
-19	543	947	-12	2488	2647	-5	1838	1679	-2	1096	1029	-5	1354	-7	560
-18	1111	1439	-10	2384	2908	-4	639	581	-1	719	727	-5	2100	-6	1308
-17	1572	1578	-9	540	712	-2	1546	1134	0	1084	284	-4	784	-5	635
-16	345	1450	-8	2205	2158	-1	319	577	1	1600	1155	-3	938	-4	1484
-15	2547	2477	-7	1090	1202	0	340	200	2	603	404	-2	1180	-3	364
-13	2329	2393	-6	1082	1391	1	1525	1505	3	729	773	-1	720	-1	991
-11	2686	2834	-5	1758	1741	4	589	457	4	802	552	0	1251	0	468
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-4	1557	1785	0	1515	1967	11	547	703	10	996	893	5	347	6	1453
-3	1359	1621	1	222	547	12	625	551	12	513	797	6	1393	7	1610
-2	354	749	2	1512	1856	13	704	707	14	723	746	8	1099	8	863
2	581	468	3	898	1186	14	886	990	15	660	670	10	394	9	1353
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5	3852	3664	6	2915	2630	18	758	815	18	1194	1035	17	550	14	494
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13	805	798	-21	583	676	-21	583	676				24	881	23	1386
			-20	891	989	-20	891	989				25	806	24	533

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-8 430	-7 592	5 745	619	-7 592	575
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2 1801	2 1155	-2 1122	984	2 1155	1217
3 467	3 952	-1 912	829	3 952	970
4 1800	4 1183	0 943	841	4 1183	1212
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-3 846	-8 939	0 1159	970	-8 939	931
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14 981	1139	18 1018	911	-13 1343	966		0	3404	3609	1	2693	2887	3	570	567
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5 1037	1201	-21 906	888	-8 1936	1551	4	1482	1791	9	1512	1421	14 1106	1277
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13 1926	19 611	626	-14 1137	385	5 719	774	3 406	508	9 1135	1242
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2 1183	1118	17 891	18 901	933	933	-1 421	421	74	5 1585	1677	-19	780	809
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8 939 991	503	1 474 430	5 961 1002	4 550 185	22 1083 1156
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11 655 723	1638	3 492 517	9 1518 1570	6 716 890	-17 674 821
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6 745	701	-5 1336	1303	9, 14,	L	-8 1191	1236	-13 575	555	13 507	617
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21 535		8 1331	8 1123	6 824	953	-2 745	705

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5 573	945 925	-6 482	477	18	857	844	9	671	717	-10 1177	1097
9 922	14 996 1100	-5 803	719	19	582	480	18	565	519	-8 1037	951
10 752 824	15 955 886	-3 1240 1312	1312	11, 9,	L		11, 12,		L	-6 836	759
11 782 811	11, 5,	-1 1667 1732	1732	-11 1601 1301			-6 754 908		9 619	557	
11, 1,	-7 805 810	0 451 91	91	-9 1325 1209			-4 938 921		12, 1,	L	
-10 864 917	-5 1416 1325	1 1924 1705	1705	-7 910 882			-2 716 702		-7 904 774		
-8 1266 1256	-3 1365 1302	2 1171 953	953	-1 793 812			7 619 538		12, 2,	L	
-6 1417 1392	-1 1088 1069	3 1382 1248	1248	1 1202 1250			11, 13,		4 681 675		
-4 1413 1317	0 430 344	4 1352 1264	1264	2 640 512			6 656 576		5 693 681		
-2 1019 950	1 739 804	5 933 939	939	3 1361 1331			8 660 877		12, 3,	L	
12 565 636	7 784 725	6 982 968	968	4 648 263			10 692 834		-6 615 463		
11, 2,	9 1417 1430	7 413 562	562	5 1625 1617			11, 14,		4 826 766		
-7 906 937	10 625 476	8 746 728	728	6 512 347			5 642 516		5 833 777		
-5 1284 1200	11 1676 1789	11 788 722	722	7 1438 1515			7 775 779		6 1159 1145		
-3 1327 1257	12 959 976	13 1116 1080	1080	8 574 553			9 812 947		7 1011 1024		
-1 796 915	13 1565 1717	15 1347 1318	1318	9 1414 1478			12, -3,		8 872 1066		
11, 3,	14 1245 1199	16 770 554	554	10 884 874			-3 658 695		9 886 1007		
-14 830 868	15 1285 1190	17 1338 1317	1317	11 812 910			-1 775 757		10 706 871		
-4 754 624	16 1192 987	18 746 740	740	11, 10,	L		0 648 670		13 591 610		
0 749 794	17 543 549	11, 8,	L	-10 1474 1384			1 727 869		12, 4,	L	
1 590 341	18 644 618	-12 1128 903	903	-8 1353 1396			2 829 878		-7 1058 909		
2 392 384	11, 5,	-10 837 776	776	-6 1225 1150			4 951 950		-5 868 797		
7 886 809	-6 949 915	-5 476 390	390	1 819 691			5 643 708		-3 532 546		
8 510 643	-4 1353 1281	-3 483 530	530	2 1087 1049			12, -2,		-1 401 309		
9 897 924	-2 1663 1745	-2 932 997	997	3 868 407			-2 650 634		5 1085 983		
10 707 847	0 1592 1512	0 1248 1288	1288	4 1186 1184			0 662 691		6 746 498		
12 681 521	1 947 808	2 1599 1579	1579	5 635 102			1 711 689		7 1599 1608		
13 636 906	2 1135 916	3 604 275	275	6 1064 1167			2 781 831		8 1128 1106		
11, 4,	3 893 865	4 1623 1636	1636	8 1083 1108			3 1008 1084		9 1368 1474		
-6 1038 1045	5 675 646	5 906 925	925	10 961 1002			5 1121 961		10 1243 1201		
-4 814 829	10 1248 1141	6 1353 1362	1362	11, 11,	L		6 796 764		11 1098 1265		
-2 657 612	12 1445 1560	7 892 943	943	-7 1347 1271			7 909 814		12 1144 1058		
1 464 385	13 564 632	8 1068 1203	1203	-5 1275 1263			12, -1,		13 943 973		
6 683 622	14 1633 1693	9 1033 1058	1058	-3 1060 994			3 593 600		14 1114 929		
8 1265 1311	15 1181 939	11 701 515	515	-1 487 466			4 800 821		16 592 636		
10 1393 1521	16 1374 1459	15 776 186	186	3 620 561			6 979				
11 1077 1120	18 1001 958	16 855 894	894	5 599 761							

-10	593	L	12, 5,	12, 7,	L	12, 10,	L	7	988	967	0	569	592	5	851	863
-8	973	702	-6	864	863	-7	1007	8	1117	1069	1	749	873	9	418	348
-6	1200	1038	-4	1123	1249	1	907	13,	4,	L	3	718	746	13,	10,	L
-4	1191	1564	-2	1514	1564	2	528	-8	749	910	6	573	456	2	503	471
-4	1156	1122	0	1397	1297	3	912	2	692	675	8	777	790	4	518	453
-2	976	819	1	942	894	4	525	4	1187	1176	9	652	612	6	517	593
-1	728	781	2	1333	1100	5	765	5	1128	938	10	963	994	13,	11,	L
6	1122	1051	3	1318	1254	7	792	6	1455	1481	11	966	853	0	467	461
7	532	252	4	982	870	9	775	7	1254	1212	12	957	937	13,	12,	L
8	1537	1562	5	1027	921	16	658	8	1192	1299	13	902	855	1	536	628
9	908	780	6	476	614	18	618	9	1255	1233	14	787	604	2	545	259
10	1634	1661	12	610	872	12,	11,	10	982	896	15	714	661	3	657	842
11	1200	1063	14	1029	1175	-4	632	11	1068	1024	13,	7,	L	4	528	504
12	1417	1469	15	805	797	3	486	13,	5,	L	-5	1084	1219	14,	1,	L
13	1120	1073	16	564	665	5	577	-9	1007	988	-3	1080	1170	1	651	798
14	1101	1057	12,	5,	L	15	573	-7	959	970	-1	980	995	14,	2,	L
15	1199	948	-5	519	593	12,	12,	-5	851	899	0	737	786	1	803	895
-7	982	954	-3	798	948	2	414	-4	922	886	1	775	768	2	1149	1032
-5	1344	1346	-1	1245	1267	4	520	-3	629	603	2	832	834	4	1047	934
-3	1391	1410	1	1259	1239	6	577	-2	799	852	3	675	687	6	810	721
-2	670	667	2	619	475	12,	13,	0	669	837	4	753	720	14,	3,	L
-1	1145	1150	3	1257	1276	1	466	2	536	597	12	750	676	0	761	833
0	936	859	4	765	735	5	740	5	945	982	13	570	608	1	747	793
1	841	799	5	1192	1334	7	798	6	712	463	13,	8,	L	2	1009	1038
2	970	1060	6	832	887	13,	-1,	7	1384	1544	-6	539	637	3	1183	1094
3	527	283	7	726	781	1	475	8	1134	1012	-4	920	984	4	1131	1125
4	804	764	16	687	612	13,	2,	9	1337	1411	-2	1027	1088	5	1244	1151
6	490	444	12,	9,	L	1	680	10	1163	1139	0	1142	1063	6	898	855
7	743	559	-8	798	712	2	763	11	1063	1138	1	779	662	7	1102	1035
9	1118	1114	-5	480	424	3	906	12	1191	1042	2	1094	1072	9	788	767
10	708	522	-2	713	793	-4	810	14	824	758	3	958	746	14,	4,	L
11	1326	1392	0	1002	933	5	997	13,	6,	L	4	958	942	-4	652	733
12	824	798	2	981	944	13,	3,	-8	1176	1143	5	686	543	1	632	677
13	1388	1461	2	610	582	2	610	-6	1141	1165	13,	9,	L	2	647	602
14	1172	1018	3	1136	1095	3	1136	-5	598	526	-3	669	701	3	971	970
15	1126	1071	4	1241	1118	4	1241	-4	1049	1046	-1	834	740	4	1152	943
16	1059	886	5	1273	1194	5	1273	-2	842	893	1	882	756	5	1112	1141
			6	1265	1170	6	1265	-1	848	950	3	918	791	6	1245	1043

3 705 711
12 484 386
14, 9, L
2 654 631
4 609 606
11 499 484

14, 4, L
7 1040 992
8 1172 1049
10 891 802
14, 5, L
-5 866 932
-3 784 939
-1 451 739
4 800 842
5 978 837
6 881 963
7 1082 932
8 866 903
9 1007 925
10 610 555
11 842 730
14, 6, L
-4 739 853
-2 836 895
0 776 877
5 494 493
6 666 580
7 559 569
8 947 724
9 577 613
10 779 749
14, 7, L
-4 796 821
-2 708 771
-1 707 734
1 791 766
3 617 505
9 498 580
14, 8, L
-1 792 779
0 621 526
1 825 775
2 672 452

CALCULATED AND OBSERVED STRUCTURE FACTORS OF $[\text{Pt}_3(\text{CO})(\text{S}_2\text{CNMe}_2)(\text{dppm})_3][\text{PF}_6]$

0, 0, L	5 1967	1827	22 3126	3022	0, 5, L	13 2228	2166	5 1381	1432
4 5374 5676	6 1511 1523	24 1218 1156	24 1218 1156	24 1218 1156	1 1015 1042	14 691	717	6 1648	1705
6 4490 4591	7 1478 1480	25 843 931	25 843 931	25 843 931	2 2982 2945	16 1379	1365	7 2802	2825
8 2032 1991	8 5023 5094	26 2485 2335	26 2485 2335	26 2485 2335	3 1704 1641	18 420	442	8 701	585
10 1905 1894	9 2568 2570	28 1291 1266	28 1291 1266	28 1291 1266	4 1564 1549	19 2432	2405	9 2050	2130
12 4676 4601	10 2084 2040	29 898 915	29 898 915	29 898 915	5 2518 2613	20 2280	2353	10 1633	1625
14 455 453	11 2338 2433	0, 4, L	0, 4, L	0, 4, L	6 3837 3904	21 1794	1806	11 1520	1481
16 5589 5440	12 3886 4069	0 806 678	0 806 678	0 806 678	7 3258 3274	22 570	654	12 1374	1353
18 989 837	14 1298 1315	1 1831 1857	1 1831 1857	1 1831 1857	9 2092 2166	23 1933	1950	13 1467	1515
20 3373 3198	15 1373 1433	2 1532 1560	2 1532 1560	2 1532 1560	10 3054 3054	24 1476	1464	14 1128	1070
22 1578 1548	16 4164 4243	3 1552 1504	3 1552 1504	3 1552 1504	11 1882 1928	25 1975	1915	15 1405	1352
24 1537 1492	17 2661 2612	4 3032 3042	4 3032 3042	4 3032 3042	12 2237 2224	0, 7, L	L	16 2292	2264
28 3460 3049	20 2660 2655	5 4061 4058	5 4061 4058	5 4061 4058	13 1100 1183	1 2021	2075	17 1771	1621
0, 1, L	21 1298 1259	6 1663 1575	6 1663 1575	6 1663 1575	14 981 932	2 1523	1538	18 1430	1389
2 7956 8304	24 2479 2407	7 1313 1263	7 1313 1263	7 1313 1263	15 522 507	3 442	533	19 1360	1313
3 4064 3916	25 880 879	8 5616 5624	8 5616 5624	8 5616 5624	17 1022 1049	4 1621	1572	21 1406	1427
4 1837 1725	26 743 800	9 2421 2369	9 2421 2369	9 2421 2369	18 1746 1781	5 4008	4086	23 1312	1298
5 469 434	27 1413 1436	10 1675 1760	10 1675 1760	10 1675 1760	19 626 605	6 821	782	24 522	512
6 3976 3904	28 1851 1986	11 1774 1921	11 1774 1921	11 1774 1921	20 560 585	7 2875	2966	0, 9, L	L
7 991 936	0, 3, L	12 2563 2653	12 2563 2653	12 2563 2653	21 2354 2280	9 2848	2945	1 2047	2078
8 325 338	1 3190 3141	13 1241 1195	13 1241 1195	13 1241 1195	22 2805 2794	10 501	643	2 976	911
10 3380 3449	2 5612 5574	14 1064 1080	14 1064 1080	14 1064 1080	23 1760 1699	11 2746	2764	3 2571	2596
11 742 756	3 5226 5131	15 1648 1695	15 1648 1695	15 1648 1695	24 752 757	12 1391	1409	4 1576	1535
12 488 462	4 406 306	16 352 406	16 352 406	16 352 406	25 975 945	14 1716	1774	5 784	772
13 762 860	6 6364 6440	17 1574 1673	17 1574 1673	17 1574 1673	26 1483 1463	15 913	980	6 406	328
14 4444 4520	7 2210 2101	18 1461 1448	18 1461 1448	18 1461 1448	27 1372 1314	17 1330	1308	7 1368	1371
15 786 955	8 2488 2514	19 651 767	19 651 767	19 651 767	0, 6, L	18 1552	1594	8 664	618
16 381 364	9 1670 1645	20 2630 2621	20 2630 2621	20 2630 2621	0 3077 3034	19 1666	1700	9 1870	1889
18 4464 4698	10 5548 5548	21 914 936	21 914 936	21 914 936	1 1390 1366	20 1167	1202	11 2014	1947
19 2045 2172	11 642 727	22 671 625	22 671 625	22 671 625	2 294 216	21 2132	2104	12 1362	1279
20 1028 1065	12 1228 1291	23 1013 1105	23 1013 1105	23 1013 1105	3 1697 1826	23 1808	1829	13 1962	1851
22 1875 1804	13 2649 2747	24 3319 3218	24 3319 3218	24 3319 3218	4 2171 2179	25 1010	927	14 1576	1502
26 2536 2353	14 2393 2536	25 807 971	25 807 971	25 807 971	5 1822 1863	0, 8, L	L	15 2415	2351
0, 2, L	15 1630 1620	26 1621 1601	26 1621 1601	26 1621 1601	6 418 390	0 2111	2151	16 519	605
0 4837 5192	17 892 977	28 1102 1048	28 1102 1048	28 1102 1048	7 3846 3985	1 1663	1650	17 1397	1347
1 4054 3922	18 2340 2321				8 2768 2719	2 1842	1826	18 772	699
2 568 541	19 1839 1858				9 3666 3754	3 2053	2185	19 1509	1446
4 4821 4705	20 1155 1153				12 1876 1901	4 865	832	20 983	943

0, 9, L	6 625 672	10 1986 1932	7 998 1125	-2 6289 6280	-8 6013 6043
23 528 577	7 1067 1025	12 367 200	8 1144 1108	-1 3248 3109	-7 1266 1367
0, 10, L	9 620 614	14 3947 3980	9 1904 1945	0 1207 1214	-6 1090 1067
1 3027 3015	10 979 914	16 1163 1067	10 457 449	1 1656 1600	-5 4705 4727
2 618 497	11 1326 1344	18 4513 4273	11 932 1061	2 3980 3823	-4 1208 1175
3 764 769	12 1043 990	20 1405 1346	12 4455 4485	3 4873 4667	-3 1110 1036
5 1842 1864	13 1160 1187	22 1112 1041	13 702 682	4 1453 1384	-2 3403 3284
6 344 309	14 551 636	24 753 759	16 5060 4960	5 2328 2203	-1 1257 1185
7 520 481	15 969 865	26 2331 2055	17 1805 2068	6 5134 5205	0 1282 1280
8 758 691	16 1174 1189	28 886 769	18 1341 1414	7 1913 1904	1 3442 3433
9 771 694	0, 13, L	1, 1, L	20 2909 2745	8 2150 2185	2 3124 2988
10 434 459	1 743 740	-28 2647 2515	21 949 1320	9 733 776	3 395 313
11 2383 2427	2 1201 1214	-27 941 1066	22 1268 1245	10 3605 3670	4 4892 4846
12 475 459	4 613 629	-26 569 661	24 1210 1377	11 534 553	5 2217 2256
13 1533 1545	5 1160 1154	-25 841 765	25 1080 1059	12 1207 1238	6 424 355
15 2255 2211	7 1075 1095	-24 1539 1371	28 2768 2670	13 1944 1980	8 5817 5959
17 2396 2407	9 845 877	-22 980 989	1, 2, L	14 4159 4271	10 2521 2503
21 800 799	11 869 909	-20 2815 2694	-29 1220 1303	15 2050 2152	11 2199 2189
0, 11, L	12 1115 1100	-18 832 783	-26 2428 2362	16 373 391	12 2288 2428
1 1694 1687	0, 14, L	-17 1400 1418	-25 871 987	18 2746 2782	13 760 853
2 919 911	3 1006 1018	-16 4906 4883	-22 1813 1944	19 2824 2794	15 1087 1076
3 1842 1866	4 697 623	-15 524 613	-21 819 817	21 765 934	16 1506 1657
4 903 882	1, 0, L	-14 794 852	-19 2373 2488	22 2689 2500	17 2707 2701
5 472 452	-28 1082 1006	-12 4359 4418	-18 2755 2725	24 642 685	18 1219 1299
7 592 547	-26 2040 1889	-11 1368 1334	-17 828 839	26 2345 2201	20 2215 2241
9 1022 1066	-22 1703 1708	-10 1458 1457	-16 385 424	29 1124 1075	21 1464 1437
12 711 700	-20 1348 1298	-9 1829 1873	-15 1751 1813	1, 3, L	22 569 533
13 2287 2272	-18 4906 4694	-8 2223 2413	-14 4114 4247	-24 2848 2725	24 2762 2704
14 995 933	-16 621 509	-7 833 829	-13 1812 1927	-21 638 771	26 1404 1355
15 1897 1903	-14 4457 4551	-6 757 762	-11 907 925	-20 2128 2169	27 646 790
17 930 904	-12 2203 2252	-5 940 1043	-10 3456 3573	-18 1103 1214	28 1657 1632
19 1466 1530	-10 2270 2228	-4 4468 4434	-9 1985 1980	-17 2089 2156	1, 4, L
0, 12, L	-8 1491 1533	-3 2266 2196	-8 1152 1134	-16 1480 1580	-28 982 1064
0 1794 1795	-6 3251 3224	-2 642 608	-7 621 606	-15 2351 2467	-26 1763 1745
1 1024 1049	-4 1633 1517	2 1625 1533	-6 4898 4875	-13 534 471	-25 1295 1255
2 677 672	4 679 676	3 2724 2713	-5 1848 1810	-12 3550 3641	-24 1221 1167
3 1159 1188	6 1151 1152	4 5393 5472	-4 1469 1460	-11 992 1119	-22 3265 3201
5 833 807	8 804 778	6 2606 2605	-3 3745 3616	-10 364 314	-21 1214 1280

1, 4, L	22 3246	3293	8 4875	4863	-1 2178	2274	-6 1541	1566	-12 1913	1796
-20 1147	23 1133	1208	9 3609	3672	0 816	885	-5 990	899	-11 1758	1693
-18 1653	24 836	776	10 1827	1889	2 3198	3219	-4 1191	1131	-9 2368	2290
-17 1396	26 2080	2096	11 1249	1330	3 1076	1084	-3 3806	3806	-8 599	531
-15 1152	27 782	831	12 2160	2168	4 389	368	-2 2389	2445	-7 1861	1808
-14 1690	28 1358	1206	13 2355	2354	5 4094	4147	-1 693	670	-6 495	487
-13 2397	1, 5, L		14 1643	1753	6 2513	2603	0 1978	2104	-5 2952	3053
-12 2731	-26 898	1043	15 747	832	7 2963	3090	1 749	800	-4 3230	3210
-11 1189	-25 1349	1415	17 1314	1401	9 2598	2693	3 1686	1764	-3 1333	1296
-10 4467	-24 2536	2510	18 979	989	10 1681	1813	4 2317	2268	-2 1942	1876
-9 2853	-23 2091	2154	19 912	943	11 3186	3088	5 1507	1651	-1 1853	1836
-7 2180	-21 587	683	20 2425	2396	13 401	395	6 1236	1208	0 940	989
-6 5286	-20 2602	2634	21 1859	1881	17 914	929	7 3021	3117	1 2016	2093
-5 965	-19 1951	1945	23 1230	1220	18 1710	1797	8 638	540	2 1619	1607
-4 2004	-18 908	877	24 2297	2218	19 1423	1509	9 3265	3316	3 2069	1966
-3 1229	-16 601	576	25 1759	1761	21 2180	2165	10 975	864	4 1533	1526
-2 2585	-15 707	733	1, 6, L		22 1768	1900	11 573	618	5 2142	2097
-1 842	-14 1524	1473	-26 836	941	23 2437	2474	12 1847	1782	6 459	257
1 1167	-13 2195	2235	-25 1269	1254	25 608	556	13 1899	1919	7 2922	2851
2 2771	-12 1573	1548	-23 1841	1807	26 832	857	16 1860	1808	8 853	820
3 1858	-11 1571	1486	-22 1489	1492	1, 7, L		17 444	425	9 1544	1528
4 2161	-10 853	814	-21 2885	2903	-25 1360	1346	18 703	713	11 2881	2888
5 336	-9 3295	3130	-20 920	952	-23 1843	1782	19 1944	2017	12 2012	2005
6 6048	-8 3888	3963	-18 1667	1579	-22 1316	1310	20 1016	1100	13 1191	1216
7 2689	-7 4047	4103	-17 1281	1326	-21 1211	1115	21 2168	2295	14 1900	1984
8 1333	-6 1238	1117	-15 652	651	-20 945	934	23 1496	1527	15 2044	2091
9 1192	-5 2196	2258	-14 985	943	-19 2363	2456	25 2020	2046	17 508	514
10 4945	-4 1788	1752	-12 454	434	-18 1812	1824	1, 8, L		18 991	1013
11 1167	-3 1934	1955	-11 2512	2505	-17 752	718	-23 864	802	19 1659	1649
12 2180	-2 1687	1563	-10 1265	1258	-16 2581	2444	-21 1178	1230	20 621	583
13 2426	0 1973	1908	-9 2947	2983	-14 675	632	-20 1839	1785	21 1057	1039
14 861	1 754	743	-8 473	359	-13 1359	1292	-19 1157	1190	22 467	420
15 1522	2 1661	1563	-7 2304	2265	-12 1137	1169	-18 997	984	23 1429	1478
17 974	3 2220	2160	-6 978	996	-11 927	956	-17 1612	1582	24 911	904
18 1517	4 4352	4385	-5 4181	4144	-10 571	633	-16 947	868	1, 9, L	
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7 2672	4 1233	1231	6 646	634	12 728	692	-14 1758	11 653	650
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19 1125	15 445	492	17 618	597	-13 1477	1529	1 1405	-28 970	919
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21 1296	17 863	837	19 1364	1359	-8 1023	1034	3 880	-24 952	893
23 2336	21 1401	1480	2, 10, L		-6 760	717	5 1118	-22 564	500
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-23 1024	-19 1178	1172	-16 831	790	-2 478	392	8 981	-16 1090	1117
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-16 1953	-15 1383	1227	-11 1891	1859	3 1803	1813	15 837	-8 426	408
-15 1587	-14 2222	2041	-10 887	859	4 2121	2068	16 777	-6 407	102
-12 514	-13 1472	1426	-9 688	722	6 809	738	2, 13, L	-4 498	436
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-9 718	-8 860	899	-6 1415	1480	9 531	576	-8 876	2 2848	2726
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-6 3099	-5 1502	1388	-2 1585	1604	13 1795	1817	-5 888	10 1371	1443
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-3 2058	-2 1955	2012	1 2540	2580	16 872	864	0 1843	16 1320	1294
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-26 1261	17 1196	1282	10 1340	1344	6 1377	1401	-1 826	757	-3 1721	1798
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-24 445	19 564	602	12 614	622	8 3856	3911	3 1617	1477	-1 3287	3200
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-19 620	23 483	499	14 1891	2218	10 739	835	5 1369	1367	1 416	530
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-16 2484	3, 2, L		17 665	599	13 1220	1241	8 927	825	5 1567	1496
-15 1564	-29 1127	1088	18 885	868	14 740	796	10 1761	1851	6 2448	2409
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-10 1695	-19 1605	1553	26 914	900	20 1203	1267	16 619	681	11 1643	1639
-9 2301	-18 388	470	29 785	815	21 531	526	17 450	502	13 902	857
-8 757	-17 782	747	3, 3, L		22 705	727	18 799	750	15 915	933
-7 1994	-16 455	398	-27 1289	1265	23 675	653	19 912	877	17 472	413
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1 840	1233	-3 1270	1186	-2 1120	1151	-2 2414	2317	6 538	454	15 908	938	
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5 1279	1232	1 1266	1270	2 1290	1235	4 1077	1022	10 548	484	19 878	878	
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15 999	1010	11 2685	2672	12 2647	2604	14 1412	1392	4, 10, L		-8 1401	1373	
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18 472	533	15 1173	1129	15 685	687	22 478	510	-17 1149	1111	-4 668	626	
19 485	553	16 661	611	18 1062	1038	4, 9, L		-16 718	646	-3 1502	1557	
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-22 529	594	-24 815	775	4, 8, L		-14 1544	1489	-6 1647	1757	8 1562	1539	
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-18 1169	1222	-20 2147	2206	-21 444	549	-10 704	671	-1 1691	1704	12 1023	1020	
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-8 1919 1882	3 1480 1494	-12 2667 2767	-23 1046 1006	18 1135	1100	11 932	993
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-5 574 619	-26 1894 1842	-10 692 731	-21 1900 1899	20 517	560	14 363	368
-4 678 663	-24 636 758	-9 1587 1584	-20 693 740	21 1722	1713	16 2345	2326
-3 745 677	-22 2655 2673	-8 3164 3165	-19 489 359	23 1362	1266	18 2001	1900
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-1 813 813	-16 1116 1041	-6 2913 2891	-17 465 601	26 1053	1062	21 930	1008
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4 694 697	-8 2902 2929	-2 2318 2259	-13 3282 3274	-26 660	652	25 882	869
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-10 766 795	18 1546 1639	10 2775 2856	-2 1055 1053	-11 651	591	-14 2594	2531
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-7 916 912	24 1822 1793	12 1094 1135	0 1268 1285	-9 1898	2014	-12 3329	3409
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-3 594 613	28 844 854	15 808 887	2 2080 2001	-7 2097	2016	-10 1096	1082
-1 822 849	5, 1, L	17 810 793	3 1527 1477	-6 1079	1035	-9 1081	1100
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4 2397 2512	-23 764 676	21 1214 1273	7 1576 1540	0 2539	2500	-5 995	1041
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5, 4, L	2 3328	3111	7 909	885	10 2864	2797	6 1893	1887	8 774	681
2 3215	3027	246	8 1914	1926	11 2036	2100	7 479	402	9 1719	1651
3 561	493	1712	9 1512	1568	12 942	973	8 3105	3046	10 2615	2598
4 4326	4145	1684	10 529	511	13 1078	1120	9 1055	1009	12 1138	1165
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6 1158	1199	2379	13 2501	2494	15 835	746	13 855	873	16 1953	1936
7 977	1109	1412	14 504	557	16 1197	1200	14 2009	2085	19 1093	1067
8 1125	1026	1039	15 1934	1948	17 1954	2028	15 949	946	5, 10, L	
9 1562	1614	2284	16 1069	1114	21 624	564	17 744	831	-20 1030	986
12 663	756	1009	17 479	388	22 686	684	18 1070	1021	-16 1193	1243
13 1577	1599	1510	19 1830	1877	24 613	652	19 1430	1434	-14 684	705
14 2770	2748	2001	20 1130	1135	5, 8, L		22 805	767	-12 1653	1587
16 1554	1586	832	25 819	816	-22 838	877	5, 9, L		-11 1589	1566
17 427	677	L	5, 7, L		-21 676	805	-22 1320	1282	-10 658	640
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22 440	809	2263	-21 1427	1418	-18 758	726	-18 1772	1794	-7 1570	1554
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-17 1984	1967	804	-10 1213	1247	-13 930	924	-12 732	680	0 1050	1059
-16 2214	2166	2045	-9 330	449	-10 734	818	-11 877	805	2 946	937
-15 1298	1314	1230	-8 732	745	-9 758	714	-10 947	999	3 660	587
-14 2158	2113	373	-6 2250	2216	-8 1810	1929	-9 1972	1962	4 897	825
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-12 1073	1044	1191	-3 806	753	-6 719	768	-7 1459	1445	6 519	493
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5, 11, L	7 1038	1001	18 593	594	9 885	920	-6 1897	1823	-21 1656	1648
-17 884	9 567	613	20 2695	2665	10 2935	3056	-5 926	983	-18 2596	2645
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-12 756	13 625	640	24 2228	2222	12 4668	4706	-3 1809	1798	-15 702	785
-11 829	5, 13, L	L	26 1885	1980	13 2593	2621	-2 1846	1815	-14 4128	4265
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-1 674	3 503	567	-18 1360	1347	23 1531	1515	7 2405	2537	-3 670	659
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11 845	-20 2076	2127	-7 277	358	-23 1681	1652	17 758	671	6 1985	1951
13 751	-18 1999	2060	-6 4051	4076	-22 646	607	18 1734	1743	7 3154	3092
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-12 2260	-6 5982	6062	-2 1546	1546	-17 534	620	22 798	733	11 1938	1945
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-9 814	-2 3025	2881	0 1779	1803	-15 1625	1654	24 1127	1123	15 380	405
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4 2153	14 3590	3611	7 1488	1477	-8 2517	2503	-25 889	812	24 774	700
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6, 3, L	18 2845	2841	14 1699	1714	14 1374	1326	11 786	727	9 1432	1394
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-26 1326 1272	6, 5, L		18 1180	1202	18 1323	1241	14 805	861	12 640	610
-23 989 964	-26 501	565	19 1701	1851	20 565	632	15 2242	2230	13 1387	1434
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-19 864 780	-24 1066	1084	26 642	663	22 710	771	19 2410	2314	16 1394	1444
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-16 3709 3679	-19 1555	1506	-22 785	837	6, 7, L		24 985	1048	20 592	587
-15 1476 1487	-18 1894	1967	-21 1347	1409	-24 1081	1078	6, 8, L		21 1290	1210
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-3 1279 1296	-6 470	431	-6 2337	2321	-8 1510	1476	-10 653	625	-9 1423	1526
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6 2328 2306	2 2977	2885	2 2200	2050	0 610	528	-2 1959	1833	1 586	646
7 737 745	3 2873	2867	4 1198	1100	1 1832	1879	-1 1160	1262	2 1511	1502
9 2560 2580	4 2971	2974	5 1966	1963	2 467	376	0 2289	2205	3 1809	1870
11 1634 1674	7 2278	2291	6 1646	1650	3 3539	3503	1 3002	2967	5 994	1033
12 3320 3298	8 381	396	7 349	338	4 785	671	3 726	742	6 918	916
13 1154 1181	9 1509	1480	9 620	514	6 1409	1375	4 721	759	7 2128	2197
14 845 944	10 821	857	10 1150	1183	7 797	768	5 2229	2242	8 1827	1853
15 1151 1275	11 1302	1328	11 2711	2696	8 2063	2111	6 1055	1106	10 459	515
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6, 9, L	6, 11, L	6, 13, L	7, 1, L	14 3174	3244	2 3354	3400
14 1446 1471	-17 831 769	-6 1594 1569	-26 615 763	15 1576 1474		3 551 497	
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-9 1038	-15 630	711	13 802	804	12 673	637	22 455	513	-17 1051
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-7 942	-12 1145	1194	11, 11, L		16 1925	1911	-24 701	668	-15 964
-6 399	-10 1315	1339	-9 1051	1097	18 1626	1648	-23 504	519	-14 566
-4 578	-8 1897	1824	-7 711	727	22 378	414	-22 793	756	-13 1046
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-2 1316	-5 793	787	-5 554	485	-24 497	535	-18 1032	984	-10 1669
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7 1379	8 1856	1821	4 779	730	-12 1376	1375	-9 1233	1222	-3 2014
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9 835	11 845	836	7 765	729	-9 463	449	-7 1478	1416	-1 1227
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-20 1203 1228	3 704	643	3 1016	973	7 476	507	19 627	717	-3 1758	1783
-18 1072 1115	4 2818	2755	4 1486	1470	8 3196	3158	14, 5, L		-2 1333	1344
-16 1991 2004	5 1590	1597	6 2729	2678	9 445	364	-19 1035	1053	-1 1075	1108
-14 3272 3336	7 790	776	7 502	425	10 427	502	-17 688	615	2 928	950
-12 369 357	8 1294	1287	8 698	709	11 1199	1176	-16 888	927	3 731	743
-10 3352 3381	9 486	456	9 742	764	12 2327	2305	-13 753	806	5 1725	1729
-6 1247 1229	10 797	832	10 2074	2053	13 749	740	-12 836	835	6 1942	1958
-4 2083 2064	12 1885	1854	11 1441	1455	15 1748	1710	-11 763	724	9 2938	2965
-2 3366 3335	13 533	560	12 916	967	16 909	842	-10 741	725	10 1381	1381
0 2853 2833	14 1787	1758	13 1589	1617	17 898	905	-9 2013	2000	11 1563	1612
2 4563 4697	15 1431	1465	14 2022	1966	18 665	566	-8 2080	2088	12 543	503
6 1977 1995	16 3184	3139	16 777	804	19 866	871	-7 692	771	13 788	737
10 1117 1174	17 927	969	17 1606	1631	20 989	977	-6 507	581	14 1364	1338
12 1132 1115	18 714	763	18 1780	1798	14, 4, L		-5 1789	1808	15 509	566
14 2971 2909	19 596	531	19 1192	1166	-19 1163	1243	-4 2007	2031	16 633	643
16 1143 1124	20 2041	1926	20 557	582	-18 1889	1884	-3 943	886	14, 7, L	
18 3058 3009	21 759	710	14, 3, L		-15 1293	1352	-2 479	423	-16 1312	1315
14, 1, L	14, 2, L		-20 2019	2034	-12 1261	1242	-1 955	879	-13 656	653
-20 750 724	-21 521	584	-19 784	735	-11 1595	1615	0 858	862	-12 1479	1483
-19 1021 1005	-20 1102	1114	-17 1651	1645	-10 1655	1726	2 969	912	-11 767	808
-18 1859 1843	-18 1829	1844	-16 1709	1713	-8 1312	1366	4 1010	984	-10 1097	1051
-17 728 772	-17 1111	1096	-13 1609	1637	-7 1592	1558	6 1404	1436	-9 2286	2389
-16 2316 2325	-16 1055	1073	-12 1413	1382	-6 2370	2386	7 2546	2547	-8 657	640
-14 1246 1204	-15 1741	1763	-11 1405	1409	-3 1335	1378	8 2597	2590	-7 893	959
-13 875 877	-14 2496	2417	-10 2060	2172	-2 2040	2116	11 2153	2184	-5 2652	2748

14, 7, L	14, 9, L	2 1926	1982	18 2804	2697	15, 3, L	-8 3227	3249
-3 1928 1921	-11 899 943	4 2932 2892	2892	19 929 931	931	-19 745 784	-7 610 611	
-2 649 630	-10 1176 1228	8 1335 1288	1288	15, 2, L	L	-18 2096 2010	-6 723 749	
-1 951 959	-9 906 925	10 817 861	861	-20 1332 1318	1318	-15 2047 2042	-5 1058 1000	
0 808 787	-7 381 387	12 2051 2082	2082	-19 1135 1172	1172	-14 1263 1267	-4 2861 2912	
2 525 576	-6 795 776	14 1932 1776	1776	-18 1252 1281	1281	-13 1123 1087	-2 1093 1143	
3 957 972	-5 912 899	16 3800 3678	3678	-17 1383 1365	1365	-12 1423 1439	-1 1250 1252	
4 2125 2122	-4 433 427	20 2102 2110	2110	-16 2178 2152	2152	-11 841 859	0 1167 1169	
6 965 1002	-2 749 737	15, 1, L	L	-14 972 965	965	-10 2248 2203	1 474 409	
7 2120 2214	-1 2136 2177	-20 924 947	947	-13 1490 1481	1481	-9 1278 1263	2 838 804	
8 898 982	0 1337 1274	-18 1390 1337	1337	-12 2678 2704	2704	-8 1585 1556	3 1169 1137	
9 774 804	1 614 551	-17 812 798	798	-11 2132 2028	2028	-7 440 415	4 1248 1234	
11 1903 1866	2 643 624	-16 1882 1861	1861	-10 1124 1059	1059	-6 3107 2970	5 913 942	
13 1230 1225	3 1767 1803	-15 1060 1050	1050	-8 2683 2563	2563	-3 835 787	6 2231 2142	
15 657 684	4 1094 1110	-14 3011 2995	2995	-6 810 849	849	-2 1735 1708	7 1359 1367	
16 1116 1122	5 1503 1559	-13 782 850	850	-4 1954 1974	1974	-1 643 634	8 2943 3011	
14, 8, L	6 1153 1192	-12 576 646	646	-3 662 692	692	1 1707 1792	11 2047 1973	
-14 1740 1758	7 725 775	-10 3320 3324	3324	-1 1593 1580	1580	2 1837 1880	12 2425 2358	
-13 500 491	11 1184 1178	-9 883 868	868	0 2653 2680	2680	3 1498 1481	14 1166 1168	
-12 811 881	14, 10, L	-7 792 810	810	1 602 694	694	4 1448 1497	15 997 1061	
-11 1126 1119	-6 491 383	-6 1765 1766	1766	2 1002 967	967	5 716 688	16 526 490	
-10 1099 1022	-3 1551 1641	-5 697 638	638	3 1447 1541	1541	6 2271 2253	18 590 621	
-8 1343 1276	1 2466 2483	-4 736 776	776	4 2810 2905	2905	7 674 712	15, 5, L	
-7 1859 1862	2 507 375	-3 1005 1080	1080	5 1564 1547	1547	8 1851 1814	-17 848 868	
-5 794 822	3 1437 1439	-2 3100 3075	3075	6 825 828	828	10 2767 2802	-15 844 837	
-3 1388 1385	4 643 694	-1 702 738	738	7 418 466	466	11 587 572	-14 581 619	
0 654 643	5 863 831	0 1339 1388	1388	8 2496 2457	2457	13 1816 1786	-12 620 623	
1 1548 1549	15, 0, L	1 804 816	816	10 855 897	897	14 1841 1786	-11 1606 1525	
2 2011 1961	-20 665 686	2 3281 3320	3320	11 372 315	315	17 1158 1122	-10 1695 1664	
4 815 824	-18 1848 1922	3 1019 1091	1091	12 2829 2830	2830	18 1036 1007	-8 1038 1040	
5 1705 1744	-16 2423 2430	6 2542 2556	2556	13 679 639	639	19 1312 1274	-7 2460 2456	
6 1450 1451	-14 1491 1467	7 494 522	522	15 1754 1698	1698	15, 4, L	-6 2523 2456	
7 857 864	-12 3816 3801	10 1883 1911	1911	16 2075 2071	2071	-17 1310 1245	-5 1496 1439	
8 1155 1176	-10 652 739	11 552 591	591	17 1195 1198	1198	-16 988 993	-3 1212 1301	
9 1259 1323	-8 2308 2164	12 1297 1205	1205	19 560 524	524	-13 1311 1295	-2 1640 1616	
12 536 574	-4 1501 1517	13 926 882	882			-12 1074 963	-1 1137 1155	
13 1587 1627	-2 1832 1843	14 2965 2931	2931			-10 2055 2057	1 436 708	
	0 3150 3090	17 1017 1053	1053			-9 1358 1349	2 655 649	

15, 5, L	15, 7, L	9 1170 1154	-12 2639 2612	-6 2295 2266	6 1444 1423
4 1485 1460	-14 1384 1493	11 1341 1331	-11 1331 1289	-5 1012 1012	8 2131 2122
5 1815 1827	-11 1406 1393	15, 9, L	-10 369 434	-3 868 839	10 458 549
6 1585 1619	-10 946 974	-8 826 787	-9 848 812	-2 1891 1762	11 986 986
8 1113 1137	-7 2600 2601	-7 764 854	-8 2006 1913	0 705 726	12 2550 2548
9 2667 2682	-6 625 608	-6 926 804	-6 743 746	1 1705 1676	14 785 772
10 2504 2563	-5 1450 1393	-4 657 635	-4 1674 1635	2 2457 2418	15 1296 1266
11 1409 1371	-3 1901 1851	-3 1471 1487	-3 493 539	3 1270 1299	16 901 872
12 629 655	-1 1210 1225	0 629 585	-2 1262 1345	4 362 421	17 1152 1099
13 966 979	0 600 512	1 2106 2124	-1 998 911	5 784 747	16, 4, L
14 1141 1112	1 853 811	2 1160 1167	0 2931 2802	6 2166 2192	-17 859 868
15 778 707	2 1565 1620	3 1063 1066	1 507 447	7 980 969	-15 1086 1114
15, 6, L	3 377 427	4 674 655	2 592 646	8 655 684	-13 757 715
-16 819 896	4 465 597	5 1013 1067	3 669 630	10 2234 2207	-12 1094 1084
-13 697 664	5 1678 1735	7 1354 1417	4 2746 2743	11 557 613	-11 1035 985
-12 1011 993	6 1290 1349	16, 0, L	5 1140 1119	12 643 527	-10 1511 1452
-11 637 652	8 796 756	-18 959 1033	8 1617 1637	13 1055 1068	-9 863 782
-9 2302 2297	9 2096 2040	-16 1846 1840	9 543 487	14 2117 2083	-8 1597 1550
-8 1277 1275	11 1470 1516	-14 3205 3231	10 723 695	17 1092 1076	-7 1170 1101
-7 1295 1301	12 847 829	-10 2549 2575	11 464 541	18 1593 1588	-6 3133 3047
-6 799 862	13 1065 1082	-8 616 681	12 1862 1808	16, 3, L	-5 936 892
-5 2560 2453	14 894 870	-6 1515 1511	13 729 676	-17 1301 1338	-3 868 924
-4 1645 1585	15, 8, L	-4 1335 1315	14 1141 1151	-16 1177 1170	-2 1423 1391
-3 2117 2079	-12 1287 1287	-2 3339 3180	15 1008 1012	-15 866 882	-1 542 502
-1 701 743	-10 849 894	0 1325 1252	16 2810 2725	-13 1451 1455	1 1616 1649
0 727 759	-9 1564 1563	2 3396 3528	17 557 477	-12 1051 1088	2 515 496
3 965 990	-7 460 515	6 2135 2014	16, 2, L	-11 1745 1786	3 1373 1398
4 1262 1279	-5 1969 2014	10 1192 1266	-18 1499 1468	-10 1575 1595	4 1800 1796
5 770 756	-4 673 620	12 1277 1271	-17 638 594	-8 2928 2996	5 989 968
7 2656 2669	-3 1094 1091	14 2271 2362	-16 705 663	-7 699 649	6 1647 1701
8 1173 1240	-2 1171 1244	16 712 708	-15 1714 1678	-4 2479 2367	7 834 894
9 1172 1228	-1 1477 1446	18 2818 2724	-14 1786 1745	-3 664 638	8 1474 1493
11 2099 2136	0 1047 1045	16, 1, L	-13 1460 1454	-1 1794 1805	9 980 989
12 1177 1214	3 1671 1697	-18 1334 1354	-12 487 697	0 1675 1683	10 2816 2791
13 1295 1301	4 1553 1615	-17 641 652	-10 2284 2274	1 1166 1150	11 932 929
	5 976 936	-16 1875 1864	-9 1399 1453	3 1533 1490	12 465 546
	6 1149 1202	-14 1136 1108	-8 858 928	4 1936 1892	13 917 919
	7 1130 1140	-13 1025 1034	-7 358 370	5 1600 1578	14 947 977

16, 4, L	1 445	447	16, 9, L	6 1604	1673	-3 564	574	-7 887	896
15 891 916	2 882	935	-2 1240 1280	7 756	700	-2 1136	1135	-6 1676	1721
16, 5, L	5 1614	1701	-1 1807 1825	10 864	899	1 1797	1735	-5 1498	1552
-15 783 739	6 1002	1016	0 989 991	12 1060	1072	2 883	848	-2 1170	1161
-13 455 445	7 464	538	1 863 807	14 1650	1654	3 1640	1635	-1 1034	1006
-12 385 504	9 2130	2109	17, 0, L	17, 2, L		4 743	800	0 685	653
-10 1339 1364	10 1126	1052	-16 1628 1626	-16 1056	1026	6 1243	1262	4 1383	1416
-9 1618 1613	11 2137	2150	-14 878 950	-15 634	624	7 955	952	5 852	883
-8 2007 2021	16, 7, L		-12 2426 2375	-13 1183	1158	8 781	991	6 1057	1019
-7 1340 1295	-12 860	906	-10 997 1036	-12 1210	1229	9 494	501	7 750	805
-6 805 872	-9 1597	1682	-6 702 708	-11 1906	1912	10 2005	1885	8 961	964
-5 1704 1607	-7 914	950	-4 1185 1198	-9 930	915	13 597	604	9 835	829
-4 2197 2246	-6 843	831	-2 2119 2111	-8 1691	1695	14 1027	992	10 1492	1475
-3 1798 1785	-5 1886	1978	0 2486 2478	-4 1402	1410	15 688	599	11 1673	1640
-2 676 746	-3 1766	1808	4 2565 2611	-3 981	914	17, 4, L		17, 6, L	
-1 688 615	-2 1055	1118	6 585 708	-2 801	751	-11 658	662	-10 918	935
0 780 853	-1 734	733	8 1106 1003	-1 1300	1271	-10 1471	1558	-9 947	1083
1 747 683	0 728	754	14 1623 1780	0 1506	1497	-8 1929	1955	-8 729	680
3 1126 1116	3 1163	1239	16 2105 2098	1 1046	1007	-7 923	958	-7 749	799
6 1793 1799	4 1042	1101	17, 1, L	3 981	974	-6 637	638	-6 673	716
7 2088 2106	7 1842	1949	-16 996 1044	4 1752	1751	-4 2211	2271	-5 881	1042
8 1860 1928	9 1605	1678	-15 1016 1023	5 1579	1633	-2 863	853	-4 585	583
9 1193 1223	10 924	934	-14 2011 1979	7 539	554	-1 874	866	-3 2011	2080
11 1509 1515	-11 1141	1156	-13 1103 1133	8 1266	1275	1 802	785	0 586	529
12 1515 1514	16, 8, L		-10 1401 1447	10 412	445	2 399	444	1 511	607
13 1409 1449	-7 1399	1473	-9 631 600	12 1486	1498	3 854	890	4 723	742
16, 6, L	-6 514	434	-8 436 333	13 655	629	5 906	865	6 908	936
-14 626 609	-5 986	1002	-7 923 890	15 1086	1068	6 1648	1704	7 1772	1763
-12 676 684	-4 1140	1074	-6 848 854	16 1433	1389	8 2061	2016	8 811	845
-11 1125 1125	-3 1335	1358	-5 1019 946	17, 3, L		9 598	619	9 1660	1646
-10 1086 1124	-1 678	734	-4 1511 1452	-15 912	917	12 1559	1545	10 793	783
-8 769 773	0 1014	979	-3 587 601	-13 1462	1436	13 1096	1086	11 407	411
-7 2063 2023	1 1114	1143	-2 1954 1965	-12 657	711	14 852	847	17, 7, L	
-6 1163 1154	2 1244	1270	0 726 691	-10 1127	1108	17, 5, L		-8 650	691
-5 2043 2046	3 547	619	1 624 469	-9 964	937	-13 427	490	-7 810	907
-3 1178 1132	5 1078	1150	2 2423 2412	-8 1046	1133	-12 1072	1017	-5 1416	1478
-2 753 844	7 1053	1098	3 867 933	-6 2379	2307	-10 818	787	-4 1104	1157
-1 1652 1601			5 327 367	-5 520	534	-8 1379	1426	-1 1005	1014

17, 7, L	4 1221 1236	18, 4, L	18, 7, L	19, 3, L	6 739 751
0 585 608	5 1054 1079	-12 646 668	0 662 711	-7 1197 1273	20, 2, L
2 999 973	6 1097 1089	-9 436 407	19, 0, L	-6 809 785	-6 806 764
5 1292 1289	7 344 295	-8 1183 1201	-10 961 1054	-5 640 736	-5 1082 1088
7 829 801	13 570 593	-6 1718 1711	-8 1278 1305	2 934 963	-2 668 729
8 550 511	14 1170 1181	-5 484 448	-4 832 839	3 1007 996	0 612 599
17, 8, L	18, 2, L	-2 656 638	-2 1160 1246	4 494 492	2 808 777
-3 838 867	-13 1270 1256	1 372 402	0 985 961	5 774 807	3 1063 1100
-2 1577 1621	-10 621 603	4 1397 1352	2 556 517	9 813 835	5 838 962
0 1024 1005	-9 519 536	6 824 837	6 1506 1526	19, 4, L	20, 3, L
1 671 674	-7 1352 1324	8 867 896	8 645 694	-4 480 663	-3 635 503
18, 0, L	-6 821 840	10 1209 1260	10 630 588	-2 715 691	0 1398 1345
-14 1470 1426	-5 1066 1120	11 728 698	19, 1, L	0 1058 1116	1 616 600
-8 808 802	-4 568 639	18, 5, L	-10 395 426	2 740 718	2 692 715
-6 425 353	-2 732 723	-10 1298 1321	-7 983 1002	6 826 906	
-4 1585 1617	1 707 779	-8 695 635	-6 846 810	7 679 576	
-2 1061 1058	2 889 895	-7 591 593	-5 765 774	19, 5, L	
2 1894 1902	3 1318 1259	-6 694 691	-4 740 822	-5 957 1021	
4 1171 1145	6 524 551	-4 1039 1084	1 407 278	-3 779 774	
6 1035 1072	7 676 651	-3 1317 1285	2 643 561	-1 608 651	
8 991 1054	9 817 815	-1 408 471	3 1096 1118	1 820 841	
10 387 392	10 570 575	2 600 558	4 1117 1069	2 671 664	
12 1275 1197	11 851 808	6 1421 1421	9 523 496	4 1170 1173	
14 526 544	18, 3, L	7 528 493	10 772 823	20, 0, L	
18, 1, L	-11 1242 1217	8 1263 1239	11 900 911	-6 1747 1802	
-13 637 651	-10 918 992	9 656 650	19, 2, L	-4 656 665	
-12 1079 1027	-8 1287 1274	18, 6, L	-11 930 973	-2 914 946	
-11 1073 1089	-5 705 690	-7 812 762	-9 835 799	0 678 630	
-10 707 707	-4 1572 1546	-5 1521 1505	-7 587 584	4 1006 959	
-9 963 999	-3 518 475	-1 1020 1016	-5 889 862	6 1037 1060	
-7 363 450	-1 849 689	1 712 742	-3 631 583	20, 1, L	
-5 423 473	1 965 1042	3 496 490	-2 353 289	-5 636 618	
-4 453 523	5 1188 1234	4 693 747	-1 672 624	-4 1423 1424	
-3 527 695	6 673 679	5 909 927	1 1138 1098	-3 543 663	
-2 1084 1097	7 805 862	6 526 581	5 1106 1072	-2 929 930	
-1 627 696	8 1031 1049		6 437 383	1 694 727	
0 1399 1431	11 557 564		7 989 1074	4 613 617	
1 1063 1000	12 816 838		9 688 634	5 706 748	

CALCULATED AND OBSERVED STRUCTURE FACTORS OF
[Pt₃(CO)(SnF₃)(dppm)₃]_{0.75}[Pt₃Cl(SnF₃)(dppm)₃]_{0.25}[PF₆]_{0.25}

h	k	l	Calcd	Obsd
0	0	0	1.0000	1.0000
1	0	0	0.0000	0.0000
0	1	0	0.0000	0.0000
0	0	1	0.0000	0.0000
1	1	1	0.0000	0.0000
2	0	0	0.0000	0.0000
0	2	0	0.0000	0.0000
0	0	2	0.0000	0.0000
2	1	1	0.0000	0.0000
1	2	1	0.0000	0.0000
2	2	2	0.0000	0.0000
3	0	0	0.0000	0.0000
1	3	0	0.0000	0.0000
0	1	3	0.0000	0.0000
1	1	3	0.0000	0.0000
2	1	3	0.0000	0.0000
3	1	3	0.0000	0.0000
4	0	0	0.0000	0.0000
2	2	2	0.0000	0.0000
3	2	2	0.0000	0.0000
4	2	2	0.0000	0.0000
5	0	0	0.0000	0.0000
3	3	0	0.0000	0.0000
1	4	0	0.0000	0.0000
0	2	4	0.0000	0.0000
2	3	2	0.0000	0.0000
3	3	3	0.0000	0.0000
4	3	3	0.0000	0.0000
5	3	3	0.0000	0.0000
6	0	0	0.0000	0.0000
4	4	0	0.0000	0.0000
2	5	0	0.0000	0.0000
0	3	5	0.0000	0.0000
1	4	3	0.0000	0.0000
2	4	4	0.0000	0.0000
3	4	5	0.0000	0.0000
4	4	6	0.0000	0.0000
5	4	7	0.0000	0.0000
6	4	8	0.0000	0.0000
7	4	9	0.0000	0.0000
8	4	10	0.0000	0.0000
9	4	11	0.0000	0.0000
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11	4	13	0.0000	0.0000
12	4	14	0.0000	0.0000
13	4	15	0.0000	0.0000
14	4	16	0.0000	0.0000
15	4	17	0.0000	0.0000
16	4	18	0.0000	0.0000
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18	4	20	0.0000	0.0000
19	4	21	0.0000	0.0000
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21	4	23	0.0000	0.0000
22	4	24	0.0000	0.0000
23	4	25	0.0000	0.0000
24	4	26	0.0000	0.0000
25	4	27	0.0000	0.0000
26	4	28	0.0000	0.0000
27	4	29	0.0000	0.0000
28	4	30	0.0000	0.0000
29	4	31	0.0000	0.0000
30	4	32	0.0000	0.0000
31	4	33	0.0000	0.0000
32	4	34	0.0000	0.0000
33	4	35	0.0000	0.0000
34	4	36	0.0000	0.0000
35	4	37	0.0000	0.0000
36	4	38	0.0000	0.0000
37	4	39	0.0000	0.0000
38	4	40	0.0000	0.0000
39	4	41	0.0000	0.0000
40	4	42	0.0000	0.0000
41	4	43	0.0000	0.0000
42	4	44	0.0000	0.0000
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-210502 9217	-1 1467 1555	-3 2249 2406	0 2031 1907	-2 3755 3696	-3 2923 2930
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3, 1, L	5, 3, L	0 1469 1314	1 2860 2722	4 4817 4699	0 4922 4780
0 1965 2815	-2 2834 2912	1 1824 1858	7, 3, L	6 6068 6224	1 788 868
3, 2, L	-1 2540 2578	2 1411 1294	-2 1952 2127	8, 0, L	2 5150 5193
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0 9380 8832	-1 3406 3429	0 883 646	-2 1376 1572	-2 1635 1662	0 3113 3116
4, 1, L	0 1467 1296	1 2747 2817	-1 3099 3082	-1 1935 1875	1 1889 1843
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-210335 10792	-4 6209 6210	-5 6630 6674	7, 5, L	-3 6000 5598	6 5045 5133
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2 2761 2843	10, 1, L	3 3818 3777	2 2419 2301	2 3487 3312	6 1879 1850
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2 1967 2024	1 3696 3614	6 2121 2164	14 2454 2398	4 2098 2126	8 1820 1784
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19, 5, L	6 2274 2621	6 1865 1491	20, 20, L	21, 18, L	10 2520 2366
1 1717 2698	19, 17, L	8 2852 2489	14 2249 2174	2 1738 1582	12 1842 1752
3 1743 1925	4 2728 2546	9 2369 2274	16 2682 2656	21, 19, L	22, 13, L
19, 6, L	6 2763 2818	10 2671 2533	18 3019 3066	4 2388 2111	8 2283 2476
1 2717 2934	19, 19, L	20, 11, L	20 2558 2642	6 2530 2636	10 2177 2095
3 2364 2416	2 1444 1441	4 1781 1677	21, 6, L	21, 20, L	11 1566 1578
5 2255 1925	4 2764 2333	6 2570 2615	3 1659 1871	11 1491 751	22, 14, L
19, 7, L	6 2797 2827	8 3110 3311	21, 7, L	21, 21, L	2 1497 1838
1 3703 3912	8 2590 2644	10 2843 2997	1 3080 3580	6 2547 2515	22, 15, L
3 3278 3321	20, 0, L	20, 12, L	3 2456 2645	8 1789 2115	4 1645 1596
9 1892 1990	0 3976 4046	8 2272 2314	21, 8, L	22, 1, L	22, 16, L
1 3168 3467	20, 1, L	12 1957 1614	3 1372 2017	1 3071 3228	0 2307 2269
19, 8, L	1 2552 2764	20, 13, L	21, 9, L	22, 3, L	16 2843 2429
19, 9, L	20, 2, L	0 2140 1907	1 4657 4634	0 2461 2632	22, 18, L
1 4845 5190	0 2132 2203	6 2416 2628	3 3687 3708	1 3221 3239	16 2299 2168
3 4116 4298	2 3262 3364	8 2957 3023	5 2266 2226	2 1634 2012	18 2293 2658
9 3298 2985	20, 3, L	10 2941 2931	21, 10, L	3 3133 3180	22, 20, L
19, 10, L	0 3383 3420	12 1949 2144	1 1758 2451	22, 5, L	20 2338 1994
3 1943 2092	1 2004 2360	20, 14, L	6 1692 1647	0 3016 3059	23, 7, L
4 2042 1742	2 1936 2366	10 1356 1425	21, 11, L	2 2390 2507	1 1984 3013
6 1775 1868	3 2212 2315	20, 15, L	1 4249 4172	22, 7, L	23, 9, L
19, 11, L	20, 4, L	2 2006 1723	3 3188 3235	4 1561 1780	1 2600 3014
1 4711 4752	0 2807 2703	4 1432 1761	21, 12, L	22, 8, L	3 2663 2561
3 4199 3915	2 1820 1895	6 2169 2055	4 1851 1778	8 2314 1908	23, 11, L
5 2512 2547	4 2739 2797	8 1997 2161	21, 13, L	22, 9, L	1 1707 2032
19, 12, L	20, 5, L	10 1583 2018	1 2374 2443	8 1788 1547	3 1667 1784
1 2450 2249	0 3123 3231				10 1635 1839
3 1591 1703	2 2547 2681				
4 2193 2189					

23, 12, L	24, 6, L	26, 4, L	29, 19, L
11 1546 1615	3 1996 2291	3 1874 2017	1 1744 2036
23, 13, L	5 1715 2079	26, 5, L	3 1748 1698
8 2091 1967	24, 7, L	0 2216 2442	30, 9, L
23, 14, L	2 1877 2275	2 2128 2496	5 1509 1297
4 1500 1430	4 2125 2173	4 2253 2335	31, 7, L
6 1955 1677	24, 8, L	26, 6, L	0 1525 1853
8 1654 1663	0 1565 1673	1 1474 1817	31, 8, L
23, 16, L	5 1658 1525	3 1807 2049	2 1713 1511
10 1632 1334	24, 10, L	5 2218 2013	4 1767 1141
23, 17, L	2 1579 1523	26, 7, L	31, 10, L
5 1630 1051	24, 12, L	0 2035 1918	2 1562 1656
23, 18, L	0 2150 2042	2 1873 2060	33, 12, L
6 1736 1816	2 2201 1940	26, 13, L	0 1701 1579
23, 19, L	12 1989 2020	11 1826 1745	34, 0, L
6 2148 1938	24, 13, L	27, 8, L	0 1614 1234
23, 20, L	11 1843 1635	7 1585 1430	
4 2245 2031	13 1725 1925	27, 9, L	
23, 21, L	24, 14, L	6 1677 1569	
4 1726 1541	0 2088 2377	27, 10, L	
6 2368 1909	24, 15, L	9 1733 1657	
23, 22, L	4 1453 1318	27, 17, L	
2 1519 1445	25, 9, L	1 2523 2293	
23, 23, L	6 1675 1724	5 1770 1590	
6 1724 1522	8 1508 1722	27, 19, L	
24, 1, L	25, 13, L	1 2346 2262	
1 3158 2934	10 1620 1551	28, 3, L	
24, 3, L	25, 19, L	0 1460 1705	
0 2838 2551	1 1894 1872	28, 9, L	
1 2634 2809	7 1520 1173	7 1547 1473	
3 2403 2640	25, 21, L	28, 11, L	
24, 4, L	5 1696 1237	11 1918 1870	
3 1634 2031	26, 1, L	29, 8, L	
24, 5, L	1 1730 1980	0 1951 1790	
0 3045 3029	26, 3, L	29, 10, L	
2 2621 2701	0 1859 2002	0 1956 1771	

CALCULATED AND OBSERVED STRUCTURE FACTORS OF
[Pt₃(SnF₃)₂(dppm)₃]

0, 0, L	10 1192 1334	0, 3, L	-17 869 837	-12 1651 1656	-1 2541 2526
4 762 752	11 1381 1541	-22 1267 1288	-16 673 658	-11 429 455	0 409 445
5 676 652	13 1440 1507	-21 594 647	-15 700 724	-8 2112 2025	1 2751 2704
6 1742 1766	15 1115 1120	-20 902 926	-14 967 963	-7 585 560	2 1528 1567
8 2828 2881	17 660 677	-18 498 586	-12 1051 1016	-6 2486 2350	3 1249 1300
10 2338 2326	23 726 707	-14 523 517	-11 894 887	-5 2142 1988	4 2445 2432
11 1135 1238	0, 2, L	-12 1765 1762	-9 2540 2406	-4 1638 1520	5 1479 1422
12 2011 2124	-21 1156 1177	-11 677 695	-7 2592 2445	-3 2905 2806	6 2539 2528
14 1602 1638	-19 716 644	-10 2885 2895	-6 1595 1406	-2 1170 1135	7 2366 2421
16 989 1035	-15 955 1034	-9 868 862	-5 2047 1891	-1 1965 2021	8 1218 1225
21 719 741	-14 510 524	-8 2625 2519	-4 2475 2349	0 2031 2025	9 2694 2810
22 784 839	-13 1679 1654	-6 1020 989	-3 1754 1665	1 1323 1263	10 442 528
0, 1, L	-12 1251 1276	-5 408 379	-2 2267 2328	2 1933 1877	11 2122 2328
-23 748 854	-11 2181 2292	-4 294 210	-1 1970 2005	3 931 945	12 975 1030
-22 1250 1226	-10 1655 1691	-3 1986 1796	0 3688 3621	5 2694 2731	13 1441 1415
-20 704 630	-9 2954 2956	-2 2664 2589	2 1077 1046	6 609 675	14 793 835
-17 573 563	-8 1077 1114	-1 419 246	3 604 578	7 2391 2629	15 600 583
-15 1012 1086	-7 1765 1718	0 2663 2788	4 2068 2162	8 1638 1795	16 543 597
-14 999 979	-6 651 654	2 1039 1174	5 1315 1432	9 1073 1142	21 593 653
-13 1698 1721	-5 971 854	3 1261 1327	6 2720 2993	10 2633 2820	0, 7, L
-12 1410 1493	-4 1791 1587	4 1509 1640	8 1533 1720	12 2253 2311	-20 537 520
-11 1866 1963	-3 2521 2222	5 1277 1407	9 1720 1785	13 608 599	-18 1392 1363
-10 2003 2104	-2 2005 1833	6 915 961	10 925 997	14 1286 1310	-17 795 823
-9 2220 2280	-1 2886 2766	7 2172 2246	11 1186 1310	15 524 569	-16 1829 1786
-8 1470 1529	0 1843 1820	8 690 694	12 471 519	20 827 833	-15 1218 1254
-7 2443 2414	1 2781 2632	9 1365 1447	13 1656 1718	0, 6, L	-14 1934 1892
-6 1039 1052	2 672 677	11 1029 1098	14 542 613	-20 390 318	-13 1222 1227
-5 1088 1071	3 1529 1638	12 668 668	15 611 624	-19 1008 936	-12 1528 1516
-4 1045 924	4 945 971	14 516 514	16 776 895	-17 1814 1861	-11 1004 1024
-3 1802 1735	5 1512 1567	15 433 557	18 1234 1187	-15 2444 2447	-10 627 629
-2 1199 1110	6 376 478	17 1475 1492	20 928 933	-13 2405 2352	-9 399 365
2 2637 2886	7 1788 1873	19 1128 1082	22 907 811	-11 1627 1648	-8 689 644
3 2110 2404	9 1927 1970	21 759 731	0, 5, L	-10 502 564	-6 2261 2270
4 1254 1308	11 842 801	0, 4, L	-18 948 950	-7 1348 1302	-5 422 408
6 396 386	14 733 738	-22 1013 1014	-16 1954 1933	-6 514 449	-4 2248 2335
7 1382 1438	16 947 965	-21 438 463	-15 541 597	-5 2472 2398	-2 1525 1587
8 1513 1628	18 977 1018	-20 616 654	-14 1744 1709	-4 1109 998	-1 813 864
9 1563 1637	20 481 499	-19 533 524	-13 1199 1125	-3 2797 2784	0 2222 2183

0, 7, L	11 490 467	10 1514 1443	8 2158 2011	2 968 968	4 1408 1456
1 1645 1607	12 628 635	12 998 861	10 2295 2135	4 721 621	5 692 738
2 1608 1582	17 779 768	16 872 832	11 779 662	7 599 582	6 469 563
3 2378 2411	19 755 889	17 1151 1025	12 1281 1218	8 706 741	11 1025 1145
4 1100 1108	0, 9, L	18 1467 1226	13 635 673	0, 16, L	1, -12, L
5 1763 1689	-17 1149 1139	19 1213 1205	16 892 758	1 826 899	-13 686 627
6 1701 1634	-15 1306 1368	0, 11, L	0, 13, L	3 980 1046	-11 1848 1668
8 1946 1831	-13 1339 1298	-9 441 511	-12 1178 1247	5 686 782	-9 2388 2173
9 746 842	-10 694 604	-8 783 771	-10 826 878	1, -15, L	-7 2110 2000
10 1500 1492	-8 515 520	-7 874 939	-9 653 716	-8 659 622	-5 1294 1397
11 1232 1240	-4 554 535	-6 793 877	-7 905 881	-7 729 623	-4 670 619
12 911 861	-2 935 918	-5 1585 1589	-6 676 752	-1 838 897	-1 1816 1801
13 1017 1017	-1 830 864	-3 1564 1570	-5 902 983	1 743 873	1 2332 2498
18 584 542	1 1304 1282	-1 942 900	-4 1067 1066	3 596 748	3 2188 2291
0, 8, L	2 666 617	1 553 587	-3 1065 1204	1, -14, L	5 1360 1491
-19 625 606	3 954 894	4 717 709	-2 948 998	-10 1026 948	7 572 606
-18 847 845	6 1360 1357	5 738 750	-1 892 883	-8 1341 1318	12 976 1011
-17 891 835	8 1844 1758	6 1188 1097	0 945 1046	-6 961 880	13 768 736
-16 1398 1376	10 1624 1519	7 1712 1515	2 658 682	-2 867 841	14 557 626
-15 977 958	12 580 522	8 1128 1002	5 837 890	-1 614 713	1, -11, L
-14 1515 1500	16 1107 1013	9 2346 2275	7 1426 1474	0 1245 1313	-16 814 714
-13 828 826	18 1397 1367	11 2153 1966	8 1047 839	1 752 738	-12 1287 1152
-12 1103 1069	20 1260 1239	17 1436 1271	9 1541 1455	2 1281 1376	-10 2259 2051
-11 507 497	0, 10, L	0, 12, L	10 1390 1295	3 626 538	-9 1071 1037
-7 732 781	-17 691 752	-14 695 767	11 958 999	4 830 976	-8 2375 2187
-6 755 732	-16 699 573	-13 801 917	12 953 839	1, -13, L	-7 1500 1432
-5 926 969	-15 878 822	-11 931 1004	0, 14, L	-11 1073 909	-6 1290 1263
-3 1741 1793	-12 755 705	-9 549 615	-5 830 930	-10 1195 1210	-5 1153 1079
-2 959 981	-9 1022 997	-8 955 993	-3 1134 1128	-9 1130 1042	-2 931 895
-1 1252 1237	-8 714 768	-6 1467 1620	-1 876 868	-8 1657 1595	0 1669 1684
0 1774 1810	-7 632 593	-4 1670 1802	1 793 829	-6 1491 1544	1 566 467
1 752 778	-6 768 770	-2 1794 1798	7 928 1044	-3 504 474	2 1891 1987
2 1808 1823	-4 584 604	-1 826 714	9 1374 1276	-2 678 778	3 598 564
4 689 622	5 645 743	0 1214 1214	11 916 878	-1 999 1012	4 1457 1514
6 619 566	6 679 598	1 796 742	0, 15, L	0 1574 1770	6 796 839
8 763 771	7 1835 1808	2 745 759	-5 638 615	1 1200 1213	14 803 840
9 731 694	8 1426 1311	3 799 654	0 590 743	2 1722 1786	
10 980 919	9 1643 1591	6 1705 1693	1 531 571	3 1265 1240	

1, -10,	L	8	978	961	1, -7,	L	-11	933	966	-4	2193	2230	3	1388	1215
-18 1076	1014	10	903	840	-17 612	699	-10	2205	2305	-3	2554	2588	4	867	849
-16 762	725	11	438	424	-15 1190	1195	-8	2045	2218	-1	2595	2654	5	1191	1034
-11 1228	1178	12	845	858	-14 610	681	-7	1211	1213	0	1705	1574	6	1194	1036
-10 1396	1301	13	482	387	-13 1446	1536	-6	999	1011	1	2461	2370	8	1720	1520
-9 1440	1232	14	737	746	-12 1123	1173	-5	2298	2328	2	2753	2625	10	1428	1320
-8 2243	2116	1, -8,	L	L	-11 1231	1382	-4	1271	1218	3	1101	1023	11	468	428
-7 1271	1171	-16 861	951	951	-10 1412	1478	-3	2173	2231	4	1476	1416	12	753	736
-6 1720	1611	-14 723	872	872	-9 1012	989	-2	3225	3067	5	806	778	13	636	580
-2 820	782	-13 1008	952	952	-8 1396	1330	-1	731	780	6	1154	1015	16	747	753
0 668	728	-11 998	989	989	-7 1188	1136	0	4139	4039	7	319	157	18	1056	1006
1 463	546	-10 707	772	772	-6 755	760	2	3483	3430	8	801	656	20	857	877
2 665	721	-9 1078	1160	1160	-4 2661	2535	3	1199	1069	10	1132	1062	21	541	540
3 876	905	-8 1088	1148	1148	-2 2931	3018	4	2126	2037	12	968	951	22	766	659
6 490	513	-7 756	671	671	-1 1909	1891	5	1294	1200	13	836	833	1, -3,	L	L
7 830	832	-6 801	785	785	0 2393	2339	7	1070	1062	14	689	627	-22	1294	1228
9 989	955	-5 801	713	713	1 2731	2722	8	1131	1041	15	1110	1111	-20	944	933
11 821	904	-3 2280	2304	2304	2 609	630	9	475	450	17	1197	1146	-18	724	763
14 387	351	-1 2495	2555	2555	3 1614	1653	10	1409	1300	19	825	886	-14	788	824
1, -9,	L	0 582	499	499	4 775	760	12	1421	1255	1, -4,	L	L	-13	474	481
-19 1062	1027	1 2079	2194	2194	5 2028	1885	14	1605	1571	-21	1319	1330	-12	1053	1134
-17 1141	1081	2 1396	1315	1315	6 1232	1177	16	1836	1741	-19	875	818	-11	865	923
-15 748	736	4 1025	1001	1001	8 1302	1259	18	1308	1272	-16	673	713	-10	762	864
-14 546	486	5 689	689	689	9 974	968	19	408	359	-15	925	918	-9	1390	1489
-11 1226	1174	6 1042	1046	1046	10 898	860	20	631	616	-14	1066	1132	-8	1263	1357
-9 2125	1980	7 1269	1254	1254	11 1432	1338	1, -5,	L	L	-13	943	1002	-7	585	515
-7 2038	1885	9 841	777	777	12 478	555	-22	796	904	-12	1563	1550	-6	820	775
-6 759	651	10 581	570	570	13 1268	1180	-21	1097	1064	-11	446	386	-4	740	764
-5 1133	996	11 683	666	666	14 749	749	-20	609	619	-10	1994	2037	-3	593	576
-4 706	617	12 880	902	902	15 1568	1431	-17	936	884	-8	1234	1331	-2	619	644
-3 745	753	13 912	859	859	16 645	733	-15	1653	1638	-7	576	641	0	903	773
-2 1307	1293	14 969	910	910	17 1441	1415	-14	835	844	-5	1574	1639	1	3707	3445
-1 880	788	15 872	896	896	19 1040	992	-13	1896	1937	-4	749	780	2	886	813
0 1324	1312	16 612	549	549	20 365	406	-11	2315	2426	-2	2112	2110	4	1423	1259
2 872	856	17 672	700	700	1, -6,	L	-9	2519	2559	-1	2782	2675	5	757	649
5 788	801	18 637	638	638	-16 1464	1388	-8	688	689	0	930	853	7	2785	2439
6 930	952				-14 2087	1997	-7	873	884	1	1443	1354	9	3155	2994
7 604	577				-12 2154	2170	-6	1812	1980	2	1889	1679	10	950	892

1, -3,	L	20	892	948	1, 0,	L	-15	525	669	-12	2111	2178	-12	814	809
11 2851	2724	22 1596	1604	-21	516	485	-14	1748	1821	-10	1721	1769	-11	1387	1428
12 1151	1094	1, -1,	L	-20	716	731	-13	2309	2433	-9	608	667	-10	583	544
13 985	1007	-19 780	804	-18	1044	1101	-12	719	872	-8	1506	1585	-9	1178	1248
16 563	954	-17 1456	1512	-17	1474	1546	-11	2952	3046	-7	1460	1472	-8	1503	1588
17 639	672	-16 944	979	-16	970	1080	-9	990	1089	-6	1224	1274	-7	1966	2023
19 1313	1300	-15 763	908	-15	1958	2026	-8	1217	1237	-5	685	656	-6	1049	1106
21 1445	1417	-14 1815	1850	-13	1813	1934	-7	563	648	-4	288	234	-5	1374	1410
1, -2,	L	-12 1239	1321	-12	1747	1863	-5	476	487	-3	295	310	-4	1637	1601
-22 529	430	-11 421	438	-11	888	1059	-4	952	920	-2	1833	1874	-3	1076	1081
-16 996	974	-10 625	637	-10	1964	2017	-3	763	811	-1	1542	1563	-2	572	617
-14 867	916	-9 664	631	-9	786	827	-2	3023	3294	0	2078	2103	-1	2059	2099
-13 616	646	-8 467	447	-6	278	326	-1	2144	2263	1	1289	1143	1	1875	1969
-9 510	570	-6 462	431	-5	714	730	2	2968	2861	2	2598	2913	2	972	924
-8 826	727	-4 1775	1823	-3	3412	3474	3	1623	1620	3	2044	2062	3	2432	2488
-7 461	476	-3 1441	1564	-2	557	424	4	1723	1711	4	474	545	4	1128	1085
-6 541	563	-2 2636	2653	-1	6043	6187	5	2294	2409	5	1332	1447	5	1473	1566
-4 362	307	-1 1894	2072	2	1077	1034	6	1116	1126	6	1110	1227	6	2195	2400
-2 1402	1360	0 3814	3597	3	3493	3234	7	811	844	7	1187	1226	8	3197	3367
-1 560	646	1 1172	1157	4	1837	1772	8	2220	2231	8	2551	2648	9	909	955
0 1724	1757	2 4147	3063	5	892	821	9	1147	1239	9	489	600	10	2002	2154
1 1324	1403	3 490	424	6	545	484	10	2349	2433	10	2446	2594	11	1124	1189
2 1506	1371	4 4091	3541	7	1380	1321	11	1883	1913	11	640	842	12	960	1009
3 1348	1130	5 789	714	8	672	679	12	1537	1641	12	1377	1414	14	946	1094
4 770	721	7 3525	3160	9	3191	3137	13	897	903	13	717	623	16	2297	2432
5 1781	1541	8 2112	2035	11	3396	3373	14	1239	1273	15	1410	1461	18	2016	2100
6 1969	1732	9 2505	2442	12	784	1101	16	917	955	16	626	641	20	1227	1223
8 3444	3235	10 3895	3754	13	2534	2545	17	501	540	17	1592	1598	22	739	660
9 1763	1698	11 809	923	15	809	811	19	391	417	18	846	941	1, 4,	L	
10 3007	2804	12 3309	3187	19	546	548	22	763	813	19	950	962	-21	865	809
11 2917	2787	14 2056	1956	21	1082	1086	1, 2,	L	1010	20	559	533	-17	452	432
12 1653	1600	17 350	443	22	613	567	-22	998	1010	1, 3,	L		-16	485	439
13 2147	2102	18 592	599	23	1083	1143	-20	630	655	-22	477	504	-15	905	954
14 743	756	20 764	749	1, 1,	L		-19	751	764	-21	705	757	-14	399	410
16 499	537	21 1021	992	-22	531	491	-17	913	970	-20	550	600	-13	1466	1505
17 665	697	22 804	879	-21	799	857	-16	762	724	-18	537	506	-11	905	858
18 424	443	23 1253	1288	-18	1212	1244	-15	713	736	-15	582	614	-10	983	960
19 906	866	-16	1764	1787	-14	1705	1712			-13	1017	1087	-9	1185	1147

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-12 651 708	5, 14, L	-4 653 714	13 468 570	6 1222 1192	-2 1767 1750
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3 1546 1490	-1 545 494	6, -10, L	2 1563 1661	-15 1451 1328	16 504 435
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14 1586	1464	4	752	771	13	449	528	-11 1934	1965	14	656	670	-8 740	697
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2 674	636	9 1169	1209	-7 852	699	6 881	845	-5 1009	1006	7 1357	1402
5 493	559	10 498	483	-6 1299	1349	8 739	815	-3 499	515	8 1963	2050
7 1308	1344	11 1885	1979	-4 692	692	11 389	407	-2 1544	1541	9 765	856
9 874	911	12 888	873	-3 879	879	12 733	763	-1 563	615	10 1499	1569
12 788	679	13 2032	1943	-2 496	517	13 730	782	0 2404	2327	11 753	744
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9, -4,	L	17 768	726	1 1262	1234	16 614	690	4 480	485	15 503	541
-9 548	605	9, -2,	L	2 2692	2673	17 610	681	5 881	876	16 648	689
-7 690	685	-12 869	851	3 753	677	9, 1,	L	6 1489	1552	17 846	857
-2 867	822	-10 1261	1252	4 1829	1728	-12 1045	967	7 1311	1329	19 762	690
-1 921	897	-9 575	560	5 696	722	-10 1669	1688	8 1410	1408	9, 4,	L
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1 818	816	-7 895	909	9 352	348	-8 1860	1891	10 1155	1266	-11 931	904
2 1467	1457	-5 1110	1110	10 511	483	-7 1061	1129	11 663	768	-7 473	366
5 471	524	-3 460	558	11 1091	1100	-6 997	958	12 663	651	-5 1081	1128
8 1122	1073	-2 1258	1281	12 602	670	-3 419	536	13 571	532	-3 2065	2021
10 1048	1042	-1 781	731	13 1777	1813	-2 964	961	18 621	659	-1 2636	2625
11 1018	993	0 2516	2396	14 816	845	-1 1302	1344	20 496	576	0 433	502
12 1062	992	1 1895	1830	15 1495	1535	0 1906	1869	9, 3,	L	1 2252	2061
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9, -3,	L	4 1165	1196	18 412	413	3 564	567	-11 452	376	4 1389	1389
-11 549	507	5 746	802	9, 0,	L	4 1402	1378	-10 465	439	5 1326	1386
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-4 588	551	12 2078	2122	-9 2627	2599	8 479	494	-5 670	666	10 685	700
-3 983	987	14 2229	2242	-7 2153	2136	9 1013	1046	-4 1292	1349	11 1326	1494
		15 391	411	-5 998	999	11 357	423	-3 375	362	13 1188	1249

9, 4, L	14 659	740	-4 1012	1009	10 1357	1299	9, 11, L	9, 14, L	9 588	559
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16 765	754	-2 1780	1789	-4 441	441	L	-5 763	8 1159	-8 1166	1163
18 601	681	-1 1564	1524	-3 811	758		-4 627	10 770	-6 803	808
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9, 5, L	1 1034	988	1 1034	1 894	838		-2 1158	0 555	1 452	439
-13 920	880	2 1533	1563	1 894	838		-1 647	1 1147	3 433	328
-12 980	945	4 631	616	4 687	590		0 1615	3 844	8 592	632
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-10 865	848	8 527	507	6 778	779		4 938	-4 577	11 643	617
-9 592	576	10 578	485	7 1594	1570		5 837	-2 1372	13 589	554
-8 525	466	12 702	699	9 2136	2143		7 1125	0 1411	10, -4, L	
-6 775	782	14 745	802	10 467	486		8 1507	2 1066	-9 784	675
-4 1848	1794	16 821	769	11 1870	1914		9 1010	4 642	-2 695	728
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1 1495	1447	-3 1379	1360	9, 10, L	L		-3 848	-6 954	8 566	575
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5 860	882	4 585	611	-2 944	922		6 508	1 967	12 486	523
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10 684	764	19 649	622	2 635	658		13 724	8 731	-3 715	674
11 1145	1130	20 728	788	3 1149	1098		9, 13, L	10, -6, L	-2 1183	1154
12 530	476	9, 8, L	L	4 441	543		-4 495	-7 1476	-1 533	531
13 1012	992	-10 866	778	6 1364	1337		-2 776	-5 921	0 1903	1887
15 786	853	-8 946	896	8 2101	2113		0 1210	-4 569	2 1807	1827
9, 6, L	L	-6 596	517	9 1420	1444		2 1087	-2 892	4 1021	1037
-12 1383	1334	-1 678	721	10 1870	1913		4 581	-1 963	7 530	574
-11 622	580	1 537	485	11 1638	1564		9 1028	0 752	8 946	1081
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-8 981	951	7 904	873	13 1085	1110			3 716	12 1955	1974
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		9 462	444					7 654	16 691	692

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-9 495	13 1016	1106	0 784	821	20 654	636	5 635	551	-3 1331	1312
-6 723	14 1196	1234	1 1711	1564	10, 3, L		6 1867	1899	-2 1081	1009
-4 867	15 538	662	2 639	554	-11 625	660	8 2189	2347	-1 1599	1548
-3 980	16 963	1044	3 1271	1305	-6 1078	1027	9 825	768	0 1365	1351
-1 2168	18 535	619	5 883	866	-5 849	861	10 2109	2234	1 1639	1556
0 1060	10, 0, L		6 613	565	-4 529	478	11 988	989	2 821	806
1 2082	-12 954	942	8 1142	1211	-3 1649	1660	12 1503	1593	3 968	1019
2 1745	-10 893	901	10 995	1020	-1 2319	2219	13 850	879	7 896	991
3 1233	-8 1217	1201	12 711	785	0 630	555	16 414	389	8 445	383
4 1741	-6 1225	1297	16 696	682	1 1884	1757	18 864	930	9 1466	1502
7 497	-5 383	492	17 640	643	3 715	687	20 909	1024	11 1063	1043
9 1242	-4 673	678	18 715	723	4 826	840	10, 5, L		13 614	590
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12 854	-1 866	828	-9 697	673	7 1998	2091	-6 605	692	-8 569	566
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15 1222	4 1693	1620	-6 925	903	10 910	999	-3 1565	1506	-2 1358	1273
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-12 765	10 768	867	-3 854	861	14 469	553	0 1466	1425	8 544	575
-11 573	12 701	751	-2 806	751	16 650	633	1 1708	1611	9 627	631
-9 664	14 591	730	-1 1614	1653	17 865	854	2 1260	1156	15 558	721
-7 1010	15 723	756	0 387	365	19 1065	1069	3 641	596	17 900	901
-5 1371	16 565	692	1 1527	1463	10, 4, L		4 936	964	19 918	848
-3 538	17 935	976	3 526	604	-12 865	804	5 888	932	10, 8, L	
-2 1413	18 466	473	5 847	796	-11 824	825	7 1215	1253	-9 860	850
-1 1338	19 703	816	6 938	1041	-10 436	264	8 1294	1313	-8 514	428
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3 2397	-8 449	624	10 1117	1232	-3 603	582	12 1433	1500	-2 484	459
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9 721	-5 650	666	14 503	446	0 2482	2486	-11 1537	1498	6 1040	1089
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10, 8,	10, 11,	11, -6,	9 1502	13 571	11 1413	1604
9 764 742	-6 1064 1030	-3 731 623	11 1852 1926	14 928 987	12 597 578	
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-6 1471 1399	4 774 729	-3 714 646	-2 1326 1305	-1 1128 1138	-5 501 548	
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7 1014 -1059	10 727 734	11, -4, L	5 959 1064	9 439 502	5 589 577	
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-2 615 589	11, -7, L	11, -3, L	-2 497 478	13 385 432	-9 528 519	
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7 1697	1770	1 489	418	11, 9,	L	10 523	508	7 1428	1529	-3 502	508
8 589	552	2 1338	1313	-5 1868	1756	11 522	493	9 1860	1991	-1 707	745
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11 1885	1958	10 1654	1684	3 1484	1414	2 941	983	13 635	657	3 559	591
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14 629	665	15 496	678	7 1670	1643	9 786	795	-6 701	694	-7 844	875
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6 478	491	11, 8,	L	13 725	834	5 578	618	12, 0,	L	-8 653	634
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9 1620	1624	-5 842	822	11, 11,	L	8 1517	1625	-3 600	608	-5 1340	1342
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-6 1434 1433	1 575 536	1 575 536	12, 11, L	L	1 806 767	10 527 421	421	2 773 737	
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-7 1186 1108	2 640 621	2 640 621	6 1374 1441	1441	1 1003 931	-1 428 383	383	7 1149 1222	
-6 624 713	3 1163 1122	3 1163 1122	8 1820 1944	1944	2 1083 1070	10 655 637	637	9 559 591	
-5 856 866	4 684 720	4 684 720	10 1436 1561	1561	3 740 739	13 564 610	610	14, 4, L	
-4 790 757	5 1381 1372	5 1381 1372	13, -1, L	L	4 1332 1349	13, 7, L	L	2 1410 1327	
-3 659 559	7 935 871	7 935 871	-3 920 915	915	5 634 536	4 445 455	455	4 1757 1738	
-2 1045 1003	14 725 663	14 725 663	-2 1252 1230	1230	6 1111 1182	13, 8, L	L	6 1404 1474	
0 963 959	12, 9, L	12, 9, L	-1 715 678	678	8 876 941	1 491 504	504	8 763 811	
7 630 632	-2 829 709	-2 829 709	0 559 606	606	10 893 1008	2 517 483	483	14, 5, L	
8 974 1048	2 1304 1258	2 1304 1258	7 942 955	955	13 477 589	3 819 759	759	2 441 430	
9 1008 1023	4 2088 1981	4 2088 1981	9 1248 1357	1357	13, 4, L	4 1188 1111	1111	3 996 949	
10 1199 1176	6 1797 1814	6 1797 1814	11 811 862	862	-3 1184 1203	6 1161 1157	1157	4 998 904	
11 977 1041	8 1089 1115	8 1089 1115	13, 0, L	L	-2 988 934	8 489 577	577	5 948 893	
12 1106 1173	12, 10, L	12, 10, L	-3 1363 1431	1431	0 1027 932	13, 9, L	L	6 930 941	
13 541 612	1 887 864	1 887 864	-1 976 1001	1001	1 868 833	2 857 820	820	7 740 745	
	3 1609 1554	3 1609 1554	2 612 570	570	2 503 486	3 1668 1585	1585	8 555 605	
	4 705 648	4 705 648	8 502 558	558	3 1197 1226	5 1868 1825	1825		

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14, 6, L
5 820 715
7 665 614

CALCULATED AND OBSERVED STRUCTURE FACTORS OF
[CpMo(PrⁱSC(CO₂Me)C(CO₂Me)=C(CF₃)C(CF₃)=C(CF₃)CCF₃)]

0, 0,	12	354	362	0, 5,	L	2	107	81	0, 10,	L	9	405	421
2 2165 1885	14	410	409	-13 154	164	3	446	436	-8 437	435	11	233	228
4 1853 1703	16	381	367	-11 537	534	5	1207	1155	-6 431	388	13	206	202
6 1099 1019	0, 3,		L	-9 805	765	7	531	538	-4 251	218	15	373	416
8 302 291	-13	208	252	-7 913	848	9	800	785	-2 184	161	0, 14,		L
10 337 315	-11	413	413	-5 1303	1253	13	366	350	0 727	687	-2	360	346
12 285 251	-9	897	878	-3 1900	1749	15	269	264	4 1055	1036	0	283	263
13 178 74	-7	276	272	-1 1145	1066	17	300	294	6 938	928	2	416	362
14 425 411	-5	1709	1621	1 567	496	0, 8,			8 740	735	4	556	493
16 533 513	-3	292	295	3 1193	1138	-10	208	191	10 484	457	6	571	562
0, 1,	-1	963	878	5 472	483	-8	403	374	12 448	450	8	194	189
-15 190 171	1	905	872	7 267	270	-6	156	158	14 282	255	12	131	152
-13 316 316	3	913	881	9 1908	1914	-4	480	429	16 224	212	0, 15,		L
-11 218 216	7	1146	1151	11 560	571	-2	779	746	0, 11,	L	1	461	450
-9 478 478	8	321	139	13 580	567	0	898	855	-5 324	310	3	327	311
-7 459 430	9	1437	1397	15 618	585	2	386	429	-3 140	103	5	274	271
-5 764 736	11	897	916	17 599	613	3	150	137	-1 153	111	7	376	365
-3 671 604	13	365	386	0, 6,	L	4	1445	1427	1 155	152	9	175	183
-1 2546 2323	15	330	313	-10 231	206	6	497	469	3 586	512	11	207	213
5 1533 1433	17	512	511	-6 194	182	7	79	84	5 133	82	0, 16,		L
7 453 461	0, 4,		L	-4 369	332	8	1183	1160	7 557	559	6	249	240
9 419 462	-12	261	269	-1 88	74	12	365	370	11 125	116	8	200	182
11 495 512	-10	154	107	0 543	509	14	172	175	15 250	247	1, -15,		L
13 294 280	-8	370	377	2 226	233	16	363	388	0, 12,	L	-8	268	262
15 536 510	-6	1038	1007	4 188	168	0, 9,			-4 631	593	-6	209	202
0, 2,	-4	1281	1288	8 620	633	-9	316	278	-2 583	539	-4	177	175
-14 476 440	-2	1094	1048	10 279	288	-7	852	800	0 814	738	-2	341	328
-12 604 581	0	1875	1706	12 512	504	-5	480	457	2 793	718	1, -14,		L
-10 88 91	1	75	54	16 400	401	-3	729	693	4 1091	1005	-11	372	366
-8 1214 1196	2	1452	1345	0, 7,	L	-1	227	185	6 293	263	-9	317	223
-6 521 445	4	114	89	-11 433	400	1	265	255	8 288	265	-8	229	224
-4 1846 1755	8	623	602	-9 717	680	3	80	83	10 603	598	-7	196	184
-2 2706 2573	9	122	105	-7 633	576	5	379	383	12 308	302	-3	146	149
0 1933 1732	10	327	322	-5 1339	1243	7	211	96	0, 13,	L	-2	145	155
2 3408 3383	14	378	371	-4 77	82	9	702	718	-3 164	153	-1	110	108
5 86 69				-3 620	612	11	194	163	-1 517	493	1	120	115
6 1599 1541				-1 1093	1007	13	287	301	1 552	503			
8 1243 1131				1 1551	1464				3 628	585			

1, -13,	L	0	657	632	-4	503	531	1, -7,	L	3	540	560	-8	197	193
-12	405	1	144	120	-3	179	198	-16	139	5	580	612	-7	1075	1127
-10	520	2	545	533	-2	693	695	-15	103	6	310	305	-6	655	662
-8	394	3	131	144	-1	131	150	-13	157	7	627	598	-5	347	303
-7	266	4	158	143	0	125	140	-8	169	9	711	705	-4	596	602
-6	269	6	326	309	1	652	668	-7	145	11	531	525	-3	1066	1073
-4	706	1, -10,		L	2	370	371	-6	649	12	141	149	-2	103	104
-3	163	-16	122	134	4	95	69	-5	530	1, -5,		L	-1	128	143
-2	527	-15	101	89	5	133	103	-4	826	-16	244	237	0	795	829
-1	181	-13	199	208	6	179	177	-3	198	-14	266	249	1	878	864
0	490	-12	350	360	8	214	211	-2	527	-13	111	152	2	435	457
1	85	-11	248	259	9	106	100	-1	438	-12	309	306	3	87	120
2	338	-10	364	362	1, -8,		L	1	324	-10	434	454	4	199	188
1, -12,	L	-9	678	638	-15	203	220	2	111	-9	163	162	5	621	629
-13	136	-8	202	204	-14	192	188	3	127	-7	245	338	6	225	239
-12	201	-7	436	450	-13	490	517	4	488	-6	345	321	7	327	341
-10	258	-5	544	495	-12	355	364	6	401	-5	494	464	9	293	284
-8	358	-3	598	579	-11	668	676	8	372	-4	1037	995	10	317	306
-7	243	-2	302	322	-10	237	236	10	153	-3	200	180	11	385	368
-6	231	-1	364	371	-9	753	723	1, -6,		-1	195	197	13	154	192
-5	454	0	155	159	-8	297	288	-17	319	0	1115	1133	14	92	102
-4	263	1	212	183	-7	1007	938	-15	581	1	771	731	1, -3,		L
-3	227	3	497	483	-6	346	334	-14	176	3	530	525	-16	443	435
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2	150	5	152	143	-4	595	552	-12	225	5	509	530	-12	413	414
3	128	7	426	400	-3	632	626	-11	556	6	826	833	-11	177	183
4	154	8	121	119	-2	280	265	-10	230	8	572	562	-10	434	427
5	151	1, -9,		L	-1	841	848	-9	1180	10	373	359	-9	256	300
1, -11,	L	-16	369	397	0	213	200	-8	411	11	87	88	-8	1206	1156
-13	293	-15	121	114	1	751	763	-7	1105	12	459	436	-7	264	250
-13	238	-14	222	226	2	357	356	-5	1135	1, -4,		L	-6	188	201
-12	479	-13	134	139	4	236	214	-4	768	-15	434	456	-5	1062	1029
-10	342	-12	464	481	5	770	765	-3	1310	-14	259	247	-4	1381	1395
-8	757	-10	185	129	6	378	373	-2	728	-13	277	282	-3	341	357
-6	329	-9	146	187	7	528	506	-1	159	-12	214	180	-2	1038	1066
-4	817	-8	245	220	8	200	204	0	397	-11	553	532	-1	1761	1737
-2	676	-7	171	166	9	705	666	1	588	-10	452	476	0	1755	1719
-1	143	-6	700	680				2	158	-9	1083	1089	1	614	582

1, -3, L	1, -1, L	-6 193 302	14 129 109	-9 237 266	4 332 346
2 1226 1254	-16 534 524	-4 420 429	15 174 176	-8 111 87	5 1386 1362
3 226 238	-15 151 163	-3 223 220	1, 2, L	-7 548 595	6 1116 1111
4 1908 1944	-14 484 452	-2 215 230	-13 235 265	-6 257 231	7 142 150
6 621 639	-13 165 134	-1 1870 1839	-12 198 138	-5 284 262	8 300 283
7 212 193	-12 792 778	1 439 425	-11 324 331	-4 94 41	9 1233 1188
8 811 829	-11 98 73	2 475 477	-9 554 589	-3 123 143	10 271 275
10 625 616	-10 556 588	3 690 735	-8 217 232	-2 658 694	11 452 428
12 459 444	-9 240 248	4 724 677	-7 512 526	-1 135 130	12 148 120
14 331 310	-8 1230 1266	5 150 159	-6 237 221	0 964 955	13 450 437
1, -2, L	-7 320 243	7 884 865	-5 1681 1746	2 559 543	15 565 541
-16 197 192	-6 602 639	8 364 355	-4 186 186	4 453 450	17 454 465
-15 321 327	-5 258 263	9 804 787	-3 104 51	5 1821 1799	1, 5, L
-14 296 324	-4 775 873	10 553 590	-2 135 128	6 1624 1588	-13 139 145
-11 533 547	-3 92 43	11 288 299	-1 1249 1252	7 163 147	-12 171 166
-10 133 98	-2 297 297	13 661 625	0 337 316	8 485 464	-10 325 332
-9 381 425	-1 252 243	14 236 238	1 2752 2778	9 957 903	-9 293 308
-8 896 885	0 2958 2914	15 549 502	2 310 324	10 332 326	-8 197 196
-7 454 478	1 252 222	1, L	3 345 327	12 532 540	-6 251 279
-6 168 127	2 1215 1259	-14 480 482	4 174 163	15 286 264	-5 290 301
-5 1467 1454	3 177 200	-13 231 237	5 576 639	16 291 273	-4 566 575
-4 1804 1869	4 122 117	-12 617 637	6 257 287	17 183 176	-1 179 248
-3 449 461	5 118 140	-10 265 273	7 627 618	1, 4, L	0 791 830
-1 621 617	6 1746 1798	-9 575 598	8 208 153	-13 281 316	1 712 732
1 877 867	7 268 283	-8 793 850	9 1271 1255	-11 447 451	2 108 119
2 185 183	8 587 616	-7 362 399	10 248 231	-10 121 182	4 984 971
3 614 642	9 162 175	-6 927 1001	11 692 683	-9 591 597	5 563 579
4 78 115	10 512 507	-3 664 670	12 94 105	-7 746 739	6 322 283
5 240 212	11 121 158	-2 719 742	13 834 775	-6 169 136	7 621 607
6 325 339	12 351 347	-1 525 541	15 531 500	-5 623 663	8 1070 1120
7 235 213	13 158 134	0 687 703	16 190 198	-4 101 116	9 154 104
8 580 567	14 254 255	2 1782 1712	17 596 581	-3 1758 1856	10 1123 1097
9 225 235	1, 0, L	3 195 205	1, 3, L	-2 283 318	11 452 491
10 258 274	-14 190 189	4 424 469	-14 255 265	-1 1320 1356	12 513 501
11 145 147	-13 137 136	5 286 279	-13 101 103	0 649 634	13 186 163
12 343 330	-10 127 101	7 344 376	-12 231 254	1 1696 1692	14 299 290
13 223 216	-8 89 115	8 1030 1025	-11 351 353	2 1621 1603	15 260 237
15 259 233	-7 889 901	9 117 83	-10 150 146	3 1044 1045	16 419 412

1, 6,	L	1, 11,	L	1, 13,	L	1, 16,	L
-11 458	464	-7 177	165	-4 152	166	5 212	212
-10 81	105	-6 434	419	-3 105	107	7 208	201
-9 411	411	-4 154	155	-2 104	111	9 266	261
-7 564	535	-2 520	528	-1 143	130	2, -14,	L
-5 825	813	0 253	227	0 157	152	-10 332	355
-4 515	506	1 155	178	1 187	187	-9 216	209
-3 1052	1117	2 411	383	4 299	289	-8 497	493
-2 376	365	3 296	327	5 242	246	-6 393	377
0 653	652	4 730	663	6 413	404	-5 115	112
1 1119	1187	6 565	541	7 117	102	-4 421	405
2 102	93	8 495	498	8 246	252	-2 256	260
4 235	221	9 231	233	9 382	379	-1 172	187
5 1011	992	10 636	600	10 115	125	2, -13,	L
6 527	510	11 433	475	11 306	313	-8 241	214
7 312	316	12 253	253	13 222	243	-6 182	174
8 229	201	13 358	354	1, 14,	L	-3 143	156
9 124	173	14 184	172	-1 336	320	-2 181	209
10 120	131	17 150	171	0 99	105	1 197	193
11 469	469	1, 12,	L	1 449	431	2 139	133
13 376	395	-5 259	257	2 238	239	2, -12,	L
15 326	295	-3 170	177	3 511	489	-13 189	165
16 170	161	-2 161	183	4 154	129	-12 417	406
17 276	293	-1 477	459	5 667	651	-11 215	202
1, 7,	L	0 159	163	7 225	222	-10 450	427
-11 129	129	1 626	620	8 277	266	-8 477	496
-10 491	491	2 133	134	9 409	403	-7 138	160
-8 620	639	3 678	657	11 507	509	-6 469	457
-6 572	595	5 784	759	12 146	150	-4 345	346
-5 306	335	7 326	324	13 440	437	-2 391	405
-4 551	560	9 719	716	1, 15,	L	0 197	181
-2 390	424	11 285	272	0 116	120	2 272	280
-1 98	107	12 104	114	4 166	176	2, -11,	L
0 616	667	13 476	446	8 226	237	-13 257	264
1 447	452	14 224	215	9 260	283	-12 128	114
2 299	182	15 280	282	10 186	188	-11 442	411
3 843	766	1, 10,	L	12 354	369	-10 255	246
4 1340	1365	17 366	374	-9	532	-9	508

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-8 456	447	0	409	398	-8	167	157	7	296	290	-3	493	481	-16 425	473
-7 296	282	1	1020	1013	-7	845	817	8	459	442	-2	2045	2130	-14 440	474
-5 638	630	2	335	331	-6	534	491	10	499	499	-1	223	242	-13 112	112
-4 262	253	3	602	601	-5	967	912	11	182	174	0	800	851	-12 542	573
-3 448	476	4	182	187	-4	134	132	12	421	449	1	673	681	-11 328	342
-1 463	448	5	625	604	-3	460	456	2, -5,	L	L	2	887	945	-10 666	691
1 195	206	7	339	323	-1	243	244	-16 220	230	230	4	857	890	-9 127	83
3 346	336	9	367	380	0	573	576	-15 193	211	211	5	98	112	-8 1154	1168
4 160	158	2, -8,	L	L	1	489	468	-12 169	189	189	6	726	726	-7 266	224
5 244	222	-15 184	191	191	2	336	345	-11 207	207	207	8	451	465	-6 365	375
2, -10,	L	-13 218	209	209	3	858	843	-10 303	288	288	10	540	532	-5 626	647
-15 265	247	-11 177	142	142	5	597	585	-9 294	328	328	11	208	215	-4 806	843
-14 277	274	-10 150	46	46	7	532	539	-7 457	474	474	12	517	518	-3 822	844
-13 188	186	-9 189	161	161	9	360	362	-6 264	252	252	13	142	143	-2 570	548
-12 287	310	-8 199	157	157	10	88	99	-5 355	365	365	14	413	425	-1 1108	1058
-11 204	221	-6 133	123	123	11	326	313	-3 412	418	418	2, -3,	L	L	0 1830	1821
-10 188	192	-5 384	383	383	2, -6,	L	L	-2 764	748	748	-10 128	123	123	2 1701	1652
-9 164	147	-4 89	74	74	-16 287	294	294	-1 99	147	147	-8 307	296	296	4 973	964
-8 162	159	-2 909	919	919	-14 302	298	298	0 857	846	846	-7 458	425	425	5 304	295
-4 149	160	-1 191	200	200	-13 268	282	282	2 629	639	639	-6 94	93	93	6 925	958
-3 254	267	0 341	336	336	-12 485	466	466	3 278	301	301	-5 1004	1001	1001	7 112	56
-1 228	246	1 414	407	407	-11 110	115	115	4 458	460	460	-3 1038	967	967	8 421	409
1 365	373	2 173	172	172	-10 593	590	590	5 149	159	159	-2 258	238	238	9 636	625
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4 169	190	4 459	451	451	-8 832	782	782	9 154	140	140	0 461	464	464	12 422	405
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-15 348	353	8 323	313	313	-5 451	449	449	2, -4,	L	L	3 459	444	444	2, -1,	L
-13 589	582	9 248	250	250	-4 1050	1043	1043	-16 288	317	317	4 330	325	325	-13 193	201
-11 669	648	10 180	194	194	-3 444	440	440	-14 468	500	500	5 252	165	165	-12 417	436
-10 261	241	2, -7,	L	L	-2 122	92	92	-12 296	287	287	7 535	530	530	-11 231	248
-9 631	571	-16 130	154	154	-1 111	114	114	-11 206	215	215	8 274	299	299	-10 218	247
-8 617	556	-15 420	448	448	0 995	1034	1034	-10 604	640	640	9 215	192	192	-9 318	355
-7 702	668	-13 507	493	493	1 283	281	281	-8 1068	1062	1062	11 238	212	212	-8 110	134
-5 658	644	-12 157	170	170	2 238	253	253	-7 107	65	65	12 236	226	226	-7 597	574
-3 1113	1106	-11 727	715	715	3 455	473	473	-6 1206	1175	1175	13 466	478	478	-6 292	282
-2 346	378	-10 269	268	268	4 702	713	713	-5 151	119	119	14 101	98	98	-5 763	789

2, -1, L	4	513	505	2, 2, L	-2	463	495	10	1055	1087	-1	90	117
-4 1054	6	297	296	-14 155	-1	1174	1202	11	128	90	0	850	897
-3 246	7	683	667	-9 302	0	360	381	12	646	659	1	589	584
-2 1226	8	370	349	-8 131	1	1394	1442	13	410	430	2	815	872
-1 382	9	606	624	-7 566	2	1008	1004	14	524	518	4	1986	2054
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1 1322	11	231	228	-5 577	4	1494	1497	16	488	503	6	1098	1144
2 255	13	137	108	-4 464	5	130	103	2, 5, L			8	1270	1244
3 951	16	137	137	-2 494	6	898	935	-11	524	556	10	1046	1017
4 449	2, 1, L			-1 1015	7	758	787	-10	200	198	12	643	629
5 1471	-15 197	198		0 137	8	327	270	-9	308	326	13	268	270
6 804	-13 417	433		1 731	9	671	646	-7	268	255	14	396	387
7 788	-12 126	25		2 353	10	102	110	-6	350	350	16	501	502
8 295	-11 601	636		3 368	11	418	420	-5	495	508	18	313	314
9 914	-10 431	426		4 1538	12	172	184	-3	129	51	2, 7, L		
10 261	-9 832	921		5 307	13	357	355	-2	262	248	-11	121	94
11 448	-7 1254	1305		6 280	14	402	421	-1	499	534	-9	178	176
13 562	-6 351	334		7 284	15	249	241	0	688	698	-7	90	89
14 192	-5 1196	1265		8 950	17	172	187	1	213	231	-2	536	530
15 456	-4 369	389		9 580	17	172	187	2	583	599	-1	317	315
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-11 534	504	-11	377	384	11	214	228	-2	1395	1351	15	556	587	1	1516	1536
-9 474	446	-10	232	223	12	193	171	-1	956	968	3, -1,		L	3	1411	1403
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2 506	505	1	124	106	-3	633	567	10	177	177	-3	664	680	15	472	461
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7 180	169	6	154	165	1	390	409	-15	133	164	1	371	379	-11	124	108
8 279	285	7	164	172	2	118	146	-14	154	171	2	273	264	-10	503	514
9 238	229	9	120	122	3	485	510	-13	271	283	3	788	801	-9	307	303
3, -7,	L	10	183	201	4	239	266	-12	266	257	4	605	625	-8	155	144
-14 490	488	11	94	94	5	548	555	-11	486	496	5	141	136	-7	510	560
-13 168	192	3, -5,		L	6	113	135	-9	619	639	6	476	453	-6	290	326
-12 554	516	-14	531	540	7	399	407	-7	1080	1098	7	472	449	-5	863	863
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-6 161	135	-9	265	224	10	195	181	-4	383	382	10	357	345	-2	249	233
-4 506	512	-8	348	345	11	349	335	-3	993	1051	13	187	204	-1	131	124
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-2 957	957	-4	923	880	3, -3,		L	-1	823	780	3, 0,		L	1	457	459
-1 482	459	-3	108	108	-14	232	242	0	215	222	-13	197	218	2	723	754
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2 1028	1052	-1	384	398	-12	298	289	2	581	589	-10	477	518	4	1762	1750

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11	265	-11	120	37	3	289	276	-10	144	165	6	816	854	-1	172	170
12	371	-10	536	562	4	691	705	-7	88	70	8	570	569	0	193	209
14	223	-9	102	98	6	907	902	-6	172	170	10	503	494	1	669	679
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13	125	134	16	363	358	-6	237	-8	334	317	5	378	0	144	20
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4	156	145	3, 15,	L	-1	250	256	-12	502	457	-6	344	-10	662	656
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7	246	-6	170	172	11	404	387	-5	679	705	8	501	501	-4	483	503
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13	130	-1	468	454	-12	326	314	0	693	687	16	440	438	3	880	936
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-9	322	5	322	323	-7	302	292	6	133	135	-8	185	189	8	1241	1255
-8	151	6	243	228	-6	734	732	7	420	423	-7	339	333	9	171	103
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7	922	-5	1281	1279	11	241	243	-6	822	808	10	330	316	-5	161	114
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6	297	-6	172	171	10	354	351	4	390	6	403	400	4	213	211
7	695	-5	243	253	11	327	323	5	358	7	175	206	5	263	277
8	505	-4	274	252	12	358	346	6	596	8	216	211	6	295	281
9	1030	-3	936	936	13	600	550	7	154	9	287	299	7	476	503
10	158	-2	241	243	14	552	509	8	214	10	518	534	8	237	226
11	363	-1	492	496	15	236	227	9	406	11	138	115	9	407	434
12	249	1	1239	1242	16	325	322	10	310	12	412	415	10	324	328
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17	182	3	933	960	4, 9,		L	12	271	14	546	614	12	222	237
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-6	425	9	452	459	2	163	172	-6	213	2	231	238	6	121	126
-5	347	10	214	246	3	965	957	-5	207	4	397	395	8	169	171
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-3	420	13	439	437	5	1118	1127	-2	210	8	541	524	11	237	265
-2	229	15	610	627	6	547	536	-1	529	9	398	394	12	315	331
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3	282	-8	265	264	10	129	133	3	532	4, 14,		L	-2	173	205
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5	270	-6	206	181	13	190	202	5	240	0	511	487	-8	266	289
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2 314	339	90	7	737	691	7	671	677	1	345	347	-9	583	559
5, -10,	L	386	8	158	165	9	428	435	3	775	746	-8	115	115
-9 284	288	231	9	238	246	10	273	273	4	421	433	-7	307	343
-4 185	202	L	5, -5,	L	L	11	414	440	5	665	636	-6	419	303
-2 208	230	268	-11	268	282	5, -3,	L	L	6	1048	1030	-5	237	235
0 352	386	508	-10	251	264	-11	183	195	7	507	491	-3	112	130
1 116	152	394	-8	205	229	-9	482	498	8	441	418	-2	313	324
2 250	254	386	-7	340	352	-7	259	262	10	341	331	-1	109	89
4 184	191	616	-6	262	264	-6	766	747	11	317	319	0	244	245
5, -9,	L	308	-5	499	451	-5	346	358	12	204	206	2	474	489
-10 522	495	532	-4	295	305	-4	516	536	5, -1,	L	L	3	1056	1016
-9 136	150	330	-3	546	548	-3	419	385	-12	252	258	4	87	24
-8 250	239	485	-2	191	222	-2	565	538	-11	222	224	5	329	343
-7 209	198	130	-1	590	612	0	694	715	-10	239	220	6	667	638
-6 635	658	177	0	213	218	1	128	109	-9	331	316	7	263	235
-5 245	268	107	3	195	182	2	853	835	-8	276	246	8	533	509
-4 348	364	309	4	228	221	3	245	226	-7	497	485	10	281	287
-3 194	199	268	5	313	292	4	96	133	-6	612	573	11	130	91
-2 444	455	307	6	106	86	5	220	203	-5	472	487	12	290	287
0 254	264	388	8	132	94	6	444	436	-4	561	537	13	122	135
2 288	295	200	9	465	457	7	499	482	-2	1426	1424	14	241	235
4 365	351	L	11	282	279	8	218	229	-1	220	216	5, 1,	L	L
5 171	153	396	5, -4,	L	L	10	536	542	0	383	404	-12	267	285
6 299	320	280	-13	252	244	11	286	274	2	1505	1580	-11	254	253
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-11 133	110	484	-11	382	380	5, -2,	L	L	5	213	188	-9	364	377
-8 137	127	167	-10	196	173	-13	238	244	6	771	771	-8	315	300
-7 281	284	560	-9	590	580	-12	244	224	7	146	35	-6	1150	1121
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-5 118	155	632	-6	150	150	-10	400	427	9	98	54	-2	873	849
-4 383	394	670	-5	326	636	-9	426	422	10	422	425	-1	460	454
-3 316	314	656	-4	326	318	-8	313	325	12	425	425	0	788	815
-2 589	605	628	-3	704	686	-6	188	180	14	472	476	1	278	270
-1 238	221	419	-1	681	703	-5	654	646	5, 0,	L	L	2	266	293
0 369	380	507	1	549	564	-3	270	300	-13	244	252	3	279	269
1 427	402	231	3	901	919	-2	193	197	-12	289	287	4	638	598
2 454	468	725	4	348	328	-1	938	979	-11	161	197	6	1216	1144

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15	218	4	108	151	5, 5,	5, 5,	L	6	182	166	-3	120	90	17	250	248
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-7	269	12	322	328	-5	294	274	13	148	142	4	151	114	-1	288	286
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-4	89	14	290	309	-3	578	554	15	501	486	6	126	115	1	321	324
-3	883	15	437	443	-2	133	103	16	101	128	7	950	918	2	139	115
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-9 213	212	7, 2,	L	14 273	272	7 489	492	0 559	563	0 490	472
-8 348	334	-9 378	374	15 301	284	8 534	509	1 172	165	1 260	274
-7 326	329	-7 134	129	7, 4,	L	12 317	310	2 739	736	2 524	513
-4 320	300	-5 370	325	-9 306	307	13 259	244	4 522	551	3 349	325
-1 344	334	-4 206	187	-8 153	157	14 277	285	5 189	188	4 342	357
0 107	125	-1 682	671	-7 297	286	16 444	419	6 285	283	5 213	196
1 293	291	0 206	200	-6 392	399	7, 6,	L	8 159	157	6 146	153
2 197	-182	1 1053	1026	-4 322	319	-8 193	169	9 125	113	7 383	391
3 715	738	2 249	249	-3 345	326	-7 94	91	10 268	268	9 574	551
4 165	137	3 664	680	-2 370	375	-6 187	163	12 239	238	11 477	461
5 808	739	4 602	581	-1 391	377	-5 202	160	14 402	402	12 229	214
7 534	502	5 818	826	0 443	425	-4 352	338	15 267	254	13 163	180
8 367	367	6 333	357	1 608	643	-3 144	94	16 400	406	14 405	408
9 409	391	7 786	738	3 528	587	-2 565	575	7, 8,	L	15 396	390
10 335	346	8 116	91	5 821	904	-1 202	161	-6 358	358	16 234	227
11 467	476	9 446	435	6 553	583	0 430	426	-5 240	204	17 195	204
13 233	233	11 241	235	7 686	671	1 278	292	-4 261	252	7, 10,	L
7, 1,	L	12 297	307	9 209	194	2 346	361	-3 131	38	-4 408	394
-10 310	341	13 222	225	10 180	186	3 433	449	-2 248	233	-3 263	241
-9 299	300	15 177	168	11 202	200	4 295	318	-1 142	98	-2 234	212
-8 321	333	7, 3,	L	12 427	414	5 190	207	0 384	380	0 437	454
-7 409	389	-9 386	384	14 169	175	6 124	145	2 133	164	1 148	167
-5 300	333	-7 201	197	16 267	292	7 466	468	4 133	149	2 175	201
-4 179	177	-5 605	588	7, 5,	L	8 192	185	7 378	376	3 306	293
-3 133	138	-3 157	165	-7 202	176	9 165	145	8 445	446	4 166	156
-2 280	279	-2 292	292	-6 162	148	10 363	334	9 174	154	5 319	307
-1 401	383	-1 386	387	-5 122	137	12 303	308	10 521	491	7 320	289
0 232	228	0 437	441	-4 421	394	14 430	427	11 473	469	8 247	210
1 170	198	1 731	717	-3 211	164	15 183	175	12 518	502	9 560	554
3 595	587	2 504	534	-2 396	394	16 385	397	13 297	285	10 213	222
5 500	466	4 261	291	-1 516	514	15 271	277	15 271	277	11 432	432
6 205	160	5 437	452	0 681	685	16 313	309	16 313	309	13 601	606

8, 3,	L	10 418	399	10 552	531	11 322	296	12 355	340	8, 12,	L	9, -4,	L
10 418	399	12 431	405	12 377	339	13 355	339	13 355	340	1 453	448	-1 255	260
11 272	265	13 381	364	13 331	316	15 500	486	15 500	486	3 473	459	1 183	203
12 510	511	14 141	143	14 141	143	8, 10,	L	8, 10,	L	4 133	136	2 240	237
14 409	383	15 295	260	15 318	313	-3 217	212	-3 217	212	5 341	335	3 229	249
15 102	17	8, 6,	L	15 318	313	-1 361	352	-1 361	352	6 172	172	4 187	162
-6 208	238	8, 8,	L	15 318	313	0 176	162	0 176	162	7 456	439	5 332	347
-4 365	376	-5 171	148	-5 171	148	1 470	442	1 470	442	9 314	300	6 240	215
-2 274	277	-4 173	140	-4 173	140	2 308	285	2 308	285	12 327	345	9, -3,	L
-1 159	151	-3 224	202	-3 224	202	3 581	562	3 581	562	8, 13,	L	-4 201	219
0 491	493	-2 187	159	-2 187	159	4 174	86	4 174	86	2 256	245	-3 223	217
1 196	188	-1 155	137	-1 155	137	5 572	574	5 572	574	6 196	168	-1 222	230
2 621	647	0 190	184	0 190	184	7 512	494	7 512	494	8 260	269	1 519	542
3 234	246	1 493	500	1 493	500	9 122	108	9 122	108	9 370	350	3 550	546
4 395	412	2 231	212	2 231	212	10 187	208	10 187	208	11 283	285	5 433	446
6 489	522	3 305	295	3 305	295	11 385	376	11 385	376	12 212	240	6 202	186
7 241	220	4 367	378	4 367	378	12 237	244	12 237	244	13 200	198	7 433	443
8 344	356	5 549	566	5 549	566	13 172	178	13 172	178	14 343	347	9, -2,	L
10 371	354	6 148	184	6 148	184	15 173	157	15 173	157	8, 14,	L	-5 291	296
12 278	273	9 520	507	9 520	507	8, 11,	L	8, 11,	L	3 184	184	-4 155	141
13 180	189	10 145	101	10 145	101	-2 186	179	-2 186	179	4 329	327	-3 375	393
14 214	224	11 303	298	11 303	298	-1 207	216	-1 207	216	5 290	283	-2 189	208
16 273	275	14 278	260	14 278	260	0 252	248	0 252	248	6 309	303	-1 242	271
8, 7,	L	15 277	269	15 277	269	1 258	232	1 258	232	8 406	388	0 282	287
-6 417	405	16 163	150	16 163	150	2 104	105	2 104	105	10 245	245	1 442	458
-5 286	284	8, 9,	L	8, 9,	L	3 289	285	3 289	285	12 260	266	2 285	314
-4 364	352	-4 235	253	-4 235	253	5 321	305	5 321	305	8, 15,	L	3 342	359
-3 173	160	-3 442	466	-3 442	466	7 181	191	7 181	191	6 177	163	5 303	327
-2 312	309	-1 325	328	-1 325	328	8 232	242	8 232	242	9 167	168	7 309	294
-1 216	198	1 191	188	1 191	188	9 300	330	9 300	330	10 347	379	9 211	210
0 353	352	2 175	152	2 175	152	11 472	470	11 472	470	9, -5,	L	9, -1,	L
1 330	324	3 407	419	3 407	419	12 210	231	12 210	231	-3 143	142	-5 118	100
2 265	269	4 207	187	4 207	187	13 542	546	13 542	546	-1 266	292	-1 325	318
3 247	255	5 369	386	5 369	386	14 156	165	14 156	165	0 249	273	0 224	221
4 212	195	6 340	343	6 340	343	15 351	343	15 351	343	1 254	253	1 154	157
6 505	544	7 247	262	7 247	262					2 310	324	3 573	590
8 415	407	8 199	199	8 199	199					3 341	361	5 528	491
9 278	262	9 343	327	9 343	327					4 287	294	7 275	246
10 611	574	11 494	468	11 494	468								

9, -1,	L	9, 2,	L	3	251	268	12	337	330	4	214	203	9, 12,	L
8 185	172	-6 336	351	4	121	93	13	237	237	5	786	834	2 478	452
9 234	215	-5 141	99	5	171	180	14	358	348	7	392	391	4 480	462
9, 0,	L	-4 516	527	6	299	297	-4	219	239	9	488	490	6 595	597
-6 194	162	-3 397	389	8	178	177	-3	247	270	10	284	291	8 340	344
-5 337	347	-2 539	519	9	228	214	-2	172	181	11	458	456	9 170	170
-4 305	302	-1 433	410	10	181	158	-2	172	181	12	292	285	10 174	162
-3 365	360	0 654	659	11	184	185	0	216	220	13	295	268	12 165	168
-2 470	472	1 393	389	12	314	333	1	199	199	15	188	166	9, 13,	L
-1 387	395	2 337	331	13	265	281	3	592	598	-1	205	188	4 417	428
0 225	215	3 164	158	14	274	278	4	280	260	0	194	176	6 286	289
1 222	223	4 337	326	9, 5,	L	L	5	496	510	0	194	176	8 279	278
2 416	435	6 412	409	-5	105	121	6	209	201	1	280	286	9 208	201
3 226	-210	8 113	124	-4	336	349	7	637	670	2	154	125	10 132	135
4 590	610	10 233	338	-2	287	282	8	317	319	4	383	397	12 242	249
5 356	345	12 320	328	0	263	270	9	644	646	5	265	291	9, 14,	L
6 318	299	9, 3,	L	4	148	132	11	631	592	6	193	195	7 220	232
8 306	319	-6 265	265	5	392	405	12	203	198	7	238	268	8 203	208
9 244	257	-5 142	126	6	131	119	13	409	384	9	261	253	10 234	253
10 298	288	-4 254	252	7	554	592	15	196	183	13	188	198	10, -3,	L
11 268	288	-2 235	239	8	223	220	-3	327	335	0	159	170	0 225	252
9, 1,	L	-1 155	113	9	410	378	-2	218	215	1	207	195	1 141	152
-6 304	309	0 504	512	11	372	378	-2	218	215	1	207	195	2 142	153
-4 138	144	1 292	296	13	321	327	-1	241	235	2	258	257	10, -2,	L
-3 255	234	3 535	538	9, 6,	L	L	0	230	239	3	290	289	0 154	150
-1 365	356	-4 320	322	-4	320	322	3	291	282	4	298	298	1 248	262
1 527	541	-2 348	362	-2	348	362	4	158	67	5	291	290	3 296	325
3 383	378	-1 154	183	-1	154	183	5	252	261	6	253	260	4 206	214
4 330	322	0 167	159	8	245	246	8	245	246	7	521	527	5 332	341
5 487	460	1 186	191	9	383	357	9	383	357	8	336	335	10, -1,	L
6 152	119	2 168	143	11	237	204	11	237	204	9	264	278	-2 368	369
7 177	142	3 215	216	13	380	364	13	380	364	10	134	131	0 406	414
8 267	270	6 251	287	15	310	321	15	310	321	11	312	324	1 158	155
9 167	158	7 160	144	9, 9,	L	L	9, 9,	L	L	12	409	411	2 338	362
10 231	209	8 241	205	-2 111	115	115	-2 111	115	115	14	146	140	3 146	148
12 331	326	-2 418	413	-1 290	286	286	-1 290	286	286	4	231	241	4 231	241
		0 360	356	1 299	291	291	1 299	291	291	5	371	374	5 371	374
		2 639	639	3 525	524	524	3 525	524	524					

10, 0,	L	290	280	10, 8,	L	10, 12,	L	11, 7,	L
0 287	295	1 160	158	-1 182	196	6 324	357	3 202	176
2 285	295	2 420	428	3 183	149	8 135	135	5 385	387
3 243	213	5 223	237	4 350	357	11, 1,	L	7 260	278
4 416	419	7 275	259	5 276	267	2 356	362	8 119	143
5 353	355	9 268	262	7 155	125	3 262	282	9 257	250
8 273	242	10 154	177	8 302	302	4 254	264	10 294	326
10, 1,	L	11 190	188	9 244	193	5 197	174	11, 8,	L
-2 457	468	10, 5,	L	10 232	231	11, 2,	L	4 396	384
0 411	407	-3 237	228	11 338	323	1 243	278	6 315	315
1 145	93	-2 137	135	12 128	124	3 211	242	8 167	163
2 349	345	0 147	135	13 379	347	6 343	330	11, 9,	L
4 437	437	1 224	232	10, 9,	L	11, 3,	L	6 211	207
5 114	109	2 245	254	1 243	253	1 366	353	7 182	154
6 335	316	7 195	188	4 455	470	2 177	156	8 361	398
9 150	134	9 292	289	5 346	360	3 264	257		
10, 2,	L	11 303	276	6 434	445	4 190	181		
-3 135	142	10, 6,	L	7 216	216	5 502	494		
0 582	597	-2 204	200	8 354	350	6 160	148		
2 181	208	-1 163	127	9 251	264	7 285	272		
4 189	187	0 288	285	10 373	389	11, 4,	L		
5 119	83	1 293	302	11 185	170	1 364	389		
6 459	447	3 246	230	12 444	465	4 194	206		
10 182	175	5 318	328	10, 10,	L	11, 5,	L		
10, 3,	L	7 185	196	2 142	120	1 357	380		
-2 336	345	8 380	407	4 397	390	3 340	326		
-1 140	109	9 331	340	6 281	290	5 448	431		
0 392	403	11 391	394	8 281	293	6 164	154		
2 276	269	12 171	190	9 222	224	7 367	369		
4 266	269	13 362	362	11 256	232	9 317	325		
6 175	163	10, 7,	L	10, 11,	L	11, 6,	L		
8 149	153	-1 232	236	4 311	319	1 163	178		
9 237	227	3 284	287	5 153	70	2 159	154		
11 226	228	5 202	215	6 404	411	4 278	273		
10, 4,	L	8 245	248	8 431	433	8 185	187		
-3 177	185	9 479	508	10 361	359	9 221	223		
-2 329	352	10 305	303	12 309	332				
-1 307	320	13 206	202						

CALCULATED AND OBSERVED STRUCTURE FACTORS OF
[CpMo(F₃CCC(CF₃)C(CF₃)=C(CO₂Me)C(SPrⁱ)C(OMe)=O)]

0, 0, L	14 132 158	0, 5, L	10 542 567	15 261 282	7 129 161
2 3426 2917	15 247 244	1 583 558	11 134 125	0, 11, L	8 194 212
6 1169 1069	16 94 118	3 810 824	16 275 301	1 199 183	9 226 241
8 687 621	17 173 185	4 118 121	0, 8, L	3 176 180	10 176 181
10 291 275	0, 3, L	5 913 885	1 763 765	4 791 804	11 281 301
12 854 794	1 421 430	7 776 714	2 560 557	6 407 394	12 178 186
14 607 572	2 1211 1190	6 781 725	3 815 814	8 373 366	0, 15, L
18 170 184	3 669 592	8 543 513	4 140 127	9 134 131	2 256 265
0, 1, L	6 1005 969	9 799 782	6 166 165	10 329 340	4 417 419
2 1149 1016	7 788 753	12 205 183	10 672 659	0, 12, L	6 649 669
3 179 172	8 755 702	13 109 133	11 407 403	0 299 315	8 440 490
4 1381 1230	9 407 396	14 139 140	12 515 530	1 591 549	9 120 131
5 684 654	11 453 447	15 120 133	14 182 183	2 496 501	10 212 241
6 1060 954	12 221 218	17 269 285	15 134 109	3 290 273	0, 16, L
7 1202 1088	13 274 286	0, 6, L	16 257 279	4 323 330	0 244 249
8 918 862	14 213 232	0 191 197	0, 9, L	6 384 381	1 278 278
9 535 508	15 147 146	1 1354 1405	1 309 303	8 436 450	2 225 227
11 106 106	16 265 297	2 385 395	2 271 277	10 296 316	0, 17, L
12 328 308	0, 4, L	3 566 562	3 530 537	11 211 217	1 162 160
13 201 193	0 1427 1447	4 397 392	4 303 313	12 212 219	3 219 240
15 198 199	1 216 209	6 190 171	5 902 901	13 168 173	5 287 276
16 148 142	2 844 845	8 269 249	6 504 513	0, 13, L	7 331 343
17 133 131	3 309 299	10 445 415	7 209 205	3 508 497	0, 18, L
18 93 109	4 837 781	11 621 626	8 267 267	4 189 181	0 193 174
0, 2, L	5 628 620	12 338 318	9 506 522	5 415 406	4 110 123
0 1075 1031	6 462 447	13 429 430	10 263 258	6 305 306	1, 0, L
1 1731 1708	7 418 377	15 238 242	11 242 245	7 210 209	-18 291 309
2 98 86	8 417 395	16 228 240	12 209 211	8 166 151	-16 339 337
3 65 44	9 390 377	17 140 161	0, 10, L	9 399 419	-14 490 468
4 367 331	10 272 250	0, 7, L	0 1030 1028	10 115 133	-12 216 232
5 1252 1146	11 380 349	2 124 138	1 606 618	11 196 219	-10 572 502
6 213 205	12 347 332	3 335 334	2 416 422	0, 14, L	-8 1855 1772
7 242 241	13 123 160	4 711 734	3 941 938	0 445 432	-4 859 843
8 230 204	14 474 452	5 410 405	9 236 241	1 349 339	-4 2665 2397
9 529 500	16 229 240	6 449 458	10 277 264	2 309 315	2 677 651
10 294 277		7 492 482	11 467 488	3 447 444	4 3097 2578
11 474 466		8 771 757	12 250 254	4 377 388	6 1075 914
13 436 405		9 162 139	13 153 176	6 197 215	8 257 206

1, 0, L	1, 2, L	-11	308	300	-3	1476	1425	5	262	249	1, 7, L
10 1114 1027	-17 359	-10 122	116	116	-2 407	396	6 488	6 488	478	-16 338	347
12 186 177	-14 165 158	-9 510 484	308	300	-1 1173 1200	396	7 350	7 350	367	-14 365	381
14 320 323	-13 325 331	-8 320 313	308	300	0 960 1019	1019	8 240	8 240	219	-12 641	603
16 175 161	-12 129 146	-7 131 135	308	300	1 127 119	119	9 478	9 478	466	-11 290	255
1, 1, L	-11 183 168	-6 544 517	308	300	2 587 603	603	10 150	10 150	141	-10 191	184
-17 151 162	-10 245 254	-5 540 490	308	300	3 507 506	506	11 384	11 384	354	-8 173	144
-16 101 79	-9 874 815	-4 1226 1173	308	300	4 385 410	410	12 354	12 354	361	-7 626	579
-15 256 233	-8 172 156	-3 1033 988	308	300	5 366 349	349	13 363	13 363	365	-6 228	240
-14 249 231	-7 451 463	-2 2337 2319	308	300	6 889 896	896	14 98	14 98	115	-4 1113	1092
-13 371 358	-6 405 381	-1 93 113	308	300	7 718 694	694	15 347	15 347	375	-3 370	363
-12 471 463	-5 106 69	0 969 997	308	300	8 575 557	557	1, 6, L	1, 6, L	L	-2 946	956
-11 307 293	-4 1144 1036	2 173 182	308	300	10 544 522	522	-15 163	-15 163	177	-1 1326	1326
-10 938 -900	-3 89 69	3 1624 1571	308	300	11 259 232	232	-12 159	-12 159	160	0 1048	1051
-9 324 301	-2 981 976	4 982 898	308	300	13 152 160	160	-11 326	-11 326	300	1 899	925
-8 702 646	-1 311 328	5 272 265	308	300	14 426 396	396	-9 623	-9 623	571	2 582	567
-7 681 626	0 220 228	6 479 448	308	300	15 176 169	169	-8 198	-8 198	170	3 647	643
-6 869 814	1 245 255	7 585 530	308	300	16 110 107	107	-7 822	-7 822	745	4 529	528
-4 303 276	2 1336 1269	8 616 588	308	300	1, 5, L	L	-5 1161	-5 1161	1131	5 416	422
-3 229 235	3 1798 1671	10 929 895	308	300	-17 271 270	270	-4 650	-4 650	639	6 298	317
-2 2467 2468	4 814 753	11 324 312	308	300	-16 208 187	187	-3 306	-3 306	295	7 598	594
-1 1302 1334	5 1836 1668	12 390 380	308	300	-15 284 256	256	-1 487	-1 487	471	8 184	191
0 2366 2427	6 501 445	13 131 122	308	300	-14 201 205	205	0 733	0 733	753	9 203	209
1 459 454	7 1192 1098	14 217 212	308	300	-13 574 563	563	1 348	1 348	364	10 614	630
2 935 884	8 218 187	15 156 153	308	300	-11 362 337	337	2 792	2 792	814	11 344	370
3 358 326	9 783 716	16 161 168	308	300	-10 318 304	304	3 1612	3 1612	1650	12 394	402
4 929 812	10 337 325	17 194 217	308	300	-9 127 131	131	4 534	4 534	565	13 318	334
5 1318 1118	11 695 671	1, 4, L	308	300	-7 305 278	278	5 997	5 997	1037	14 265	290
6 657 544	13 114 124	-18 283 284	308	300	-5 1366 1339	1339	6 305	6 305	309	16 139	158
9 563 520	14 159 183	-16 297 303	308	300	-4 663 646	646	7 819	7 819	845	1, 8, L	L
10 155 151	15 193 199	-12 221 216	308	300	-3 525 519	519	8 408	8 408	415	-12 405	395
11 589 542	17 204 222	-10 509 494	308	300	-2 757 750	750	9 464	9 464	472	-11 345	328
12 603 559	1, 3, L	-9 152 157	308	300	-1 953 972	972	12 377	12 377	376	-10 521	491
13 128 114	-16 454 454	-8 872 860	308	300	0 794 789	789	15 210	15 210	236	-9 222	206
14 189 195	-15 269 263	-7 842 793	308	300	1 915 934	934	16 180	16 180	173	-8 575	551
16 135 154	-14 515 496	-6 280 283	308	300	2 338 338	338				-7 714	722
17 134 120	-13 108 134	-5 793 763	308	300	3 504 503	503				-6 742	765
	-12 324	-4 151	308	300	4 417	417				-5 543	558

1, 8,	1, 10,	14, 276	1, 14,	-4, 187	208	16, 184	184
-3 351 351	-11 557	1, 12,	-11 255	-3 197	201	2, 1,	L
-1 221 205	-10 262	-12 208	-10 176	1 196	195	-18 267	285
0 209 182	-9 314	-11 101	-9 322	3 156	156	-17 160	173
1 302 299	-8 312	-10 388	-8 372	4 158	112	-13 202	201
2 575 619	-7 588	-9 231	-7 429	5 291	264	-12 138	145
3 678 705	-6 362	-8 364	-6 380	6 208	157	-11 332	289
4 640 645	-5 688	-7 177	-5 385	7 188	190	-10 791	754
6 923 929	-4 251	-6 570	-4 501	8 188	168	-9 825	817
7 406 421	-3 146	-5 207	-3 241	9 138	140	-8 449	429
8 709 719	-2 205	-4 394	-2 200	1, 17,	L	-7 411	393
9 188 208	-1 493	0 192	-1 279	-5 250	262	-6 845	735
10 240 254	2 267	2 563	0 164	-3 418	422	-5 211	195
12 230 242	3 273	3 468	3 126	-1 430	436	-4 381	399
14 148 163	4 594	4 375	4 607	1 186	189	-3 623	660
16 217 244	5 146	5 282	6 658	3 167	140	-2 566	605
1, 9,	6 175	6 218	7 340	4 159	140	-1 754	780
-16 194 196	7 622	8 409	8 220	6 154	142	0 857	888
-15 236 225	9 208	9 248	9 180	1, 18,	L	1 1330	1361
-13 613 609	15 196	1, 13,	1, 15,	4 267	252	2 128	59
-11 306 292	1, 11,	-13 279	-10 166	2, 0,	L	3 1016	1021
-8 227 220	-14 252	-11 240	-7 175	-18 273	278	4 861	873
-7 465 437	-12 388	-9 217	-4 258	-16 327	317	5 530	503
-6 146 152	-10 232	-8 261	-2 518	-14 404	412	6 112	93
-5 319 321	-8 405	-7 227	0 530	-12 846	826	7 799	732
-4 393 376	-5 135	-5 385	2 490	-10 661	625	8 1070	932
-3 1084 1121	-4 513	-4 157	3 143	-8 413	360	9 726	643
-2 221 237	-2 553	-3 477	4 185	-6 501	435	10 398	360
-1 1135 1165	-1 102	-1 397	6 136	-4 1289	1299	11 217	197
0 265 259	0 645	1 673	7 163	-2 507	515	12 145	121
1 1115 1129	1 503	3 429	8 198	0 879	969	13 279	271
3 169 184	2 354	4 404	10 254	2 483	504	15 224	226
4 166 142	3 197	5 199	1, 16,	4 820	785	16 281	288
9 693 720	7 301	9 283	-9 184	6 1256	1136	17 207	222
10 277 304	8 433	11 252	-8 219	8 1263	1094	2, 2,	L
11 326 367	9 259	277	-7 178	10 1319	1135	-17 407	413
13 164 171	10 251	242	-6 318	12 434	381	-16 183	190
15 265 306	12 107	133	-5 213	14 389	391	-15 446	452

2, 10,	L	537	2, 15,	L	3, 0,	L	6 637	660	3, 3,	L
3 302	278	184	-10 267	293	-18 128	160	7 240	242	-17 207	214
6 321	319	194	-8 313	323	-14 348	346	8 953	906	-16 285	285
7 613	594	201	-6 433	456	-12 1078	1020	9 602	580	-14 458	451
8 352	352	333	-4 400	408	-10 646	615	10 377	357	-13 580	563
9 345	348	282	0 385	373	-8 553	511	11 325	304	-9 409	424
10 221	221	144	2 469	469	-6 1259	1322	12 229	211	-8 204	215
11 231	235	276	3 148	130	-4 209	215	13 501	485	-7 251	254
12 205	209	157	4 453	425	-2 428	453	14 279	280	-6 697	732
13 308	336	L	6 459	431	0 601	690	15 95	99	-5 896	942
-13 194	201	218	8 213	218	2 907	978	16 100	110	-4 577	619
-9 377	390	L	2, 16,	L	4 184	212	3, 2,	L	-3 543	579
-7 410	408	214	-6 203	214	6 1337	1289	-15 349	349	-2 530	550
-5 173	165	223	-5 225	223	8 599	532	-13 582	569	-1 786	832
-3 233	238	147	-4 163	147	10 133	121	-11 762	723	0 863	909
4 135	139	384	-3 381	384	12 263	255	-10 249	246	1 1144	1149
5 584	537	269	-2 277	269	14 429	417	-9 805	772	2 499	509
6 164	169	350	-1 352	350	16 372	392	-8 398	387	3 1046	1071
7 488	484	404	0 415	404	3, 1,	L	-7 955	984	4 995	983
-12 218	245	340	1 354	340	-18 232	233	-5 308	328	5 281	305
-9 133	151	256	3 274	256	-17 251	254	-4 282	277	6 893	892
-7 159	183	L	2, 17,	L	-16 133	167	-2 1512	1555	8 652	635
-5 413	419	298	-7 285	298	-15 404	421	-1 274	304	9 543	517
-4 280	272	233	-5 245	233	-14 618	627	0 261	271	10 492	473
-3 342	356	230	1 232	230	-13 336	334	1 132	148	11 220	214
-2 327	328	175	2 162	175	-12 605	589	2 454	442	12 409	415
-1 577	584	320	3 350	320	-10 112	125	3 218	211	13 318	315
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1 298	297	L	2, 18,	L	-6 1667	1691	5 998	1005	3, 4,	L
2 401	378	224	-5 195	224	-5 612	652	6 198	194	-17 178	199
7 204	193	254	-4 216	254	-4 1265	1350	7 620	632	-12 511	526
9 176	176	244	-3 226	244	-3 920	978	10 151	148	-10 687	640
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11 171	169	190	-1 185	190	0 1162	1218	13 367	366	-8 1057	974
-2 495	498	168	0 188	159	1 578	601	15 378	381	-7 304	309
-4 572	567	145	1 145	168	2 659	691	-6 1306	1431	-6 664	672
-2 495	498	145	4 456	472	3 282	296	-5 664	672	-4 133	145

3, 4,	L	10 182	168	-1 123	144	-8 201	180	10 191	207	8 200	177
-3 333	332	11 419	397	0 570	588	-7 559	563	12 233	252	10 133	141
-2 335	328	13 263	266	1 151	144	-5 469	471	13 154	171	12 144	141
-1 515	550	14 233	232	2 614	643	-3 1056	1090	3, 11,	L	3, 13,	L
0 302	313	3, 6,	L	3 219	205	-2 251	250	-14 264	276	-11 177	169
1 711	700	-14 257	251	4 339	332	-1 663	662	-12 265	273	-9 206	230
2 1407	1427	-13 325	314	5 173	178	0 203	216	-11 297	312	-7 415	433
3 424	413	-11 693	660	6 276	285	1 410	428	-10 140	130	-5 343	349
4 1268	1260	-10 315	295	8 227	244	2 263	267	-8 334	328	-3 683	674
6 843	846	-9 448	390	9 179	185	3 273	307	-7 269	252	-2 327	335
7 300	313	-8 467	469	10 564	576	5 648	620	-6 610	625	-1 601	578
8 263	251	-7 919	912	12 326	346	7 617	607	-5 122	123	1 432	422
9 173	172	-6 455	450	3, 8,	L	8 170	172	-4 939	957	5 241	247
12 192	179	-5 980	1053	-15 116	135	9 449	438	-3 252	253	7 304	280
14 285	281	-3 183	211	-14 198	182	10 106	115	-2 738	728	9 370	348
16 305	334	-2 159	139	-12 389	386	11 504	508	-1 352	347	11 366	398
3, 5,	L	-1 499	510	-10 546	516	13 179	181	0 207	203	3, 14,	L
-17 221	222	0 694	653	-9 416	395	3, 10,	L	1 311	295	-12 193	225
-15 445	459	1 820	857	-8 481	450	-13 195	206	2 257	269	-10 213	225
-14 245	235	2 550	516	-6 745	783	-12 179	174	6 304	295	-9 246	246
-12 460	439	3 760	787	-4 597	600	-11 311	337	8 422	413	-8 254	270
-11 338	351	4 313	306	-3 197	209	-10 379	375	10 479	498	-6 463	464
-9 369	395	5 729	732	-1 603	647	-9 534	510	12 356	386	-2 242	259
-8 149	128	7 488	464	0 314	333	-8 158	155	3, 12,	L	-1 337	304
-7 819	819	8 162	196	1 511	514	-4 156	139	-14 116	136	0 444	452
-5 290	318	11 228	217	2 550	571	-3 188	203	-12 133	142	1 193	162
-4 705	730	13 255	251	3 483	456	-2 300	304	-11 300	333	2 444	437
-3 472	479	15 240	254	4 787	742	-1 392	396	-10 366	372	3 421	393
-2 409	429	3, 7,	L	5 697	661	0 265	291	-9 346	375	4 368	348
-1 234	248	-16 366	345	6 600	593	1 281	260	-8 298	292	5 601	533
0 1133	1161	-14 311	282	7 222	234	2 297	275	-2 367	372	6 378	359
1 533	509	-10 327	318	8 286	300	3 587	543	-1 187	203	7 195	174
2 378	370	-7 187	206	13 230	257	4 192	173	0 481	471	3, 15,	L
3 476	495	-6 201	218	3, 9,	L	5 727	678	2 267	244	-6 277	280
4 383	359	-5 309	328	-15 324	339	6 274	272	3 215	180	-4 390	376
5 688	699	-4 775	823	-13 251	253	7 304	307	4 657	615	-3 170	163
7 391	380	-3 269	287	-11 326	330	8 450	457	5 449	414	-2 503	508
8 523	514	-2 1012	1094	-9 177	144	9 400	384	6 461	450	0 459	433

3, 15,	L	10	260	248	-5	961	1018	15	98	97	0	968	1012	-11	221	221	221
8 236	236	12	443	431	-3	919	975	-17	155	167	1	626	681	-10	545	502	502
3, 16,	L	-15	203	215	-2	411	416	-16	307	312	2	716	717	-8	399	433	433
-9 250	253	-14	702	685	-1	1211	1278	-14	305	292	3	998	971	-6	775	840	840
-7 314	341	-13	532	508	0	472	455	-9	280	270	4	620	614	-5	287	314	314
-5 314	325	-12	1063	1013	1	682	697	-8	690	732	6	328	336	-4	177	160	160
-4 162	177	-11	477	414	2	166	167	-7	378	390	8	105	95	-2	226	274	274
0 303	301	-10	597	530	4	587	620	-6	1189	1287	10	246	244	0	139	182	182
1 240	229	-9	545	578	6	323	323	-5	250	277	11	412	433	1	354	356	356
2 271	249	-8	665	686	7	422	425	-4	1095	1147	12	141	139	2	571	592	592
3 304	285	-7	695	731	8	105	122	-3	831	852	13	233	241	3	209	215	215
4 356	331	-6	113	137	9	397	405	-2	1262	1291	14	224	233	4	776	773	773
5 243	222	-5	291	300	11	369	371	-1	204	212	15	305	317	5	181	186	186
6 214	200	-4	199	206	12	157	172	0	868	879	4, 6,	L	L	6	461	463	463
3, 17,	L	-3	487	494	13	324	324	0	271	259	-17	273	283	7	410	425	425
-7 162	179	-2	608	613	14	262	279	2	375	345	-15	334	325	8	165	156	156
-5 326	333	-1	546	567	-15	210	228	3	94	83	-12	178	144	10	206	205	205
-3 388	387	0	1286	1336	-13	426	467	4	177	204	-8	259	249	12	248	267	267
-1 343	347	1	726	757	-12	225	242	6	419	418	-7	764	806	14	186	200	200
1 153	169	2	291	311	-10	521	518	8	679	677	-5	630	684	4, 8,	L	L	L
3, 18,	L	3	296	310	-9	805	812	10	281	281	-3	1221	1313	-16	306	298	298
0 228	218	4	474	496	-8	874	901	11	202	214	-2	230	226	-14	281	268	268
1 177	181	5	1004	1038	-7	439	493	12	479	490	-1	330	350	-12	309	297	297
2 319	306	6	656	675	-5	439	452	14	183	169	0	237	242	-9	269	254	254
4, 0,	L	7	372	374	-4	154	140	4, 5,	L	L	2	167	159	-8	301	325	325
-18 132	135	11	145	155	-3	426	434	-14	237	232	4	758	759	-7	211	234	234
-16 493	519	12	194	198	0	351	363	-13	141	164	5	423	431	-6	255	249	249
-14 604	599	13	365	369	1	86	101	-12	192	198	6	196	174	-5	248	266	266
-10 207	191	14	376	381	2	501	518	-11	823	775	7	652	639	-4	624	650	650
-8 1057	1058	15	160	190	3	828	830	-9	182	185	8	261	262	-3	458	451	451
-6 1954	2007	4, 2,	L	L	4	644	663	-8	427	476	9	616	630	-2	686	705	705
-4 1484	1570	-17	202	199	5	1012	1043	-7	413	423	10	358	347	0	207	216	216
-2 1694	1833	-15	401	396	6	247	257	-5	411	461	11	438	466	1	667	651	651
0 1038	1057	-13	211	200	8	92	119	-4	214	233	4, 7,	L	L	2	425	407	407
2 436	440	-8	138	117	12	190	198	-3	572	566	-16	220	234	3	167	174	174
4 231	258	-7	509	552	13	226	223	-2	121	134	-13	100	75	4	437	407	407
6 426	397	-6	137	146	14	246	252	-1	377	382	-12	343	313	5	346	313	313

4, 8,	L	0	374	367	-1	204	179	4, 15,	L	5, 1,	L	0	254	256
6	241	1	469	463	1	174	177	-10	250	-18	229	1	619	683
7	162	2	398	386	3	222	228	-8	367	-17	239	2	244	229
8	280	3	176	169	4	252	228	-2	204	-16	418	3	1092	1140
9	271	5	340	318	5	187	204	0	462	-15	272	5	771	815
10	385	6	204	187	6	160	146	2	525	-12	193	6	189	170
12	265	7	354	340	8	320	318	4	465	-10	152	11	269	273
14	155	8	434	441	9	265	255	6	200	-9	318	13	222	229
4, 9,	L	9	478	469	10	350	357	4, 16,	L	-8	689	5, 3,	L	L
-15	197	10	290	302	11	235	224	-7	264	-7	482	-17	252	255
-13	209	11	185	193	4, 13,	L	L	-6	184	-5	900	-16	276	264
-11	644	11	185	193	-13	176	173	-5	221	-4	1034	-14	172	169
-10	161	-14	198	189	-11	358	372	-4	258	-3	1158	-10	213	222
-9	374	-12	472	491	-9	409	398	-3	277	-2	827	-9	610	619
-5	367	-10	627	606	-8	213	226	-1	321	-1	373	-8	586	612
-3	263	-9	127	135	-7	360	346	0	172	0	262	-7	214	199
-2	727	-8	247	253	-6	110	122	4, 17,	L	1	362	-6	1131	1189
-1	376	-4	198	223	-1	171	168	-1	228	2	239	-5	301	322
0	329	-2	430	422	0	224	221	0	191	4	388	-4	1063	1138
1	316	-1	119	116	1	388	371	1	354	5	378	-3	532	556
3	902	0	493	479	3	691	651	3	300	6	734	-2	229	263
5	272	2	655	603	5	521	489	5, 0,	L	7	241	-1	211	197
7	185	3	399	379	7	284	275	-14	400	8	146	3	322	318
9	178	4	609	577	8	188	191	-12	465	10	441	5	507	541
11	356	5	207	214	4, 14,	L	L	-10	650	11	250	6	556	566
13	138	6	363	347	-6	393	413	-8	427	13	115	7	238	248
4, 10,	L	8	293	285	-5	307	284	-6	282	14	176	8	486	487
-15	148	10	155	158	-4	309	301	-4	1187	5, 2,	L	9	108	103
-13	359	12	216	238	-3	420	403	0	369	-15	313	10	161	141
-11	343	12	216	238	-2	444	428	2	218	-13	331	11	181	167
-9	362	-13	199	197	-1	388	375	4	1029	-11	335	12	379	403
-7	314	-12	293	289	0	292	272	6	465	-10	132	5, 4,	L	L
-6	232	-9	371	375	1	269	255	10	125	-9	574	-16	266	275
-5	701	-8	201	211	5	161	147	12	380	-5	91	-14	359	329
-4	594	-5	228	225	7	196	198	14	249	-4	766	-13	458	462
-3	652	-4	391	378	8	212	216	14	249	-3	518	-12	227	244
-2	306	-3	297	285	9	333	348	14	249	-2	697	-11	372	385
-1	307	-2	253	256	9	333	348	14	249	-1	1371	-10	581	608

5, 4,	L	8	164	166	-2	656	675	5, 9,	L	-7	218	237	5, 14,	L
-9	169	9	412	412	-1	228	223	-15	210	-6	483	455	-11	112
-8	412	10	400	431	0	787	801	-14	149	-5	325	326	-10	386
-7	354	11	234	234	1	205	184	-13	260	-4	581	564	-9	284
-5	118	12	93	114	2	124	138	-11	166	-3	158	148	-8	202
-3	305	13	156	155	4	133	139	-7	519	-2	330	328	-7	371
-2	745	5, 6,	L	L	5	425	413	-6	173	0	165	174	-3	158
-1	324	-13	428	451	6	343	336	-5	1095	1	183	194	-2	338
0	708	-12	359	378	7	217	213	-1	288	2	275	274	0	344
1	905	-11	757	786	8	522	532	0	147	4	333	334	1	437
2	938	-9	335	366	9	249	259	1	120	6	409	384	2	294
3	169	-8	231	268	10	258	272	2	198	7	271	265	3	218
4	1232	-7	159	204	11	321	356	3	472	8	409	412	4	236
5	255	-5	209	241	12	256	288	5	133	5, 12,	L	L	5	255
8	178	-4	309	337	5, 8,	L	L	6	222	-13	148	153	7	231
9	342	-3	419	427	-16	168	198	7	651	-12	314	313	5, 15,	L
12	416	-1	299	316	-14	247	224	9	314	-11	216	209	-8	233
14	252	0	302	328	-13	360	348	12	145	-9	193	200	-6	286
5, 5,	L	1	1172	1201	-12	401	385	5, 10,	L	-8	210	199	-4	274
-17	190	2	291	317	-10	297	299	-14	227	-2	218	210	-2	248
-16	179	3	934	924	-9	152	150	-13	437	-1	263	261	6	313
-15	415	4	235	209	-8	295	304	-12	324	0	345	316	5, 16,	L
-13	152	5	141	150	-7	370	394	-11	462	1	243	228	-1	257
-11	256	6	318	301	-6	417	426	-9	525	2	635	600	0	192
-10	109	7	268	263	-4	206	216	-7	523	3	202	165	1	235
-9	329	9	224	217	-2	778	785	0	386	5	201	187	2	188
-8	155	11	308	325	0	719	710	1	616	6	242	211	3	361
-7	756	13	277	307	1	215	210	2	371	5, 13,	L	L	4	149
-5	1148	5, 7,	L	L	2	681	652	3	422	-7	399	412	5, 17,	L
-4	768	-16	165	170	4	484	489	4	299	-5	469	458	-3	308
-3	648	-15	191	192	6	460	472	6	208	-4	196	191	6, 0,	L
-2	507	-14	370	359	8	261	272	7	154	-3	410	386	-16	195
-1	565	-9	359	381	9	159	146	9	166	-1	432	421	-14	282
0	201	-8	292	312	10	247	257	11	117	3	265	246	-12	605
1	181	-7	363	363	12	228	247	5, 11,	L	5	282	259	-10	225
2	339	-6	869	899	12	228	247	-14	240	7	442	399	-8	380
6	361	-4	626	674	12	228	247	-10	269	9	356	350	-6	655
7	572	-3	312	350	12	228	247	-8	283	308			-4	843

6, 0,	L	187	12	242	251	2	936	960	0	633	646	5	432	407
0 132	133	339	13	117	137	3	587	573	1	364	346	9	144	134
2 454	451	248	6, 4,		L	5	503	501	2	477	477	11	146	145
4 1495	1494	511	-16	219	206	9	156	169	3	441	434	6, 10,		L
6 507	528	127	-15	172	193	10	405	423	4	718	705	-10	217	246
8 561	578	732	-14	316	311	11	159	156	5	159	172	-9	256	267
10 447	478	623	-13	191	199	13	239	275	6	447	443	-8	226	181
6, 1,	L	394	-10	550	584	6, 6,		L	12	286	309	-7	687	693
-15 225	213	136	-8	523	553	-15	203	225	6, 8,		L	-6	222	224
-13 328	322	415	-7	319	339	-11	257	265	-15	122	154	-5	573	564
-12 321	335	432	-6	952	992	-9	407	431	-8	521	525	-4	591	628
-7 162	152	194	-5	268	298	-8	501	502	-7	408	400	-3	565	575
-6 204	196	125	-4	796	864	-7	607	638	-6	655	684	-2	440	423
-4 492	516	162	-3	434	439	-6	188	178	-4	349	345	1	459	451
-3 755	808	L	-1	310	316	-5	123	156	-3	166	187	2	270	249
-1 1048	1079	152	1	340	334	-2	481	502	-2	570	544	3	187	168
0 519	551	338	2	281	307	-1	179	183	-1	179	183	4	389	386
1 658	686	167	3	180	167	0	215	238	0	438	423	5	471	452
2 841	852	470	4	579	553	2	167	169	1	323	312	6	307	297
3 429	438	110	6	613	596	5	394	359	5	170	134	7	431	418
4 193	205	502	7	237	231	6	283	295	6	288	274	8	195	194
5 300	297	327	8	605	636	7	309	315	7	319	295	9	155	161
6 118	108	346	9	147	170	8	108	102	8	198	196	6, 11,		L
7 185	195	162	10	370	390	9	320	338	10	266	277	-12	241	252
9 291	293	166	6, 5,		L	11	306	342	11	113	126	-11	149	137
10 346	363	173	-15	255	251	12	123	146	6, 9,		L	-10	297	309
11 205	199	514	-13	497	516	6, 7,		L	-15	286	291	-9	236	247
13 166	158	1179	-12	354	364	-14	328	343	-13	378	388	-8	158	156
6, 2,	L	312	-11	580	598	-12	267	297	-11	261	271	-6	214	206
-17 289	279	670	-10	237	241	-11	317	343	-10	294	295	-5	160	166
-15 227	224	355	-9	557	580	-10	415	421	-9	476	475	-4	157	165
-13 292	300	641	-6	313	306	-9	231	244	-8	250	246	-3	195	209
-12 286	287	734	-5	269	288	-8	312	340	-6	207	225	-2	403	400
-11 282	296	257	-4	436	437	-7	328	327	-5	237	226	0	815	789
-9 670	663	310	-2	424	454	-5	207	219	-3	219	195	1	216	223
-8 210	209	354	-1	845	889	-4	180	187	-1	749	743	2	578	563
-7 532	557	235	0	420	410	-2	621	662	-1	223	221	3	146	145
-5 900	965	291	1	1023	1008	-1	210	191	3	280	289	4	341	316

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8 180	173	-5 241	236	8 173	176	176	-2	217	213	-9 528	552	552	-10 367	367
6, 12,	L	-4 212	189	9 226	232	232	-1	219	231	-8 355	370	370	-8 378	397
-9 242	261	7, 0,	L	11 116	156	156	0	246	269	-7 259	263	263	-7 300	336
-8 402	415	-16 158	146	7, 2,	L	L	1	194	231	-6 591	626	626	-6 384	404
-7 365	364	-14 532	529	-17 194	183	183	2	511	508	-5 306	301	301	-5 338	330
-6 426	402	-12 450	462	-15 356	357	357	4	321	324	-3 334	329	329	-4 210	202
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-4 442	432	-8 806	776	-11 314	307	307	7	362	383	1 175	186	186	2 231	214
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-2 610	607	-4 826	824	-8 143	150	150	9	135	142	3 422	432	432	6 332	327
3 244	219	-2 443	474	-7 267	255	255	10	194	232	5 570	564	564	7 193	185
5 257	228	0 860	833	-6 341	350	350	11	103	118	6 278	284	284	8 307	319
6 436	398	2 631	622	-5 175	192	192	7, 4,	L	86	7 216	199	199	9 145	157
7 214	189	4 653	646	-4 207	207	207	-15 123	513	513	8 280	280	280	10 320	350
8 121	116	6 391	388	-3 474	510	510	-14 541	657	657	9 320	328	328	7, 8,	L
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-11 271	264	10 113	134	-1 906	898	898	-10 451	455	455	7, 6,	L	L	-14 248	254
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-2 172	171	-17 131	142	1 670	675	675	-7 228	229	229	-13 186	164	164	-11 254	256
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1 502	459	-14 136	113	3 231	241	241	-4 458	479	479	-9 329	340	340	-4 208	208
2 118	106	-11 152	140	4 237	229	229	-3 443	434	434	-7 184	194	194	-3 536	529
3 336	299	-9 189	187	5 372	371	371	-2 562	571	571	-6 331	344	344	-2 393	399
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-9 174	160	-6 550	551	9 342	367	367	2 758	755	755	-2 392	402	402	2 397	381
-7 363	379	-5 359	358	11 189	203	203	3 530	540	540	-1 267	273	273	3 359	365
-6 178	178	-4 280	276	7, 3,	L	L	4 318	312	312	0 235	251	251	4 173	170
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4 188	175	1 531	549	-12 275	264	264	10 259	269	269	7 116	89	89	7, 9,	L
5 262	220	2 595	633	-10 540	542	542	12 223	243	243	9 311	305	305	-11 172	198
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0 321	311	5 318	334	-6 489	495	495	-12 302	314	314	-5 502	506	506	-5 502	506
2 301	291	6 296	315	-5 309	311	311	-11 269	286	286	-4 203	236	236	-4 203	236

B, 9,	L	B, 13,	L	9, 2,	L	-4	589	562	0	223	231	-1	171	174
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3	218	1	416	-9	329	2	421	405	6	417	391	-12	374	363
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-10	214	-5	249	-4	120	9, 5,		L	-5	201	224	-6	222	221
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7	287	-9	463	-6	432	9, 6,		L	3	259	249	-4	517	511
B, 11,	L	-8	490	-5	154	-14	212	168	5	347	333	-3	313	288
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5	239			-5	189	-6	452	440	-2	171	185	2	234	223

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3 506 476		-11 358	341	-8 253	235	-9 317	308	-8 260	249
5 247 230		-9 370 365		-6 484 490		-7 181 215		-5 187 194	
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1 475 449		10, 8,	L	3 178 169		-9 193 165		-3 194 179	
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$$[\text{CpMoI}(\text{CO})(\text{C}_4\text{Ph}_4)] \cdot \text{CH}_2\text{Cl}_2$$

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9 1087 1106	3 1936 2010	1 1493 1564	0, 16, L	6 2675 2710	12 578 558
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11 2167 2095	5 1934 1963	3 2049 2140	5 1839 1859	9 706 672	1, 5, L
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9 1595 1624	1, 11, L	16 847 916	22 929 952	1 1219 1192	14 1385 1404
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11 897 897	6 1469 1371	1, 17, L	2, 2, L	7 1670 1746	17 485 466
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13 675 706	8 1291 1319	3 953 895	2 393 438	9 1633 1671	19 503 469
16 705 798	10 1066 1021	5 663 709	3 3726 3659	10 988 968	22 477 499
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19 628 666	15 748 624	1, 19, L	6 509 400	14 693 615	0 998 1012
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15 1589	9 1158	1170	0 1030	24 1071	1035	8 490	489	21 954	936
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1 525	2, 11, L		2, 16, L	3 1053	1137	16 888	969	2 1448	1554
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13 701	22 727	651	2 2276	3, 3, L		9 1245	1162	3, 8, L	
14 575	2, 12, L		4 1210	1 2691	2738	10 1742	1639	1 1139	1102
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14 499	4 684	632	3, 1, L	18 423	400	1 2368	2496	17 447	618
15 1030	8 1138	1138	1 1215	19 480	616	2 1711	1772	21 1177	1125
17 950	2, 14, L		2 1003	21 888	938	3 2127	2174	3, 9, L	
18 479	5 1277	1274	4 657	25 813	777	4 1415	1403	1 2069	2101
19 628	6 1016	1006	6 853	27 780	797	5 328	374	2 939	1026
20 864	12 955	927	9 1073	3, 4, L		7 1216	1242	4 873	928
22 521	14 1031	970	10 2233	1 2859	2895	8 697	674	5 740	811
2, 10, L	16 504	399	12 1605	2 864	859	9 1983	1999	7 2192	2205
0 1839	18 654	629	13 1110	3 1138	1056	10 762	780	11 988	1039
1 961	20 920	975	15 759	4 1373	1427	11 1352	1408	13 1237	1209
3 1804			16 1167	5 688	665	13 947	1002	14 614	575
5 2008			18 1137	6 928	922	16 715	621	19 1090	1154

3, 10, L	3, 14, L	B	692	605	4, 4, L	4, 7, L	23	607	579
2 1055 1075	4 1174 1142	9 644 473			1 2333 2223	0 1148 1101	4, 10, L		
4 1021 1028	6 1209 1189	10 884 916			2 1063 1037	3 1400 1447	0 1034 967		
6 1170 1112	11 925 821	12 1500 1548			3 2714 2708	4 540 578	1 708 751		
8 1458 1470	12 548 535	14 1782 1733			4 785 763	5 2020 2044	2 774 747		
12 812 754	3, 15, L	18 996 967			5 2106 2191	6 1078 1092	3 912 955		
13 900 955	1 823 729	20 1313 1265			8 760 756	7 1162 1259	8 880 942		
14 867 847	5 1129 1120	22 1005 974			9 788 702	10 1042 1029	9 1147 1154		
16 826 790	7 905 902	4, 2, L			10 1136 1138	12 991 983	11 973 936		
18 499 591	3, 16, L	1 1041 1030			11 1268 1284	13 506 568	14 499 451		
19 840 754	3 537 466	2 505 547			13 1127 1013	15 1262 1326	15 979 897		
3, 11, L	11 1216 1157	3 2115 2072			14 617 701	17 894 863	17 1205 1268		
1 854 785	13 726 705	4 1025 1003			4, 5, L	18 747 713	19 563 609		
2 682 725	3, 17, L	5 1049 1057			0 1966 1941	25 534 476	4, 11, L		
4 734 708	7 983 1100	9 974 986			1 708 729	4, 8, L	0 1571 1560		
6 466 618	8 523 502	10 1118 1167			2 1314 1392	1 564 472	2 1090 1028		
7 937 937	13 738 686	11 1158 1209			3 2274 2311	4 579 654	5 493 480		
9 817 703	3, 18, L	12 1220 1253			4 871 745	5 556 530	8 1072 1099		
10 673 753	11 622 689	14 824 766			5 1765 1734	6 927 981	9 710 852		
13 904 918	4, 0, L	20 1122 1102			6 1406 1370	7 1102 1131	10 1124 1184		
15 712 678	0 2306 2322	23 876 940			7 908 898	9 1707 1695	12 754 785		
16 541 563	2 831 828	4, 3, L			8 405 464	11 996 1002	4, 12, L		
19 910 902	4 2994 2864	0 1449 1472			12 798 717	12 632 559	0 1974 2058		
21 620 666	6 2288 2209	1 965 957			13 658 543	14 859 887	1 512 464		
3, 12, L	10 801 920	2 653 651			15 889 861	15 1022 962	2 1567 1563		
2 957 1024	12 3010 2879	3 2594 2555			4, 6, L	17 939 985	3 1142 1092		
4 1485 1577	14 1723 1747	5 1785 1796			0 1264 1252	20 920 928	6 1159 1097		
6 1615 1602	18 1079 992	6 708 770			1 636 583	22 824 831	8 1512 1571		
7 428 475	20 1244 1244	7 677 736			2 990 1008	4, 9, L	9 678 670		
12 849 885	22 1081 1098	8 913 844			3 686 761	6 742 729	4, 13, L		
14 767 734	26 522 536	9 821 790			5 1669 1627	9 852 824	0 1258 1272		
16 759 715	4, 1, L	10 814 818			6 1578 1607	10 1453 1448	2 1066 1056		
3, 13, L	0 1979 2028	11 1037 1083			7 973 986	12 756 695	4 515 508		
4 1571 1592	1 476 438	13 854 861			8 1174 1168	14 488 448	6 863 841		
6 729 713	2 1606 1622	14 823 881			12 785 741	15 1036 993	8 1015 959		
10 1131 1076	3 638 647	20 996 1027			14 1179 1278	17 1093 1193			
16 827 868	5 1215 1261	22 793 767			18 723 739	18 729 689			
18 740 711	7 582 638				20 746 802	20 681 659			

4, 14, L	14 1934	1957	12 581	578	18 645	589	5, 11, L	24 435	535
0 1166 1154	16 1822 1880		13 435 446		19 781 799		2 1531 1518	6, 1, L	
5 597 482	18 1092 1014		15 858 770		21 600 653		4 1640 1647	0 2244 2188	
6 1018 1062	20 424 295		5, 5, L		23 585 645		5 887 843	1 365 357	
8 932 949	24 937 992		1 2085 2143		24 871 858		6 1108 1056	2 1052 1032	
9 633 632	26 792 822		3 935 924		5, 8, L		7 641 757	4 1328 1254	
14 479 386	5, 2, L		4 1024 1056		1 707 708		11 754 719	6 1802 1687	
15 632 540	2 406 314		6 748 669		5 539 601		13 577 692	8 1041 1105	
4, 15, L	5 673 700		7 1961 2019		6 542 532		5, 12, L	11 649 594	
2 744 690	6 669 740		9 1822 1827		8 451 481		2 1059 1012	12 808 699	
9 570 681	7 1047 1082		11 1404 1454		10 824 735		4 1185 1148	14 537 495	
11 471 444	8 719 753		12 1110 1122		11 689 664		6 589 675	15 488 518	
15 617 695	9 1186 1114		16 536 512		13 834 775		7 815 802	24 468 481	
4, 16, L	13 827 873		17 513 375		16 837 760		10 714 721	6, 2, L	
15 756 756	16 1052 1046		24 670 623		22 597 548		5, 13, L	0 967 1055	
4, 17, L	18 727 501		25 467 462		5, 9, L		2 1304 1248	1 833 760	
7 727 605	5, 3, L		5, 6, L		2 679 608		4 1923 1853	2 588 535	
4, 19, L	1 2691 2807		1 847 936		3 591 684		6 1353 1327	3 1522 1496	
0 741 724	2 775 683		2 944 1038		4 583 494		10 746 740	4 1075 1051	
5, 0, L	3 1065 1068		3 1132 1060		5 1210 1195		5, 14, L	5 1903 1920	
2 448 546	4 994 999		4 679 786		6 763 755		4 834 751	7 793 740	
4 1214 1171	5 351 438		7 691 744		7 775 838		5, 15, L	9 1187 1213	
8 1815 1792	6 1685 1625		10 802 819		11 1583 1580		2 806 831	10 1085 1057	
10 1694 1592	7 2556 2545		13 489 417		13 1463 1450		4 1156 1095	11 813 784	
12 612 600	8 1804 1685		14 446 396		14 510 529		6 598 557	12 1661 1591	
14 1262 1327	9 1987 1913		16 992 1072		16 1393 1426		11 482 566	14 781 815	
16 942 946	11 489 506		18 606 664		18 691 656		13 874 835	18 706 682	
18 530 544	12 639 639		5, 7, L		19 1044 1150		6, 0, L	20 1107 1173	
22 769 669	13 646 686		1 811 814		21 807 889		0 913 834	22 676 708	
24 839 866	14 1048 1011		3 894 853		5, 10, L		2 639 702	6, 3, L	
5, 1, L	15 501 507		4 682 715		1 618 672		4 803 832	0 1121 987	
1 797 760	16 796 772		8 594 650		2 748 659		6 564 584	1 1216 1199	
4 535 545	5, 4, L		9 1101 1065		5 542 652		10 1738 1555	2 350 314	
5 557 588	1 1310 1224		10 1114 1175		13 507 601		12 2481 2353	3 1894 1950	
6 1667 1691	3 880 855		11 996 1085		19 559 624		14 1213 1233	5 791 810	
8 1866 1870	5 561 668		12 664 674				18 1273 1248	6 634 552	
9 785 673	7 1699 1574		13 1034 985				20 1273 1403	7 852 789	
10 2039 2013	9 996 975		16 1361 1369				22 967 1087	9 1171 1118	

6, 3,	L	6, 1246	1177	6, 10,	L	7, 0,	L	7, 4,	L	10, 700	742
10, 682	652	7, 809	749	0, 1217	1322	2, 1459	1507	4, 488	494	16, 876	802
11, 1015	1070	8, 759	769	3, 473	461	4, 2460	2529	5, 718	842	19, 761	723
15, 920	817	11, 445	436	4, 593	466	6, 1239	1202	7, 707	852	21, 513	505
17, 934	977	12, 762	765	8, 974	1011	10, 907	941	9, 428	493	7, 8,	L
6, 4,	L	13, 685	625	9, 883	890	14, 508	519	11, 665	679	1, 1492	1480
0, 807	866	14, 826	818	10, 1033	968	22, 653	607	13, 1202	1226	4, 864	929
1, 1201	1237	15, 873	915	11, 822	891	7, 1,	L	15, 522	490	6, 577	530
2, 1094	1130	17, 830	839	12, 587	477	1, 857	877	16, 649	601	7, 682	654
3, 2063	2093	6, 7,	L	15, 732	715	4, 551	554	19, 1203	1197	9, 779	806
4, 575	539	0, 1306	1369	17, 1066	1071	6, 537	565	21, 931	904	11, 611	641
5, 1845	1957	2, 614	617	6, 11,	L	7, 798	782	7, 5,	L	14, 520	491
6, 404	433	3, 687	569	0, 1076	976	8, 509	697	1, 1614	1637	7, 9,	L
7, 1187	1202	4, 505	484	3, 922	883	9, 693	744	2, 1325	1242	6, 832	837
8, 678	631	6, 598	631	4, 693	701	10, 1211	1250	4, 1520	1481	8, 739	724
10, 673	692	8, 870	938	5, 971	945	14, 1100	1115	7, 975	1077	13, 1076	1149
11, 1088	1027	20, 1016	1127	6, 849	807	16, 1304	1273	9, 1230	1284	14, 763	679
12, 473	401	6, 8,	L	7, 451	303	18, 1066	1123	11, 759	800	16, 726	739
13, 856	983	0, 689	766	9, 939	806	22, 584	641	22, 424	416	19, 845	921
23, 430	166	2, 708	603	6, 12,	L	7, 2,	L	7, 10,	L	7, 10,	L
6, 5,	L	3, 576	611	0, 1588	1500	2, 1061	1057	1, 522	577	1, 1100	1062
1, 607	527	5, 565	616	2, 1206	1205	4, 1598	1553	2, 1570	1587	5, 830	875
6, 573	594	7, 924	1005	6, 980	902	5, 1488	1567	4, 1000	949	6, 772	822
8, 728	753	9, 1057	1007	8, 1450	1432	6, 762	755	9, 816	857	7, 1174	1187
9, 476	508	11, 671	699	10, 573	606	7, 786	775	10, 928	900	7, 11,	L
11, 507	419	12, 798	828	6, 13,	L	11, 571	528	11, 895	909	2, 877	941
12, 708	755	14, 1028	961	4, 482	506	12, 571	512	12, 586	615	4, 1028	981
14, 956	903	15, 1043	992	6, 861	882	13, 920	943	15, 601	667	12, 532	564
17, 1078	1093	17, 870	915	12, 740	739	15, 881	873	16, 858	911	14, 813	770
18, 548	617	19, 595	617	14, 756	727	19, 900	876	18, 1001	994	7, 12,	L
20, 588	715	20, 735	614	6, 14,	L	7, 3,	L	19, 668	755	8, 772	791
6, 6,	L	6, 9,	L	0, 1188	1199	1, 2346	2375	22, 683	635	10, 818	873
0, 1219	1159	1, 811	850	2, 971	956	4, 786	746	7, 7,	L	16, 936	917
1, 857	877	3, 1558	1486	8, 1131	1140	5, 884	817	1, 797	661	7, 13,	L
2, 934	871	5, 803	846	10, 656	637	6, 652	662	2, 827	813	2, 567	547
3, 1283	1285	9, 651	692	15, 595	641	7, 1355	1361	4, 702	746	4, 1050	1171
4, 749	715	11, 511	521	15, 595	641	9, 876	823	8, 1008	981	6, 1067	1164
5, 1693	1707	15, 651	640					9, 686	720		

7, 14, L	8, 5, L	8, 11, L	9, 6, L	10, 1, L	11, 1, L
1 654 586	0 1164 1100	0 933 923	1 648 663	0 1487 1449	4 682 619
8 865 815	4 829 728	5 775 747	2 1274 1213	2 1016 1028	6 790 821
7, 16, L	6 891 814	12 1005 1058	3 448 461	6 1236 1215	11, 2, L
1 856 884	7 1429 1439	14 797 771	4 1294 1269	8 1002 1051	2 542 513
8, 0, L	8 808 863	8, 13, L	8 850 751	9 470 550	4 825 814
12 541 654	9 922 865	4 762 785	9 668 764	10, 2, L	11, 3, L
14 604 535	12 795 713	10 490 554	10 1076 1036	0 839 771	1 770 706
20 848 857	15 909 949	8, 15, L	11 1002 912	2 669 614	6 570 647
8, 1, L	17 1200 1195	3 789 748	13 788 804	8 694 745	11, 4, L
0 1708 1795	8, 6, L	9, 0, L	16 823 744	10, 3, L	7 636 702
2 1441 1409	0 584 730	2 1633 1637	9, 8, L	2 777 785	11, 5, L
6 1067 1119	3 610 533	4 2066 2016	1 1206 1244	9 999 1051	9 640 662
7 604 505	6 482 412	6 1415 1407	2 777 690	15 863 854	11, 6, L
8 1364 1439	14 551 604	10 1565 1564	3 710 739	10, 4, L	4 666 815
10 1172 1100	8, 7, L	12 854 901	4 1280 1183	0 744 690	12, 0, L
8, 2, L	0 1480 1346	9, 1, L	6 743 613	10, 5, L	0 421 491
3 585 614	3 717 829	2 920 742	7 1020 1053	7 662 686	12, 3, L
20 517 511	4 478 372	9, 2, L	9 836 788	8 623 644	1 616 535
8, 3, L	5 840 771	1 506 437	10 914 1011	12 602 673	3 863 988
0 602 547	6 1149 1089	4 1226 1412	9, 9, L	14 551 453	
2 1072 1160	7 1248 1098	5 712 875	1 758 730	15 550 639	
3 665 597	8 980 894	6 1426 1343	9, 10, L	10, 7, L	
4 651 647	9 789 758	7 674 613	1 734 706	0 726 735	
7 964 896	12 523 527	10 600 634	4 515 521	2 749 823	
8 835 868	15 482 358	11 987 883	6 752 738	3 817 751	
9 1247 1281	17 429 460	12 892 756	7 1084 1265	6 808 868	
10 753 744	8, 8, L	13 1214 1134	9 676 646	8 769 829	
11 598 714	4 597 597	19 796 901	9, 12, L	11 449 456	
13 574 392	8, 9, L	9, 4, L	8 847 954	10, 9, L	
15 1464 1422	0 786 715	5 718 648	10 713 869	3 985 938	
17 1196 1235	2 916 856	7 1142 1111	9, 14, L	5 709 828	
8, 4, L	3 1214 1343	8 481 467	1 827 683	11 558 650	
0 945 907	5 1291 1288	10 628 705	10, 0, L	11, 0, L	
3 1012 1025	6 732 735	11 1469 1448	0 1375 1377	2 896 801	
5 854 804	11 715 815	13 1571 1541	2 535 455	4 1117 1104	
6 418 380	13 483 417	14 582 588	6 683 586	8 854 741	
13 691 752		17 585 651	8 642 575	10 751 766	

CALCULATED AND OBSERVED STRUCTURE FACTORS OF
[WBr(CO)₂(SC₆F₅)(C₇H₈)]

0, 0, L	7 561 556	9 502 453	3 1070 1050	0 147 149	13 420 428
2 3423 3365	10 906 928	10 495 492	4 440 437	1 1989 1987	1, 3, L
4 1773 1766	12 435 480	11 349 366	5 1064 1102	2 652 653	-12 286 246
6 2302 2403	14 392 427	12 482 450	7 1048 1060	3 1738 1758	-10 267 279
8 1592 1604	0, 4, L	13 330 311	9 908 912	4 390 401	-9 409 360
10 576 574	0 1231 1242	0, 7, L	0, 12, L	5 1043 1039	-6 476 444
12 846 829	1 383 382	2 870 843	1 933 927	7 853 853	-5 848 834
14 537 537	2 1100 1155	3 681 703	3 747 773	8 657 581	-4 471 459
0, 1, L	3 564 544	5 763 707	5 702 752	9 1016 1064	-2 572 560
1 608 562	4 1733 1747	7 622 572	7 881 950	10 423 391	-1 837 878
2 158 172	5 593 588	8 573 582	0, 14, L	11 617 613	1 831 792
3 372 346	6 2299 2397	9 693 646	1 1063 1036	12 358 311	3 1272 1280
4 1097 1041	7 633 683	10 400 413	3 673 728	13 623 606	5 499 507
5 1046 .999	8 1471 1495	11 478 391	5 537 729	1, 2, L	6 322 318
6 883 861	9 760 798	0, 8, L	1, 0, L	-13 867 980	7 900 953
8 1247 1283	10 988 997	0 1521 1543	-13 481 571	-12 284 349	8 542 556
10 887 911	11 493 528	1 1984 1966	-11 838 912	-11 901 962	9 1380 1405
11 274 269	12 847 846	2 1377 1444	-9 1538 1645	-10 255 287	10 646 624
12 479 468	0, 5, L	3 1488 1526	-7 1988 2018	-9 1310 1333	11 1105 1105
14 436 462	2 378 403	4 596 563	-5 2863 2638	-8 286 246	12 381 350
0, 2, L	3 286 288	5 1281 1262	-3 3745 3175	-7 2205 2086	13 554 562
0 2781 2901	4 557 483	6 845 795	-1 3100 2662	-6 509 466	1, 4, L
2 2758 2610	5 703 659	7 1052 1034	1 2418 2995	-5 2400 2221	-13 772 808
3 1257 1171	6 407 367	8 506 506	3 509 541	-3 2403 2250	-12 501 517
4 1875 1874	7 439 408	9 418 430	5 281 285	-2 475 466	-11 862 875
5 426 423	8 392 358	10 343 311	7 669 701	-1 2856 3011	-10 677 647
6 1986 1969	10 532 508	0, 9, L	9 1144 1136	0 1594 1663	-9 807 819
7 919 936	12 448 435	1 334 291	11 479 467	1 2515 2589	-8 812 768
8 1509 1516	0, 6, L	2 527 515	1, 1, L	2 1110 1185	-7 1123 1109
9 642 651	0 960 949	4 479 484	-13 394 434	3 1924 1957	-6 936 924
10 934 951	1 2103 2106	6 628 653	-9 632 636	4 1842 1814	-5 1421 1408
11 404 438	2 906 871	7 392 340	-8 307 326	5 1204 1252	-4 661 696
12 647 689	3 1416 1415	8 544 616	-7 514 491	6 596 644	-3 441 457
14 579 609	4 782 716	9 470 491	-5 542 527	7 1277 1299	-2 887 907
0, 3, L	5 1360 1267	0, 10, L	-4 403 329	8 724 742	-1 687 650
1 491 493	6 1315 1264	0 932 956	-3 1085 965	9 1022 1046	0 970 1044
3 284 277	7 945 864	1 1094 1119	-2 625 592	10 339 408	1 2307 2502
4 766 769	8 1159 1099	2 758 725	-1 1026 1071	11 436 478	2 1052 1116

1, 4,	L	1061	1, 8,	L	-4	685	750	-8	1783	1730	-3	1131	1138
3 2059	2252	-5 1113	-12 582	572	-2	874	887	-6	2413	2190	-2	872	930
4 1026	1025	-3 1309	-11 416	427	-1	705	707	-4	3223	2776	-1	786	786
5 1657	1702	-2 1479	-10 786	740	0	1406	1291	-2	731	650	0	1697	1703
6 1164	1141	-1 1593	-9 564	580	1	781	765	0	3029	2930	1	1734	1738
7 1637	1685	0 2065	-8 1116	1112	2	1412	1427	2	977	985	2	1512	1517
8 915	923	1 1779	-7 978	935	4	894	852	4	534	498	3	995	1020
9 816	884	2 1005	-6 1598	1591	6	1221	1211	6	891	884	4	515	520
10 574	543	3 1708	-5 1308	1259	8	872	795	8	1189	1087	5	932	918
11 384	423	4 468	-4 1825	1660	10	509	519	-2, 1,	L	L	6	780	818
13 361	323	5 1333	-3 855	895	1, 11,	L	L	-10	271	298	8	890	873
1, 5,	L	6 487	-2 1915	1854	5	412	290	-8	429	430	10	493	464
-5 432	387	7 1224	-1 1192	1272	6	488	453	-6	1882	1691	-11	266	247
-4 639	-620	8 559	0 1151	1074	8	744	752	-5	878	816	-10	312	306
-2 698	617	9 758	1 937	997	1, 12,	L	L	-4	2136	2000	-6	280	298
-1 754	697	12 441	2 1280	1263	-8	596	673	-3	1181	1151	-5	421	455
0 504	520	1, 7,	4 980	951	-6	571	631	-2	2913	2903	-4	765	774
2 1477	1444	-10 420	6 667	687	-2	738	781	0	2085	2026	-3	611	645
3 730	749	-9 438	7 530	513	0	1401	1301	1	299	276	-2	531	579
4 1125	1158	-8 278	10 529	458	2	1112	1143	2	1118	1171	-1	305	311
5 542	590	-6 268	1, 9,	L	4	896	950	4	1637	1589	0	1268	1202
6 661	646	-4 647	-7 354	404	6	846	741	5	292	259	1	737	788
7 950	922	-2 919	-1 579	598	8	850	809	6	1193	1136	2	2124	2124
8 664	595	-1 615	0 458	441	1, 13,	L	L	7	370	426	3	845	860
9 1195	1135	0 1203	1 978	1010	-1	422	408	8	985	922	4	2389	2595
10 502	480	1 456	2 742	752	4	540	451	9	366	369	5	1558	1546
11 963	909	2 1080	3 1046	1129	1, 14,	L	L	10	1489	1404	6	1839	1868
12 429	339	3 832	4 852	841	-4	780	810	12	1218	1093	7	872	893
13 664	584	4 1030	5 458	421	0	904	846	14	753	640	8	1268	1325
1, 6,	L	5 956	6 732	695	2	669	654	-14	649	722	9	624	601
-13 672	630	6 622	8 599	579	4	573	505	-12	1065	1146	10	1329	1333
-12 554	536	7 548	10 781	716	1, 15,	L	L	-12	1045	1080	11	501	468
-11 665	600	9 478	-10 712	739	2	516	554	-10	1079	850	12	879	850
-10 777	698	10 433	-8 682	693	2, 0,	L	L	-8	1141	1162	2, 4,	L	L
-9 782	746	11 679	-7 870	866	-14	862	931	-6	2250	2190	-13	387	417
-8 1250	1152		-6 1099	1076	-12	1313	1362	-5	721	691	-12	735	778
-7 1432	1320		-5 875	904	-10	1420	1421	-4	1810	1786	-11	488	507
-6 1732	1669												

-10 760 714	2, 4, L	11 731 746	3 502 492	3, 2, L	-13 810 891
-9 468 431	-13 523 477	2, 10, L	2, 15, L	-13 810 891	-11 988 1004
-8 956 987	-12 608 554	-9 656 651	-1 527 500	-10 272 233	-9 1037 1027
-7 482 488	-11 837 803	-7 627 660	1 574 464	-8 296 301	-7 1246 1245
-6 1918 1924	-10 509 552	-6 315 388	3, 0, L	-6 703 704	-5 1410 1429
-5 1113 1071	-9 1090 1020	-5 1146 1159	-13 748 836	-4 1436 1426	-3 1270 1283
-4 1831 1866	-8 801 776	-4 411 304	-11 1008 1080	-2 331 354	-1 1643 1669
-3 619 625	-7 1186 1144	-3 995 1122	-9 704 690	0 626 635	1 907 886
-2 1147 1344	-6 1145 1081	-1 1117 1121	-7 476 469	2 442 433	3 960 1021
-1 967 1068	-5 1261 1202	1 1329 1345	-5 1602 1441	5 605 565	6 820 782
0 2740 2859	-4 995 996	3 1198 1154	-3 2433 2144	6 820 782	7 226 153
1 941 991	-3 754 746	5 454 412	-1 1252 1049	3, 3, L	-11 553 556
2 3080 3262	-2 1410 1396	7 520 532	1 1194 1172	-9 421 437	-8 314 317
3 702 808	0 1345 1317	2, 11, L	3 1884 1856	-7 386 382	-6 227 228
4 902 953	1 899 866	-1 474 476	5 1018 1044	-5 1056 1082	-4 838 857
5 566 579	2 1594 1486	1 769 837	7 652 619	-3 1635 1755	-2 1143 1186
6 701 738	3 396 354	3 1266 1259	9 1057 907	-1 2040 2069	0 1336 1318
7 622 640	4 969 936	5 1212 1206	3, 1, L	1 2588 2609	2 1688 1700
8 589 564	5 375 320	7 1111 1028	-11 787 801	2 1688 1700	3 1794 1924
2, 5, L	6 353 401	9 919 876	-9 1041 1038	3 1794 1924	4 932 969
-5 621 578	7 787 705	2, 12, L	-7 936 888	5 1673 1779	6 709 776
-3 772 747	8 437 369	-7 710 783	-5 2001 1901		
-2 339 329	9 722 648	-5 844 943	-4 504 466		
-1 367 362	11 414 352	-3 965 1020	-3 2111 2045		
0 896 787	2, 7, L	-1 1098 1054	-2 2025 1956		
1 399 401	-7 671 680	0 430 342	0 745 742		
2 1616 1635	-5 1249 1238	1 1110 1076	1 440 443		
3 823 795	-4 900 930	2 352 223	2 1010 1060		
4 1945 1949	-3 1184 1196	3 800 765	3 2041 2049		
5 464 448	-1 1253 1269	2, 13, L	4 1092 1142		
6 1537 1441	0 544 575	1 660 660	5 1672 1752		
7 329 344	1 994 915	3 796 795	7 1338 1305		
8 1366 1251	2 1034 954	5 713 764	9 1599 1577		
9 476 486	3 388 399	7 697 603	10 287 275		
10 1168 1096	4 993 985	2, 14, L	11 1319 1269		
12 823 692	5 605 630	-5 572 748	13 970 894		
	6 767 725	-3 545 655			
	7 513 405	-1 519 525			

3, 3, L	0 1222	1029	-1 821	804	6 1189	1123	3, 14, L	-8 288	264
7 1484 1554	1 2279	2337	0 383	384	7 512	513	2 465	-7 633	636
8 325 274	2 1151	1120	1 750	712	8 949	862	4, 0, L	-6 697	722
9 888 957	3 2286	2327	2 531	555	9 555	488	-12 673	-5 237	283
11 912 959	4 1310	1250	3 1209	1188	10 854	737	-10 517	-4 1098	1130
13 691 675	5 1662	1600	4 652	585	3, 10, L		-6 345	-3 318	275
3, 4, L	6 1080	1025	5 1105	1073	-10 515	524	-4 1139	-2 202	232
-13 589 582	7 868	864	6 821	827	-8 620	694	-2 1020	-1 434	447
-12 366 400	8 908	788	7 685	618	-6 843	911	0 827	0 169	161
-11 711 699	9 885	801	8 994	882	-4 927	1016	2 988	1 471	492
-10 517 522	10 864	742	9 671	588	-2 947	955	4 740	2 451	471
-9 1048 1079	11 819	709	10 1095	934	0 649	582	8 860	3 922	940
-8 707 703	12 625	616	11 788	770	2 530	554	10 761	4 334	351
-7 1275 1358	3, 6, L		12 865	742	3, 11, L		12 669	6 524	510
-6 683 676	-12 576	557	3, 8, L		-6 497	532	4, 1, L	12 466	450
-5 1516 1538	-11 844	781	-11 552	557	-4 741	841	-12 614	4, 3, L	
-4 605 629	-10 558	551	-10 540	573	-2 954	1026	-10 998	-12 409	457
-3 1415 1567	-9 615	615	-9 373	380	0 1388	1370	-8 938	-11 291	300
-2 706 756	-8 618	540	-8 360	427	2 1263	1201	-7 427	-10 1013	1035
-1 919 932	-7 1115	1037	-6 683	735	4 1164	1103	-6 1466	-8 1193	1263
0 346 290	-5 1512	1482	-4 542	578	6 1042	976	-5 885	-7 556	568
1 1116 1157	-3 794	860	-2 297	335	8 688	674	-4 1390	-6 984	1024
2 456 447	-1 529	468	0 562	577	3, 12, L		-3 189	-5 1259	1293
3 844 872	1 1202	1157	2 973	994	-6 773	819	-2 1930	-4 1859	1952
4 286 305	2 771	773	3 524	552	-4 910	972	-1 577	-3 1718	1789
5 704 736	3 436	458	4 412	342	-2 683	728	0 1663	-2 2609	2632
7 389 453	4 831	809	7 379	395	0 588	581	2 1901	-1 1186	1196
3, 5, L	6 416	452	3, 9, L		3, 13, L		3 314	0 2031	1960
-12 408 331	3, 7, L		-8 618	674	-6 466	532	4 2464	1 900	831
-10 515 429	-10 775	764	-7 446	533	-4 631	753	6 2150	2 1473	1459
-9 601 560	-9 487	447	-6 457	452	-2 808	913	8 1220	3 597	583
-8 706 695	-8 822	846	-5 420	491	0 766	739	10 1073	4 1660	1816
-6 1021 979	-7 338	288	-4 930	990	2 917	908	12 947	5 413	423
-5 782 711	-6 875	860	-2 1049	1105	4 1011	955	4, 2, L	6 1494	1528
-4 822 825	-5 808	837	0 991	1030	6 755	739	-12 632	7 360	323
-3 1595 1665	-4 1010	1025	2 1059	1000			-11 331	8 886	876
-2 1017 1039	-3 1287	1351	4 1445	1387			-10 780	10 772	781
-1 2259 2266	-2 530	568	5 633	665			-9 436	12 766	782

4, 4,	L	4, 6,	1	544	616	3	800	744	-4	1045	1070	2	343	400
-12 710	731	-12 575	2	420	434	5, 0,		L	-3	370	361	5, 5,		L
-11 479	498	-10 623	4	358	414	-11 353	300		-2	604	603	-11 769	725	
-10 794	835	-8 430	8	422	304	-9 310	237		0	280	220	-10 612	572	
-9 528	525	-6 646	4, 9,		L	-7 542	515		1	776	781	-9 1331	1217	
-8 876	902	-5 407	-9 569	576		-5 617	651		3	430	419	-8 663	624	
-6 845	896	-4 498	-7 665	639		-3 1661	1508		5	797	817	-7 1193	1127	
-5 492	544	-3 885	-5 845	844		-1 456	426		7	1057	1096	-6 591	574	
-4 776	813	-1 270	-3 1099	1187		1 576	544		9	725	734	-5 1235	1222	
0 862	865	1 811	-1 1209	1229		3 590	656		11	484	495	-4 675	676	
2 257	284	3 991	1 1110	1159		5 713	744		5, 3,		L	-3 1275	1247	
4 568	638	4, 7,	2 464	492		7 1710	1679		-12	318	281	-2 1302	1254	
6 834	838	-11 555	3 899	900		9 645	676		-11	842	875	-1 1171	1180	
-12 313	291	-10 585	4 1143	1196		11 506	629		-10	548	553	0 1209	1201	
-11 562	490	-9 881	5 1022	992		5, 1,	L		-9	1328	1352	1 849	744	
-10 763	719	-8 670	6 743	732		-13 417	446		-8	544	563	2 1212	1108	
-9 527	473	-7 585	7 1052	994		-11 731	713		-7	1200	1223	3 562	500	
-8 1236	1107	-6 512	8 481	429		-10 261	276		-6	520	566	4 1106	1106	
-7 572	588	-5 599	9 667	664		-9 1322	1236		-5	1043	1096	5 762	820	
-6 1196	1131	-4 979	4, 10,	L		-8 257	199		-4	410	410	6 701	780	
-5 484	472	-3 970	-9 568	587		-7 1245	1172		-3	2083	2220	7 720	731	
-4 1757	1816	-2 799	-7 493	464		-6 592	604		-2	506	513	9 581	557	
-2 2281	2210	-1 902	-5 591	560		-5 1643	1578		-1	1683	1737	10 431	426	
-1 853	838	0 942	-3 356	396		-4 534	556		0	291	345	11 679	623	
0 2197	2040	1 1475	1 278	336		-3 2274	2160		1	1109	1074	5, 6,	L	
1 1005	897	2 1117	4, 11,	L		-2 331	313		3	1127	1132	-4 401	386	
2 1233	1139	3 1315	-7 691	802		-1 2333	2199		5	1318	1284	-2 529	534	
3 978	990	4 1156	-5 1066	1150		1 2233	2187		7	1038	1089	-1 629	689	
4 935	959	5 1378	-3 1360	1413		2 205	196		9	765	793	1 512	475	
5 1141	1127	6 942	-1 1084	1170		3 2334	2491		11	632	659	4 313	353	
6 638	642	7 1429	1 888	908		5 2041	1930		12	371	329	6 703	697	
7 802	800	8 723	3 1073	1022		7 1324	1324		5, 4,		L	8 912	811	
8 475	458	9 1106	5 886	909		9 734	715		-5	325	374	9 650	571	
9 703	624	11 798	4, 13,	L		11 470	458		-3	318	295	11 408	361	
10 784	695	-5 434	-5 678	804		5, 2,	L		-2	542	554	5, 7,	L	
11 669	551	-4 364	-3 861	881		-11 462	454		-1	860	846	-9 630	614	
12 662	607	-1 614	-1 833	826		-9 530	538		0	393	377	-8 645	638	
		-1 614	1 722	695		-6 506	512		1	1024	998	-7 613	604	

5, 7,	L	5, 10,	4	623	670	9	566	530	6, 6,	L	9	402	545
-6 620	566	0 538	5	494	507	-8	357	396	-6 545	560	6, 9,	L	
-5 449	479	2 430	6	696	727	-6	614	676	-1 712	721	-9	682	672
-4 981	1003	5 468	7	436	425	-6	217	255	0 642	639	-8	375	429
-3 978	961	5, 11,	6, 2,	L	L	-5	588	637	1 1077	1037	-7	853	830
-2 1397	1275	-6 886	7	480	545	-4	588	637	3 1036	940	-5	700	716
-1 1094	1082	-4 871	-6	354	371	-1	260	259	5 713	678	-4	681	657
0 1611	1483	-2 909	-5	495	507	0	403	406	6 538	617	-3	911	862
1 995	980	0 632	-3	278	333	2	700	695	8 770	779	-2	723	743
2 1546	1441	2 489	-2	852	880	4	395	365	10 617	568	-1	778	798
3 754	777	4 615	0	1157	1161	6	627	587	6, 7,	L	0	670	715
4 990	1067	6 712	1	302	306	7	462	448	-10 480	418	1	821	828
5 710	700	5, 13,	2	1047	1044	8	1074	1013	-9 809	784	3	571	536
6 900	867	-4 616	3	277	309	9	423	380	-8 806	859	5	654	645
7 482	488	-2 889	4	949	1022	10	855	844	-7 894	901	6, 10,	L	
8 691	640	0 556	6	1053	1071	6, 5,	L		-6 554	568	0	469	481
5, 8,	L	2 557	8	793	792	-11	511	448	-5 1094	1127	1	560	535
2 374	378	6, 0,	10	747	721	-10	603	616	-4 614	620	2	539	471
4 616	623	-10 365	11	387	337	-9	509	449	-3 1529	1452	7	477	492
5 644	704	-6 286	6, 3,	L		-8	957	872	-2 1173	1121	6, 11,	L	
6 568	597	-4 524	-12	630	656	-7	599	536	-1 1386	1391	-5	501	554
7 727	673	-2 700	-10	910	941	-6	541	576	0 1044	1015	-3	553	496
8 611	573	0 1957	-9	452	473	-5	951	896	1 977	943	-1	374	442
5, 9,	L	2 1731	-8	1106	1103	-4	866	850	2 487	485	1	579	617
-8 935	951	4 1233	-7	401	431	-3	865	818	3 712	673	3	618	752
-6 816	877	6 1363	-6	996	1029	-2	822	844	4 586	541	5	571	687
-4 1068	1148	8 1408	-4	885	956	-1	934	909	5 476	450	6, 12,	L	
-2 1257	1152	10 576	-2	1071	1086	0	889	867	6, 8,	L	3	379	366
-1 935	931	6, 1,	-1	350	414	1	1411	1208	-5 346	321	7, 0,	L	
0 1137	1144	-12 505	0	1018	968	2	571	520	-1 593	665	-9	481	426
1 738	740	-10 1007	1	420	410	3	601	547	0 749	718	-7	760	682
2 744	874	-8 1580	2	924	1000	4	996	931	1 1106	1078	-5	1129	1031
3 697	699	-6 1639	3	785	807	6	1113	1080	2 1095	989	-3	1426	1359
4 1012	1034	-4 1306	4	1033	1047	7	547	528	3 667	610	-1	1427	1328
5 823	899	-3 506	5	639	609	8	560	582	5 653	683	1	1618	1540
6 1061	965	-2 2266	6	1163	1134				6 681	696	3	1383	1530
8 725	725	0 2061	7	637	603				7 552	635	5	683	653
		2 1009	8	850	825				8 564	532	7	823	803

7, 0, L	531	966	2	827	788	-1	319	298	-1	639	668
9 759 669	1204	678	4	762	700	0	368	332	0	1134	1114
11 636 562	683	668	7, 9, L						1	739	765
-11 712 667	1033	236	-7 599 592			-10	495	486	2	1488	1553
-9 1319 1246	731	L	-6 531 588			-8	597	577	3	647	668
-7 1459 1360	731	355	-4 558 559			-6	823	806	4	1274	1359
-5 1164 1115	599	515	-2 449 468			-4	942	935	5	455	443
-3 1026 987	572	539	0 494 469			-2	925	897	6	1016	975
-2 229 206	402	1026	4 352 438			-1	479	500	7	638	599
-1 1039 988	L	361	7, 10, L			0	1080	1074	8	855	834
0 369 327	415	787	-2 418 359			1	704	715	-8, 5, L		
1 474 452	560	447	0 509 495			2	1238	1303	-8	604	544
2 657 695	289	696	1 481 530			3	650	722	-7	374	329
3 305 333	311	837	2 907 865			4	846	917	-6	619	547
4 310 319	261	778	4 700 653			5	551	504	-4	748	790
7 386 370	409	772	6 545 659			6	779	793	-2	700	709
-9 286 227	979	591	7, 11, L			7	411	405	0	726	693
-7 631 619	704	872	-2 477 533			8	767	786	-8, 6, L		
-5 1170 1207	1066	785	0 771 839			9	361	250	-7	618	587
-3 829 862	393	781	7, 12, L						-6	642	599
-2 282 300	921	L	0 425 541			-10	383	326	-5	908	826
-1 1102 1113	563	761	-8, 0, L			-8	588	587	-4	863	815
1 1727 1787	1106	639	-10 564 521			-6	837	876	-3	818	781
3 1152 1223	528	720	-8 963 882			-5	568	600	-2	633	597
5 830 842	1065	424	-6 1136 1062			-4	377	373	-1	461	471
6 271 314	399	773	-4 1335 1301			-3	478	458	0	739	700
7 1240 1217	10 409 399	558	-2 1084 1081			-2	508	505	1	391	400
8 602 622	L	574	0 1017 980			-1	599	601	2	936	922
9 985 935	425	661	2 852 876			0	672	665	3	539	517
10 450 425	574	L	4 756 842			1	386	386	4	1009	904
-11 462 512	532	573	6 479 467			2	272	301	5	598	519
-9 665 670	591	436	8 1002 961			5	357	318	6	739	678
-7 891 936	592	725	-8, 1, L						8	641	579
-5 566 568	490	581	-10 517 523			-6	681	718	-8, 7, L		
-3 793 832	807	737	-8 718 673			-5	469	537	-8	535	440
	679	697	-6 660 626			-4	840	896			
	920	759	-5 334 317			-3	433	491			
	427	666	-3 358 346			-2	847	882			

B, 8,	L	5	713	721	9, 8,	L	10, 4,	L	11, 3,	L
-7 516	489	7	763	771	-4 536	568	-6 756	801	-3 458	473
-6 559	588	9, 3,	L	L	-2 511	513	-5 426	475	-1 544	557
-5 754	741	-7 375	376		2 658	710	-4 791	891	1 412	426
-4 691	562	-6 413	431		3 427	392	-3 403	447		
-3 655	685	-5 301	284		4 817	709	-2 749	770		
-2 445	474	5 487	494		10, 0,	L	-1 346	323		
0 392	420	9, 4,	L		-6 574	591	0 503	513		
1 671	622	-7 614	621		-4 886	836	2 414	443		
2 614	587	-6 432	337		-2 987	947	3 339	332		
3 559	537	-5 887	930		0 1034	1000	4 370	418		
5 554	552	-4 501	487		2 908	898	10, 5,	L		
B, 10,	L	-3 1123	1145		4 900	1037	-2 506	503		
-3 581	581	-2 599	616		6 670	704	0 398	387		
-1 635	702	-1 1116	1124		10, 1,	L	4 504	455		
1 793	867	0 476	490		1 302	298	10, 6,	L		
3 1009	925	1 985	983		4 422	457	-4 554	581		
9, 0,	L	2 543	547		6 532	614	-2 502	498		
-9 685	650	3 931	966		10, 2,	L	-1 426	492		
-7 679	667	4 433	451		-7 358	351	1 707	646		
-5 890	840	5 639	675		-6 568	548	2 301	351		
-3 887	900	9, 5,	L		-5 352	344	3 713	666		
-1 493	462	-7 549	531		-4 889	886	11, 0,	L		
1 820	813	3 574	572		-3 380	360	-3 990	959		
3 1078	1241	5 634	559		-2 936	908	-1 759	735		
5 946	950	7 407	337		0 685	658	1 504	560		
7 855	792	9, 6,	L		2 784	859	11, 1,	L		
9, 2,	L	-5 873	870		4 802	861	-3 363	392		
-9 619	603	-4 522	510		6 451	432	-1 598	569		
-7 628	672	-3 894	851		10, 3,	L	1 637	661		
-5 1046	1074	-1 572	520		-3 267	272	3 545	648		
-4 524	549	0 288	211		-2 554	565	11, 2,	L		
-3 1038	1093	1 814	794		-1 311	346	-3 744	776		
-2 577	576	2 339	386		0 418	449	-1 617	627		
-1 925	933	3 751	709		1 328	355	3 425	481		
0 553	518	4 780	733		4 550	551				
1 1117	1145	5 587	459		6 325	409				
3 1059	1189	6 588	555							

CALCULATED AND OBSERVED STRUCTURE FACTORS OF
[Ru(SC(NH₂)₂)₆][CF₃SO₃]₂

0, 0,	L	8	735	744	13	248	250	0, 7,	L	7	139	143	11	176	184
2	781	826	9	465	14	281	283	1	100	8	258	257	0,	14,	L
3	424	393	10	512	15	309	314	2	258	10	56	53	0	152	144
4	356	364	11	54	16	110	117	3	139	12	155	149	1	290	281
6	43	56	12	445	17	114	115	4	422	14	76	71	2	151	150
7	751	758	13	310	0,	5,	L	5	130	0,	10,	L	4	108	111
8	480	472	14	367	1	475	460	6	217	0	489	490	5	48	43
9	146	143	15	273	2	217	211	7	208	1	779	743	6	178	178
10	563	552	17	143	3	740	728	8	180	2	410	408	7	166	163
11	153	157	0,	3,	4	198	197	9	238	3	213	213	0,	15,	L
12	304	300	1	95	5	37	41	10	138	137	4	196	2	44	47
14	217	213	3	988	6	156	170	11	63	74	5	125	3	100	95
15	208	212	4	169	7	267	282	12	113	112	6	324	321	93	95
17	106	105	5	341	8	218	228	13	44	52	7	188	189	1,	L
0, 1,	L	20	7	210	9	381	387	14	44	52	8	285	283	42	42
1	17	20	8	82	10	171	168	15	60	54	9	166	164	66	64
2	27	8	9	255	11	59	49	0,	8,	L	11	109	106	182	182
3	788	788	10	127	12	49	52	0	361	359	12	53	58	14	468
4	725	729	11	160	13	158	159	1	262	260	13	146	148	250	254
5	151	145	12	112	14	88	84	2	167	160	0,	11,	L	533	524
6	419	427	13	222	0,	6,	L	3	89	86	1	76	70	460	457
8	142	146	14	69	0	672	657	4	318	317	2	272	274	377	369
9	233	236	16	114	1	454	447	5	506	501	3	208	201	447	445
10	139	137	17	60	2	589	590	6	679	655	4	151	148	120	122
11	128	133	0,	4,	3	529	543	7	551	545	5	97	94	701	710
12	42	29	0	818	4	346	342	8	394	396	7	43	57	673	673
13	153	158	1	901	5	756	739	9	326	327	8	127	129	798	802
15	44	38	2	55	6	607	599	11	136	134	12	81	81	945	978
17	38	42	3	578	7	468	460	12	180	182	0,	12,	L	847	844
0, 2,	L	4	184	192	8	342	336	13	350	341	0	275	275	738	729
0	911	914	5	706	9	70	73	14	273	275	1	416	399	728	708
1	37	38	6	553	10	117	124	15	243	243	2	196	197	899	923
2	416	412	7	704	12	75	72	0,	9,	L	4	67	72	559	539
3	468	472	8	453	13	301	298	1	136	135	6	331	330	879	927
4	171	159	9	127	14	282	285	2	85	83	7	112	111	309	299
5	201	212	10	193	15	298	299	3	312	309	8	277	270	273	271
6	410	404	11	41	16	147	154	5	145	151	9	230	226	764	769
7	965	983	12	256	0	238	246	6	238	246	10	148	152	475	450

1, 1,	12	49	50	1, 4,	L	-12	101	105	1	55	47	11	197	204
8 928	13	75	67	-15 51	61	-9	286	294	3	83	90	12	326	323
9 323	15	44	48	-14 166	165	-8	570	579	4	158	135	13	256	267
10 185	1, 3,	L	L	-13 71	69	-7	834	819	5	406	418	14	269	266
11 144	17	62	62	-12 107	111	-6	968	931	6	123	139	15	178	164
12 118	-16	130	131	-11 75	78	-5	819	790	7	323	319	1, 8,	L	L
13 186	-15	248	252	-10 247	263	-4	520	520	8	134	136	-13	70	67
14 178	-14	566	568	-9 98	105	-3	40	50	9	61	67	-12	73	73
15 332	-13	331	327	-8 112	108	-2	721	705	10	49	49	-10	155	155
16 158	-12	537	543	-7 405	421	-1	714	716	11	225	223	-9	150	149
17 130	-11	371	371	-6 306	301	0	740	726	14	114	108	-7	273	295
1, 2,	-10	116	114	-5 616	610	1	885	891	1, 7,	L	L	-6	293	283
-17 88	-9	492	487	-4 270	276	2	554	557	-16	145	146	-5	83	89
-14 188	-8	399	392	-3 85	88	3	528	504	-15	188	192	-4	325	324
-12 70	-7	477	467	-2 260	273	4	443	437	-14	245	253	-3	52	46
-11 398	-6	487	511	-1 1248	1300	5	99	92	-13	300	310	-2	71	62
-10 67	-5	68	68	0 697	693	6	566	571	-12	127	123	0	174	172
-9 522	-4	200	214	1 859	859	7	134	133	-11	139	138	1	302	296
-8 408	-2	28	31	2 763	760	8	323	323	-10	203	209	2	112	110
-7 128	-1	260	276	4 147	118	9	162	152	-9	309	301	3	155	156
-6 66	0	982	999	5 319	321	12	41	43	-8	447	453	4	242	247
-5 121	1	858	847	6 148	147	13	247	238	-7	618	618	6	360	352
-4 77	2	576	571	7 101	108	14	328	326	-6	538	548	9	79	80
-3 873	3	428	430	8 99	108	15	263	258	-5	188	182	10	191	191
-2 400	4	112	112	9 469	462	16	200	203	-4	98	111	12	108	108
-1 305	5	461	475	10 40	39	1, 6,	L	L	-2	92	102	14	62	60
0 468	6	471	477	11 106	105	-14	98	100	-1	229	218	1, 9,	L	L
1 60	7	735	743	12 75	70	-11	250	259	0	625	608	-14	197	197
2 383	8	874	883	13 60	59	-10	70	79	1	786	795	-13	289	283
3 153	9	472	467	14 141	135	-9	228	215	2	723	706	-12	155	152
4 670	10	242	242	15 242	248	-8	464	481	3	204	208	-10	218	210
5 444	11	428	429	16 55	54	-7	116	114	4	463	457	-9	129	128
6 184	12	169	166	1, 5,	L	-6	265	274	5	130	122	-8	338	337
7 419	13	312	308	-17 116	124	-4	153	168	6	390	400	-7	371	350
8 172	14	220	217	-16 132	136	-3	204	186	7	302	288	-6	529	522
9 229	15	270	271	-15 103	107	-2	685	669	8	466	458	-5	387	380
10 101	16	155	160	-14 143	146	-1	572	585	9	274	269	-4	104	110
11 319				-13 105	102	0	488	483	10	236	219	-3	202	211

1, 9, L	1, 11, L	1, 13, L	2, 0, L	-11	202	197	-1	616	588
-1 327 312	-12 130 129	-10 196 191	-18 104 104	-10	272	273	0	661	636
0 482 473	-10 136 129	-9 46 48	-15 231 236	-8	143	139	1	417	389
1 378 383	-9 71 61	-8 115 109	-14 287 294	-7	321	321	2	594	602
2 642 625	-8 247 250	-7 91 78	-13 597 590	-6	204	184	3	752	759
3 203 207	-7 316 322	-6 65 65	-12 340 335	-5	327	342	4	441	442
4 203 196	-6 392 381	-5 172 164	-11 458 457	-4	212	208	5	235	248
5 276 270	-5 449 447	-3 205 202	-10 145 139	-3	238	242	6	1072	1088
6 207 213	-4 208 210	-1 200 200	-8 598 605	-2	289	298	7	828	829
7 431 424	-3 362 370	0 287 267	-7 783 785	-1	381	363	8	530	530
8 193 185	-2 295 299	1 90 95	-6 1028 1064	0	522	499	9	584	586
9 287 287	-1 416 412	2 248 242	-5 844 849	1	575	557	10	106	100
10 173 167	0 423 402	4 84 76	-4 728 687	2	613	592	11	147	142
11 106 106	1 308 310	5 144 144	-3 1044 1126	3	873	882	13	84	76
12 155 157	2 299 301	6 192 197	-2 232 234	4	149	140	14	309	307
13 167 178	3 102 106	7 385 385	-1 1242 1248	5	303	302	15	201	203
1, 10, L	4 109 113	8 235 236	0 158 123	6	340	335	16	243	250
-12 47 46	5 151 151	9 260 277	1 548 537	7	68	61	2, 3, L		
-11 50 66	6 116 121	1, 14, L	2 486 473	8	195	196	-18	50	54
-10 214 220	7 336 331	-8 76 73	3 287 251	9	278	265	-17	74	84
-9 148 145	8 90 92	-7 75 81	4 432 451	10	296	297	-16	102	100
-8 113 114	9 273 265	-5 82 77	5 376 354	12	56	49	-14	106	112
-7 98 106	10 62 65	-4 53 55	6 564 576	2, 2, L			-13	191	197
-6 103 102	1, 12, L	-1 59 57	7 387 375	-18	165	167	-12	166	162
-4 256 270	-11 75 73	0 81 88	8 329 343	-16	216	214	-11	169	166
-3 131 134	-8 68 66	1 139 138	9 565 550	-15	383	377	-10	128	133
-2 59 56	-7 99 97	2 83 83	10 75 81	-14	369	372	-9	202	216
0 218 215	-6 180 176	3 63 53	11 66 67	-13	536	530	-8	136	131
1 136 134	-5 144 130	1, 15, L	12 73 69	-12	255	258	-7	543	528
2 63 68	-3 53 47	-4 130 123	13 166 165	-11	482	477	-5	236	224
4 332 330	-2 151 148	-3 230 228	14 404 401	-10	185	183	-4	130	120
5 65 61	-1 109 102	-2 180 175	15 289 280	-9	90	92	-3	847	824
6 160 156	0 121 112	-1 227 233	16 348 348	-8	484	486	-2	406	407
7 197 198	1 97 93	0 240 238	2, 1, L	-6	763	763	-1	28	9
8 52 55	2 161 151	1 106 110	-18 55 59	-5	745	739	0	251	238
10 175 173	3 72 73	2 193 197	-15 65 53	-4	936	936	1	490	503
11 56 54	4 43 28		-13 275 280	-3	366	382	2	507	507
12 81 90			-12 269 264	-2	722	714	3	349	359

2, 3, L	11	240	232	-10	380	377	-3	115	120	8	504	505	-6	444	441
4	242	260	135	-9	339	350	-2	182	188	9	350	354	-5	444	445
5	138	139	114	-8	498	495	-1	103	103	10	368	370	-4	367	352
6	40	37	253	-7	629	621	0	171	170	11	194	193	-3	239	248
7	201	210	185	-6	604	593	1	436	435	12	96	94	-2	493	490
8	183	181	L	-5	501	517	2	56	42	13	64	72	-1	170	177
11	170	164	122	-4	180	176	3	174	184	2, 9,	L	L	0	435	446
12	51	53	110	-3	405	421	4	102	96	-14	139	148	1	495	484
13	177	179	212	-2	85	89	6	212	208	-13	129	123	2	352	350
15	72	68	219	-1	558	526	7	232	228	-12	190	189	3	101	94
2, 4,	2, 4,	-10	291	0	802	783	9	176	167	-11	179	181	4	62	55
-17	40	32	163	1	1018	978	10	146	156	-10	42	35	5	95	92
-16	71	79	147	2	939	892	12	74	78	-8	49	50	7	176	172
-15	111	117	434	3	182	178	14	115	115	-7	157	153	8	301	306
-14	241	258	43	4	82	82	2, 8,	L	L	-6	50	45	9	123	122
-13	422	413	95	5	222	218	-15	117	118	-5	194	199	10	142	146
-4	146	142	142	6	259	262	-14	222	217	-4	210	207	12	97	96
-11	422	429	403	7	333	330	-13	224	220	-3	70	74	2, 11,	L	L
-10	122	132	84	8	328	331	-12	294	293	-2	266	272	-10	62	60
-9	158	151	175	9	386	392	-11	170	177	0	405	417	-8	144	143
-8	575	575	496	10	252	258	-10	231	239	1	141	134	-7	100	100
-7	529	531	392	11	271	272	-9	136	129	2	484	482	-5	69	80
-6	976	959	337	12	237	244	-8	247	246	3	323	331	-3	96	98
-5	314	332	389	13	153	149	-7	315	315	4	152	154	-2	136	137
-4	398	371	201	14	201	205	-6	328	330	5	213	208	-1	184	176
-3	77	85	219	15	135	145	-5	298	307	6	179	169	0	110	108
-2	262	256	141	15	135	145	-4	180	189	8	237	244	1	163	159
-1	537	497	357	15	102	101	-3	98	94	9	262	252	2	38	46
0	975	983	129	-14	103	98	-2	74	65	11	90	74	3	79	83
1	743	726	66	-13	131	131	-1	162	162	2, 10,	L	L	4	89	92
2	831	838	80	-12	136	140	0	629	611	-14	161	165	6	139	141
3	421	446	153	-11	110	105	1	713	698	-13	187	182	8	81	73
4	396	389	48	-10	239	238	2	516	515	-12	220	217	10	74	77
6	154	141	L	-8	182	189	3	455	447	-11	107	116	2, 12,	L	L
7	391	388	182	-7	154	164	4	317	324	-10	200	195	-11	65	57
8	355	355	267	-6	86	84	5	354	346	-9	286	282	-9	93	85
9	629	610	313	-5	259	260	6	282	291	-8	160	161	-7	331	318
10	209	202	221	-4	414	428	7	374	373	-7	447	445	-6	192	190

2, 12,	L	284	290	-11	153	145	-4	245	235	4	407	394	10	347	336
-5	361	185	183	-10	65	59	-3	388	385	5	319	322	11	181	181
-4	269	215	213	-8	309	315	-2	350	367	6	172	164	12	118	115
-2	306	315	607	-7	218	224	-1	283	264	7	102	87	13	87	83
0	169	172	37	-6	258	267	0	923	917	8	52	43	14	159	160
1	175	173	261	-5	123	130	1	870	854	9	234	236	3,	6,	L
2	135	139	238	-4	381	379	2	1046	1047	10	186	192	-14	141	149
3	280	271	824	-3	213	228	3	727	742	11	144	144	-12	241	239
4	80	81	596	-2	168	183	4	658	662	12	232	226	-10	44	40
5	334	332	1113	-1	250	236	5	405	425	13	71	75	-9	241	239
6	224	224	1078	0	722	695	6	260	258	14	50	59	-8	295	299
7	299	306	406	1	316	302	7	388	393	15	102	98	-7	145	149
8	298	296	220	2	237	238	8	526	519	3,	5,	L	-6	54	60
9	123	130	157	3	220	216	9	454	435	-16	90	95	-5	144	153
2, 13,	L	606	595	4	30	18	10	410	403	-15	116	122	-3	491	502
-9	88	85	619	5	447	453	11	76	73	-14	232	228	-2	83	99
-5	74	70	558	6	110	110	12	133	129	-13	412	419	-1	535	522
-1	43	40	1055	7	223	229	14	141	138	-12	542	545	0	316	324
4	56	55	735	8	313	314	15	205	201	-11	564	560	1	210	203
2, 14,	L	412	407	9	44	40	3,	4,	L	-10	331	354	3	123	120
-7	180	187	331	10	80	87	-18	105	104	-9	306	304	4	119	116
-6	174	173	59	12	42	36	-14	40	44	-8	297	294	5	370	362
-5	273	282	402	15	100	95	-13	106	115	-7	598	604	8	268	281
-4	189	182	317	3,	3,	L	-11	301	301	-6	480	471	10	80	74
-2	161	157	259	-18	161	163	-10	266	277	-5	872	840	11	41	47
0	158	163	305	-17	196	196	-9	84	80	-4	824	806	13	94	87
1	217	212	153	-16	92	89	-8	280	273	-3	845	819	3,	7,	L
2	122	126	286	-15	283	284	-7	254	241	-2	268	274	-15	66	66
3	217	224	137	-14	156	160	-6	157	151	-1	232	253	-14	267	269
5	253	254	295	-13	432	416	-5	644	626	0	616	624	-13	380	382
6	204	211	305	-12	517	515	-4	553	540	1	224	221	-12	393	394
2, 15,	L	298	L	-11	168	169	-3	378	359	2	104	104	-11	364	372
-3	82	83	46	-10	172	166	-2	772	743	3	300	284	-10	264	253
-2	83	86	56	-9	105	105	-1	424	401	4	258	263	-8	140	149
3, 1,	L	73	71	-8	72	66	0	144	139	5	73	67	-7	170	174
-18	158	163	202	-7	271	279	1	76	73	7	522	530	-6	160	149
-17	180	178	104	-6	415	417	2	139	132	8	525	521	-5	353	354
-16	94	102	284	-5	956	950	3	114	107	9	328	332	-4	438	433

3, 7, L	3, 9, L	7 283	263	3, 13, L	-7 403	408	0 172	171
-3 396 400	-14 204 200	9 82 78	78	-8 87 81	-6 670 697	697	1 164 157	
-2 425 423	-13 280 291	10 102 102	102	-7 92 98	-5 857 858	858	2 96 102	
-1 592 567	-12 249 261	3, 11, L	L	-6 348 346	-4 448 479	479	3 566 546	
0 576 570	-11 356 352	-13 169 177	177	-5 71 72	-2 96 79	79	4 250 255	
1 696 682	-10 157 158	-12 73 68	68	-4 188 186	-1 119 114	114	5 226 228	
2 455 442	-9 36 30	-11 220 209	209	-1 217 207	0 63 65	65	6 138 138	
3 434 436	-8 75 77	-10 113 126	126	0 69 70	1 619 625	625	7 228 240	
5 285 290	-7 328 321	-9 271 268	268	1 450 452	2 741 744	744	8 57 48	
6 206 207	-6 300 307	-8 313 305	305	2 351 352	3 894 899	899	11 77 71	
7 462 470	-5 360 360	-7 332 344	344	3 352 352	4 244 218	218	4, 2, L	
8 462 470	-4 395 378	-6 602 587	587	4 324 329	5 45 39	39	-18 217 221	
9 521 499	-3 207 215	-5 265 283	283	5 137 144	6 115 113	113	-17 222 219	
10 316 326	-2 243 243	-4 456 447	447	6 215 222	7 92 97	97	-16 75 79	
11 123 124	-1 361 358	-3 78 79	79	3, 14, L	8 411 408	408	-15 43 43	
3, 8, L	0 387 406	-1 79 75	75	-5 89 85	9 598 578	578	-14 178 172	
-15 85 83	1 468 460	1 357 358	358	-4 127 123	10 492 486	486	-13 62 64	
-13 118 106	2 354 358	2 322 327	327	-3 106 101	11 526 507	507	-12 487 473	
-12 131 126	3 293 294	3 200 203	203	-2 114 115	12 258 256	256	-11 474 480	
-11 47 48	4 217 216	4 131 128	128	-1 96 91	13 268 256	256	-10 46 41	
-10 56 58	6 221 230	6 226 214	214	0 81 78	14 93 89	89	-9 336 329	
-9 272 270	7 230 232	8 194 187	187	3, 15, L	15 196 197	197	-8 110 116	
-8 115 113	8 366 372	9 170 174	174	-4 214 210	4, 1, L	L	-7 507 507	
-7 176 182	9 307 322	10 128 132	132	-3 112 104	-18 78 84	84	-6 497 505	
-6 176 189	10 186 194	3, 12, L	L	-1 75 86	-17 106 103	103	-5 668 665	
-4 118 122	11 106 100	-12 87 81	81	1 171 179	-16 137 136	136	-4 797 810	
-3 344 351	3, 10, L	-11 43 40	40	4, 0, L	-15 107 108	108	-3 387 375	
-2 251 256	-13 177 176	-10 67 73	73	-19 272 265	-14 141 134	134	-2 332 349	
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3 329 334	-3 160 157	-6 95 89	89	-16 102 113	-11 60 63	63	1 1016 1010	
4 236 231	-2 86 83	-5 204 210	210	-14 251 269	-10 252 247	247	2 747 739	
5 42 39	-1 81 83	-4 116 122	122	-13 323 305	-8 56 61	61	3 990 998	
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9 144 146	1 234 233	-2 79 75	75	-11 777 761	-4 250 238	238	5 328 317	
10 144 139	3 166 163	-1 118 114	114	-10 434 410	-3 370 378	378	6 168 171	
13 55 50	4 124 118	3 74 69	69	-9 824 827	-2 771 748	748	8 441 431	
	5 119 118	8 74 81	81	-8 644 651	-1 88 84	84	9 351 347	

4, 2,	L	-13	280	285	0	88	96	8	425	433	-9	94	95	3	268	280
10	271	-12	549	545	1	165	162	9	334	318	-8	81	75	4	113	106
11	276	-11	364	380	3	200	194	10	331	334	-7	204	200	5	266	278
12	76	-10	153	141	4	373	359	11	180	183	-6	372	386	6	67	68
13	87	-9	325	317	5	165	162	13	71	68	-5	382	372	7	139	138
4,	3,	-8	174	180	6	72	71	4,	7,	L	-4	727	717	8	77	69
-16	88	-7	506	507	7	200	209	-16	108	117	-3	532	510	9	134	133
-12	352	-6	476	488	8	165	170	-14	95	89	-2	505	503	4,	10,	L
-11	188	-5	665	675	9	50	56	-13	140	143	-1	224	219	-14	105	108
-10	94	-4	828	830	11	75	71	-12	50	65	0	331	330	-13	301	305
-9	161	-3	293	321	12	57	57	-11	105	102	1	421	439	-12	313	313
-8	125	-2	392	401	13	64	64	-10	40	36	2	554	549	-11	329	328
-7	539	-1	286	262	14	111	114	-9	67	71	3	570	573	-10	385	373
-6	602	0	174	181	4,	6,	L	-8	45	42	4	481	472	-9	140	144
-5	337	1	543	544	-17	87	91	-7	513	499	5	264	261	-8	196	189
-4	303	2	584	573	-16	132	134	-6	57	70	6	152	157	-7	209	209
-3	247	3	717	737	-15	178	184	-5	536	546	7	346	336	-6	385	380
-2	671	4	268	262	-14	328	329	-4	161	164	8	269	279	-5	493	508
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2	223	9	338	341	-10	277	268	1	103	102	-15	162	174	0	42	45
3	275	10	328	321	-9	67	66	2	89	90	-14	90	87	1	221	233
4	193	11	173	179	-8	130	134	4	78	79	-13	46	46	2	275	268
5	58	13	109	111	-7	205	193	5	110	108	-12	66	60	3	117	131
6	50	14	123	119	-6	675	676	6	348	346	-11	153	148	4	179	172
7	180	4,	5,	L	-5	739	723	7	170	169	-10	73	77	5	118	119
8	322	-16	104	104	-4	827	801	8	82	74	-9	63	66	6	86	81
9	178	-14	72	84	-3	270	268	9	109	118	-8	59	56	7	175	183
10	174	-13	100	106	-2	226	227	11	107	108	-7	153	149	8	106	117
13	94	-12	242	241	0	392	385	12	74	72	-6	220	218	9	273	271
14	125	-10	130	123	1	204	190	4,	8,	L	-5	81	82	10	235	242
-18	132	-9	65	40	2	221	221	-16	49	45	-4	186	190	4,	11,	L
-17	145	-7	353	338	3	262	243	-15	90	91	-3	215	219	-13	62	54
-16	114	-6	401	414	4	256	261	-14	163	169	-2	293	301	-11	147	143
-15	85	-4	63	69	5	191	183	-13	266	275	-1	530	534	-9	58	37
-14	339	-2	308	322	6	422	414	-12	240	246	1	63	68	-8	59	52
		-1	97	83	7	292	293	-11	158	153	2	130	126	-7	248	245

4, 11,	L	0	211	207	5, 2,	L	-6	791	799	3	118	117	9	363	362
-6	173	167	1	218	-18	113	108	-5	717	709	4	112	10	259	260
-5	143	149	2	354	-17	52	50	-4	896	897	5	228	11	272	274
-1	113	121	3	138	-14	124	114	-3	1035	1036	6	446	12	176	181
3	98	91	4	162	-13	77	88	-2	168	182	7	121	13	127	138
5	104	102	5, 1,	5, 1,	-11	85	95	-1	647	643	8	110	5, 6,	L	L
6	115	116	-19	173	-10	186	182	0	443	456	9	96	-17	69	76
7	100	103	-18	235	-9	50	52	1	618	606	10	111	-15	106	108
8	105	103	-17	140	-8	47	49	2	783	775	11	65	-14	92	96
9	133	130	-16	103	-7	388	390	3	670	658	12	120	-11	248	244
4, 12,	L	-13	538	544	-6	146	142	4	498	499	5, 5,	L	-10	110	100
-12	123	122	-12	531	-5	304	302	6	116	107	-18	263	-9	174	170
-11	145	144	-11	794	-4	236	248	7	230	236	-17	195	-8	85	89
-10	258	253	-10	626	-3	404	402	8	112	108	-16	245	-7	241	247
-8	200	185	-9	481	-2	170	174	9	327	319	-15	163	-6	208	222
-7	101	101	-8	152	-1	201	206	10	198	197	-14	226	-5	173	178
-6	207	210	-7	148	0	510	525	11	244	242	-13	367	-4	490	503
-5	294	292	-6	481	1	90	79	12	50	49	-12	428	-3	38	34
-3	381	371	-5	268	2	453	464	5, 4,	5, 4,	L	-11	534	-2	96	94
-1	164	164	-4	297	3	144	145	-18	91	83	-10	362	-1	87	78
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3	171	175	0	733	8	72	60	-13	145	139	-6	606	5	198	205
4	284	281	1	530	10	101	105	-12	77	73	-5	606	7	109	116
5	106	111	2	474	11	66	65	-11	43	50	-4	525	8	40	43
7	135	136	3	260	12	76	72	-10	352	358	-3	159	9	68	63
8	90	94	4	558	5, 3,	L	L	-9	334	328	-2	137	10	86	84
4, 13,	L	5	152	157	-18	242	240	-8	459	457	-1	201	5, 7,	L	L
-10	57	44	6	161	-17	97	97	-7	418	429	0	225	-17	146	162
4, 14,	L	7	343	332	-16	133	129	-6	265	249	1	228	-16	172	177
-8	143	148	8	415	-13	255	250	-5	121	120	2	468	-15	139	141
-7	114	128	9	611	-12	282	289	-4	65	67	3	579	-14	106	107
-6	191	193	10	407	-11	438	429	-3	147	151	4	639	-13	129	122
-5	186	186	11	402	-10	168	158	-1	56	54	5	392	-12	162	167
-3	164	164	12	108	-9	220	214	0	84	93	6	249	-11	224	224
-2	72	72	13	43	-8	346	335	1	192	194	7	185	-9	85	68
-1	149	151	14	115	-7	166	155	2	124	134	8	200	-8	258	255

5, 7,	L	6	106	107	-2	84	81	5, 13,	L	1	404	413	13	52	55
-7 329	311	7 124	120	120	0	216	212	-10 162	163	2 446	462	6,	2,	L	
-6 586	580	8 54	54	54	1	244	252	-9 121	119	3 614	637	-19	87	89	
-5 799	789	5, 9,	L	L	3 127	114	114	-7 141	143	4 576	529	-18	158	161	
-4 819	795	-15 102	101	101	4 70	72	72	-5 238	237	5 600	610	-17	214	213	
-3 307	316	-14 164	169	169	5 90	91	91	-4 292	288	6 384	345	-15	110	112	
-2 180	172	-13 131	139	139	7 121	121	121	-3 295	293	7 582	586	-14	60	60	
-1 164	157	-12 260	241	241	5, 11,	L	L	-2 403	410	8 451	454	-13	164	161	
1 335	335	-11 221	228	228	-13 141	142	142	-1 148	159	9 331	338	-12	549	544	
2 635	622	-10 201	201	201	-12 348	332	332	0 343	332	10 399	373	-11	405	402	
3 698	692	-9 141	143	143	-11 306	307	307	1 245	251	11 167	175	-10	587	579	
4 510	504	-8 146	145	145	-10 361	349	349	2 240	235	12 151	150	-9	186	177	
5 292	293	-7 246	241	241	-9 204	199	199	3 182	191	6, 1,	L	-8	259	259	
6 218	211	-6 295	294	294	-7 238	236	236	5 85	93	-19 54	56	-7	191	196	
8 64	61	-5 427	437	437	-5 278	270	270	5, 14,	L	-18 55	49	-6	346	349	
9 123	121	-4 491	480	480	-4 160	160	160	-7 83	90	-16 45	41	-5	976	993	
10 148	148	-3 374	365	365	-3 41	40	40	-4 56	44	-15 44	54	-4	723	734	
11 104	111	-2 233	227	227	-2 253	253	253	1 79	71	-14 90	86	-3	842	858	
12 75	81	-1 130	133	133	0 204	203	203	6, 0,	L	-13 149	148	-2	853	856	
5, 8,	L	0 79	83	83	1 166	172	172	-19 90	96	-12 135	131	-1	229	233	
-15 134	131	1 298	300	300	2 193	204	204	-18 236	239	-11 90	92	0	569	569	
-13 120	108	2 441	435	435	3 253	249	249	-17 325	329	-10 103	99	1	30	23	
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-10 73	79	4 203	202	202	5 229	223	223	-15 388	394	-8 211	209	3	506	492	
-9 87	84	5 270	267	267	7 128	118	118	-14 409	397	-7 101	101	4	283	286	
-8 77	71	6 94	106	106	8 234	240	240	-13 335	321	-6 257	267	5	434	434	
-7 47	35	7 71	76	76	5, 12,	L	L	-12 795	784	-5 387	395	7	116	109	
-5 331	337	8 158	164	164	-9 131	119	119	-11 393	384	-3 136	135	8	107	108	
-4 151	142	9 160	165	165	-8 98	100	100	-10 391	384	-2 212	218	9	210	200	
-3 288	279	10 181	175	175	-7 206	201	201	-9 111	108	-1 270	301	10	284	270	
-2 57	39	5, 10,	L	L	-6 143	145	145	-7 73	77	0 298	296	11	124	119	
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6, 3,	1	250	250	6, 6,	L	-4	198	210	6, 9,	L	6, 11,	L
-10 221	2 564	572	-17 298	301	-3 282	281	-14 141	140	-13 92	83	-13 92	83
-9 291	3 614	603	-16 172	172	-2 287	272	-12 111	113	-11 165	163	-11 165	163
-8 81	4 315	338	-15 113	117	0 170	161	-11 75	73	-10 70	66	-10 70	66
-6 286	5 376	365	-13 195	202	1 83	84	-9 158	160	-9 75	69	-9 75	69
-5 185	6 86	98	-12 327	317	3 301	297	-8 57	47	-8 112	110	-8 112	110
-4 95	7 192	182	-11 428	428	4 56	52	-7 178	179	-7 77	82	-7 77	82
-3 306	8 132	138	-10 421	416	5 254	262	-6 112	110	-6 71	72	-6 71	72
-2 50	9 218	214	-9 400	389	6 118	126	-5 173	175	-5 86	87	-5 86	87
-1 525	10 308	306	-8 292	288	7 102	103	-4 162	169	-1 113	113	-1 113	113
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2 203	12 201	208	-6 175	178	11 77	82	0 279	284	3 132	130	3 132	130
4 571	6, 5,	L	-5 232	247	6, 8,	L	1 280	267	4 137	141	4 137	141
5 211	-16 98	96	-4 104	101	-16 106	108	2 88	98	5 175	172	5 175	172
7 75	-15 99	96	-3 327	342	-12 95	95	4 85	95	6 44	51	6 44	51
8 90	-14 96	92	-2 126	111	-11 316	314	7 57	63	6, 12,	L	6, 12,	L
9 119	-13 71	68	-1 349	346	-10 457	450	9 63	74	-11 236	233	-11 236	233
10 220	-12 172	170	0 59	54	-9 432	431	6, 10,	L	-10 154	151	-10 154	151
11 79	-11 81	80	1 342	340	-8 297	290	-14 190	184	-9 143	134	-9 143	134
6, 4,	-10 117	111	2 573	561	-7 270	261	-13 211	205	-7 65	58	-7 65	58
-18 225	-9 312	312	3 623	606	-6 365	352	-12 219	218	-6 273	279	-6 273	279
-17 282	-8 98	97	4 508	478	-5 545	540	-11 310	302	-5 147	153	-5 147	153
-16 111	-7 145	156	5 281	276	-4 457	442	-10 147	149	-4 463	452	-4 463	452
-15 198	-6 310	323	6 93	83	-3 491	472	-9 142	128	-3 307	315	-3 307	315
-13 171	-5 267	267	7 74	78	-2 289	289	-7 88	87	-2 269	271	-2 269	271
-12 352	-4 108	113	8 157	156	-1 257	246	-6 254	252	-1 160	169	-1 160	169
-11 474	-3 100	104	9 292	294	0 312	307	-4 356	348	1 228	235	1 228	235
-10 522	-2 151	163	10 325	323	1 482	500	-3 159	165	2 91	95	2 91	95
-9 400	-1 152	138	11 220	229	2 495	491	-2 142	140	3 178	177	3 178	177
-8 349	3 162	160	6, 7,	L	3 327	326	-1 59	43	4 189	193	4 189	193
-7 107	4 442	443	-17 116	113	4 226	237	1 183	177	5 102	108	5 102	108
-6 151	5 60	69	-13 174	169	5 63	56	2 175	185	6, 13,	L	6, 13,	L
-5 633	8 134	135	-11 155	149	7 121	113	3 359	347	-2 60	62	-2 60	62
-4 566	9 62	60	-10 101	94	9 109	109	4 352	368	6, 14,	L	6, 14,	L
-3 534	10 144	138	-8 407	400	10 165	167	5 193	193	-6 168	173	-6 168	173
-2 253			-7 158	159			6 185	189	-5 164	162	-5 164	162
-1 107			-6 66	64			7 121	128	-4 302	307	-4 302	307
0 358			-5 57	53			8 211	215	-3 231	232	-3 231	232

6, 14, L	176	64	57	-1	218	234	7, 5, L	-6	251	250	10	199	204
-2 170	176	-12 328	327	0	101	96	-18 197	-5	166	172	7, 8, L		
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7, 1, L		-8 327	318	2	89	74	-16 199	-3	272	274	-12 109	105	
-19 184	182	-7 297	318	3	494	483	-14 130	-2	227	239	-10 85	78	
-18 205	207	-6 38	61	4	322	323	-13 129	-1	489	488	-9 211	206	
-17 378	366	-5 138	136	5	188	193	-12 381	0	73	74	-8 86	103	
-16 335	329	-4 314	311	6	206	194	-11 357	1	167	169	-6 192	213	
-15 163	169	-3 527	546	8	137	135	-10 510	4	69	76	-5 205	205	
-14 292	294	-2 322	315	9	113	112	-9 439	5	95	106	-4 130	129	
-13 54	50	-1 516	524	10	265	259	-8 132	6	90	89	-3 115	118	
-12 289	301	0 134	126	11	254	245	-7 220	7	47	41	-2 128	132	
-11 437	440	1 61	65	7, 4, L			-6 171	10	77	80	0 148	148	
-10 145	139	2 28	45	-17 91	96		-4 169	7, 7, L	1	178	1	178	
-9 588	580	3 153	130	-16 105	107		-3 324	-17 148	153	153	2 69	76	
-8 71	68	4 59	37	-15 109	104		-2 529	-16 116	112	112	4 95	93	
-7 305	305	5 113	98	-14 114	113		-1 341	-13 90	90	90	7 108	106	
-6 203	197	6 136	124	-13 310	317		0 408	-12 333	323	323	7, 9, L		
-5 375	399	10 72	68	-12 106	102		1 466	-11 400	402	402	-15 99	90	
-4 881	884	11 59	59	-11 243	237		2 639	-10 543	538	538	-13 134	130	
-3 525	530	12 119	114	-9 178	176		3 528	-9 405	398	398	-12 168	174	
-2 393	428	7, 3, L		-8 130	131		4 394	-8 291	292	292	-11 309	313	
-1 184	167	-18 120	119	-7 308	305		5 283	-7 187	192	192	-10 433	427	
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3 964	956	-15 77	83	-2 195	200		8 211	-4 425	427	427	-6 138	138	
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7 123	119	-10 437	435	2 252	256		7, 6, L	0 263	269	269	-2 346	343	
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12 88	88	-6 566	577	7 93	98		-13 66	4	209	219	2 245	237	
7, 2, L		-5 630	619	8 141	137		-12 89	5	130	129	3 122	118	
-17 124	119	-4 947	960	9 130	127		-10 252	7	43	40	4 265	261	
-14 35	29	-3 569	592	-8 80	86		-8 80	8	154	160	5 196	207	
-13 144	151	-2 557	550	-7 319	318		-7 319	9	210	207	6 83	85	

7, 9, L 92	7, 12, L 70	-1 554	570	B, 2, L	81	6 298	283
7 90	-10 68	0 214	213	-18 254	-5 105	7 145	152
B 100 103	-6 50 68	1 546	551	-17 217 210	-4 83	8 94	99
7, 10, L	-2 85 82	2 588	591	-16 310 306	-3 41	9 182	184
-12 83 76	-1 138 148	3 530	507	-15 164 171	-2 62	10 130	132
-10 63 70	0 80 81	4 730	726	-13 142 141	-1 79	B, 5, L	
-9 202 200	1 114 109	5 384	371	-12 221 226	0 161	-15 188	179
-8 126 125	2 89 95	6 388	385	-11 572 562	1 384	-13 106	105
-7 70 69	7, 13, L	9 117 119	119	-10 470 470	2 97	-11 186	184
-5 115 119	-9 106 96	10 97 102	102	-9 786 792	3 93	-10 122	121
-4 111 94	-8 186 188	11 263	265	-8 835 846	4 152	-9 198	203
-3 193 192	-7 179 187	B, 1, L		-7 142 137	5 250	-8 40	45
-2 178 171	-6 266 268	-16 50 46	46	-6 430 441	6 99	-7 125	130
-1 96 97	-5 322 329	-15 63 61	61	-5 185 185	7 79	-5 256	257
1 88 102	-4 288 290	-14 105 104	104	-4 302 299	B, 4, L	-4 131	133
4 85 86	-3 333 333	-13 131 124	124	-3 409 413	-18 188	-3 91	106
7, 11, L	-2 105 106	-12 181 180	180	-2 222 231	-17 196	-2 116	135
-13 146 143	-1 94 91	-11 292 280	280	-1 487 484	-16 282	-1 73	58
-12 96 105	2 54 66	-9 83 78	78	0 95 97	-15 166	0 97	96
-11 184 187	B, 0, L	-8 224 223	223	1 105 107	-14 118	1 208	218
-10 202 204	-19 154 165	-7 121 129	129	2 415 406	-13 158	2 253	251
-8 148 145	-18 367 372	-6 304 311	311	3 362 373	-12 222	3 125	126
-6 60 49	-17 341 331	-5 361 388	388	4 442 436	-11 426	4 77	65
-5 165 173	-16 450 432	-4 241 247	247	5 77 77	-10 368	5 197	203
-4 137 136	-15 301 282	-3 228 227	227	6 184 182	-9 371	6 77	67
-3 374 372	-14 82 74	-2 113 125	125	9 185 177	-8 160	7 59	53
-2 123 126	-13 81 80	-1 80 87	87	10 171 168	-6 213	8 69	63
-1 141 134	-12 216 209	0 75 80	80	B, 3, L	-5 316	10 54	43
0 64 66	-11 268 275	2 221 219	219	-18 47 45	-4 480	B, 6, L	
1 137 140	-10 130 124	3 50 48	48	-17 72 58	-3 642	-17 173	176
2 308 311	-9 408 423	4 72 65	65	-16 89 95	-2 470	-16 211	216
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4 358 351	-7 242 228	7 61 61	61	-14 57 53	0 119	-14 190	199
5 261 265	-6 583 571	8 125 125	125	-13 122 122	1 103	-13 208	204
6 147	-5 257 256	9 107 103	103	-12 104 105	2 195	-12 180	176
	-4 444 453	10 133 135	135	-11 167 173	3 210	-11 270	266
	-3 428 427	11 103 98	98	-9 291 285	4 464	-10 292	302
	-2 203 227			-7 133 128	5 350	-9 194	196

B, 6,	L	B, 8,	L	-1	51	54	-4	129	139	-11	163	162	2	132	135
-7	101	-16	106	0	99	98	-3	96	98	-10	116	119	3	151	150
-5	248	-15	109	3	52	51	-2	155	160	-9	230	233	4	176	181
-4	571	-14	132	6	41	41	-1	62	66	-8	58	59	5	361	354
-3	764	-13	194	7	83	85	0	190	186	-7	345	358	6	249	251
-2	623	-12	351	8, 10,	L	L	1	107	106	-6	225	218	7	247	239
-1	377	-11	417	-14	111	111	2	143	149	-5	170	174	8	138	134
0	186	-10	526	-11	60	54	9, 1,	L	65	-2	70	65	9	165	169
1	101	-9	347	-10	211	215	-18	154	154	-1	170	179	9, 4,	L	L
2	103	-8	211	-9	136	137	-17	268	262	0	46	53	-18	73	79
3	220	-7	254	-7	186	180	-16	152	157	1	131	119	-17	234	227
4	308	-6	104	-5	159	161	-15	209	205	3	58	43	-16	126	124
5	374	-5	218	-4	275	266	-14	57	46	4	98	92	-13	113	111
6	266	-4	319	-3	349	334	-12	262	271	6	164	159	-11	317	315
7	249	-3	432	-2	356	343	-11	225	222	7	129	129	-10	60	68
8	157	-2	394	-1	123	124	-10	643	647	8	81	83	-9	127	134
9	145	-1	220	0	244	242	-9	436	446	9	81	84	-8	170	173
B, 7,	L	0	94	1	166	171	-8	536	548	9, 3,	L	L	-7	310	308
-16	130	1	65	2	202	203	-7	69	67	-18	147	158	-5	91	92
-14	141	2	42	3	271	276	-6	68	72	-17	313	309	-4	375	387
-11	129	3	89	4	263	261	-5	269	270	-16	254	258	-3	374	374
-10	69	4	143	5	273	282	-4	66	66	-15	343	338	-2	180	183
-8	172	5	219	6	146	143	-3	556	553	-14	203	198	-1	323	324
-7	136	6	110	8, 11,	L	L	-2	532	539	-13	233	234	1	124	114
-6	99	7	138	-10	94	86	-1	593	606	-12	282	269	2	253	243
-5	225	8	114	-8	101	103	0	584	590	-11	266	264	3	388	390
-4	59	B, 9,	L	-4	50	54	1	79	82	-10	544	539	4	177	164
-3	261	-15	90	-3	85	85	2	435	443	-9	529	533	5	256	247
-2	51	-13	120	-2	94	91	3	267	277	-8	515	512	9, 5,	L	L
-1	98	-12	76	-1	91	109	4	210	209	-7	241	243	-17	241	239
0	50	-11	139	0	120	117	5	274	273	-6	95	97	-16	228	224
1	68	-10	165	8, 12,	L	L	6	141	138	-5	251	251	-15	239	243
2	221	-7	154	-10	284	270	7	114	115	-4	174	168	-11	115	119
3	58	-6	232	-9	256	257	9	160	161	-3	486	479	-10	267	254
6	142	-5	178	-8	217	212	9, 2,	L	43	-2	362	373	-9	274	272
8	96	-4	227	-7	250	250	-17	44	43	-1	165	168	-8	250	241
		-3	42	-6	90	91	-13	102	94	0	81	81	-7	231	239
		-2	122	-5	157	153	-12	93	96	1	165	163	-6	122	132

9, 5,	L	191	-7	175	175	9, 12,	L	248	260	-9	248	260	7	232	225
-5	294	104	-6	159	153	-5	77	173	164	-8	173	164	8	213	214
-4	365	178	-5	86	90	-4	124	126	202	-7	186	202	10,	3,	L
-3	513	252	-4	321	312	-2	107	106	182	-6	175	182	-18	82	80
-2	518	498	-3	260	258	-1	62	59	86	-5	89	86	-16	52	47
-1	485	477	-2	265	271	10,	0,	L	175	-3	168	175	-15	136	126
0	423	414	-1	177	160	-18	45	38	123	-2	125	123	-12	60	59
1	191	189	0	63	72	-17	193	185	31	-1	38	31	-11	94	89
2	211	214	1	120	123	-16	179	178	78	1	81	78	-10	48	46
3	344	350	-4	354	347	-15	96	93	129	2	125	129	-9	152	152
4	379	374	3	262	260	-14	138	136	106	3	111	106	-8	111	117
5	370	369	5	209	214	-13	61	64	71	4	72	71	-7	116	119
6	205	204	9, 10,	10,	L	-12	277	272	71	5	69	71	-6	124	135
7	125	131	-12	111	114	-11	105	112	54	6	73	54	-5	377	381
-13	67	67	-9	53	47	-10	177	166	79	7	75	79	-4	102	102
-12	40	36	-8	138	139	-9	587	593	L	10,	2,	L	-2	151	155
-11	79	80	-7	80	73	-8	91	95	224	-17	238	224	-1	137	134
-10	144	150	-6	143	141	-7	299	284	281	-16	271	281	0	104	99
-9	208	210	-5	98	89	-6	178	192	288	-15	288	288	1	262	271
-8	115	114	-2	170	169	-5	347	346	325	-14	339	325	2	116	117
-7	255	271	2	81	76	-4	363	370	170	-13	167	170	5	66	63
-6	63	65	9, 11,	11,	L	-3	646	657	325	-12	323	325	7	62	67
-5	93	90	-11	166	175	-2	1008	999	288	-11	289	288	8	71	73
-4	166	160	-10	161	168	-1	638	645	415	-10	413	415	10,	4,	L
-3	105	98	-9	203	205	0	358	358	655	-9	662	655	-17	234	227
-2	213	217	-8	130	131	1	115	111	203	-8	208	203	-16	263	263
-1	111	114	-7	97	93	2	88	89	125	-7	127	125	-15	208	212
0	123	116	-6	121	115	3	88	98	95	-6	84	95	-14	216	210
2	106	116	-4	220	224	5	278	272	99	-4	105	99	-13	46	42
3	38	41	-3	217	218	6	224	225	235	-3	240	235	-12	118	121
4	146	148	-2	304	300	7	131	130	462	-2	460	462	-11	97	93
6	117	112	-1	293	294	8	164	172	262	-1	259	262	-10	238	242
7	81	83	0	188	193	10,	1,	L	299	0	297	299	-9	440	441
-16	262	266	-15	70	68	-15	70	68	209	1	203	209	-8	354	360
-15	271	274	-14	77	71	-14	77	71	153	3	151	153	-7	335	339
-14	244	253	-12	57	62	-12	57	62	128	4	125	128	-6	267	263
			-11	201	195	-11	201	195	310	5	311	310	-5	162	159
			-10	221	216	-10	221	216	304	6	308	304	-4	246	245

10, 4,	L	-2	311	294	5	211	214	11, 1,	L	1	112	111	-6	89	91
-3	251	-1	259	259	10,	9,	L	-17	118	2	123	125	-5	81	87
-2	448	0	190	188	-12	86	80	-16	276	3	164	163	-4	92	97
-1	310	1	165	167	-11	72	73	-15	268	4	102	104	-3	250	250
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1	157	3	242	254	-9	135	143	-13	255	11, 3,	L	L	-1	225	224
2	113	4	266	263	-8	147	144	-12	43	-17	197	196	0	246	247
3	275	5	197	209	-7	134	129	-11	110	-16	359	348	1	76	82
4	248	6	161	165	-5	56	53	-10	89	-15	329	328	2	72	65
5	278	10, 7,	L	L	-4	74	73	-9	354	-14	300	295	11, 5,	L	L
6	235	-12	44	34	-3	73	69	-8	392	-13	212	206	-16	97	103
7	82	-11	97	94	-1	109	103	-7	281	-12	55	62	-15	93	93
10, 5,	L	-9	171	173	1	83	77	-6	475	-11	140	139	-14	73	77
-15	83	-8	140	142	10, 10,	L	L	-5	210	-10	115	110	-13	80	78
-12	66	-7	117	124	-11	95	92	-4	354	-9	262	272	-12	52	53
-11	83	-6	231	235	-10	154	163	-3	383	-8	253	250	-11	229	232
-9	126	-4	73	80	-9	149	154	-2	368	-7	81	80	-10	336	339
-7	76	-3	58	56	-8	247	255	-1	363	-6	167	169	-9	505	501
-5	239	-2	77	80	-7	177	178	0	119	-4	90	92	-8	555	553
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-1	178	2	97	95	-5	196	195	2	58	-2	241	239	-6	262	259
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1	218	10, 8,	L	L	-3	356	344	4	205	0	202	200	-4	127	127
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-16	251	-13	116	120	0	174	183	7	152	3	146	147	-1	457	449
-15	190	-11	202	197	1	99	101	11, 2,	L	4	257	259	0	303	306
-14	131	-10	276	284	2	113	109	-13	159	5	296	288	1	248	242
-13	65	-9	283	284	10, 11,	L	L	-12	165	6	337	340	2	129	134
-12	59	-8	372	376	-8	90	82	-11	173	11, 4,	L	L	3	61	58
-10	186	-7	220	228	-7	54	56	-10	211	-17	65	65	4	78	74
-9	341	-6	79	80	-6	136	126	-7	57	-16	115	117	11, 6,	L	L
-8	441	-1	148	148	-5	84	88	-6	72	-14	65	62	-16	42	35
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-4	367	3	200	205	-3	39	51	-8	73	-8	73	80	-11	110	111
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11, 6,	L	11, 9,	-2	104	108	-5	140	136	-5	151	149	-3	160	170
-5 44	44	-12 63	-1	88	84	-4	54	59	-4	63	61	-2	262	252
-4 116	114	-11 92	0	169	178	-3	156	158	-3	225	229	-1	293	300
-3 68	69	-10 180	2	87	183	-2	202	206	-2	274	273	0	259	258
-2 65	61	-9 204	3	99	103	-1	262	262	-1	317	311	1	117	113
1 125	134	-8 237	4	112	116	0	350	353	0	308	302	12, 7,	L	L
3 130	130	-7 240	5	152	156	1	217	221	1	129	135	-14	85	81
11, 7,	L	-6 95	12, 1,	L	L	2	220	221	2	133	130	-12	98	93
-15 211	208	-5 66	-15	113	112	3	180	182	4	49	62	-10	92	95
-14 124	124	-3 121	-14	143	141	4	217	216	12, 5,	L	L	-9	97	88
-13 101	96	-2 208	-13	141	139	5	257	255	-15	72	61	-6	94	97
-12 94	91	-1 218	-12	119	117	12, 3,	L	L	-14	67	66	-5	93	96
-11 153	157	0 259	-11	60	50	-15	89	84	-12	46	56	-4	97	100
-10 249	246	1 153	-9	53	58	-14	55	66	-11	56	55	-3	110	114
-9 322	316	11, 10,	-8	57	43	-11	83	86	-10	66	65	-1	121	120
-8 320	320	-8 56	-6	52	51	-10	124	119	-9	74	72	12, 8,	L	L
-7 168	160	-6 76	-5	82	79	-9	108	113	-8	124	119	-12	116	117
-3 94	97	-5 122	-4	58	66	-8	157	161	-7	82	85	-11	73	82
-2 183	182	11, 11,	-3	43	42	-6	50	55	-6	50	43	-10	64	68
-1 343	340	-7 252	-2	110	111	-5	199	206	-5	117	112	-9	105	112
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2 135	156	-4 156	1	115	115	-2	97	91	-2	108	102	-3	127	129
3 117	124	12, 0,	2	73	74	1	51	53	1	59	62	-2	279	276
4 120	125	-17 155	4	58	54	2	62	62	2	76	74	-1	317	314
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-14 55	44	-15 297	12, 2,	L	L	12, 4,	L	L	-15	207	205	12, 9,	L	L
-12 79	73	-14 261	-17	224	220	-16	115	126	-14	185	185	-10	49	50
-11 95	86	-13 200	-16	216	215	-15	249	255	-13	197	197	-9	68	79
-9 72	73	-12 72	-15	304	300	-14	217	209	-12	184	185	-7	93	97
-8 65	59	-10 279	-14	227	225	-13	219	219	-11	172	172	-2	152	140
-7 63	65	-9 400	-13	161	166	-12	145	145	-10	242	245	-1	83	85
-6 56	58	-8 577	-11	72	70	-11	70	72	-9	287	288	12, 10,	L	L
-5 72	72	-7 533	-10	121	117	-10	200	196	-8	317	316	-6	191	187
-3 58	60	-6 435	-9	148	142	-9	257	263	-7	237	248	13, 1,	L	L
-2 88	91	-5 348	-8	252	260	-8	369	354	-6	153	146	-16	107	110
2 96	100	-4 165	-7	288	293	-7	288	285	-5	161	159	-15	177	182
		-3 260	-6	272	276	-6	169	171	-4	93	93	-14	241	245

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-13 113	111	-6 201	-6 201	201	201	-10 60	64	-4 121	115	115	-5 162	157	157	-11 102	97
-12 138	142	-5 133	-5 133	132	132	-5 103	105	-3 51	57	57	-4 107	110	110	-10 136	129
-11 145	140	-4 139	-4 139	136	136	-3 105	99	0 68	65	65	-3 93	84	84	-9 116	116
-10 170	176	-3 131	-3 131	130	130	-2 91	95	14, 2,	L	L	-2 70	69	69	-8 193	185
-9 414	404	-2 276	-2 276	274	274	0 56	52	-14 178	174	174	-1 138	143	143	-7 243	237
-8 372	361	-1 291	-1 291	285	285	13, 7,	L	-13 198	197	197	14, 5,	L	L	-6 293	291
-7 415	416	0 317	0 317	315	315	-12 125	129	-12 113	109	109	-10 74	71	71	-5 278	274
-6 251	253	1 274	1 274	271	271	-8 81	85	-11 94	94	94	-9 55	53	53	-4 132	137
-5 92	90	2 186	2 186	183	183	-7 158	155	-9 116	118	118	-8 106	108	108	15, 4,	L
-3 70	63	13, 4,	13, 4,	L	L	-6 170	173	-8 204	201	201	-7 69	66	66	-10 160	157
-2 134	137	-13 68	-13 68	73	73	-5 172	169	-7 181	181	181	-6 47	46	46	-9 139	132
-1 215	219	-12 91	-12 91	90	90	-4 179	184	-6 298	287	287	-4 77	77	77	-6 76	75
0 284	283	-8 106	-8 106	109	109	-3 197	192	-5 176	167	167	14, 6,	L	L	-5 49	42
1 266	273	-7 182	-7 182	188	188	-2 213	214	-4 170	173	173	-11 65	67	67		
2 146	150	-6 248	-6 248	250	250	-1 200	203	-3 184	185	185	-9 119	120	120		
3 130	128	-5 114	-5 114	118	118	13, 8,	L	-2 209	205	205	-8 173	183	183		
13, 2,	L	-4 147	-4 147	143	143	-6 44	43	-1 315	320	320	-7 287	273	273		
-16 92	94	-3 58	-3 58	49	49	-4 55	54	0 253	255	255	-6 262	263	263		
-15 98	96	0 65	0 65	59	59	14, 0,	L	14, 3,	L	L	-5 144	145	145		
-14 60	54	13, 5,	13, 5,	L	L	-14 274	276	-14 95	96	96	14, 7,	L	L		
-12 43	48	-14 333	-14 333	329	329	-13 337	327	-11 38	25	25	-7 94	89	89		
-11 54	51	-13 290	-13 290	286	286	-12 281	283	-10 101	95	95	15, 1,	L	L		
-10 80	79	-12 265	-12 265	266	266	-11 319	313	-9 107	105	105	-12 267	263	263		
-5 35	36	-11 130	-11 130	131	131	-10 209	203	-8 168	170	170	-11 106	106	106		
-3 92	95	-10 146	-10 146	149	149	-9 198	193	-7 131	132	132	-10 69	68	68		
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-1 70	69	-8 205	-8 205	202	202	-7 129	132	-4 101	95	95	-7 94	93	93		
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2 56	56	-6 275	-6 275	277	277	-1 207	203	-2 81	82	82	-5 224	221	221		
3 65	67	-5 225	-5 225	224	224	0 196	193	-2 14, 4,	L	L	-4 99	98	98		
13, 3,	L	-4 147	-4 147	152	152	14, 1,	L	-13 205	203	203	-3 162	161	161		
-15 148	150	-3 88	-3 88	84	84	-12 37	43	-12 106	109	109	15, 2,	L	L		
-14 222	216	-2 164	-2 164	165	165	-11 71	70	-11 102	103	103	-8 88	84	84		
-13 91	90	-1 112	-1 112	118	118	-9 52	49	-10 76	73	73	-6 69	69	69		
-12 77	79	0 94	0 94	95	95	-8 71	70	-9 161	159	159	-5 47	40	40		
-9 161	158					-7 96	96	-8 204	202	202	-3 39	24	24		
-8 197	192					-6 85	81	-7 237	234	234					

C₄Cl₄O₂S at 295K

0, 0,	L	0, 6,	L	-11	85	101	0 1593	1350	10	117	132	1 113	133
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4 1242	1274	1 144	139	-9	88	99	2 240	256	12	81	82	3 164	165
6 978	1068	2 50	36	-8	525	540	3 151	133	13	91	87	4 179	206
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10 221	221	4 182	193	-5	637	637	5 426	392	-13	7, 7,	L	7 63	53
12 145	151	5 417	449	-4	690	677	6 266	224	-12	64	65	1, 11,	L
16 38	46	6 162	143	-3	918	901	7 521	516	-11	72	69	0 92	95
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1 920	933	9 164	172	0	728	613	10 201	204	-8	99	41	-14	244
2 1341	1323	10 75	75	1	636	586	11 193	200	-7	245	101	-12	75
3 718	711	11 50	42	2	81	54	12 61	50	-4	117	258	-10	598
4 87	97	12 143	139	3	213	215	13 113	98	-3	205	239	-8	947
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6 398	411	14 52	50	5	147	148	1, 5,	L	-1	186	206	-2	732
7 1060	1067	0, 8,	L	6	450	428	-14	74	0 295	282	2 1249	0 254	260
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9 477	472	1 76	95	8	247	235	-12	36	3 184	186	10 528	12 212	227
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11 354	361	5 214	211	11	250	244	-9	296	5 185	156	16 59	2, 2,	L
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13 150	135	7 128	139	13	139	133	-7	423	7 110	116	-15	65	67
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-11 153 131	-3	682	670	9	64	74	-13	149	159	-4	522	517	10	86	84
-10 309 296	-2	710	664	10	93	78	-11	189	188	-3	36	49	12	48	70
-9 601 615	-1	604	573	11	83	86	-10	68	66	-2	806	766	14	46	51
-8 60 66	0	718	653	12	63	58	-9	200	226	-1	234	238	5, 7, L	5, 7, L	L
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1 1519 1488	10	211	192	-3	227	219	0	302	317	8	185	188	-3	169	146
2 166 151	12	100	99	-2	76	75	1	383	371	9	145	148	-1	336	349
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6 254 252	4, 6, L	63	66	2	251	234	4	161	139	13	124	124	3	109	130
7 439 417	-13	96	100	4	146	145	5	52	48	14	35	37	4	58	62
8 204 210	-12	96	100	5	109	128	6	192	194	5, 5, L	5, 5, L	L	5	280	273
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11 215 198	-9	100	107	8	95	84	9	260	268	-11	96	107	8	83	78
12 200 204	-8	48	50	9	79	74	11	292	304	-10	156	168	10	48	40

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-5 38	1	31	11	-13	68	71	-2	75	85	-8	54	45	8	200	199
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-6 239	226	-6 200	200	-6	0	308	306	-1	161	162	-6	133	128	-4	68	56
-5 486	501	-5 190	206	-5	1	334	332	0	172	154	-4	259	258	-3	94	85
-4 482	506	-4 212	224	-4	4	220	231	1	205	215	-2	598	594	-2	165	171
-2 578	591	-3 94	84	-3	5	260	280	2	216	210	0	543	561	-1	565	576
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4 258	256	2 211	198	2	13, 3,	102	L	7	78	73	10	98	100	9	100	95
5 483	523	3 95	101	3	-12	102	105	8	100	100	14, 2,	L	L	14, 6,	L	L
6 214	218	4 60	55	4	-10	56	54	10	132	130	-13	90	87	-10	101	104
9 80	85	5 60	59	5	-8	357	371	11	49	69	-12	77	52	-9	79	79
12, 4,	L	6 116	109	6	-6	72	79	13, 7,	L	L	-11	196	180	-8	233	223
-13 207	181	8 54	34	8	-5	55	54	-10	60	56	-10	255	249	-7	67	63
-10 98	122	10 52	50	10	-4	400	403	-9	115	110	-9	494	480	-6	55	69
-9 66	49	11 49	48	11	-3	368	391	-8	57	36	-8	161	174	-5	60	71
-7 536	563	12, 8,	L	12	-1	463	460	-7	43	67	-7	61	67	-3	274	294
-6 182	146	-8 129	142	-8	2	98	90	-6	77	98	-6	222	213	-2	179	166
-4 345	348	-7 98	108	-7	3	51	39	-5	112	113	-5	214	227	-1	86	95
-3 37	58	-5 159	154	-5	4	271	273	-4	48	45	-4	459	458	0	68	68
-2 588	592	-4 54	50	-4	5	286	293	-3	152	134	-3	556	565	3	102	105
-1 219	241	-2 123	138	-2	6	248	235	-2	237	225	-2	327	334	5	102	111
1 192	209	-1 184	179	-1	7	77	70	0	142	132	-1	303	300	9	91	94
2 156	151	0 54	63	0	8	55	62	1	117	107	0	134	133	14, 8,	L	L
3 121	133	1 60	50	1	9	67	46	2	135	140	1	319	329	-6	120	128
4 120	121	2 108	103	2	10	80	88	4	154	147	3	179	174	-5	64	67
6 156	147	8 75	76	8	12	44	33	13, 9,	L	L	5	56	52	-3	63	69
7 205	225	13, 1,	L	13	13, 5,	97	L	-5	97	78	11	96	105	0	115	115
8 59	47	-13 55	51	-13	-12	97	91	-2	94	91	12	85	81	4	77	72
11 101	101	-11 122	130	-11	-11	74	89	0	88	89	14, 4,	L	L	6	68	61
12 42	41	-9 124	154	-9	-10	102	92	1	131	130	-12	83	91	15, 1,	L	L
13 83	77	-8 44	25	-8	-9	42	8	2	91	90	-11	164	168	-13	113	121
12, 6,	L	-6 49	58	-6	-8	77	78	3	64	53	-10	192	164	-12	168	169
-11 68	68	-5 229	233	-5	-7	131	150	14, 0,	L	L	-9	55	40	-10	138	145
-10 54	59	-4 324	322	-4	-6	192	202	-14	88	77	-8	159	148	-9	100	83
-9 232	256	-3 65	66	-3	-6	66	60	-12	282	254	-7	258	246	-8	92	95
-8 82	92	-2 346	366	-2	-4	36	12	-10	239	256	-6	233	231	-6	151	147

15, 1,	L	3	193	193	-6	51	81	16,	8,	L	8	76	69	18,	2,	L
-5 152	152	4	64	64	-5	169	159	-5	83	94	17,	5,	L	-11	52	43
-3 357	355	6	63	63	-4	120	119	-4	108	97	-9	56	52	-10	61	47
-1 115	119	8	133	124	-3	294	295	-3	91	89	-8	79	71	-9	124	106
0 300	320	15,	7,	L	-2	269	268	1	67	80	-7	65	80	-8	159	146
2 61	54	-8	58	46	-1	88	79	2	74	70	-5	118	127	-7	317	314
3 403	409	-7	141	129	0	162	171	17,	1,	L	-3	48	46	-6	135	150
4 215	209	-6	171	145	1	253	291	-11	158	146	-2	63	57	-4	143	131
5 240	232	-5	155	146	2	129	110	-9	83	94	-1	179	183	-3	106	110
6 123	136	-2	103	113	3	282	288	-8	141	142	0	45	50	-2	192	185
7 106	96	-1	171	162	4	154	146	-5	78	63	1	106	116	-1	406	409
8 43	51	0	240	238	5	90	85	-4	138	143	3	104	113	0	270	268
10 66	52	1	63	40	7	68	87	-3	83	88	4	122	121	1	182	187
11 37	43	2	53	54	8	124	139	-2	119	128	5	94	76	3	188	203
15, 3,	L	3	69	78	16,	4,	L	-1	35	31	6	84	69	4	58	68
-12 80	84	4	69	68	-11	207	188	0	252	271	10	80	81	5	42	31
-11 169	177	6	67	67	-8	58	73	1	59	59	17,	7,	L	6	59	45
-7 155	164	8	56	61	-5	421	426	2	102	111	-7	73	73	7	60	48
-4 129	126	15,	9,	L	-3	104	83	3	136	132	-5	105	102	18,	4,	L
-2 238	235	-1	87	84	-2	178	185	4	98	89	-3	85	83	-9	181	178
-1 163	172	1	63	47	-1	53	40	5	61	85	-2	53	46	-8	166	147
0 165	171	16,	0,	L	0	105	112	7	95	104	0	88	86	-7	71	63
1 281	293	-12	112	92	1	152	156	8	161	165	1	69	65	-6	112	110
2 482	492	-10	231	211	3	117	94	9	129	119	2	68	62	-5	65	69
3 44	57	-8	127	111	11	45	49	17,	3,	L	4	73	73	-4	102	99
5 139	120	-6	350	316	16,	6,	L	-10	53	56	6	85	76	-3	149	158
6 138	148	-4	777	817	-7	202	217	-8	80	75	18,	0,	L	-2	34	60
8 86	71	-2	558	556	-5	129	128	-7	33	11	-12	151	136	0	91	101
11 78	61	0	352	348	-4	157	144	-6	260	266	-10	249	222	1	178	179
15, 5,	L	2	348	331	-3	74	74	-4	91	94	-8	287	299	2	59	29
-11 122	137	4	69	83	-2	121	129	-2	309	304	-4	86	82	3	53	55
-10 77	91	6	71	70	-1	138	140	0	146	130	-2	239	221	5	80	87
-9 212	209	8	104	128	1	93	96	1	280	282	0	154	150	9	75	66
-7 235	223	16,	2,	L	2	122	123	2	90	81	2	271	284	18,	6,	L
-4 142	118	-11	63	58	3	122	122	4	152	145	4	106	122	-8	80	80
-2 101	91	-9	35	41	4	104	95	5	49	54	6	75	106	-6	183	171
-1 45	34	-8	276	272	7	62	71	6	84	80	8	207	209	-4	49	52
2 59	59	-7	282	281				7	162	163				-3	38	37

18, 6, L	19, 5, L	5, 100	88	22, 2, L	24, 2, L
-1 87 104	-7 132 135	7 66	57	-7 72 62	-3 119 114
0 188 165	-5 138 125	20, 6, L	L	-5 152 155	-2 61 64
1 51 35	-4 65 55	-5 141 153	L	-4 72 82	25, 1, L
5 74 79	-3 49 47	-4 71 51	51	-1 55 64	-4 63 58
19, 1, L	0 103 103	1 65 64	64	0 49 34	0 61 62
-11 81 86	3 38 42	2 78 75	75	1 107 101	
-10 128 128	6 44 49	4 60 72	72	2 115 118	
-9 84 61	8 74 68	21, 1, L	L	3 102 111	
-8 139 143	19, 7, L	-9 145 142	142	5 54 60	
-7 56 58	-4 121 111	-8 74 64	64	6 46 47	
-4 77 75	0 67 75	-6 113 109	109	22, 4, L	
-3 118 113	2 116 113	-3 95 80	80	-6 73 56	
-2 107 95	4 56 52	-2 122 129	129	-5 60 55	
-1 246 248	20, 0, L	1 90 83	83	-1 87 93	
0 79 63	-8 166 161	9 57 53	53	2 62 73	
1 122 116	-4 181 170	21, 3, L	L	23, 1, L	
2 108 127	-2 400 436	-5 99 80	80	-7 94 79	
4 74 71	0 239 246	-4 126 135	135	-6 109 110	
5 160 145	2 110 95	-2 104 106	106	-1 101 100	
6 62 59	4 122 112	0 183 179	179	0 49 33	
7 79 65	20, 2, L	1 54 50	50	1 73 81	
19, 3, L	-8 81 84	3 139 146	146	2 74 60	
-10 119 123	-6 163 154	4 70 66	66	23, 3, L	
-9 132 141	-5 230 216	21, 5, L	L	-6 41 45	
-8 63 69	-4 92 102	-6 64 58	58	-5 85 82	
-7 84 82	-1 78 101	-5 73 77	77	-4 89 74	
-6 74 56	0 60 72	-2 76 68	68	-2 75 67	
-4 83 69	1 103 101	5 47 53	53	-1 49 53	
-3 103 91	4 50 26	6 68 65	65	1 84 77	
-2 57 66	5 67 68	22, 0, L	L	4 57 56	
-1 96 96	6 39 35	-8 120 102	102	5 58 60	
2 141 147	20, 4, L	-6 165 168	168	24, 0, L	
3 130 134	-3 197 197	0 203 195	195	-6 69 75	
4 135 136	-1 121 119	2 54 46	46	0 106 133	
7 48 35	0 70 71	4 45 55	55	2 46 46	
8 61 55	1 70 60			6 50 34	
	3 64 45				

CALCULATED AND OBSERVED STRUCTURE FACTORS OF
C₄Cl₄O₂S at 150K

0, 0,	2 1326	L	10 297	317	0, 10,	L	12 457	472	15 110	110	-9	89	65
2 1326	1354		11 93	106	0 192	171	14 243	240	17 162	171	-8	296	298
4 1319	1342		12 136	128	1 267	251	15 476	472	18 89	94	-7	495	514
6 1249	1307		13 55	50	2 248	255	16 174	173	1, 5,	L	-6	125	114
8 246	263		14 177	185	3 59	54	17 169	166	-17 162	169	-4	119	109
10 525	539		15 284	288	4 226	206	1, 3,	L	-15 189	187	-3	425	442
12 384	402		16 213	214	5 269	293	-17 356	368	-14 182	198	-2	93	91
0, 2,	L		17 82	77	6 267	282	-16 54	54	-13 109	122	-1	392	399
0 135	102		0, 6,	L	7 138	151	-15 140	147	-11 35	24	0	590	594
1 1084	1077		0 927	979	1, 1,	L	-14 228	223	-10 150	172	1	56	70
2 1442	1430		1 79	84	-18 206	218	-13 253	259	-9 638	689	2	87	90
3 667	664		2 79	75	-16 235	241	-12 84	75	-8 732	777	3	289	284
4 172	196		3 1231	1296	-15 63	75	-11 671	679	-7 780	815	4	477	485
5 1077	1959		4 281	276	-14 113	107	-10 263	271	-6 147	160	5	236	208
6 476	494		5 554	577	-13 320	330	-9 512	532	-5 232	255	6	455	497
7 1402	1406		6 280	286	-12 110	113	-8 598	627	-2 291	304	7	208	211
8 735	780		7 190	206	-11 110	116	-7 207	209	-1 188	201	8	209	192
9 803	833		8 180	170	-10 975	1020	-6 381	388	0 158	149	9	195	178
10 328	346		9 300	322	-9 38	33	-5 201	207	1 702	713	11	205	203
11 609	634		10 309	312	-8 702	735	-4 1086	1069	3 59	47	12	513	475
12 366	382		11 189	188	-7 461	465	-3 130	132	4 403	398	1, 9,	L	L
13 351	339		12 346	330	-6 47	28	-2 384	376	5 1125	1148	-9	94	92
14 85	81		13 133	141	-5 780	805	-1 1538	1514	6 693	679	-8	323	340
15 238	241		15 115	114	-4 699	730	0 1622	1560	7 394	394	-7	621	616
16 92	89		16 119	105	-3 1023	1010	1 403	378	8 497	479	-5	81	78
17 123	115		0, 8,	L	-2 239	248	2 330	350	9 92	91	-4	194	173
18 130	136		0 283	279	-1 987	977	3 132	94	10 265	274	-3	524	527
0, 4,	L		1 129	136	0 635	597	4 798	767	11 334	316	-2	33	6
0 207	201		2 240	247	1 552	506	5 538	493	12 242	260	0	49	18
1 1510	1537		3 479	496	2 72	35	6 277	229	13 302	299	-1	671	681
2 762	769		4 73	83	3 203	208	7 931	934	14 462	450	1	216	235
3 436	429		5 382	381	4 308	304	8 221	214	15 203	202	2	336	356
4 810	827		6 375	364	5 274	259	9 434	442	16 80	86	3	353	350
5 227	222		7 271	266	6 492	493	10 230	226	17 122	116	4	429	455
6 155	171		8 141	133	8 361	352	11 364	371	1, 7,	L	5	415	427
7 779	786		9 106	91	9 732	744	12 39	36	-12 152	158	6	43	73
8 719	731		10 100	98	10 59	28	13 331	330	-11 105	110	7	155	159
9 63	53		11 172	175	11 457	471	14 310	321	-10 360	384	9	49	46

1, 11, L	-4	85	59	-2	553	592	5	660	667	4	412	409	3, 3, L
-1 160 162	-3 2214	2284	-1 269	-1 592	592	592	6 355	373	373	5 214	201	409	-18 119
0 174 185	-2 1712	1789	0 506	0 535	535	535	7 138	129	129	6 63	89	201	-17 118
1 73 87	-1 1397	1382	1 1176	1 1124	1124	1124	8 86	69	69	3, 1, L	L	89	-16 147
2, 0, L	0 24	28	2 371	2 380	380	380	9 194	198	198	-18 131	133	133	-15 220
-18 87 84	1 1570	1541	3 441	3 443	443	443	10 222	210	210	-17 116	131	131	-14 210
-16 91 111	2 526	523	4 1303	4 1273	1273	1273	11 87	101	101	-15 288	288	288	-13 197
-14 465 497	3 1025	983	5 343	5 345	345	345	12 231	185	185	-14 463	480	480	-12 514
-12 135 121	4 1618	1543	6 169	6 147	147	147	13 237	228	228	-13 457	449	449	-11 443
-10 946 1078	5 1285	1291	7 974	7 964	964	964	2, 8, L	L	L	-12 205	208	208	-10 310
-8 1181 1234	6 373	370	8 438	8 433	433	433	-11 87	92	92	-11 210	234	234	-9 439
-4 2534 2770	7 216	225	9 63	9 45	45	45	-9 96	86	86	-10 285	309	309	-8 372
-2 791 788	8 105	106	10 150	10 162	162	162	-8 344	362	362	-9 312	336	336	-7 130
0 237 229	9 31	6	11 271	11 275	275	275	-7 415	437	437	-8 173	181	181	-6 201
2 1299 1330	10 296	319	12 99	12 93	93	93	-6 366	381	381	-7 970	1016	1016	-5 871
4 872 876	11 417	431	13 325	13 310	310	310	-5 376	393	393	-6 487	476	476	-4 443
6 2025 2084	12 85	89	14 94	14 81	81	81	-4 254	255	255	-5 36	21	21	-3 149
8 106 113	13 137	132	15 95	15 102	102	102	-2 191	188	188	-4 1260	1273	1273	-2 97
10 793 806	14 60	52	16 89	16 91	91	91	-1 374	368	368	-3 456	445	445	-1 960
12 496 529	15 58	66	17 193	17 182	182	182	0 308	308	308	-2 989	961	961	0 123
14 273 266	16 167	177	2, 6, L	L	L	L	1 353	345	345	-1 195	203	203	1 237
16 151 145	17 87	85	-13 245	245	245	245	2 370	370	370	0 540	532	532	2 403
18 78 79	18 74	82	-12 120	119	119	119	3 436	436	436	1 1424	1354	1354	3 534
2, 2, L	2, 4, L	L	-11 377	390	390	390	6 57	61	61	2 643	586	586	4 74
-18 152 165	-16 81	51	-10 174	150	150	150	7 206	213	213	3 44	39	39	5 613
-17 149 142	-15 242	229	-9 129	149	149	149	8 541	568	568	4 538	507	507	6 796
-16 234 237	-14 271	269	-7 617	653	653	653	9 206	224	224	5 706	718	718	7 441
-15 232 225	-13 90	79	-6 54	71	71	71	10 170	176	176	7 684	693	693	8 352
-14 208 224	-12 184	159	-5 245	257	257	257	11 102	88	88	8 253	249	249	9 110
-13 250 240	-11 339	328	-4 314	336	336	336	2, 10, L	L	L	9 420	423	423	10 723
-12 30 30	-10 244	244	-3 45	35	35	35	-5 148	150	150	10 333	348	348	11 455
-11 82 87	-9 144	143	-2 302	297	297	297	-4 128	111	111	11 405	421	421	12 182
-10 410 429	-8 173	195	-1 637	653	653	653	-3 123	102	102	12 602	601	601	13 415
-9 117 102	-7 393	404	0 497	496	496	496	-2 438	433	433	13 123	124	124	14 118
-8 255 275	-6 92	88	1 370	388	388	388	-1 112	121	121	14 72	71	71	15 260
-7 270 264	-5 928	985	2 139	133	133	133	0 113	121	121	16 81	78	78	16 159
-6 118 131	-4 344	336	3 481	466	466	466	1 415	440	440	17 64	69	69	17 171
-5 297 316	-3 753	791	4 363	353	353	353	2 161	175	175	18 171	171	171	18 171

4, 10,	L	14	356	365	5, 5,	L	-8	509	531	6, 0,	L	0	1356	1355
-1	127	125	16	178	167	91	-7	280	255	-18	61	1	226	187
0	288	289	17	317	304	251	-6	620	624	-16	262	2	137	112
1	159	148	18	65	69	218	-5	362	361	-14	58	3	519	543
2	130	105	5, 3,	L	L	267	-4	278	267	-12	495	4	559	536
3	158	147	-16	147	169	267	-3	228	207	-10	138	5	997	964
6	270	253	-15	437	401	110	-2	72	26	-8	747	6	1086	1086
5, 1,	L	-14	64	53	-10	278	-1	548	560	-6	2091	7	1259	1262
-17	234	234	-13	105	119	36	0	64	69	-4	383	8	481	482
-16	142	140	-12	65	53	38	1	554	566	-2	3129	9	95	93
-15	144	153	-11	252	258	831	2	424	420	0	511	10	102	104
-14	284	291	-10	49	29	286	3	204	218	2	586	11	332	338
-13	348	370	-9	554	553	656	4	60	54	4	719	12	106	111
-11	354	371	-8	93	83	145	5	438	439	6	598	13	339	348
-10	46	53	-7	779	815	540	7	162	144	8	610	14	101	90
-9	214	220	-6	802	815	213	8	244	243	10	293	15	118	110
-8	1223	1271	-5	873	897	576	9	97	95	12	76	18	69	73
-7	39	6	-4	742	748	47	10	159	135	14	252	6, 4,	L	L
-6	389	392	-2	892	886	356	11	171	154	16	187	-17	236	249
-5	1029	986	-1	209	224	626	12	67	71	18	63	-16	169	175
-4	80	82	0	353	342	795	5, 9,	L	L	6, 2,	L	-15	118	113
-3	191	149	1	941	896	63	-9	57	36	-17	189	-14	142	97
-2	937	878	2	1171	1121	659	-7	207	207	-16	102	-13	164	136
-1	583	555	3	371	348	981	-6	450	476	-15	339	-12	264	265
0	283	281	4	176	155	560	-5	520	521	-14	437	-11	308	293
1	558	522	5	689	658	302	-2	274	252	-13	619	-10	507	488
2	1664	1496	6	533	515	203	-1	444	441	-12	84	-9	126	102
3	304	345	7	366	336	167	1	674	688	-11	310	-7	80	85
4	99	53	8	377	373	92	2	75	50	-10	236	-6	391	382
5	105	86	9	459	450	196	3	114	94	-9	142	-5	951	985
6	224	208	10	85	75	143	4	373	391	-8	753	-4	162	137
7	127	139	11	405	397	227	5	432	412	-7	159	-3	764	808
8	601	608	12	162	162	137	6	295	303	-6	53	-2	128	121
9	469	488	13	329	326	L	7	124	122	-5	194	-1	1201	1181
10	196	185	15	287	282	116	8	36	21	-4	464	0	899	938
11	625	636	16	131	138	188				-3	95	1	652	650
12	370	374	18	80	69	375				-2	408	2	240	270
13	141	131				113				-1	1277	3	836	818

6, 4,	L	9	83	70	-14	75	90	-10	517	521	-6	96	102	6	465	433
4 564	535	10	64	45	-13	153	155	-9	251	233	-5	181	184	7	103	108
5 524	513	11	105	107	-12	524	555	-8	294	314	-4	353	375	8	91	99
6 755	738	12	122	115	-11	377	382	-7	131	130	-3	185	167	9	123	130
7 159	136	6, 8,	L	L	-10	168	165	-6	755	780	-2	194	207	10	214	208
8 404	389	-9	123	113	-9	66	69	-5	591	601	-1	1388	1371	11	223	215
9 594	585	-8	60	48	-8	469	515	-4	83	95	0	313	328	7, 9,	L	19
10 185	184	-6	169	178	-7	304	314	-3	798	794	2	299	308	-7	40	19
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-15 242	256	4	434	436	2	1117	1014	6	266	269	11	347	328	3	351	372
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-11 376	392	10	402	404	6	97	109	10	130	139	7, 7,	L	L	8	81	75
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-9 449	451	-5	146	148	8	319	340	13	213	208	-11	148	125	-18	133	123
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-6 151	167	-2	146	126	12	252	254	16	75	87	-8	474	436	-12	275	314
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18, 0, L	14	115	115	-1	250	265	-8 185 195	19, 7, L	6	140	136	6	140
-12 391 425	18, 4, L	18	4, L	0	508	500	-7 102 93	-3	97	86	8	130	127
-10 579 600	-14 65	63	63	1	96	84	-6 153 152	-2	55	40	9	94	109
-8 688 770	-13 140	127	127	2	87	72	-4 88 64	0	265	272	10	229	233
-6 111 134	-12 133	134	134	3	71	74	-3 336 329	1	144	146	11	109	124
-4 158 165	-11 64	61	61	4	180	179	-2 137 137	2	402	400	13	109	104
-2 415 394	-10 128	116	116	5	203	211	-1 167 167	3	32	33	20, 4, L	20, 4, L	55
0 428 453	-9 574	534	534	6	129	132	0 30 6	20, 0, L	-14	79	84	-12 81	55
2 521 521	-8 442	368	368	19, 1, L	1, L	1, L	2 292 298	-14	79	84	-11 58	69	69
4 415 429	-7 139	122	122	-14 85	85	87	3 458 471	-12	196	216	-10 207	225	225
6 44 57	-6 199	202	202	-12 118	117	117	4 374 378	-10	64	91	-9 265	264	264
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12 304 311	-3 353	358	358	-9 172	143	143	7 268 264	-4	587	593	-5 50	42	42
18, 2, L	-1 158	147	147	-8 232	227	227	8 225 231	-2	966	1003	-4 188	180	180
-12 232 227	0 203	197	197	-7 134	133	133	9 156 168	0	617	619	-3 565	550	550
-11 62 49	1 436	443	443	-4 177	180	180	10 155 149	2	258	257	-1 322	320	320
-10 110 77	2 145	125	125	-3 296	298	298	11 66 71	4	373	362	0 129	128	128
-9 313 280	3 125	126	126	-2 144	131	131	12 105 113	6	166	173	1 156	149	149
-8 382 352	5 223	237	237	-1 571	600	600	13 197 183	8	50	47	2 174	184	184
-7 628 610	6 134	131	131	0 84	80	80	19, 5, L	12	90	88	3 165	164	164
-6 345 358	7 60	45	45	1 123	129	129	-10 205 193	20, 2, L	376	376	5 384	376	376
-5 177 175	9 298	292	292	2 275	297	297	-9 296 298	-12	180	170	7 113	116	116
-4 322 318	10 82	76	76	3 248	256	256	-8 68 82	-11	411	407	10 87	84	84
-3 239 243	11 141	144	144	4 269	277	277	-7 317 312	-10	98	110	11 60	49	49
-2 447 449	12 106	122	122	5 492	493	493	-6 98 88	-8	221	217	20, 6, L	20, 6, L	49
-1 960 970	13 252	248	248	6 314	316	316	-5 418 412	-7	179	172	-10 147	119	119
0 540 541	18, 6, L	18	6, L	7 135	123	123	-4 129 123	-6	443	430	-9 115	135	135
1 456 469	-11 146	183	183	8 69	78	78	-3 75 74	-5	592	566	-8 167	163	163
2 32 24	-10 88	65	65	10 55	38	38	-1 137 144	-4	251	253	-7 33	39	39
3 489 510	-9 78	81	81	11 186	191	191	0 325 322	-3	145	144	-6 56	41	41
4 113 119	-8 202	219	219	12 81	68	68	2 173 179	-2	76	82	-5 495	501	501

20, 6,	L	290	259	22, 0,	L	444	446	-1	338	344	-2	106	106
-4 286	265	-5 318	311	-10 401	366	-6 197	159	1	242	255	-1	49	44
-3 88	86	-3 60	35	-8 397	406	-5 193	178	2	160	153	1	79	79
-2 166	147	-2 307	319	-6 483	539	-4 81	85	3	32	20	2	206	232
-1 88	80	-1 82	86	-4 188	187	-3 118	113	4	30	38	3	46	46
1 232	224	0 500	509	0 567	559	-2 162	161	5	189	152	4	170	179
2 292	291	1 159	153	2 374	400	-1 293	298	6	154	192	5	89	89
3 147	139	3 553	564	4 151	157	1 241	240	7	284	299	7	101	108
4 206	206	4 259	259	6 104	121	2 228	236	8	77	89	8	173	182
5 95	92	6 205	210	8 87	109	3 203	209	11	124	127	24, 0,	L	L
6 216	227	7 225	226	10 245	257	4 99	90	23, 3,	L	L	-10	217	226
21, 1,	L	9 328	334	12 100	76	5 120	123	-10	108	102	-8	192	186
-12 328	323	11 95	97	22, 2,	L	6 71	82	-9	251	219	-6	200	208
-11 104	402	12 107	105	-12 92	110	7 53	69	-8	244	244	-2	259	261
-9 304	281	21, 5,	L	-11 76	73	22, 6,	L	-7	136	144	0	381	408
-8 143	133	-11 234	242	-10 166	166	-9 180	205	-6	164	172	2	167	164
-6 233	225	-10 155	150	-7 287	254	-8 106	89	-5	177	166	4	181	166
-5 88	85	-9 68	63	-6 163	148	-7 125	133	-4	255	250	6	245	247
-4 72	72	-8 148	133	-5 436	432	-6 109	113	-3	116	109	8	214	217
-3 173	165	-7 81	85	-4 270	283	-4 291	306	-2	134	129	10	163	169
-2 210	210	-6 155	137	-3 51	45	-3 184	189	-1	251	254	24, 2,	L	L
-1 116	115	-5 193	201	-2 143	140	-1 73	22	0	147	139	-10	77	60
0 71	77	-4 114	109	-1 124	137	0 106	112	1	284	275	-9	249	242
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2 163	169	-2 215	213	1 394	404	2 367	387	3	70	67	-7	99	87
3 43	48	-1 64	59	2 388	391	3 202	210	4	137	139	-6	227	226
4 183	198	0 98	103	3 323	340	5 116	129	5	323	333	-5	156	154
5 35	27	1 225	234	4 69	70	6 109	117	6	155	158	-4	228	220
6 51	41	2 112	112	5 175	183	23, 1,	L	8	157	161	-3	420	426
7 184	196	3 91	94	6 145	144	-12 82	78	9	71	66	-2	240	251
8 231	249	4 133	143	8 81	72	-10 128	124	10	132	131	-1	219	231
9 366	382	5 165	175	9 227	231	-8 256	240	23, 5,	L	L	0	59	64
10 185	184	6 274	297	10 173	159	-7 214	187	-8	200	186	4	80	77
21, 3,	L	7 185	193	22, 4,	L	-6 282	282	-7	304	303	5	127	128
-11 174	159	8 65	50	-11 80	69	-5 70	67	-6	120	109	6	143	134
-10 92	104	9 104	113	-10 89	91	-4 77	80	-5	250	266	7	53	59
-9 111	95	10 59	60	-9 104	110	-3 83	81	-4	144	148			
-6 117	101			-8 36	21	-2 111	119	-3	226	236			

24, 4,	L	75	4	37	35	26, 2,	L	27, 3,	L
-9 72			5	60	42	-8 81	81	-6 162	163
-8 203		212	6	112	135	-6 32	44	-5 102	111
-6 112	105		7	96	104	-5 174	165	-4 207	211
-5 146	153		8	81	70	-3 75	88	-3 202	189
-3 46	39		9	180	195	-2 117	130	-2 195	192
-2 94	91		25, 3,		L	0 51	44	-1 84	81
-1 240	254		-9 216	200		2 92	84	0 41	39
0 100	100		-7 140	126		3 160	167	1 171	172
1 235	246		-6 70	55		4 192	196	2 140	136
2 88	94		-3 254	239		5 265	282	3 164	155
3 85	88		0 244	253		6 92	102	4 110	107
5 82	74		2 230	235		7 103	112	5 120	115
7 293	297		3 178	177		26, 4,	L	28, 0,	L
8 51	32		5 248	253		-5 273	286	-6 269	254
24, 6,	L		6 222	228		-4 125	118	-2 160	151
-6 51	53		8 116	119		-3 161	169	0 145	152
-5 105	121		25, 5,	L		-2 84	93	2 129	149
-4 93	89		-6 223	221		-1 136	137	28, 2,	L
-3 243	274		-4 65	65		0 64	74	-4 149	154
-2 282	293		-3 241	258		1 140	147	-3 106	113
0 60	53		-2 95	104		2 55	63	-2 93	89
3 115	132		-1 90	106		3 183	190	-1 177	181
4 197	214		0 209	216		4 66	82	0 133	140
25, 1,	L		1 59	51		5 87	96	1 125	136
-10 244	240		2 59	59		27, 1,	L	2 64	68
-9 85	75		3 106	123		-7 90	67	3 66	59
-7 235	221		4 115	125		-6 183	177	4 65	57
-6 148	135		5 73	84		-5 249	228	29, 1,	L
-5 114	104		26, 0,	L		-4 168	162	-2 81	77
-4 184	176		-8 255	238		-3 85	79	-1 61	59
-3 150	151		-6 177	154		-2 107	112	0 109	102
-2 104	99		-4 376	390		-1 163	159		
-1 46	34		0 187	195		1 197	196		
0 201	192		2 373	372		2 41	26		
1 249	249		4 171	185		3 93	98		
2 101	101		8 104	86		4 187	181		
3 221	224								

